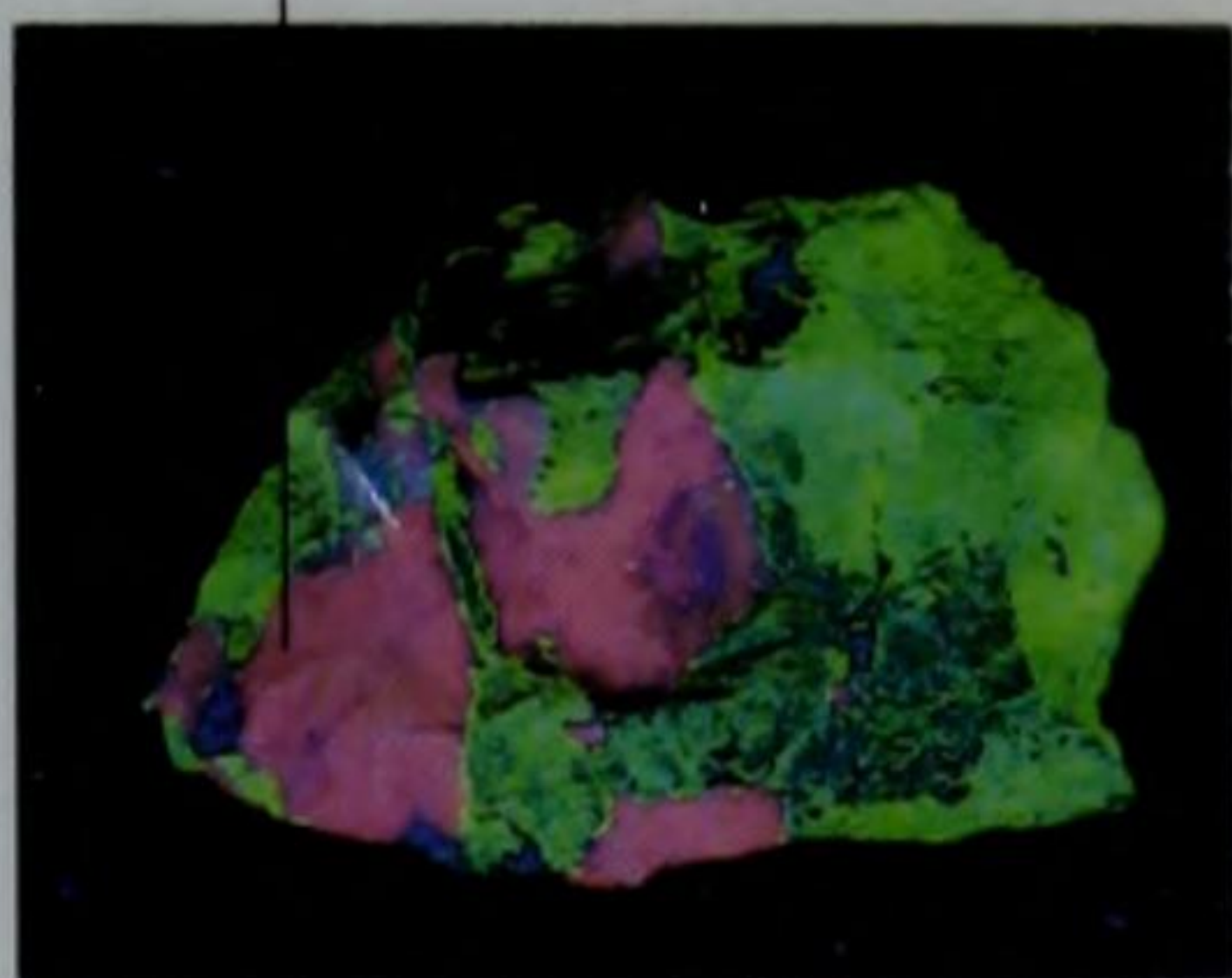


# THE VISUAL DICTIONARY of

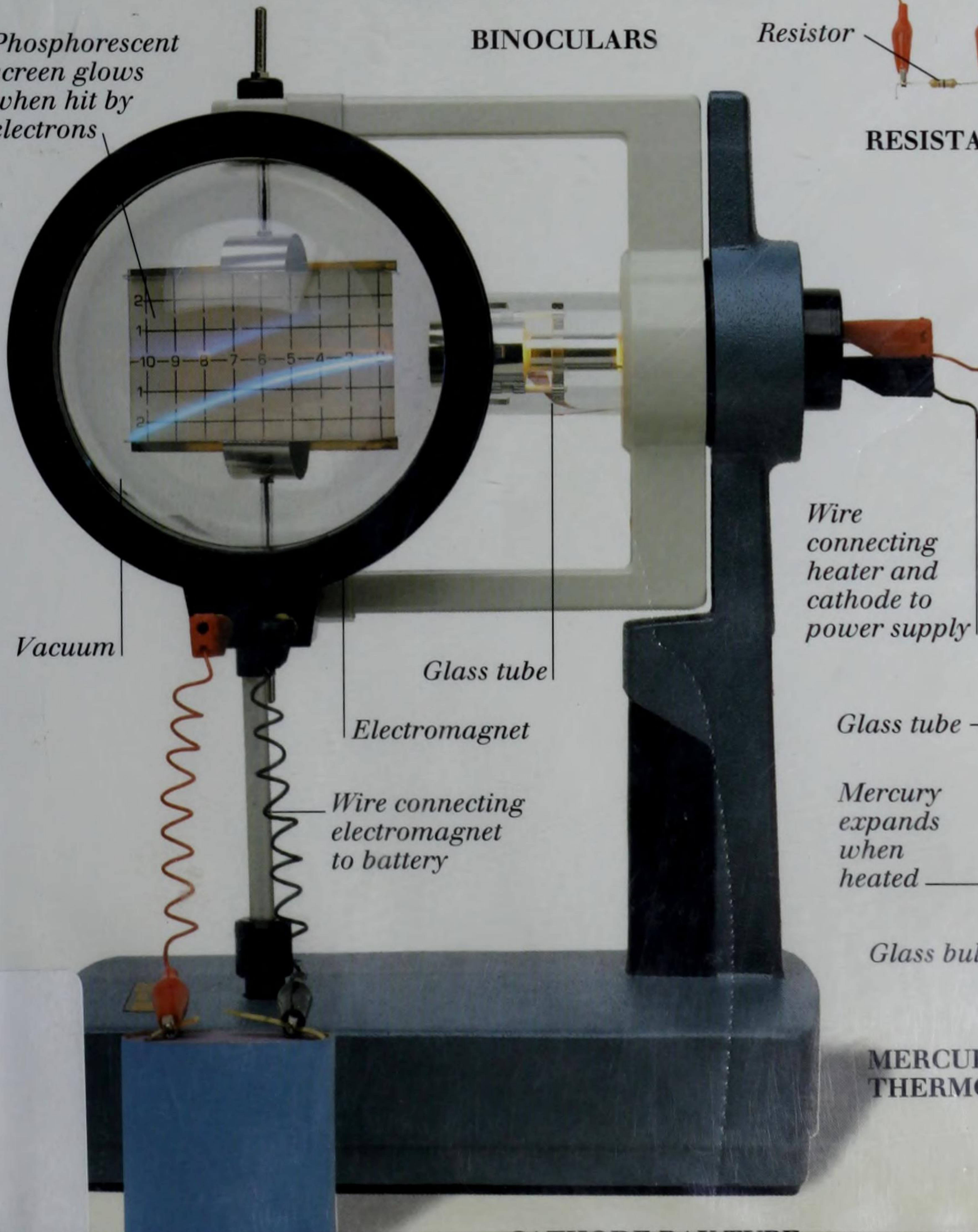
# PHYSICS

In ultraviolet light, rock  
appears brightly coloured



WILLEMITE

Phosphorescent  
screen glows  
when hit by  
electrons



Vacuum

Glass tube

Electromagnet

Wire connecting  
electromagnet  
to battery

BINOCULARS



Objective  
lens

Eyepiece

Resistor

RESISTANCE



Reading:  
0.089 A

Ammeter

Battery

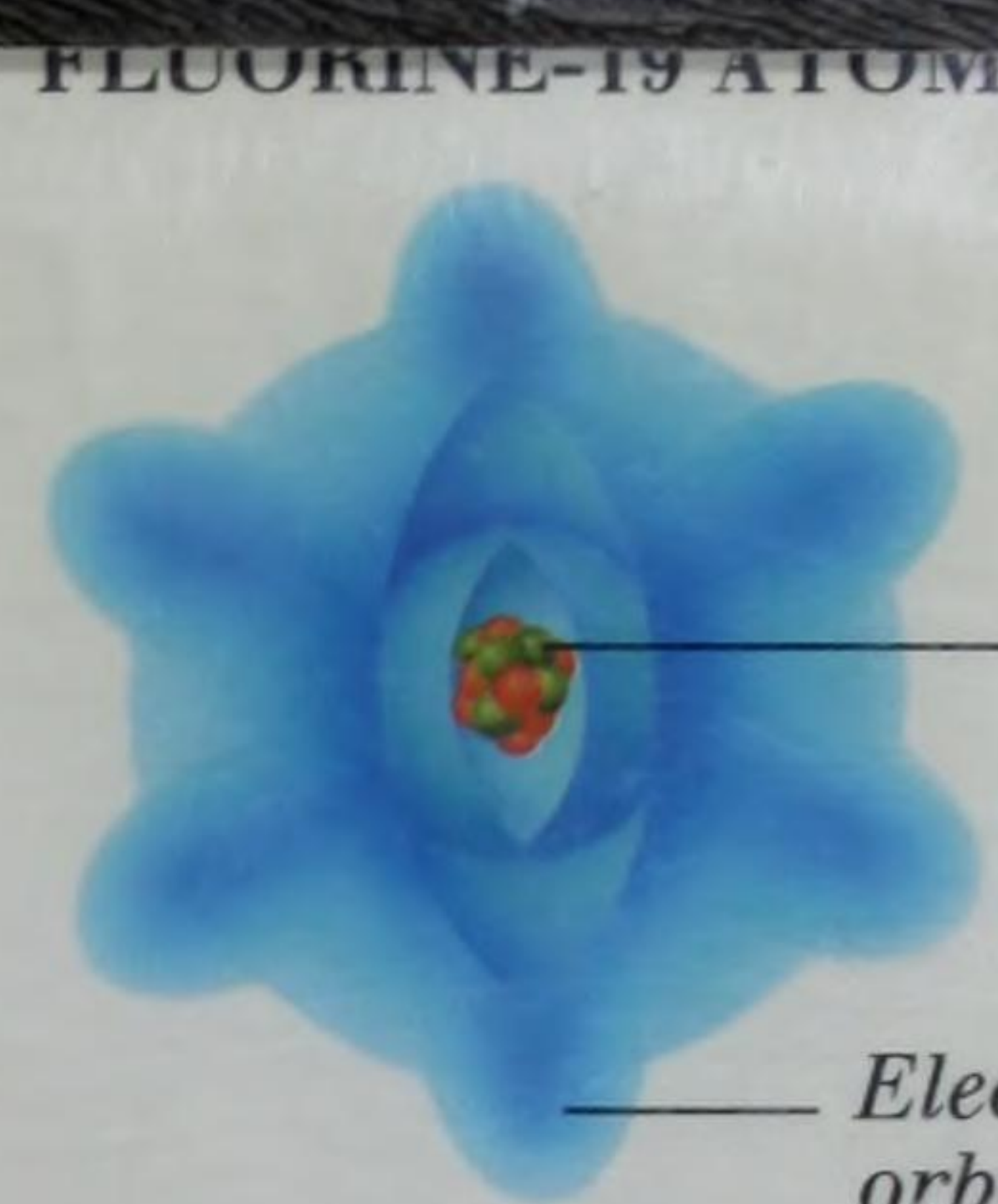
Wire  
connecting  
heater and  
cathode to  
power supply

Glass tube

Mercury  
expands  
when  
heated

Glass bulb

MERCURY  
THERMOMETER

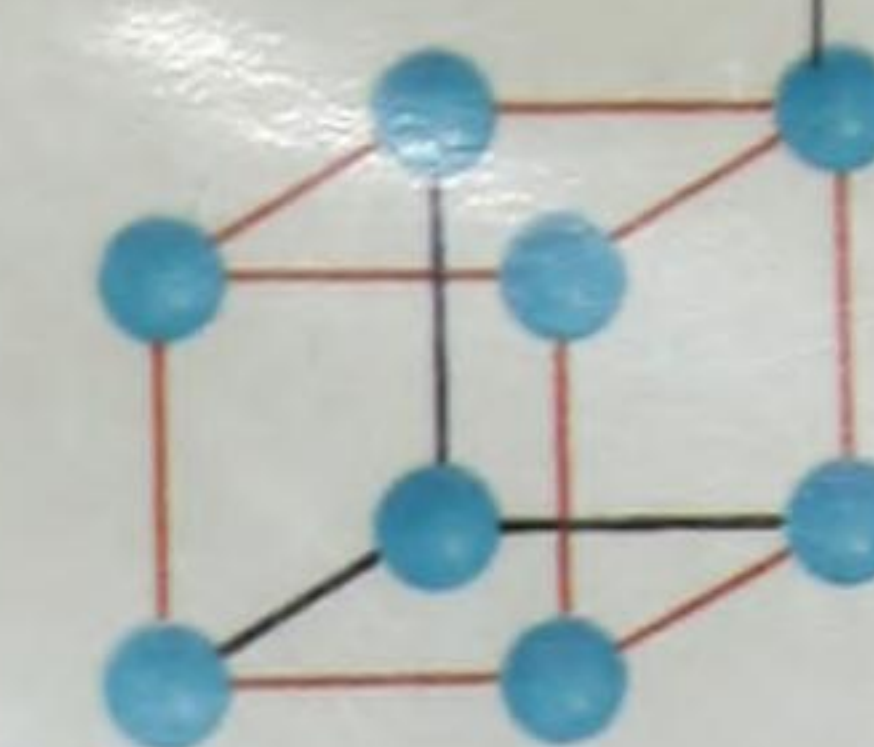


FLUORINE-19 ATOM

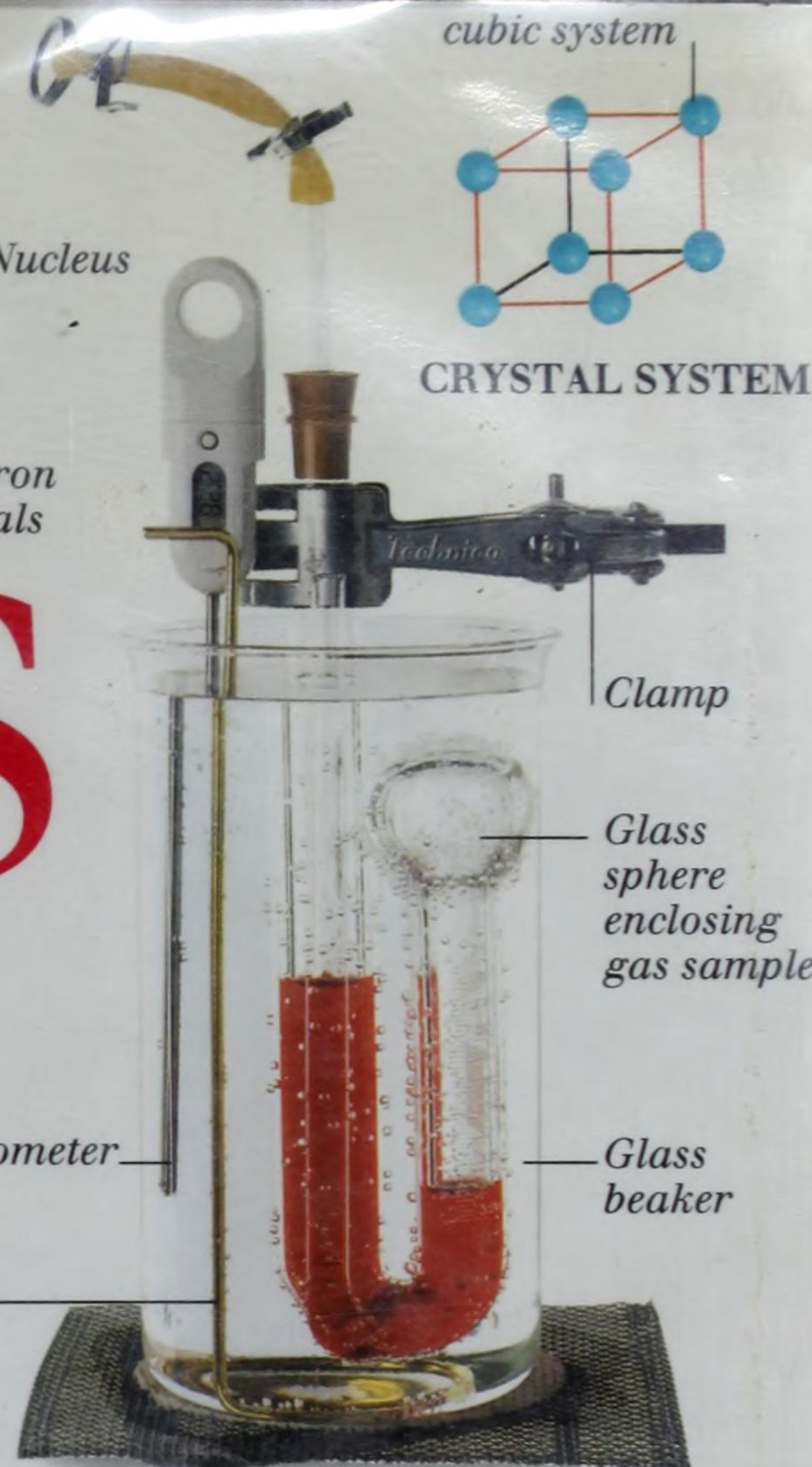
Nucleus

Electron  
orbitals

cubic system



CRYSTAL SYSTEM



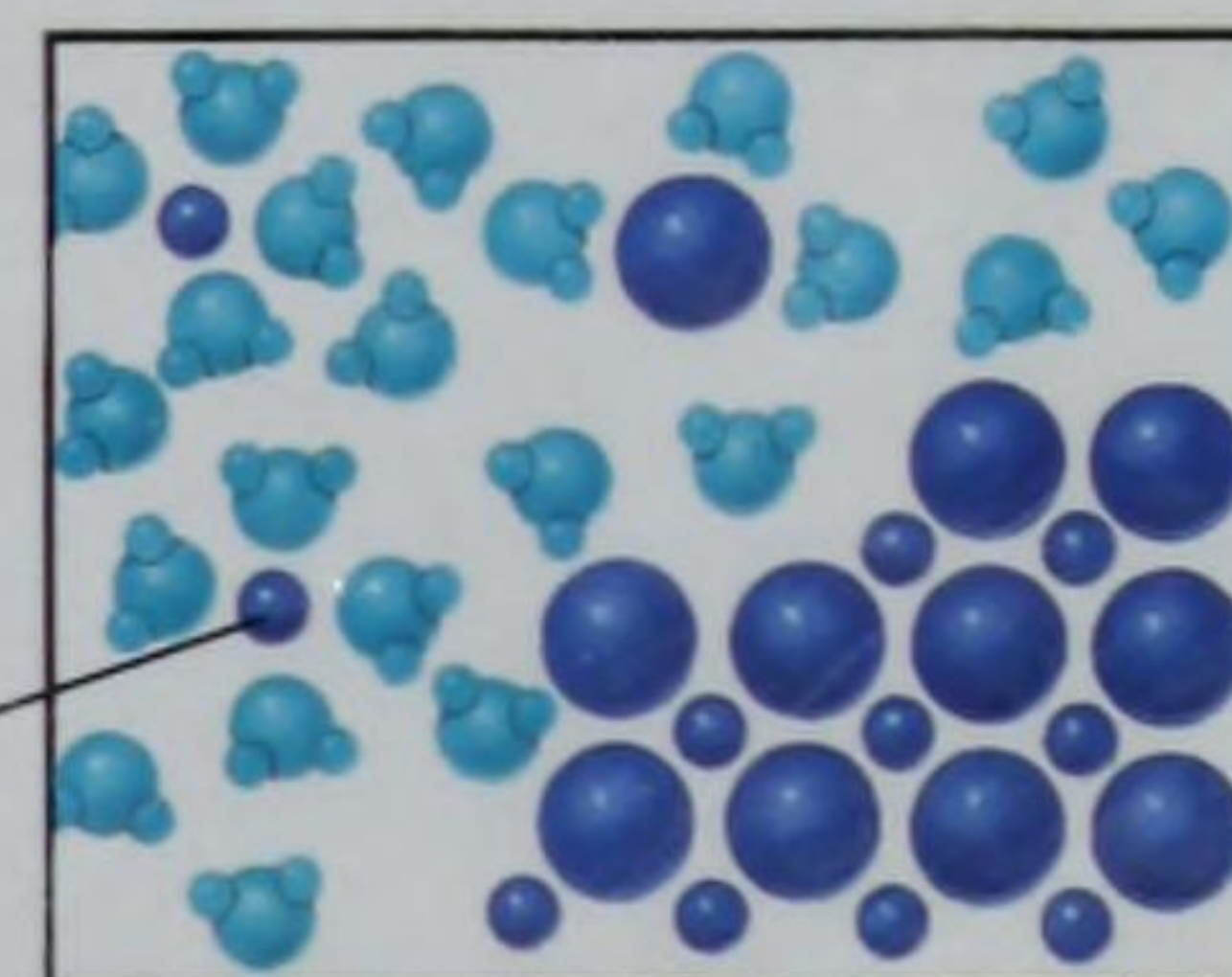
Clamp

Glass  
sphere  
enclosing  
gas sample

Glass  
beaker

CHARLES' LAW

Atoms  
break  
away

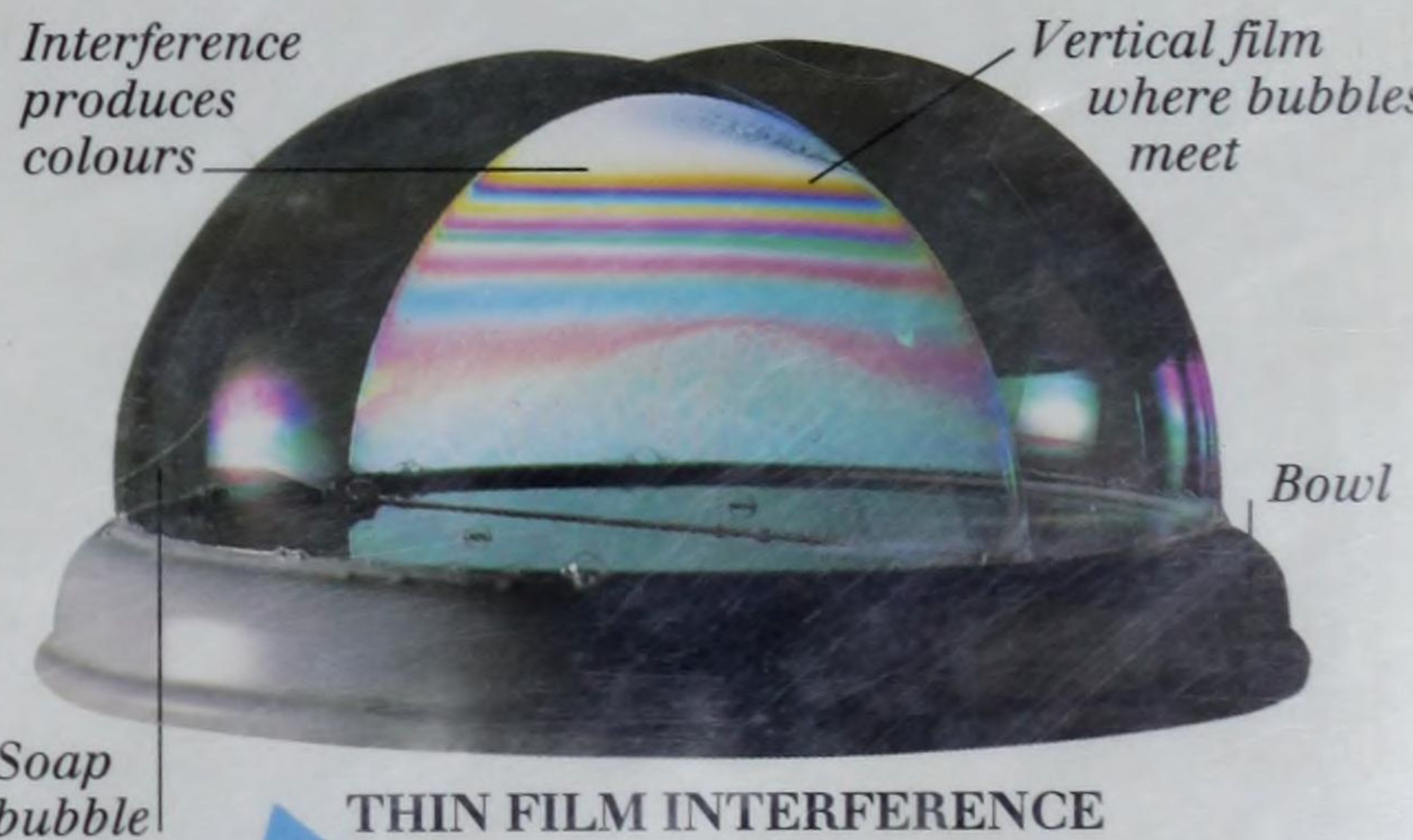


DISSOLVING SOLID



GREEN LED

Interference  
produces  
colours

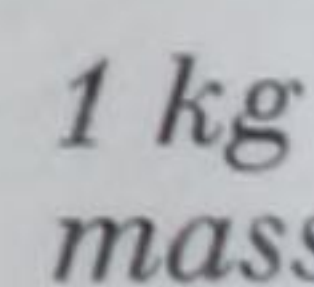


THIN FILM INTERFERENCE

Reaction force

Soap  
bubble

1 kg  
mass



Gravity

Reaction force

Newton  
meter

2.4 N force  
required to  
hold mass  
on slope

FORCES ON A  
SHALLOW SLOPE

Newton  
meter

2.4 N force  
required to  
hold mass  
on slope

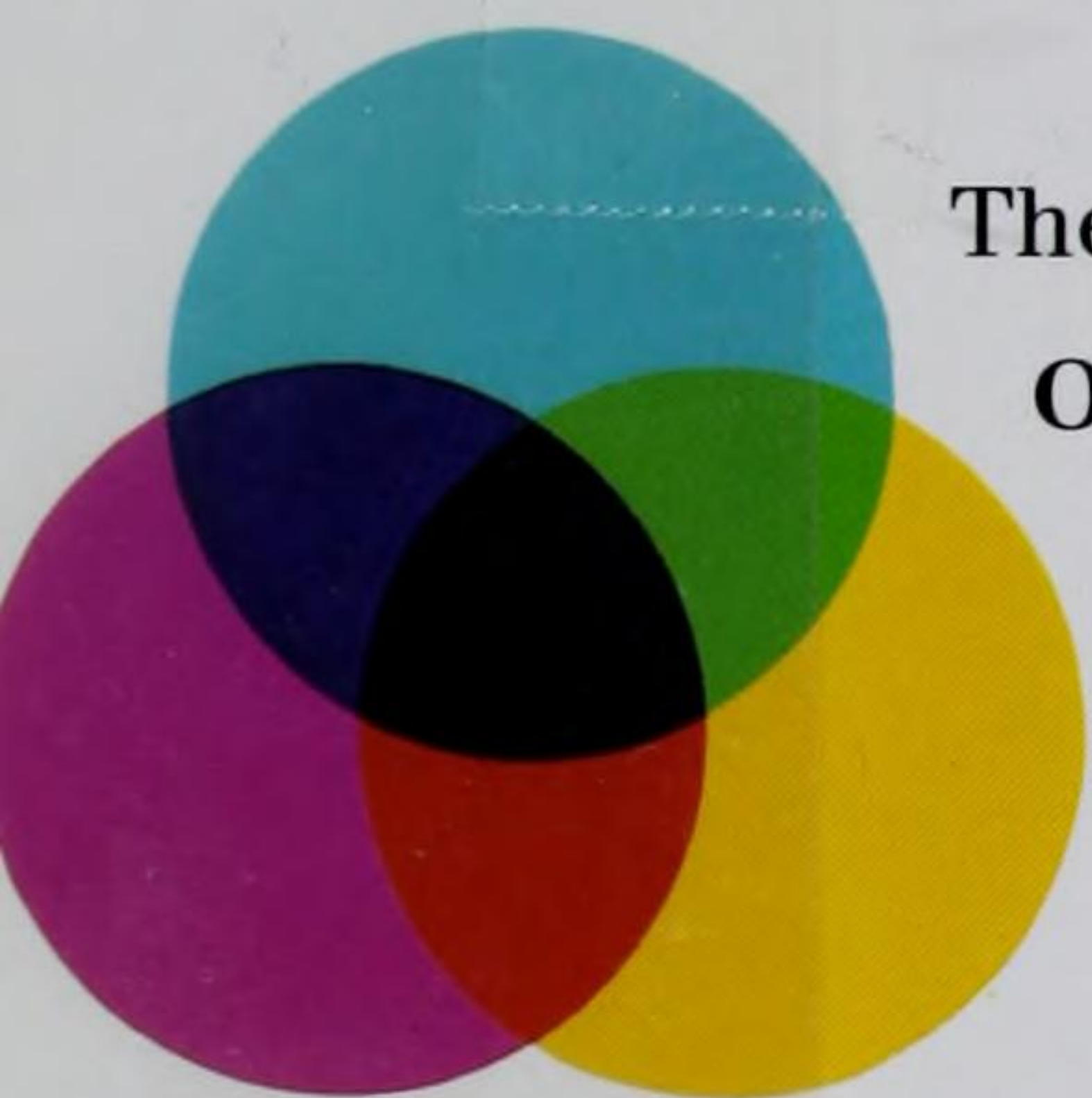
FORCES ON A  
SHALLOW SLOPE



EYEWITNESS VISUAL DICTIONARIES

# THE VISUAL DICTIONARY *of* PHYSICS

Here is an entirely new kind of dictionary — one that is packed with superb, full-color photographs and illustrations, and with thousands of scientific terms.



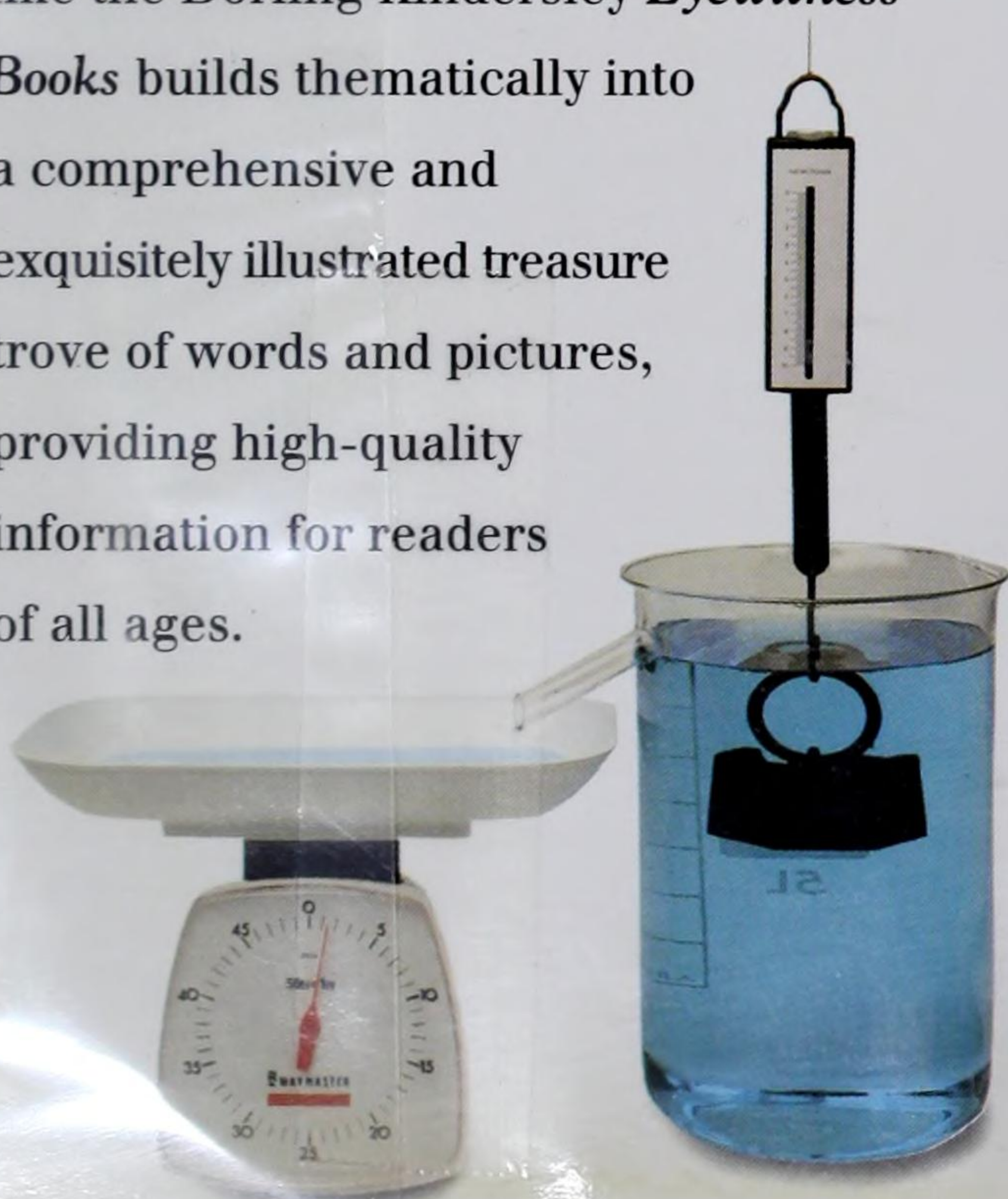
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OF PHYSICS** will give you  
instant access to the  
specialized vocabulary  
relating to physics in  
a way that is clear,

informative, and easy to understand.

If you have heard of a particular scientific process, but don't understand how it works, then turn to the labels around the illustrations.

Alternatively, if you know a term but don't know exactly what it refers to, then the comprehensive index will direct you to the illustration that bears the name.

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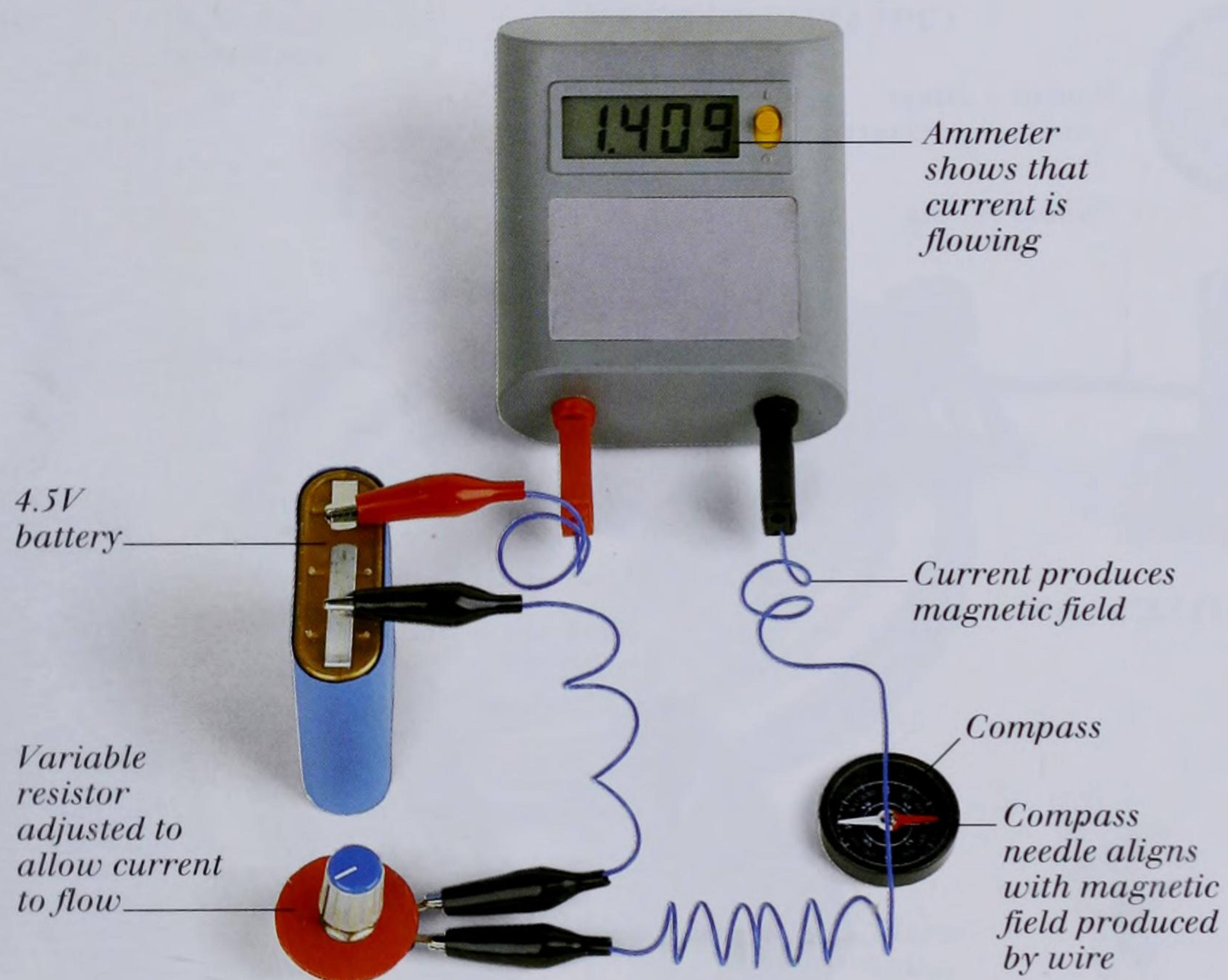




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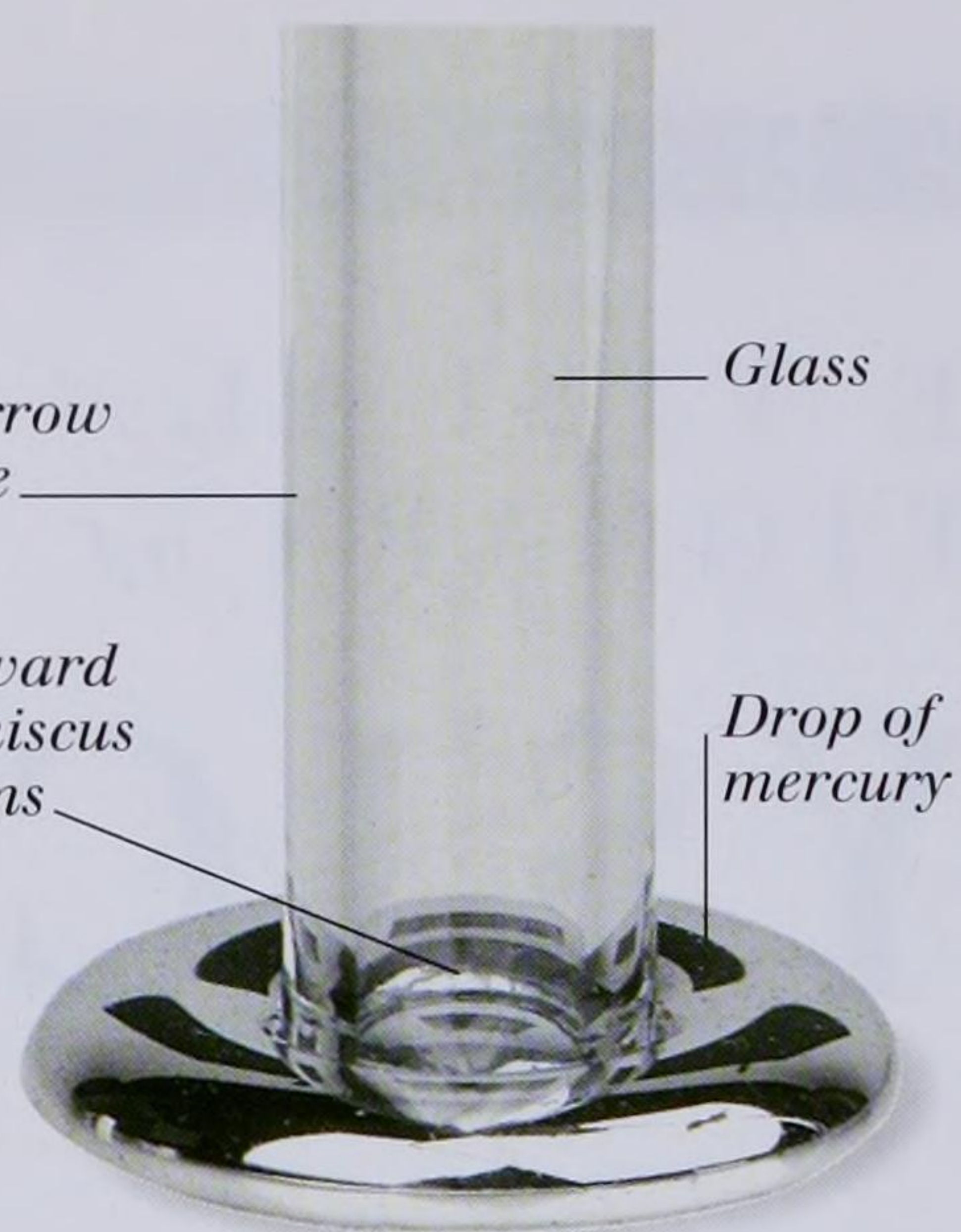
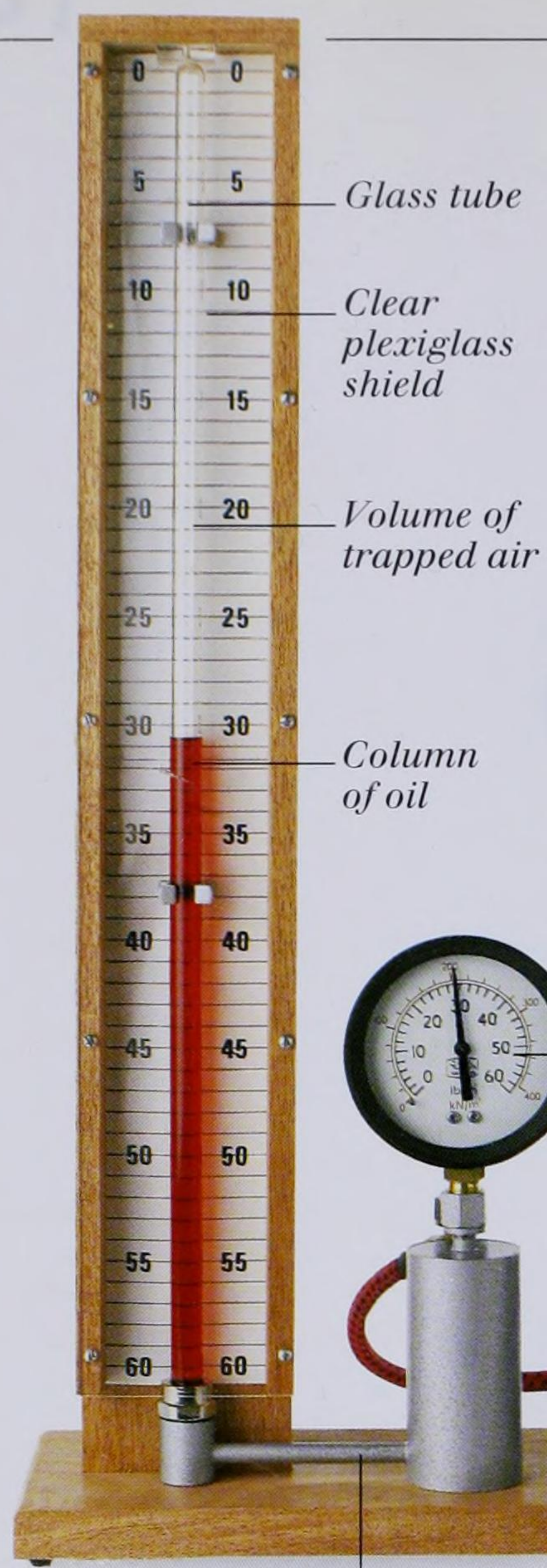


THE VISUAL  
DICTIONARY of  
**PHYSICS**



**ELECTROMAGNETISM AFFECTING A COMPASS NEEDLE**

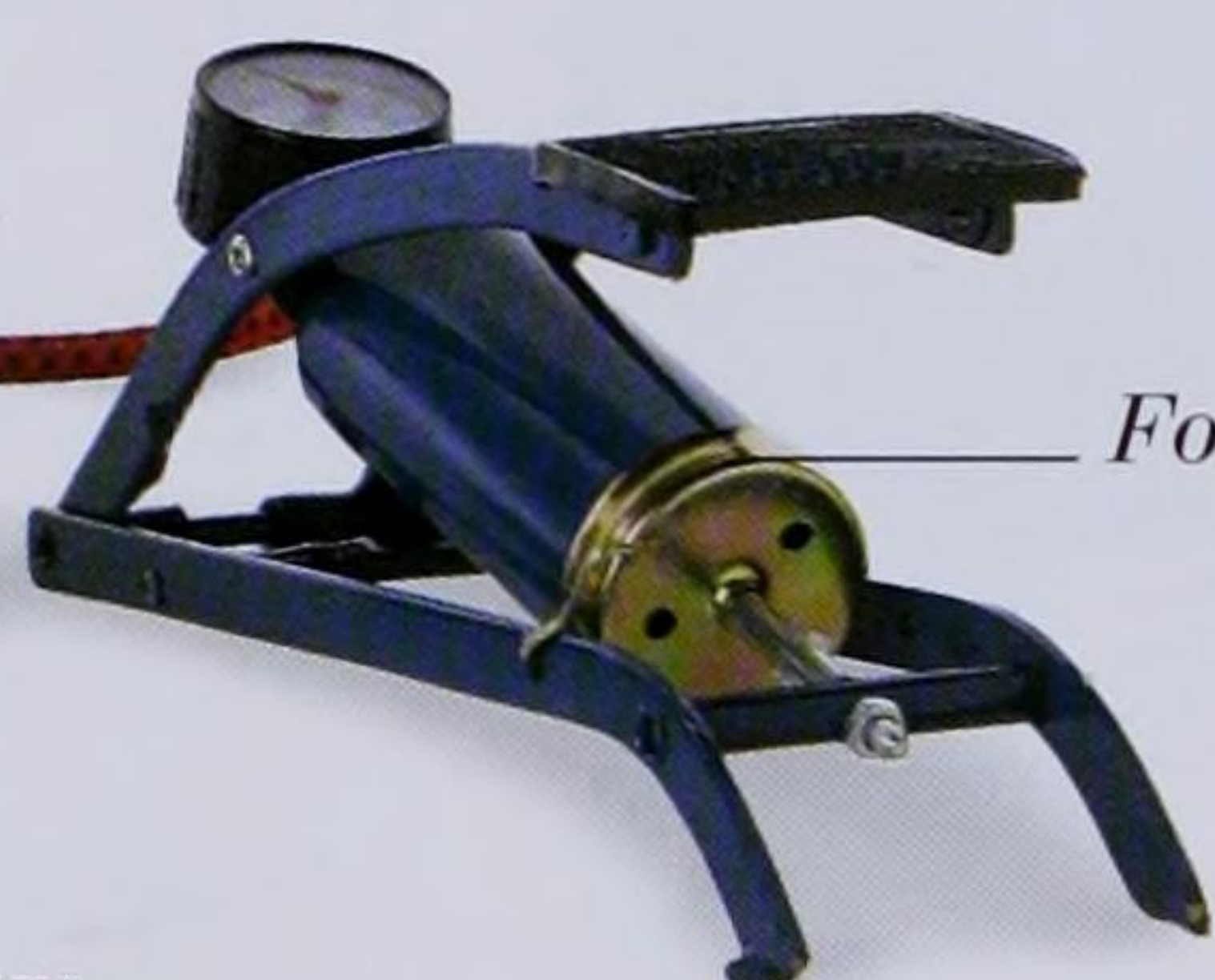




UPWARD MENISCUS



Rubber tubing



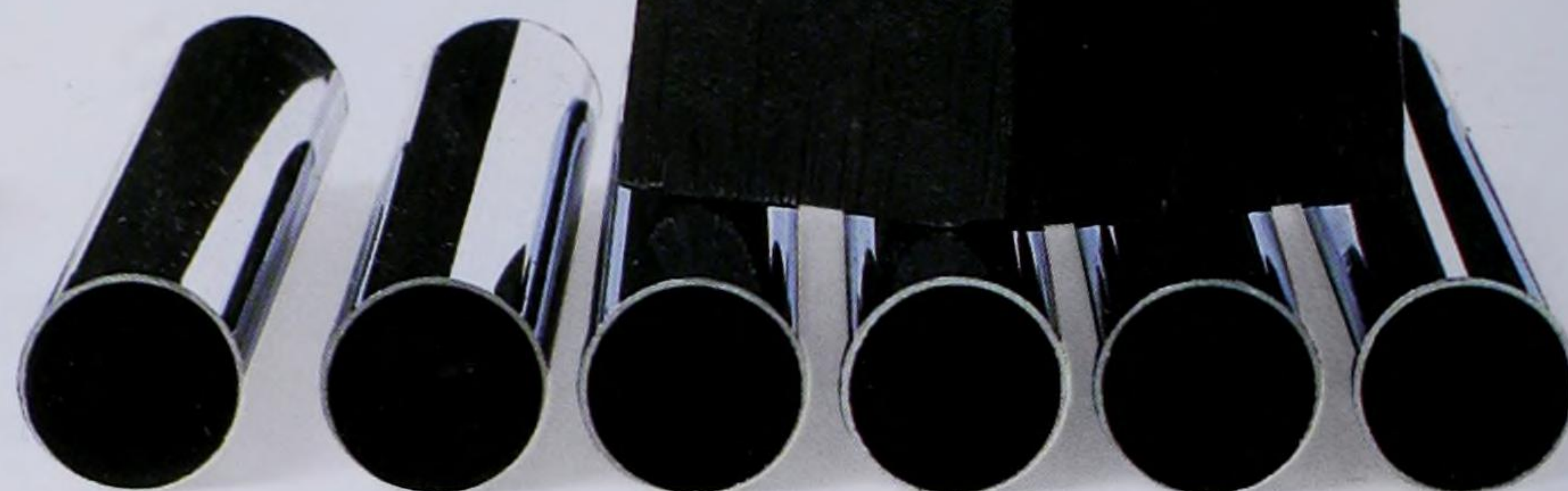
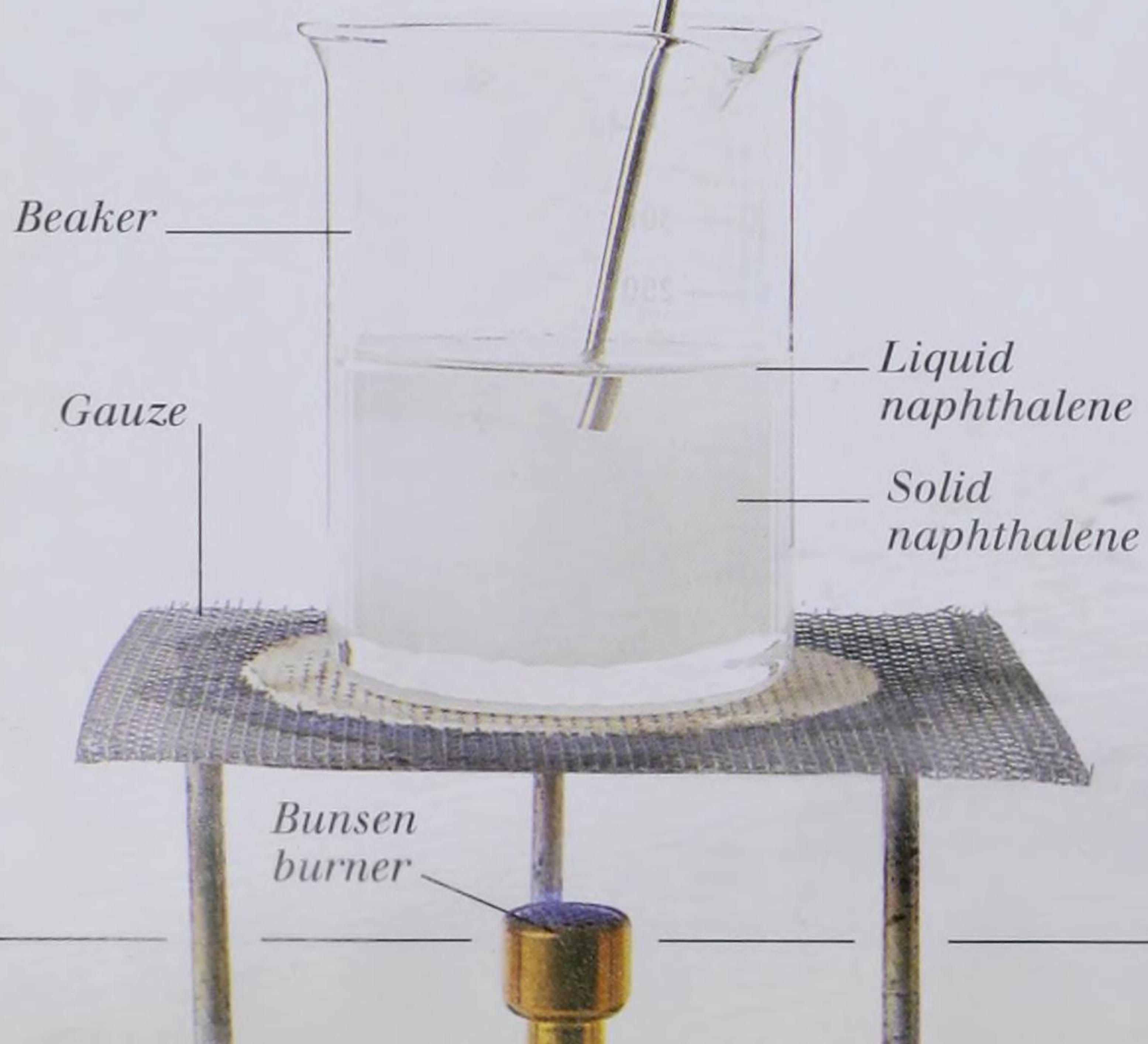
Connecting pipe

BOYLE'S LAW

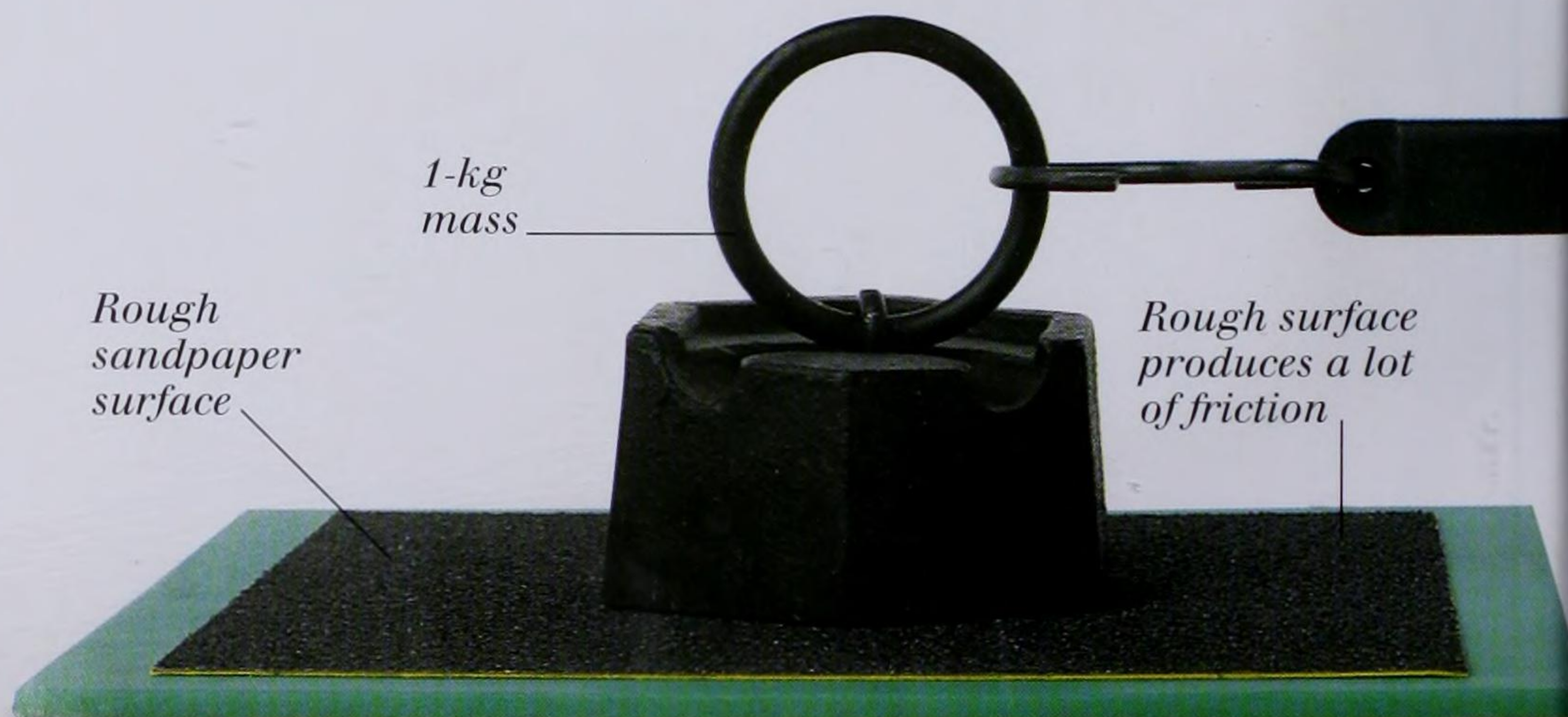
LATENT HEAT

Thermometer reads 80.5°C

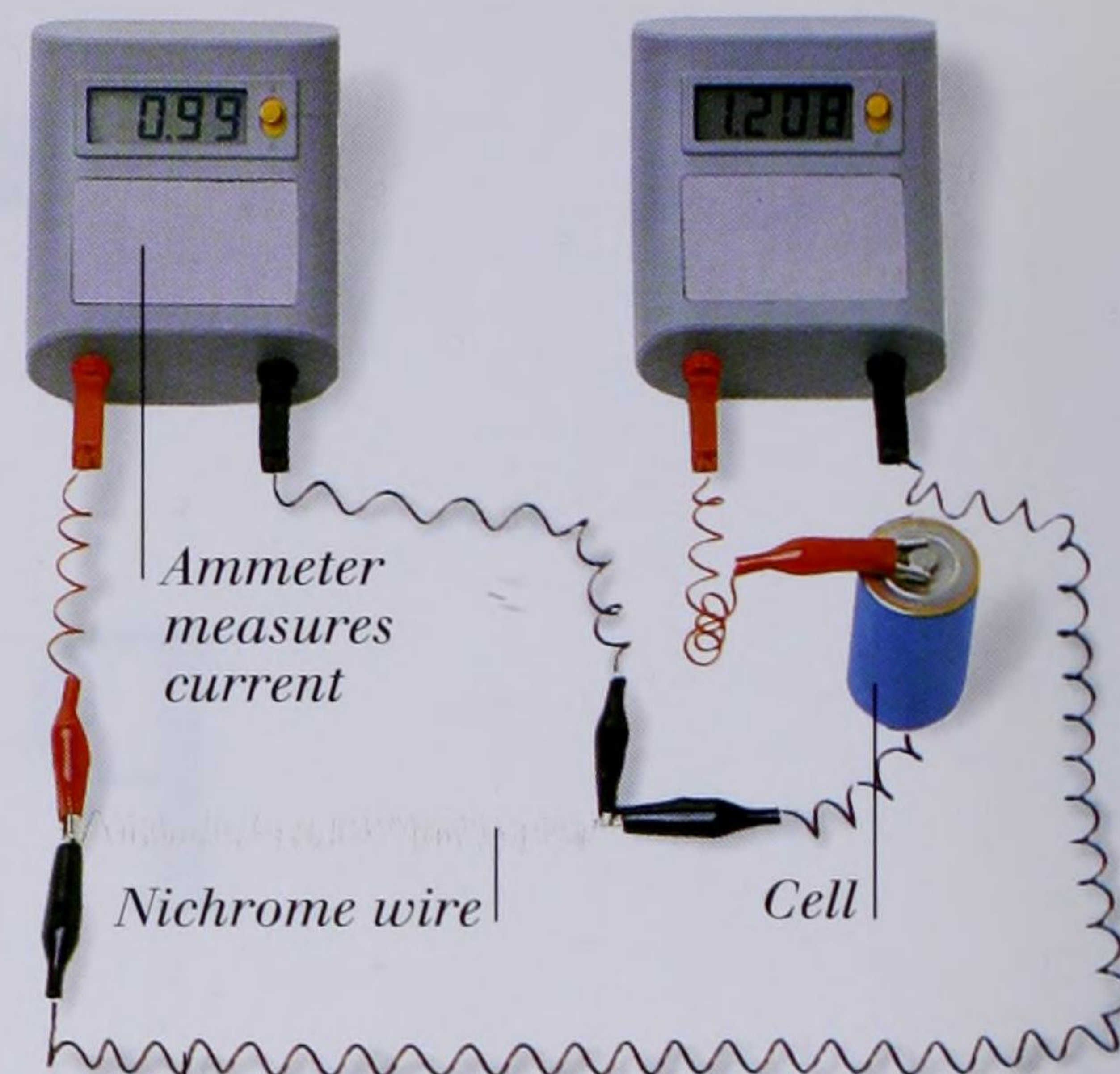
Temperature stays the same during melting



ROLLERS



FRICTION BETWEEN SURFACES



Thick connecting wires have little resistance

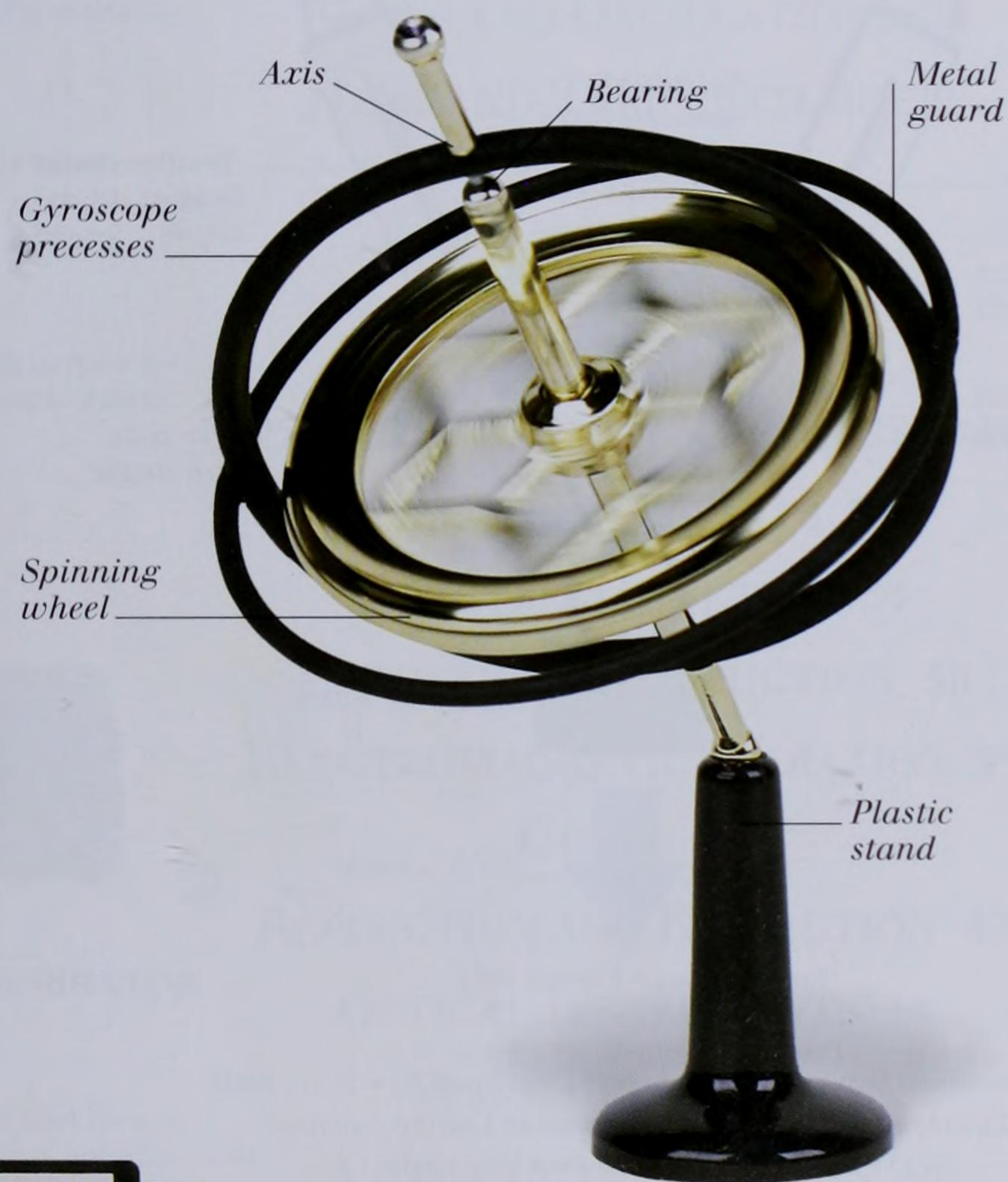
OHM'S LAW



EYEWITNESS VISUAL DICTIONARIES

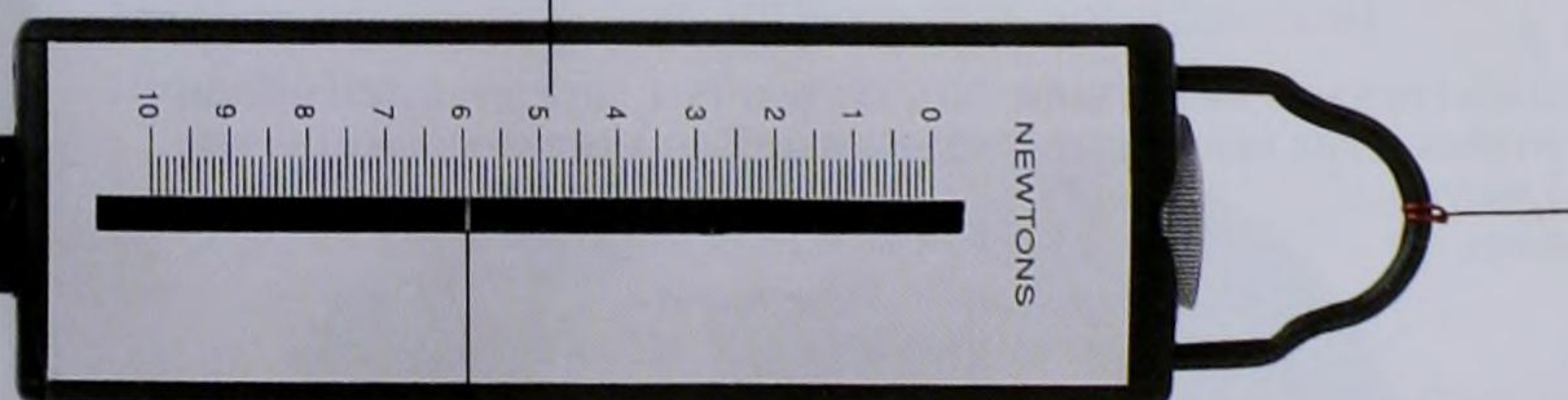
THE VISUAL  
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**PHYSICS**

written by  
Jack Challoner



GYROSCOPE

Newton  
meter



Newton meter measures  
limiting friction



DK PUBLISHING, INC





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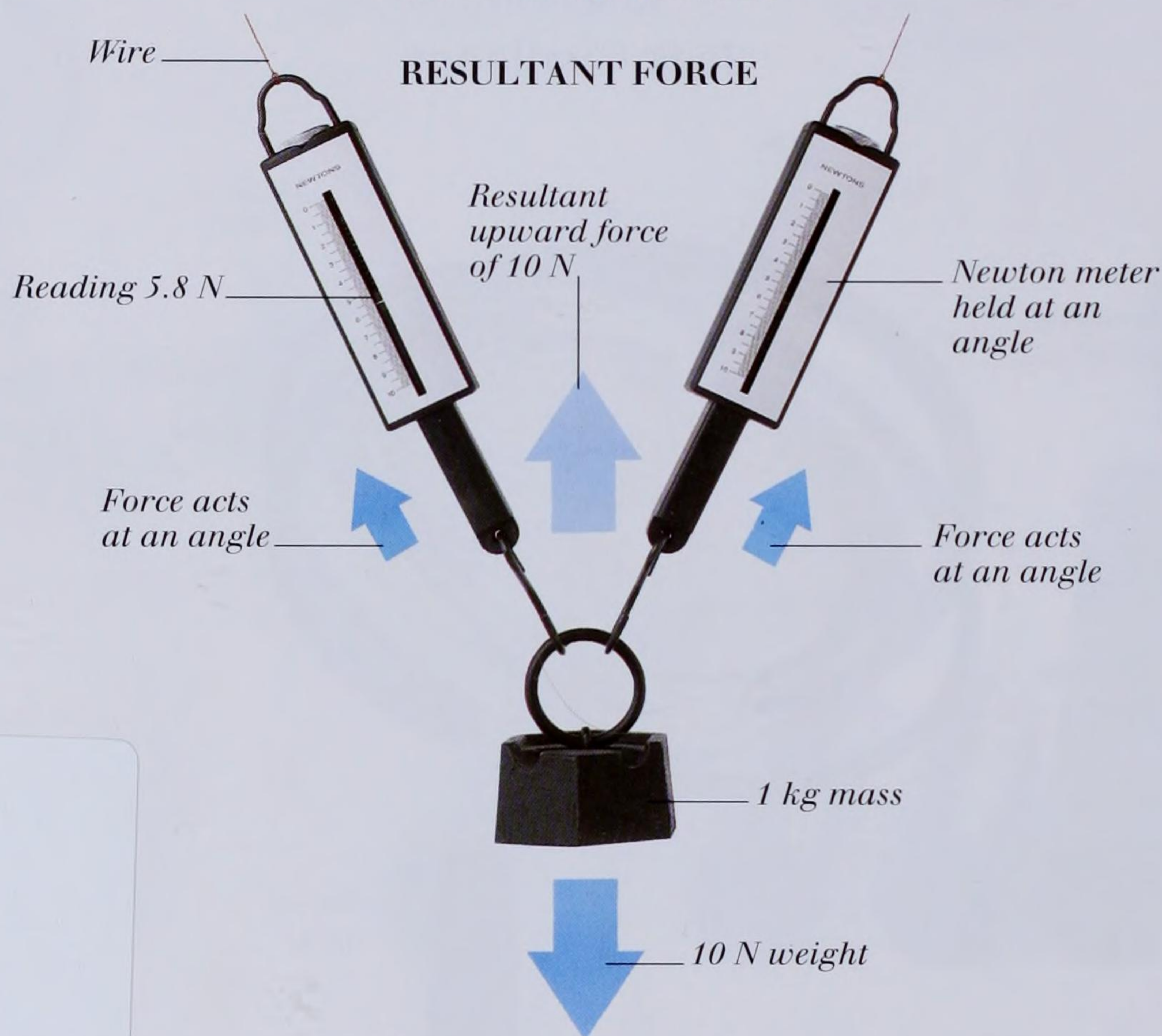
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PHOTOGRAPHY ANDY CRAWFORD

ILLUSTRATIONS CHRIS LYON, JANOS MARFFY

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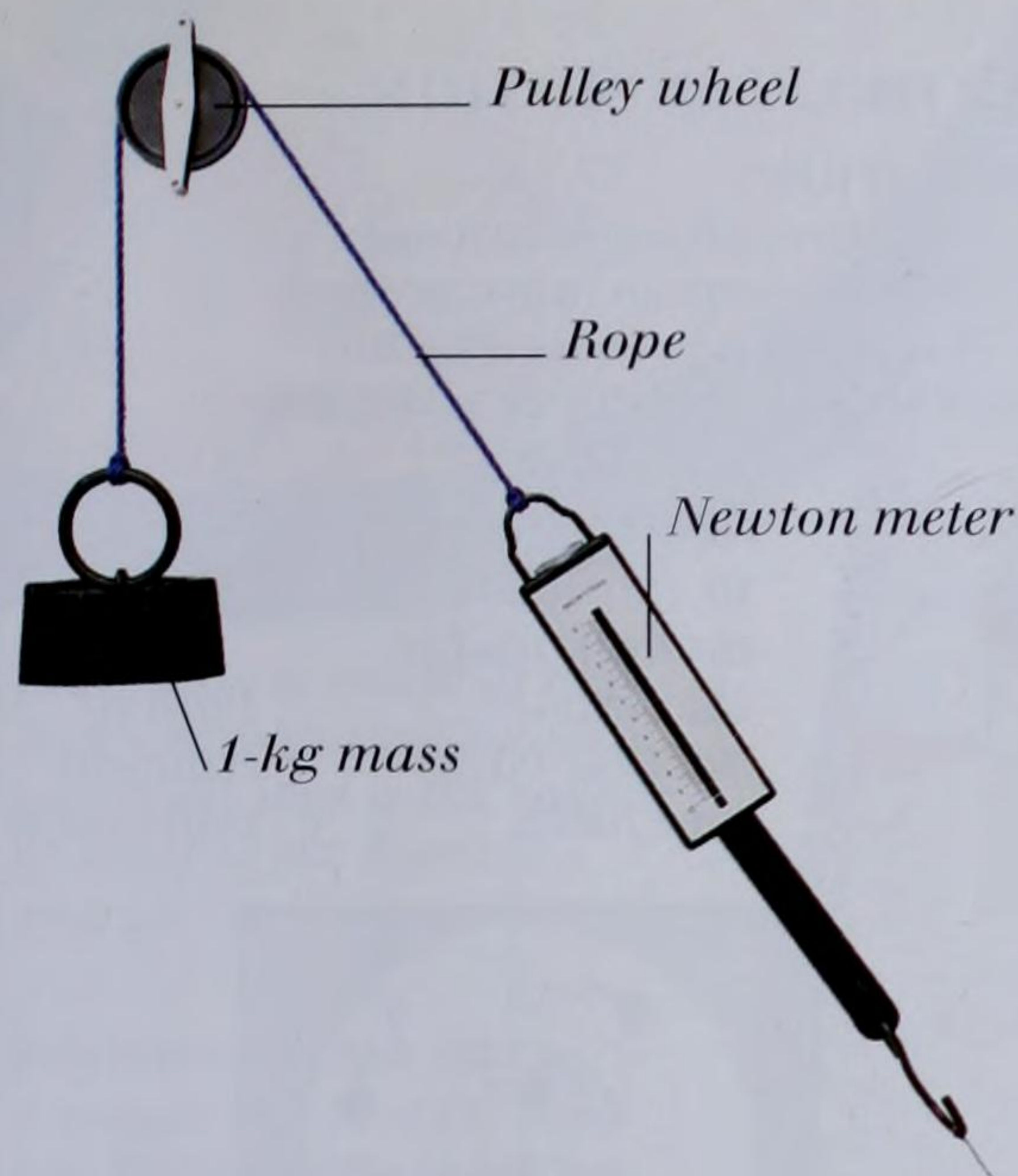
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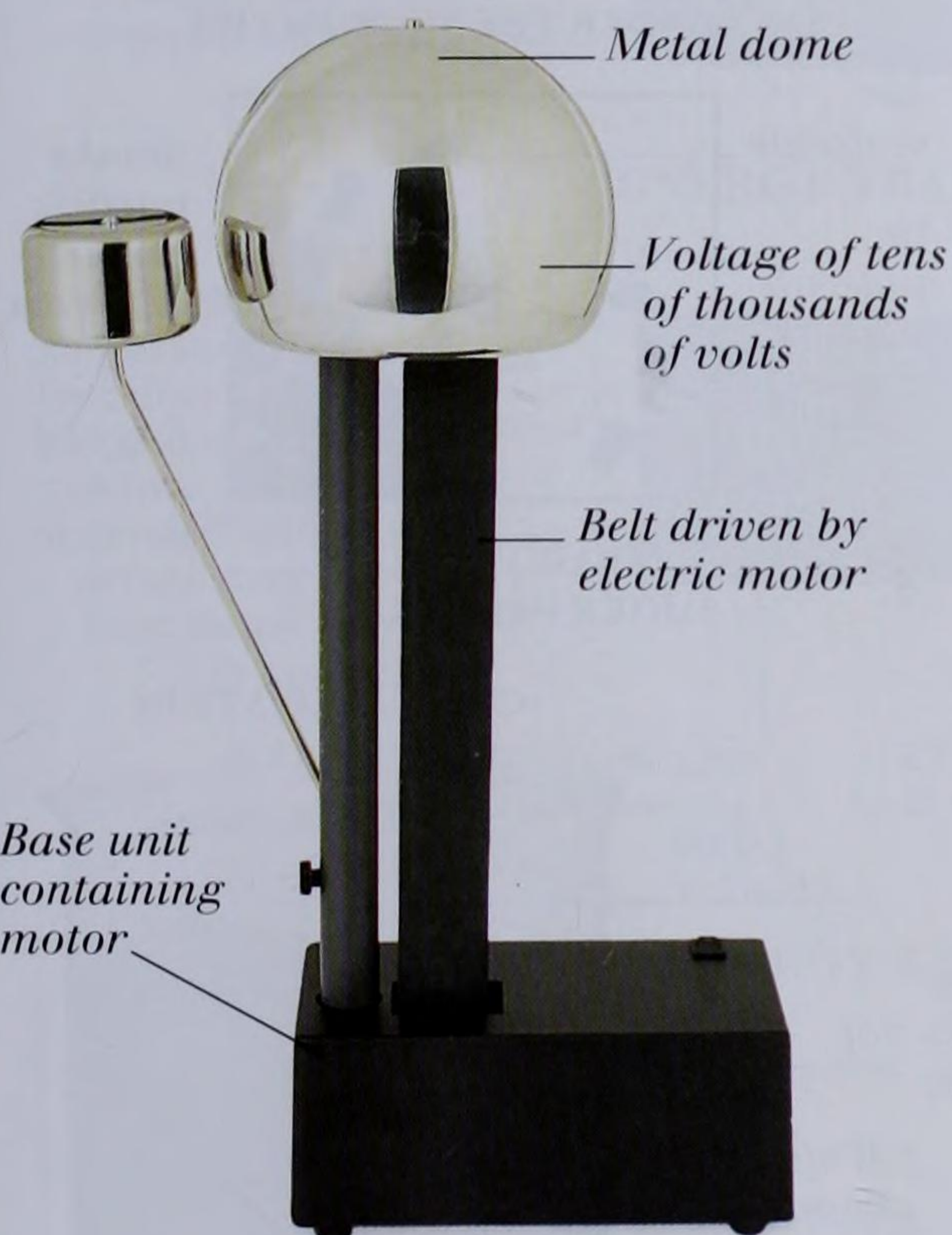
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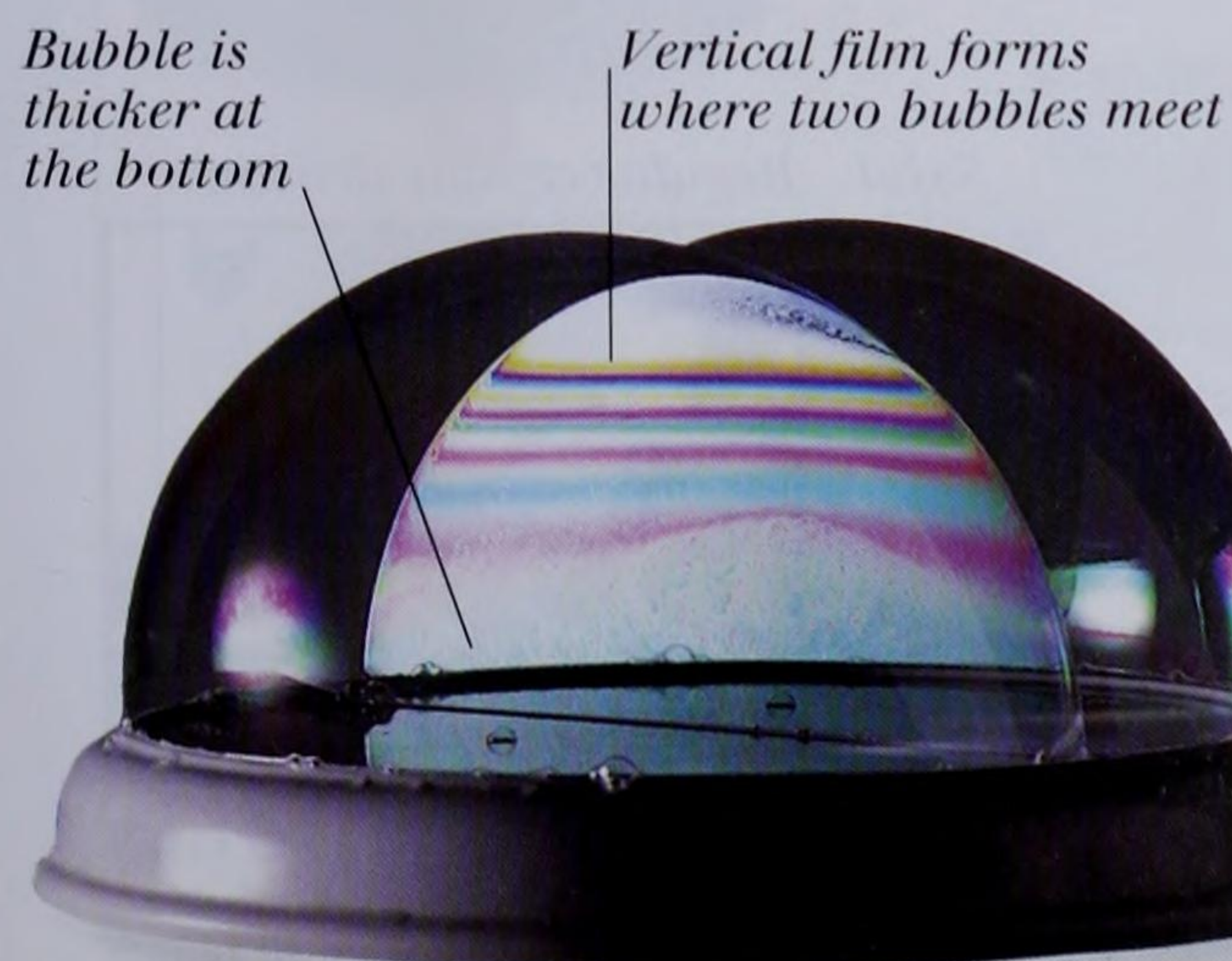




**SIMPLE PULLEY**



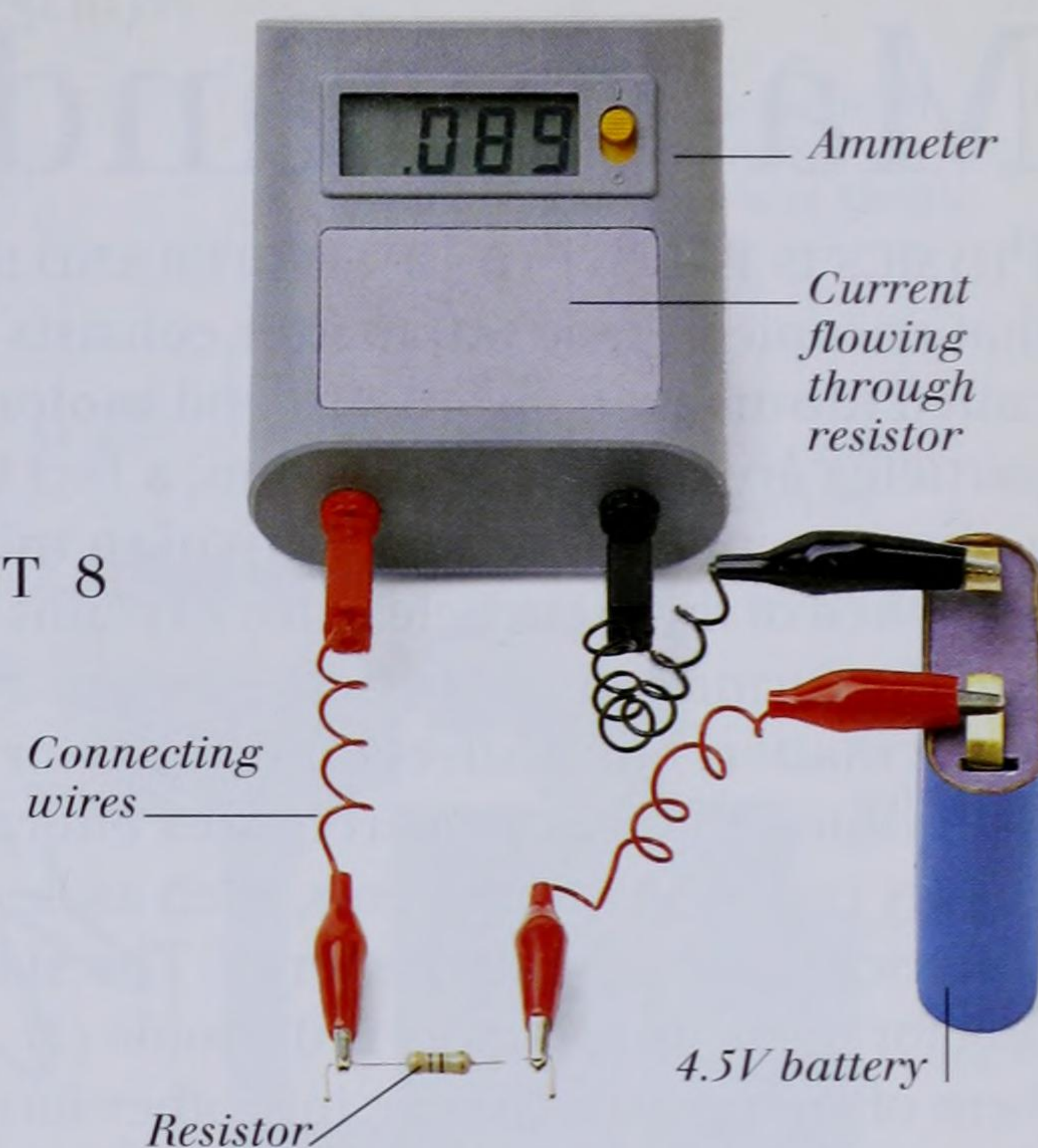
**VAN DE GRAAFF GENERATOR**



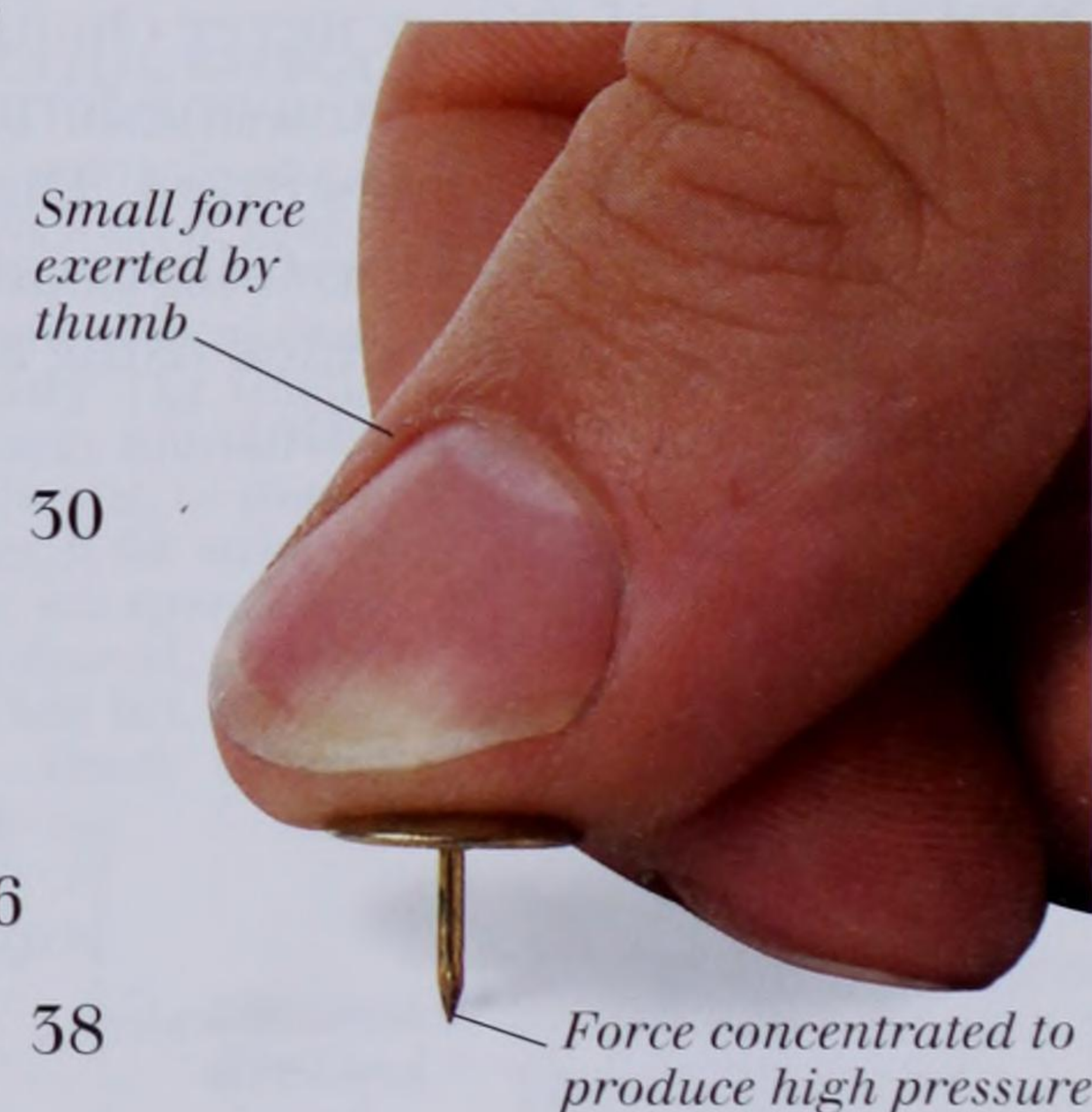
**SOAP BUBBLE**

# Contents

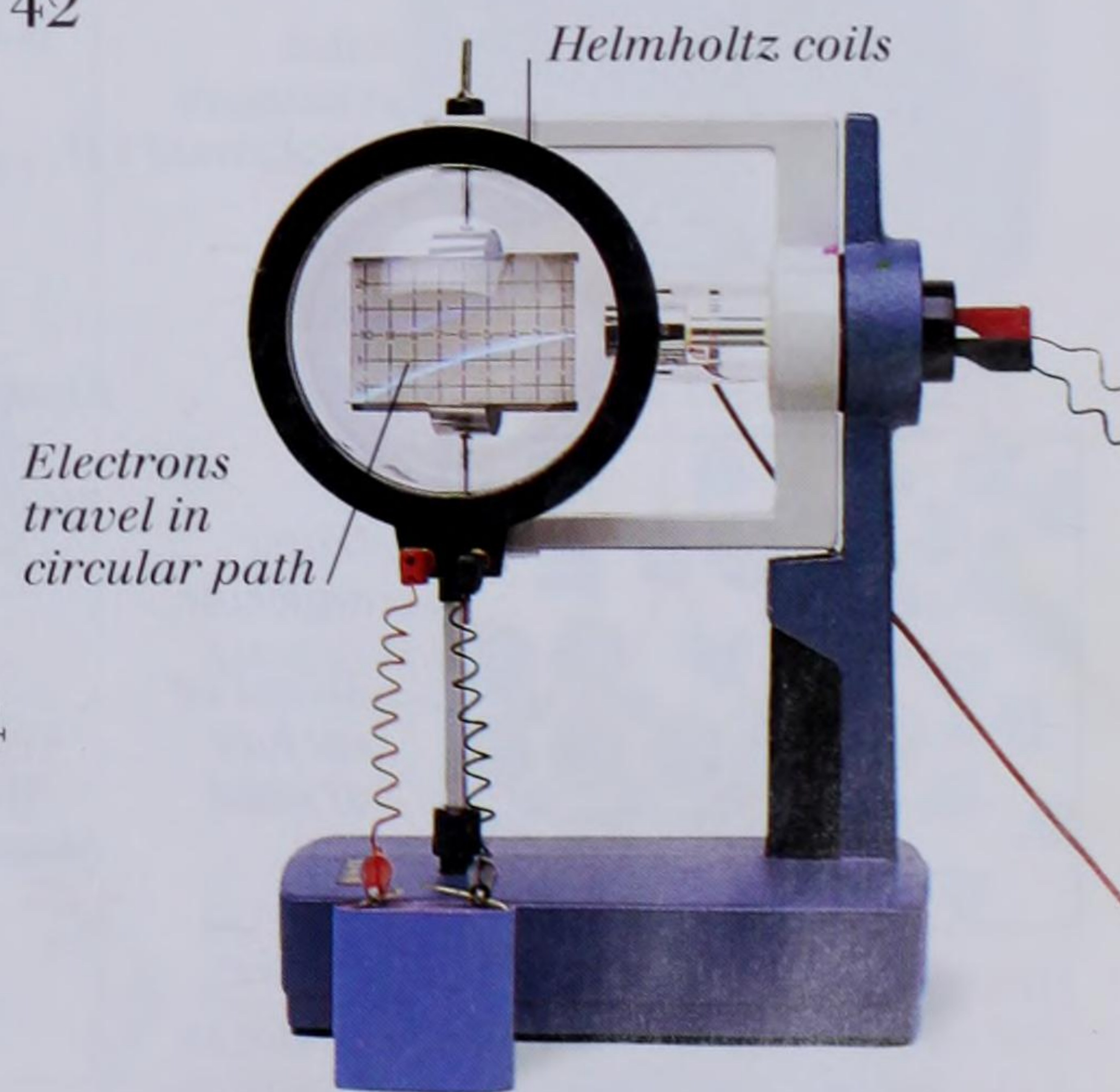
MATTER AND ENERGY	6
MEASUREMENT AND EXPERIMENT	8
FORCES 1	10
FORCES 2	12
FRICTION	14
SIMPLE MACHINES	16
CIRCULAR MOTION	18
WAVES AND OSCILLATIONS	20
HEAT AND TEMPERATURE	22
SOLIDS	24
LIQUIDS	26
GASES	28
ELECTRICITY AND MAGNETISM	30
ELECTRIC CIRCUITS	32
ELECTROMAGNETISM	34
ELECTRICITY PRODUCTION	36
ELECTROMAGNETIC RADIATION	38
COLOR	40
REFLECTION AND REFRACTION	42
OPTICAL INSTRUMENTS	44
WAVE BEHAVIOR	46
ATOMS AND ELECTRONS	48
NUCLEAR PHYSICS	50
PARTICLE PHYSICS	52
FORMULAS AND APPENDIX	54
GLOSSARY	56
INDEX	60
ACKNOWLEDGMENTS	64



**RESISTANCE**



**DRAWING PIN**



**CATHODE RAY TUBE**

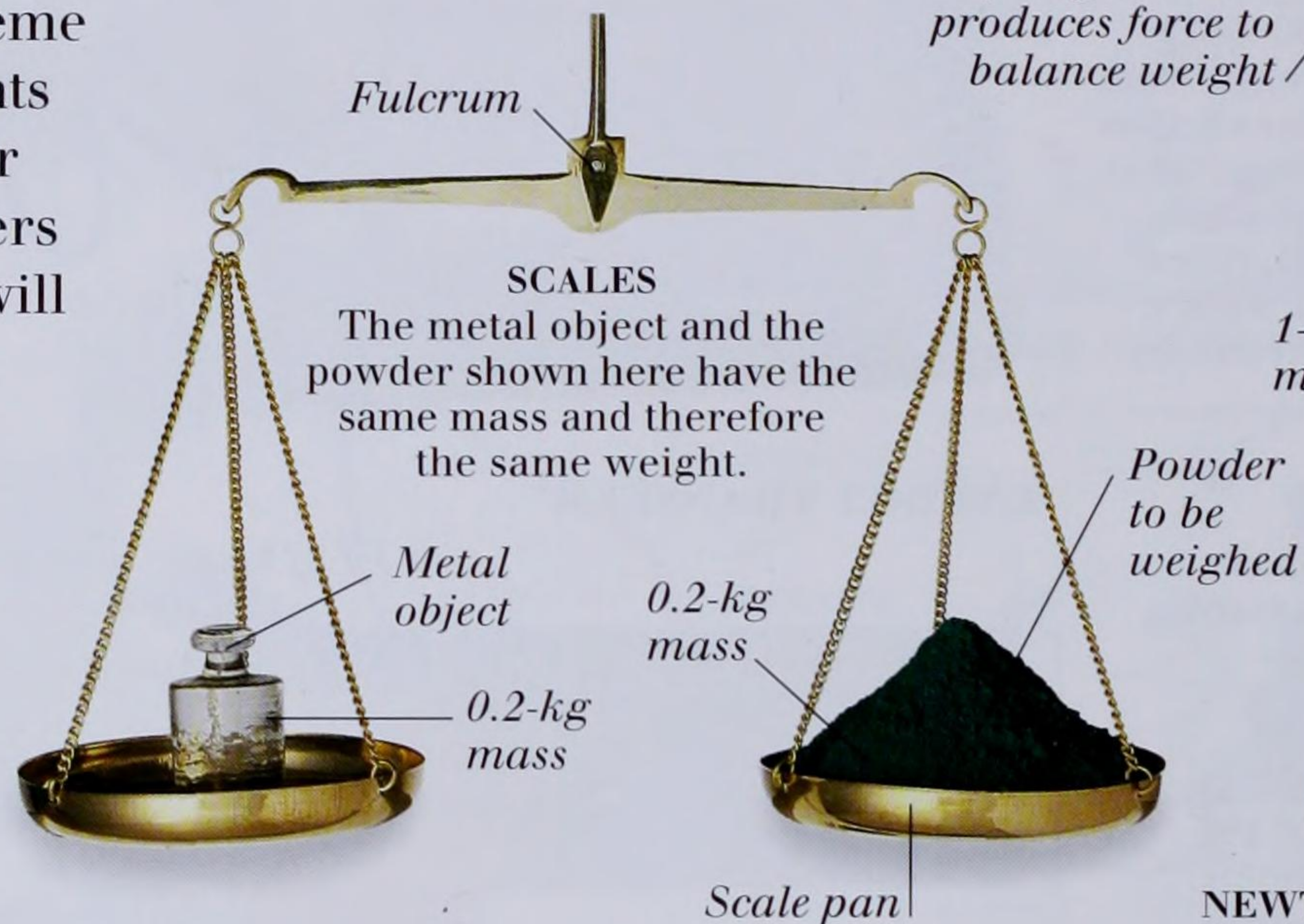
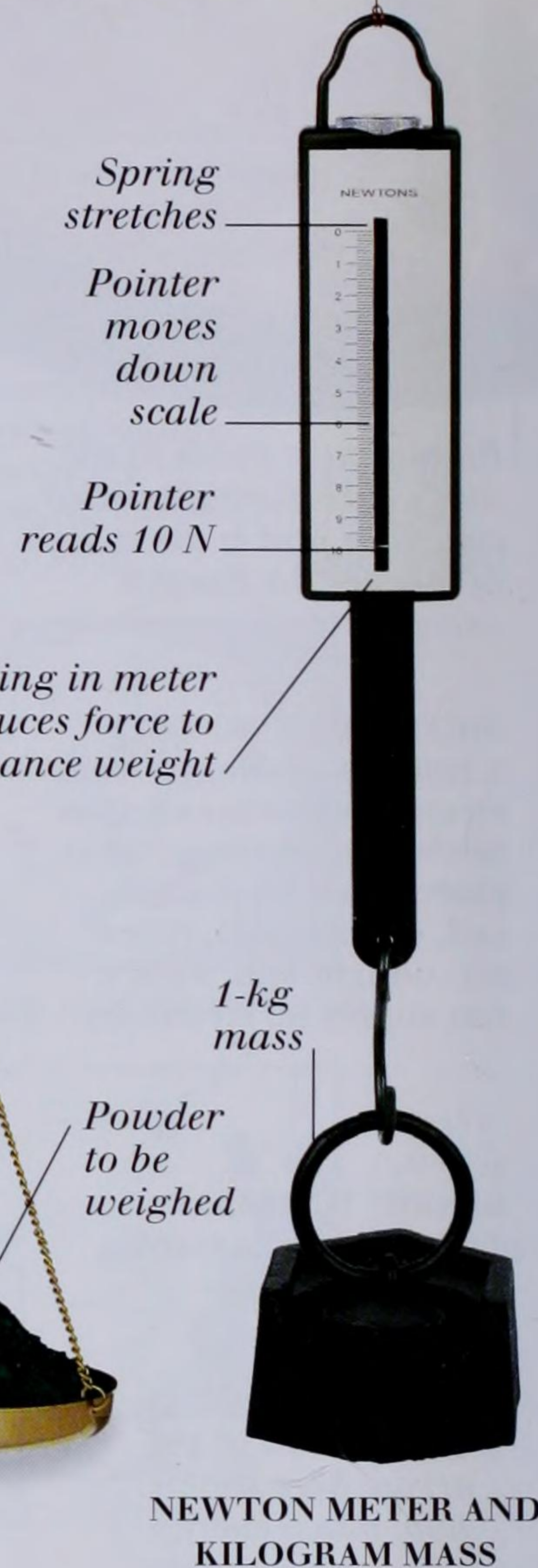


# Measurement and experiment

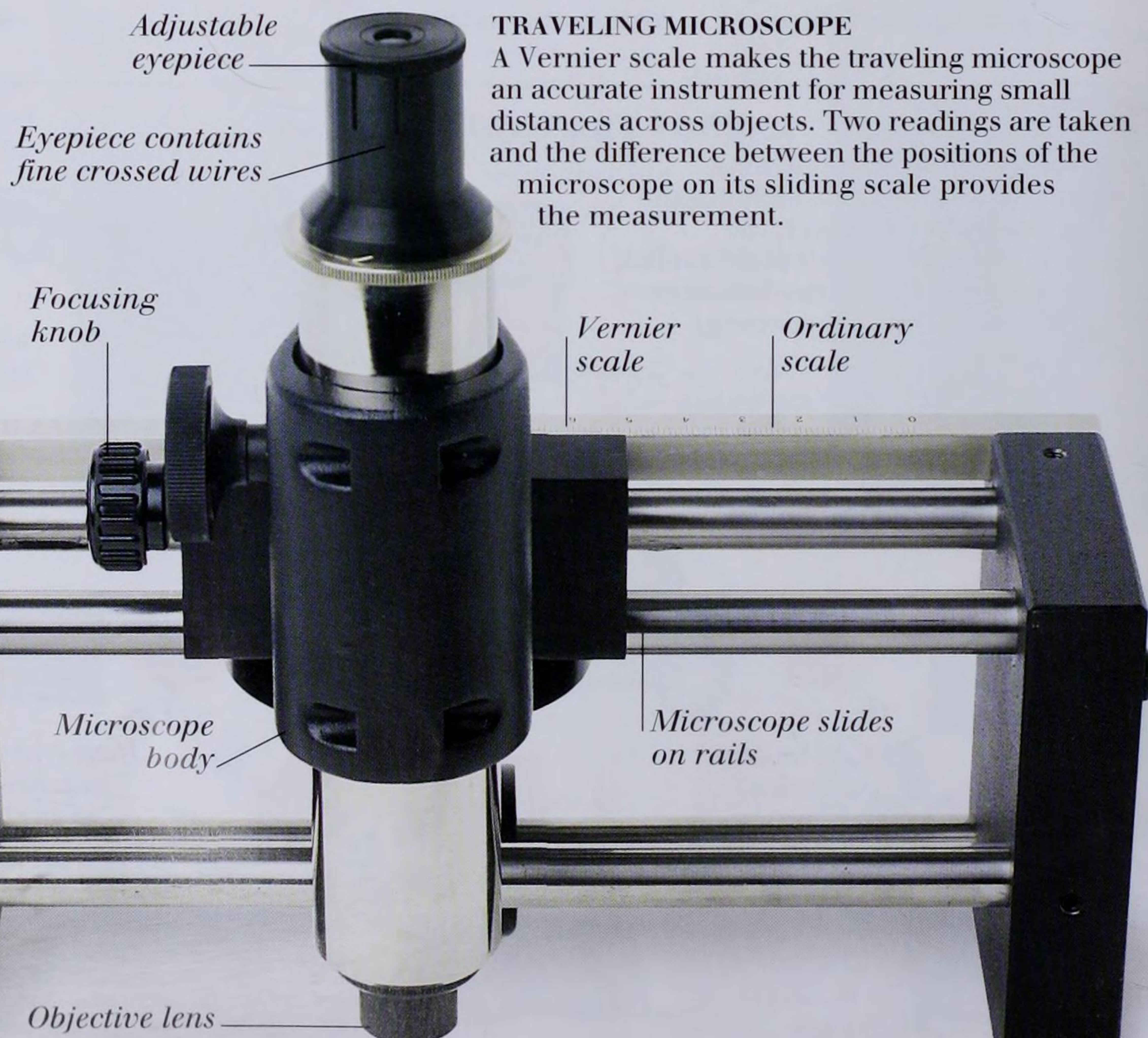
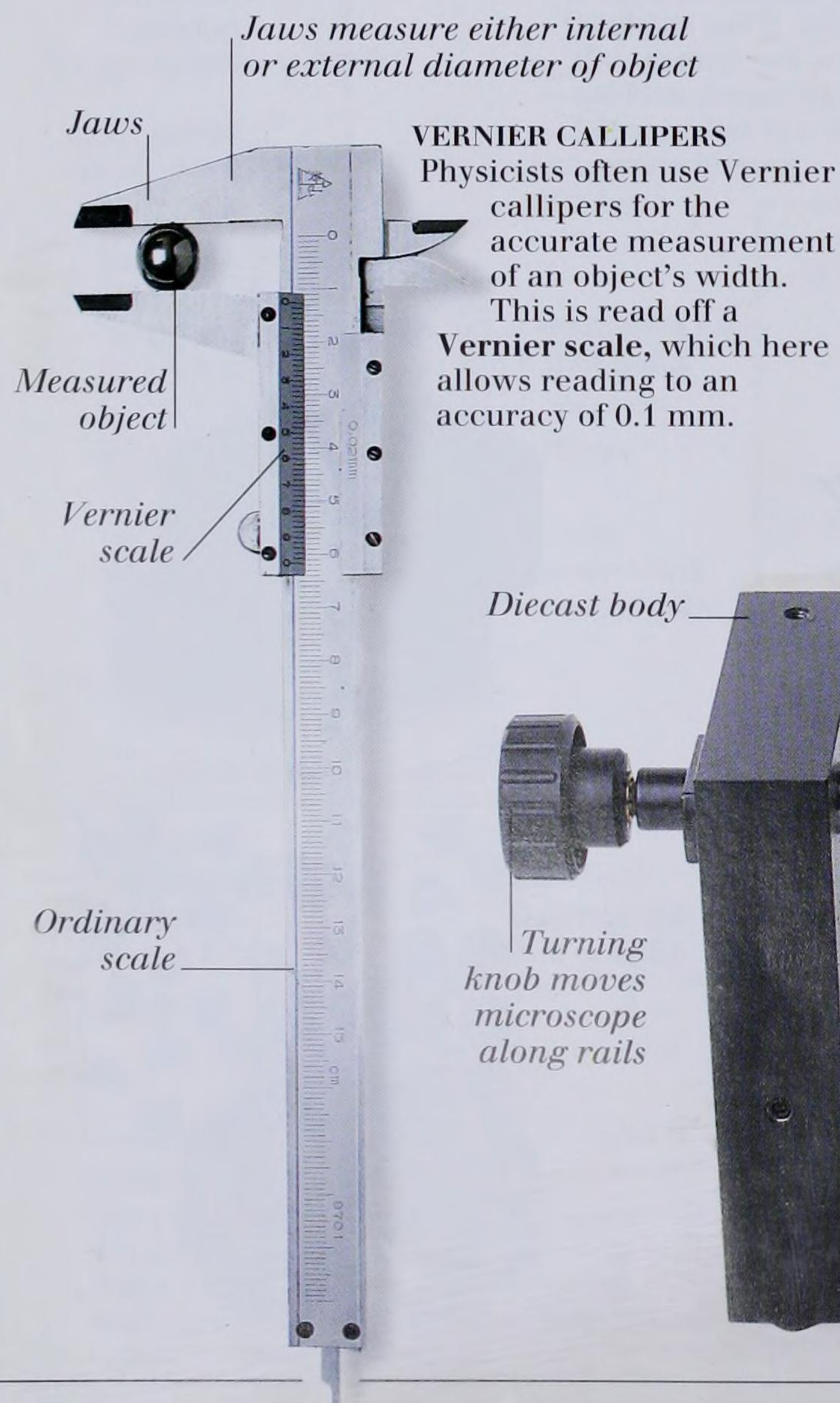
THE SCIENCE OF PHYSICS IS BASED on the formulation and testing of theories. Experiments are designed to test theories and involve making measurements: of **mass**, length, time, or other quantities. In order to compare the results of various experiments, standard units are necessary. The kilogram (kg), the meter (m), and the second (s) are the fundamental units of a system called **SI units** (Système International). Physicists use a variety of instruments for making measurements. Some, like the Vernier callipers, traveling microscopes, and thermometers are common to many laboratories, while others will be made for a particular experiment. The results of measurements are interpreted in many ways, but most often as graphs. Graphs provide a way of illustrating the relationship between two measurements involved in an experiment. For example, in an experiment to investigate falling objects, a graph can show the relationship between the duration and the height of the fall.

## MASS AND WEIGHT

Mass is the amount of matter in an object, and is measured in kilograms. **Gravitational force** gives the mass its **weight**. Weight is a force, and is measured in newtons (see pp.10–11), using a newton meter like the one shown on the right. It is common to speak of weight being measured in kilograms, but in physics this is not correct.



## MEASURING DISTANCE

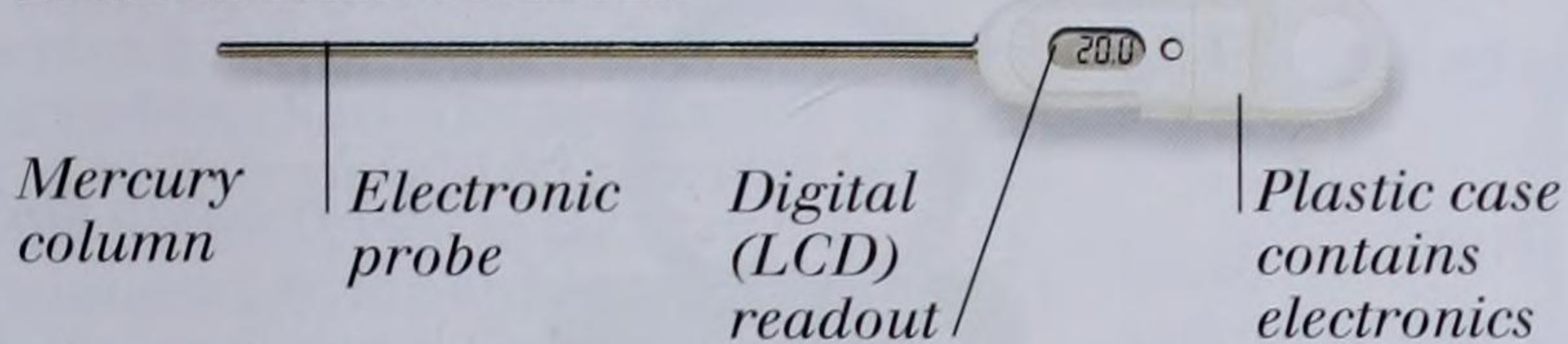




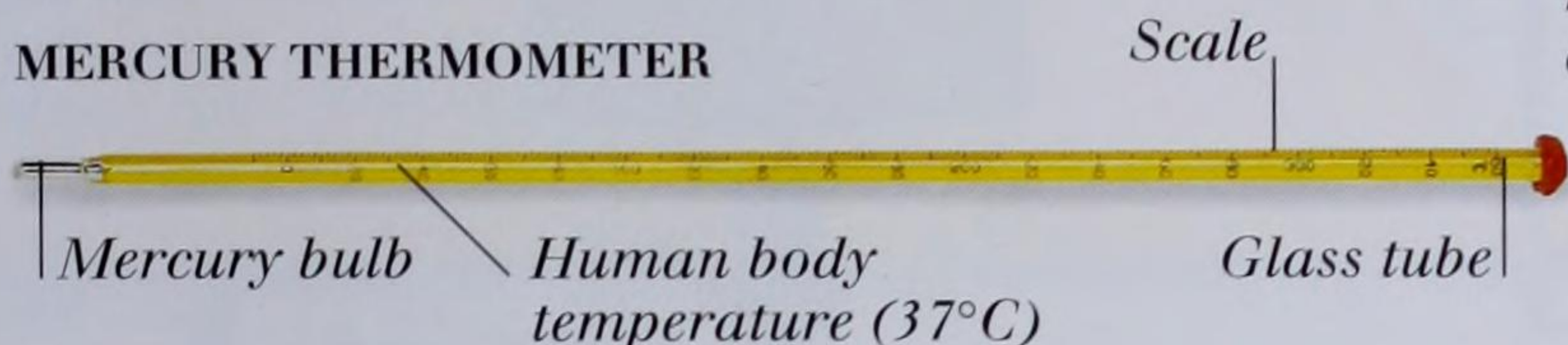
## THERMOMETERS

There are two types of thermometer commonly used in modern physics. The mercury thermometer has a glass bulb containing mercury that expands as the temperature rises, while the digital thermometer contains an electronic probe and has a digital readout.

### DIGITAL THERMOMETER



### MERCURY THERMOMETER



### MAGNIFIED VIEW OF MERCURY THERMOMETER

Glass bulb

## INTERPRETING DATA

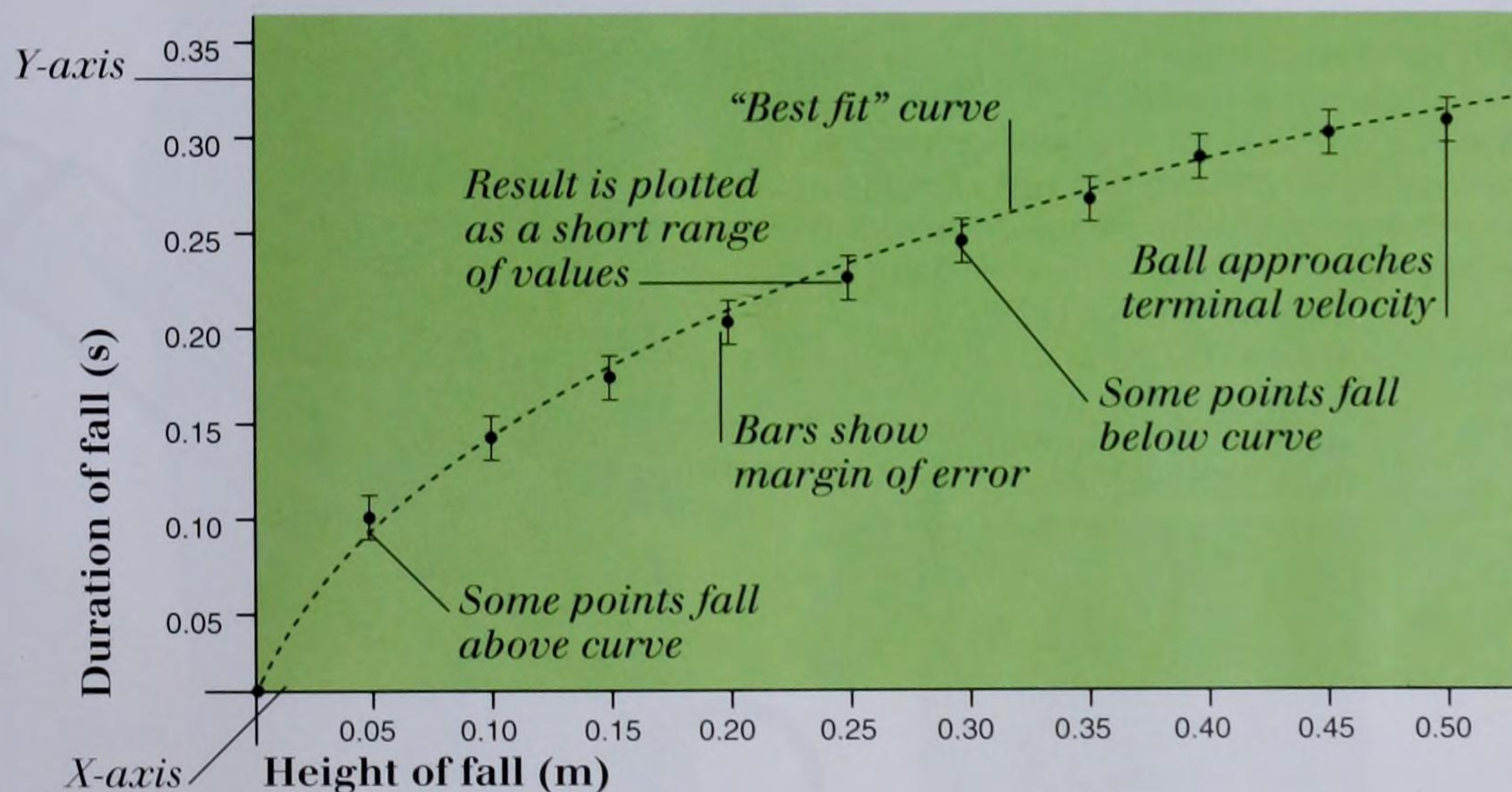
### TABLE OF RESULTS FOR A FREEFALL EXPERIMENT

A steel ball is dropped from a variety of heights and the duration of each fall is timed. The results of these measurements are entered into a table.

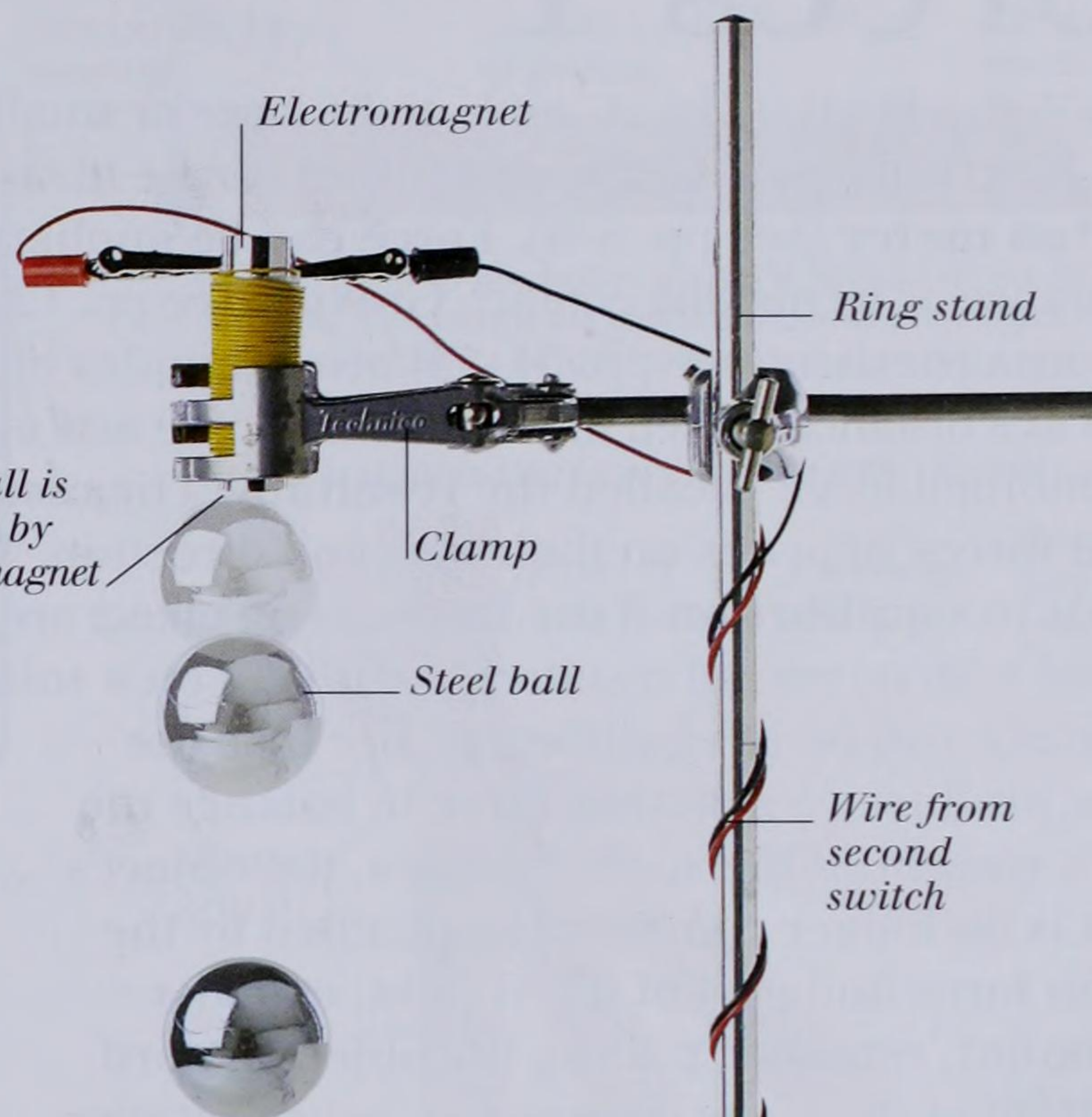
HEIGHT (m)	0	0.05	0.10	0.15	0.20	0.25	0.30	0.35	0.40	0.45	0.50
TIME (s)	0	0.10	0.14	0.17	0.21	0.22	0.24	0.26	0.27	0.30	0.31

### RESULTS OF A FREEFALL EXPERIMENT IN GRAPH FORM

A **graph** allows us to identify visually the relationship between the time and the height of the fall. There is an element of uncertainty or error in every result obtained, so each is plotted on the graph as a short range of values forming an **error bar** instead of a point. The curve is drawn so that it passes through all the bars.

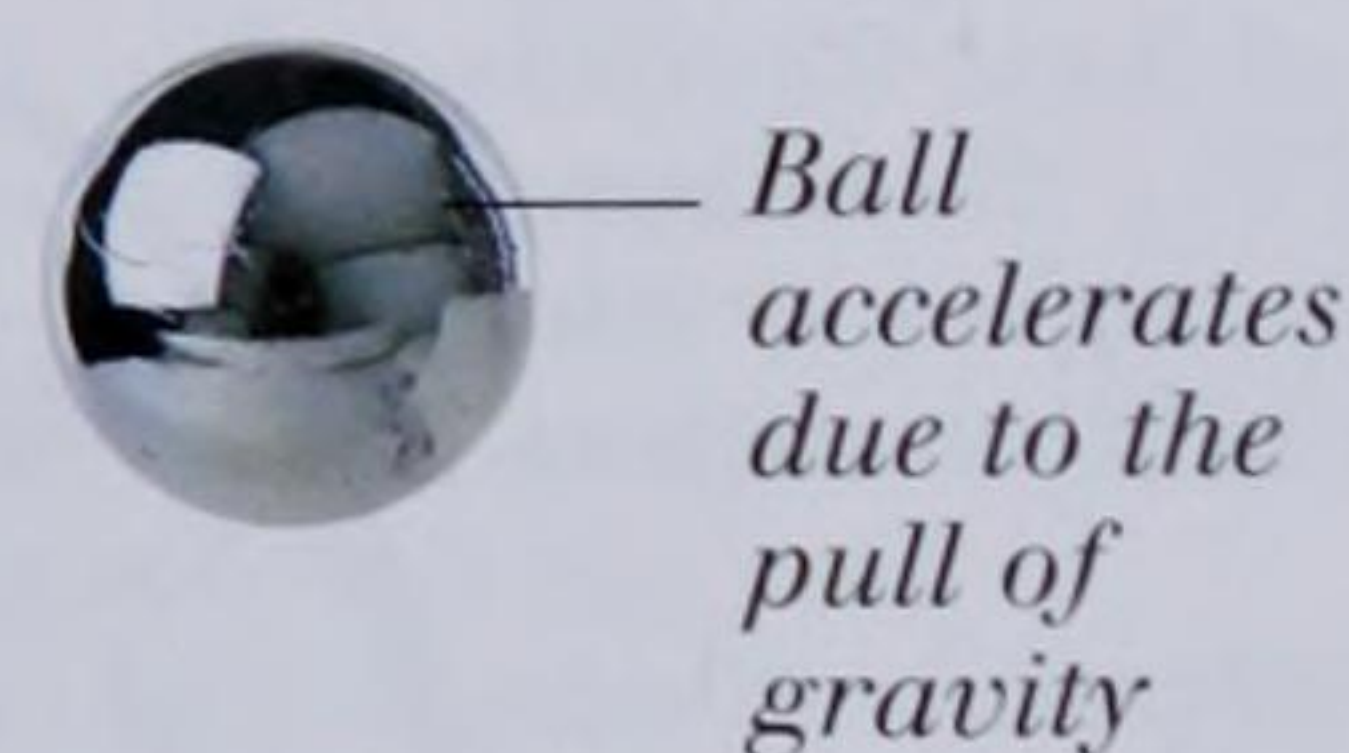


## FREEFALL EXPERIMENT



### APPARATUS FOR TIMING THE FALL OF AN OBJECT

A switch turns off the electromagnet, releasing the ball while simultaneously starting the timer. As the ball hits the ring stand base, a second switch is activated, and the timer stops. Times of falls from various heights are measured, and plotted on a graph (see left).



As ball hits base, second switch is activated

Second switch (two contacts normally held apart)

Ring stand base

Wire to battery

Electronic timer

Electronic timer stand

Digital display

Wire from second switch

Wire from first switch

Switch

Alligator clip

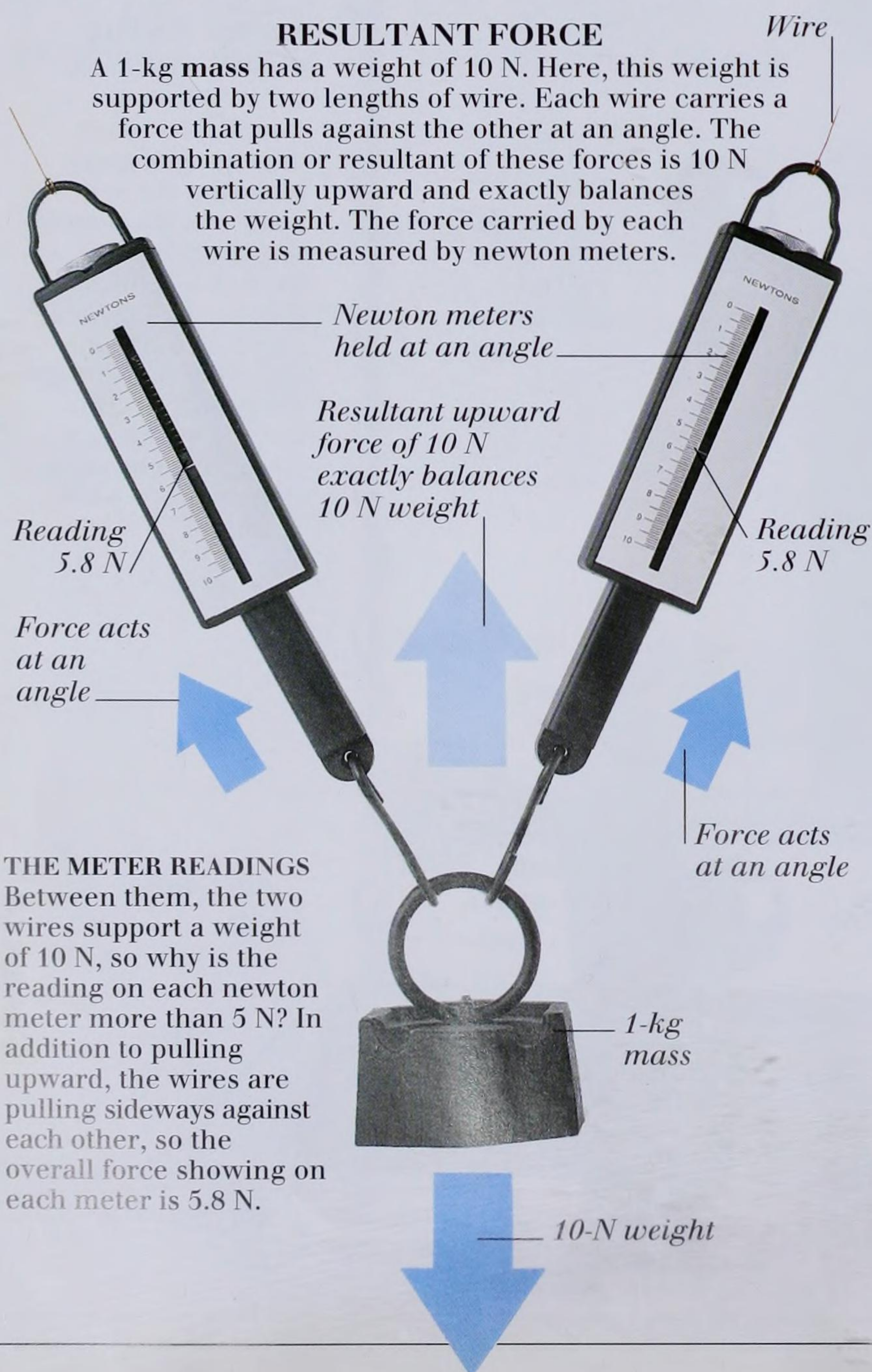


# Forces 1

A **FORCE** IS A PUSH OR PULL, and can be large or small. The usual unit of force is the newton (N), and forces can be measured using a **newton meter** (see pp. 8–9). Force can be applied to objects at a distance or by making contact. **Gravity** (see pp. 12–13) and **electromagnetism** (see pp. 34–35) are examples of forces that can act at a distance. When more than one force acts on an object, the combined force is called the **resultant**. The resultant of several forces depends on their size and direction. The object is in **equilibrium** if the forces on an object are balanced with no overall resultant. An object on a solid flat surface will be in equilibrium, because the surface produces a **reaction** force to balance the object's **weight**. If the surface slopes, the object's weight is no longer completely cancelled by the reaction force and part of the weight, called a **component**, remains, pulling the object toward the bottom of the slope. Forces can cause rotation as well as straight line motion. If an object is free to rotate about a certain point, then a force can have a turning effect, known as a **moment**.

## RESULTANT FORCE

A 1-kg mass has a weight of 10 N. Here, this weight is supported by two lengths of wire. Each wire carries a force that pulls against the other at an angle. The combination or resultant of these forces is 10 N vertically upward and exactly balances the weight. The force carried by each wire is measured by newton meters.

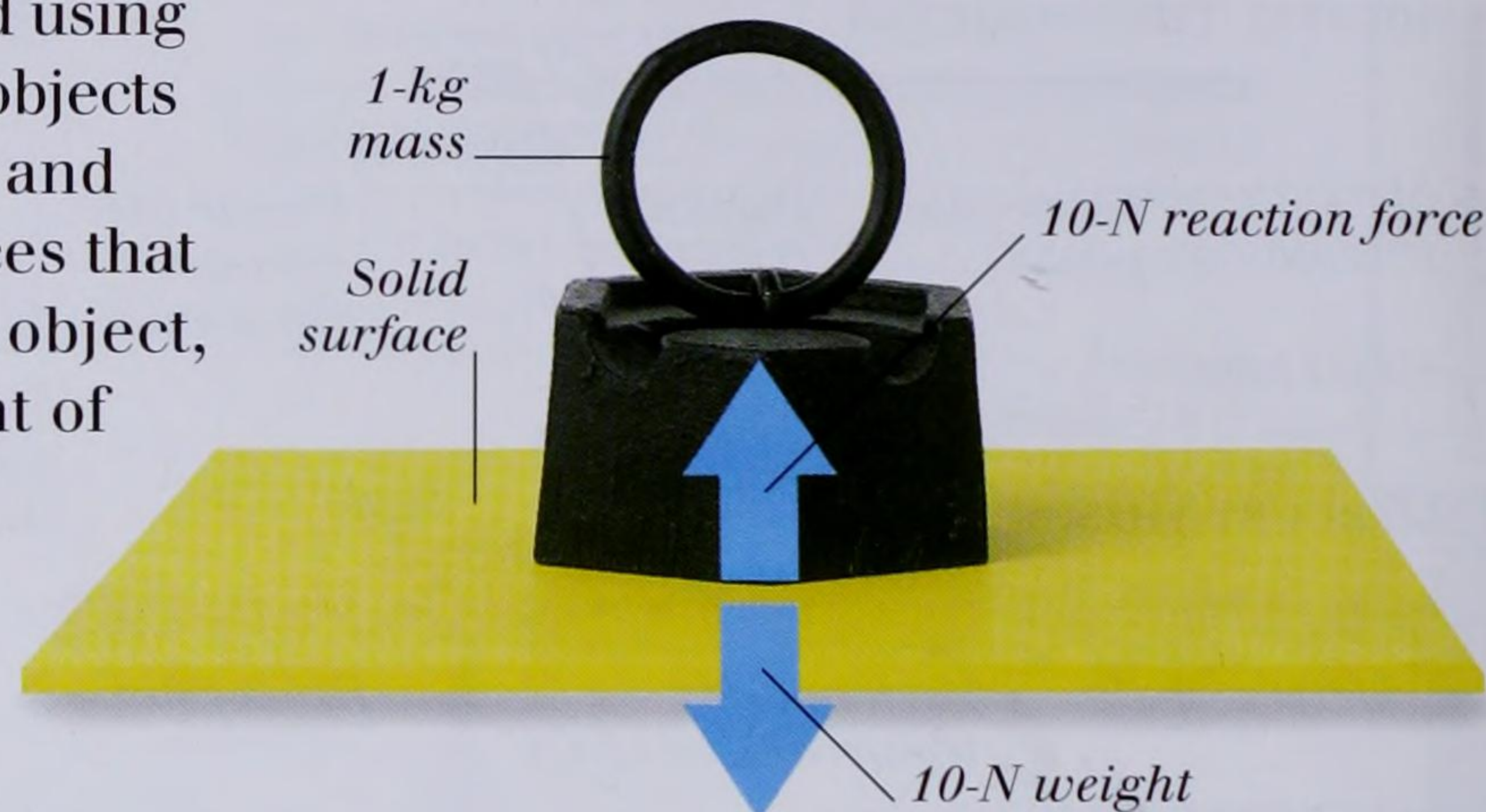


**THE METER READINGS**  
Between them, the two wires support a weight of 10 N, so why is the reading on each newton meter more than 5 N? In addition to pulling upward, the wires are pulling sideways against each other, so the overall force showing on each meter is 5.8 N.

## REACTION FORCES

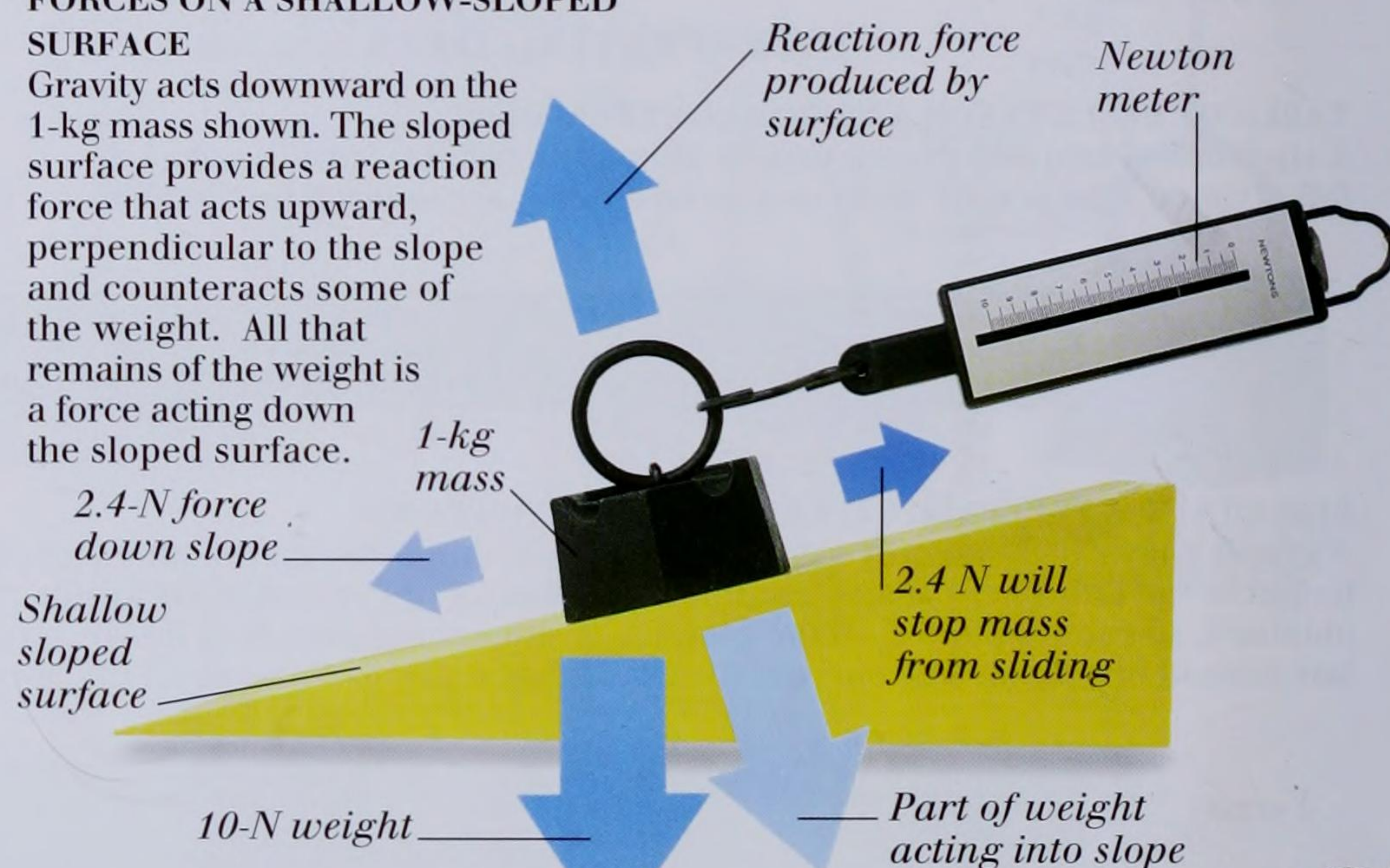
### FORCES ON A LEVEL SURFACE

A table provides a force called a reaction, which exactly balances the weight of an object placed upon it. The resultant force is zero, so the object does not fall through the table.



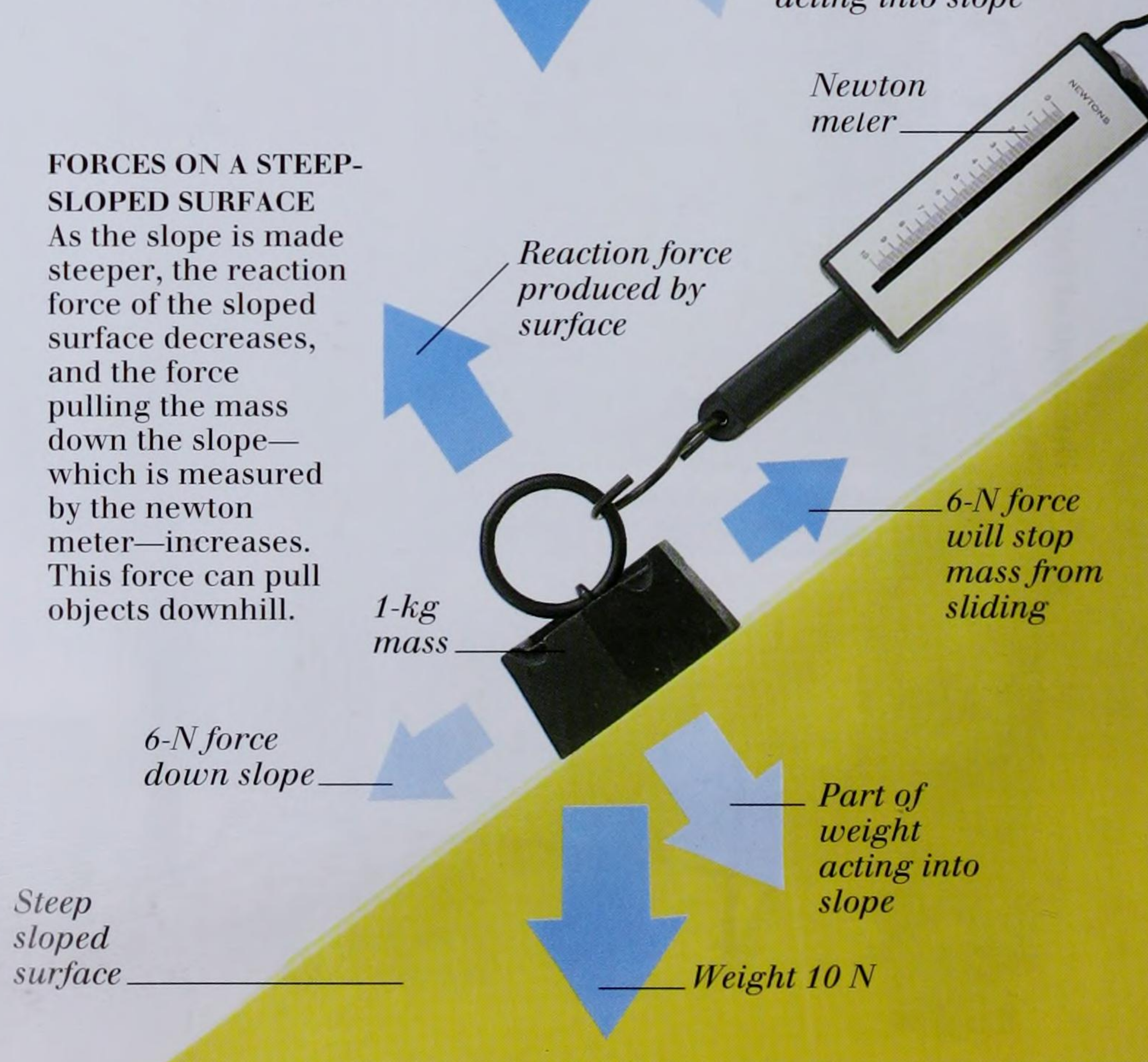
### FORCES ON A SHALLOW-SLOPED SURFACE

Gravity acts downward on the 1-kg mass shown. The sloped surface provides a reaction force that acts upward, perpendicular to the slope and counteracts some of the weight. All that remains of the weight is a force acting down the sloped surface.



### FORCES ON A STEEP-SLOPED SURFACE

As the slope is made steeper, the reaction force of the sloped surface decreases, and the force pulling the mass down the slope—which is measured by the newton meter—increases. This force can pull objects downhill.

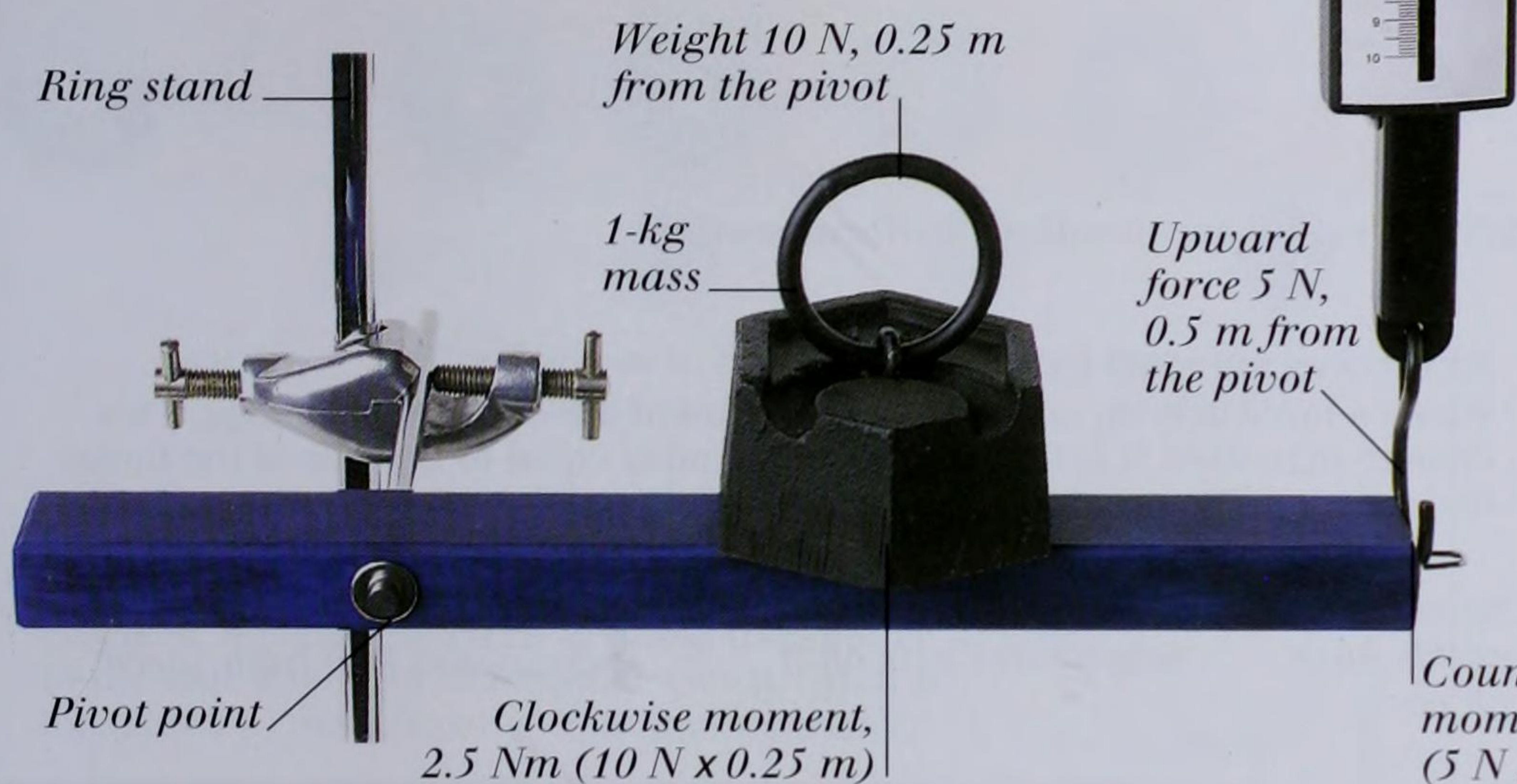




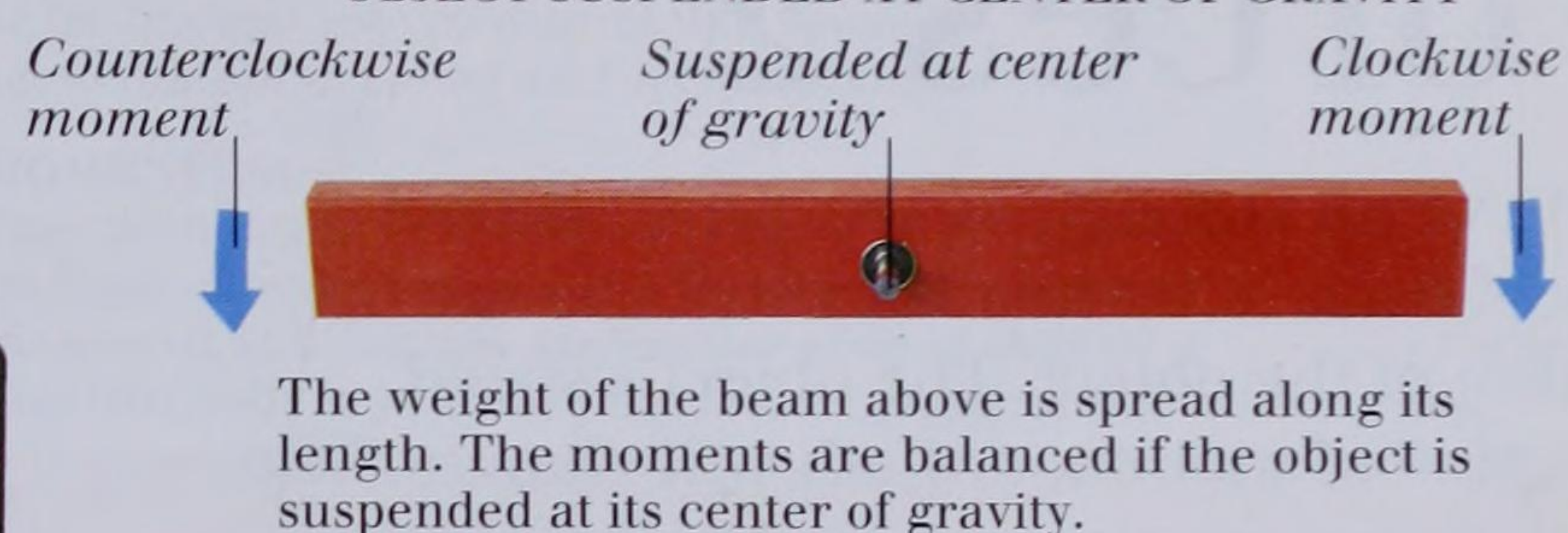
## TURNING FORCES

### TURNING FORCES AROUND A PIVOT

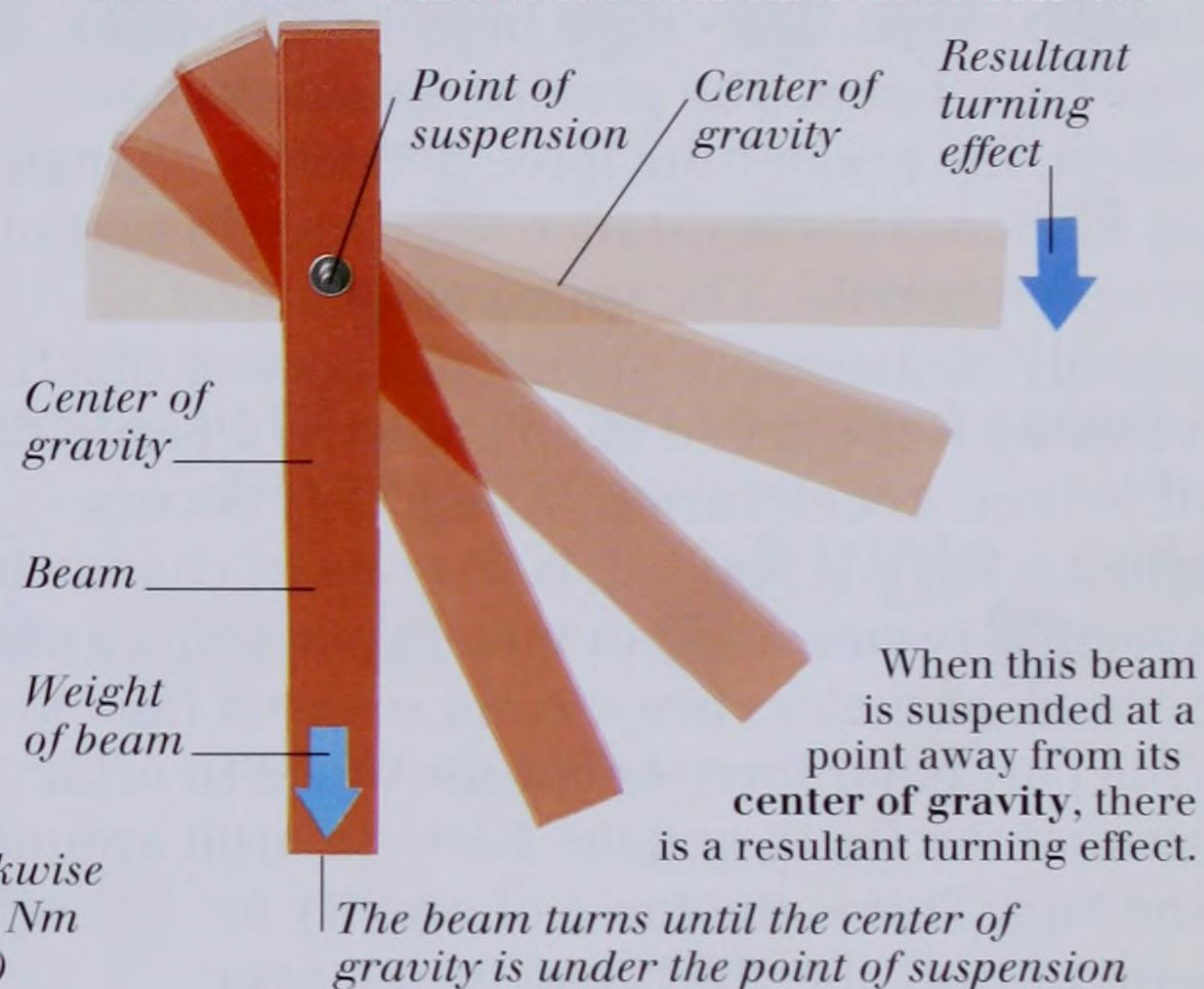
A force acting on an object that is free to rotate will have a turning effect, or turning force, also known as a moment. The moment of a force is equal to the size of the force multiplied by the distance of the force from the turning point around which it acts (see p. 54). It is measured in newton meters (Nm). The mass below exerts a weight of 10 N downward on a pivoted beam. The newton meter—twice as far from the pivot—measures 5 N, the upward force needed to stop the beam turning. The clockwise moment created by the weight and counterclockwise moment created by the upward pull on the newton meter are equal, and the object is therefore in equilibrium.



### OBJECT SUSPENDED AT CENTER OF GRAVITY



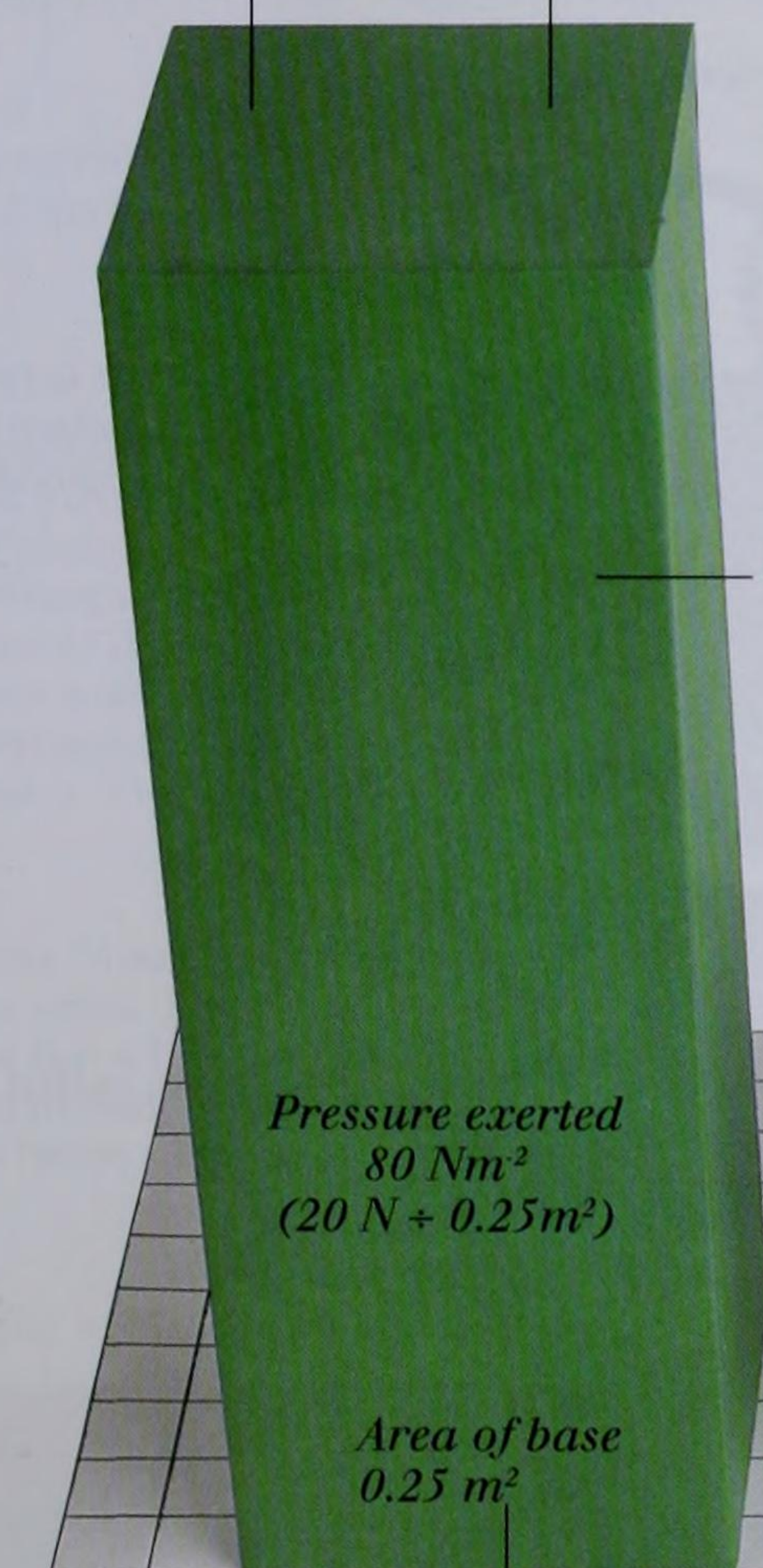
### OBJECT SUSPENDED AWAY FROM CENTER OF GRAVITY



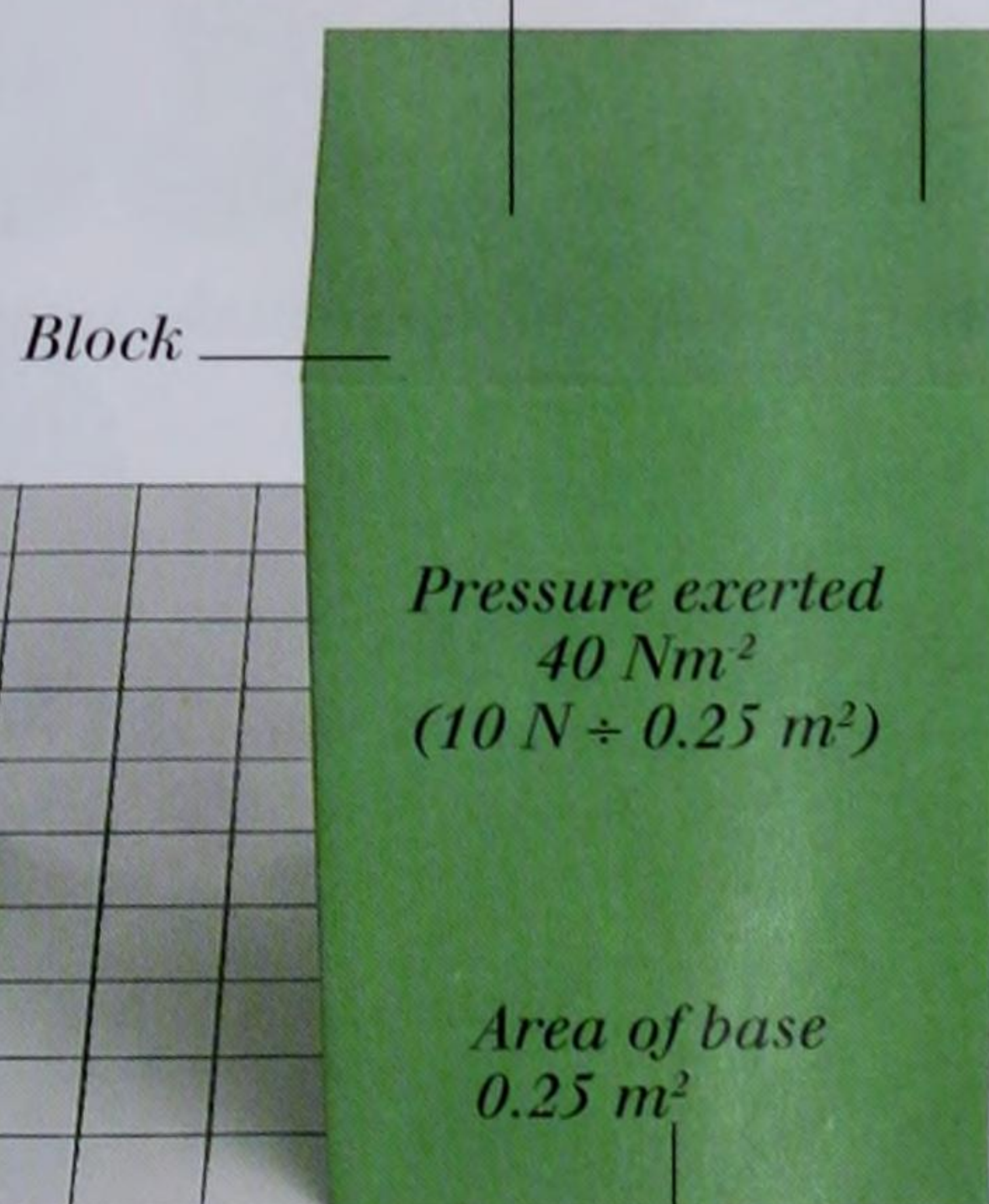
## PRESSURE

Why can a drawing pin be pushed into a wall, and yet a building will not sink into the ground? Forces can act over large or small areas. A force acting over a large area will exert less **pressure** than the same force acting over a small area. The pressure exerted on an area can be worked out simply by dividing the applied force by the area over which it acts (see p. 54). Pressure is normally measured in units of newtons per square meter ( $\text{Nm}^{-2}$ ). A drawing pin concentrates force to produce high pressure, whereas the foundations of a building spread the load to reduce pressure. Gases also exert pressure (see pp. 28–29).

Mass of block 2 kg  
Weight of block 20 N

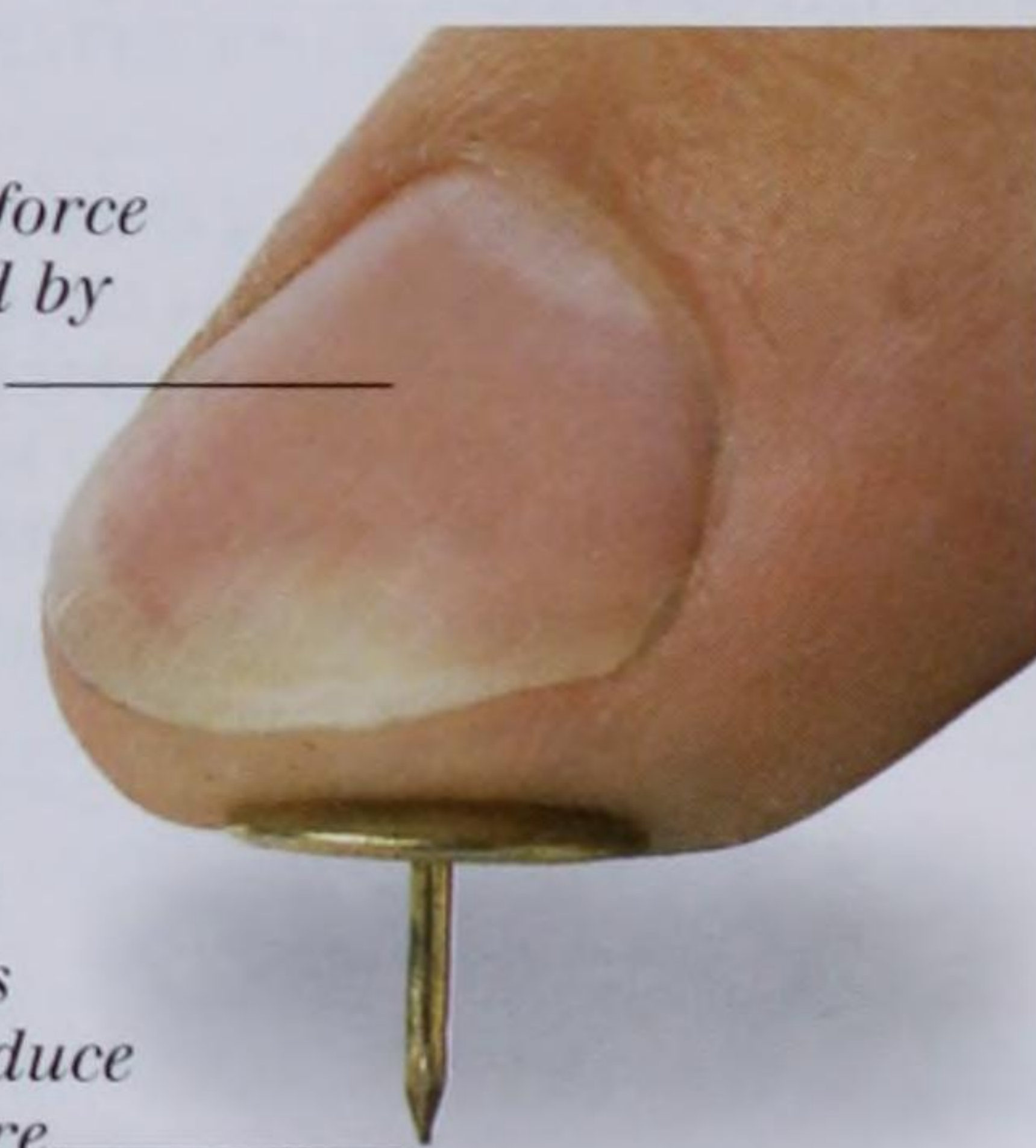


Mass of block 1 kg  
Weight of block 10 N

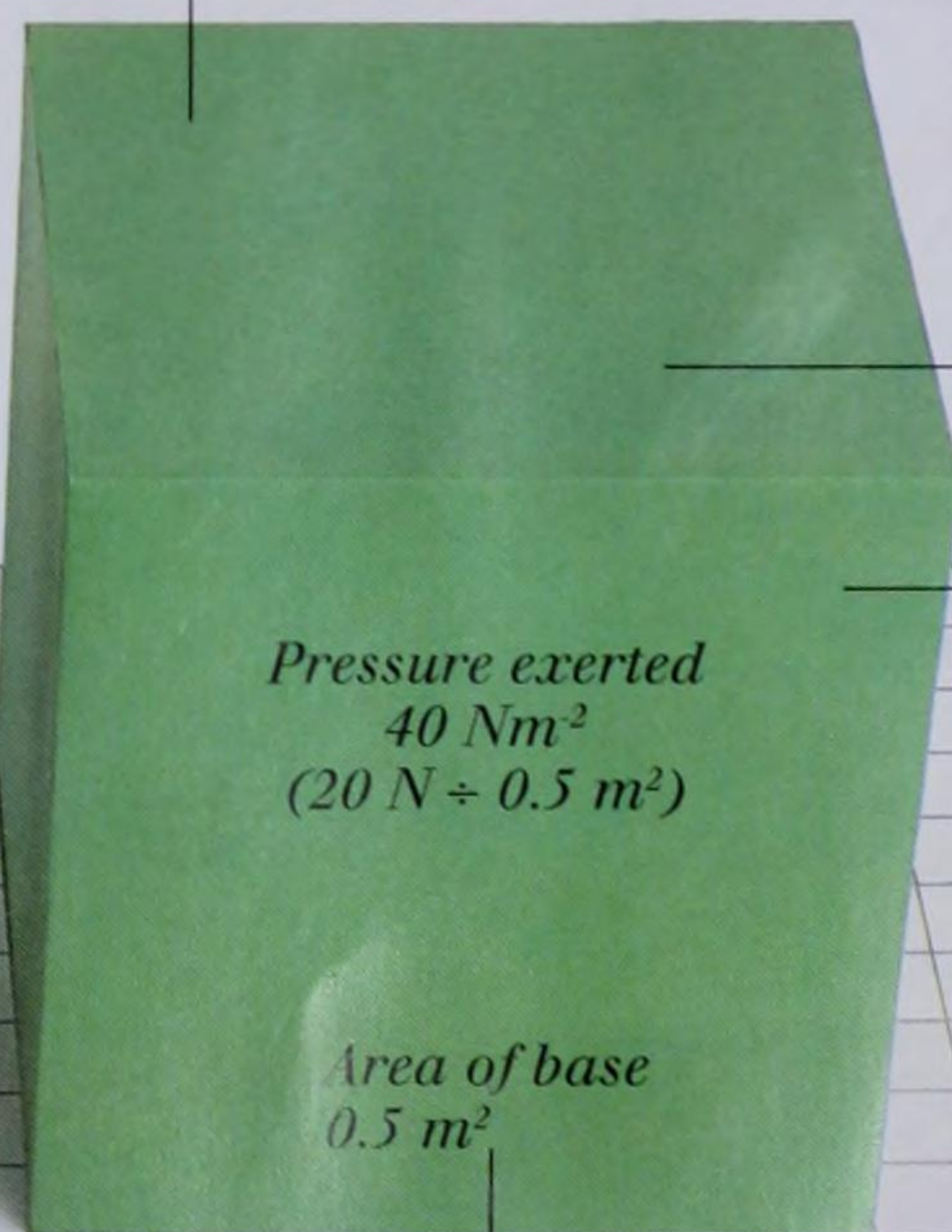


Small force exerted by thumb

Tiny area at pin point concentrates force to produce high pressure



Block



Grid with squares of area 0.01  $\text{m}^2$

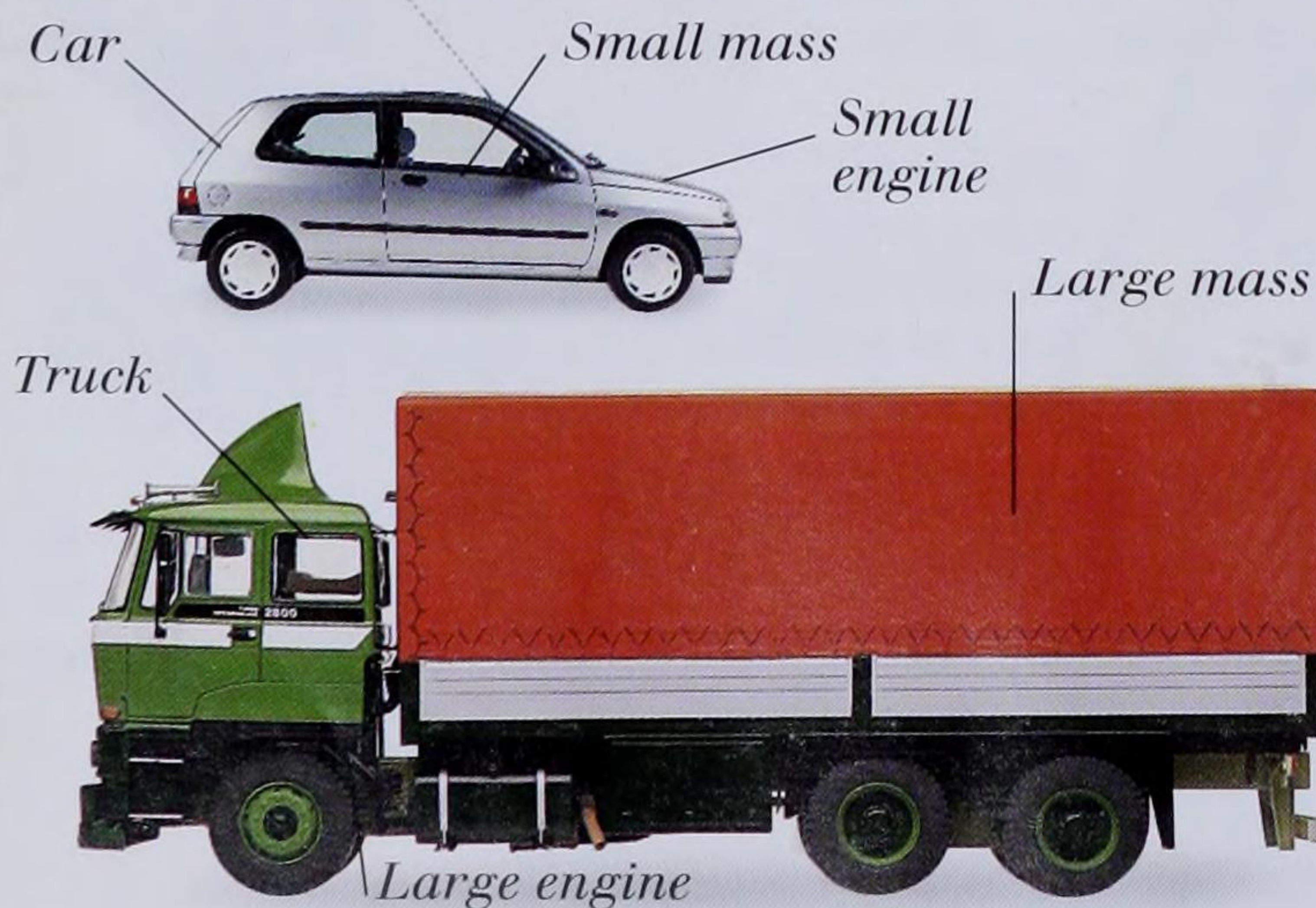


# Forces 2

WHEN THE FORCES ON AN OBJECT do not cancel each other out, they will change the motion of the object. The object's speed, direction of motion, or both will change. The rules governing the way forces change the motion of objects were first worked out by Sir Isaac Newton. They have become known as Newton's Laws. The greater the mass of an object, the greater the force needed to change its motion. This resistance to change in motion is called **inertia**. The **speed** of an object is usually measured in meters per second ( $\text{ms}^{-1}$ ). **Velocity** is the speed of an object in a particular direction. **Acceleration**, which only occurs when a force is applied, is the rate of change in speed. It is measured in meters per second per second, or meters per second squared ( $\text{ms}^{-2}$ ). One particular force keeps the Moon in orbit around the Earth and the Earth in orbit around the Sun. This is the force of **gravity** or **gravitation**; its effects can be felt over great distances.

## NEWTON'S SECOND LAW IN ACTION

Trucks have a greater mass than cars. According to Newton's second law (see right) a large mass requires a larger force to produce a given acceleration. This is why a truck needs to have a larger engine than a car.



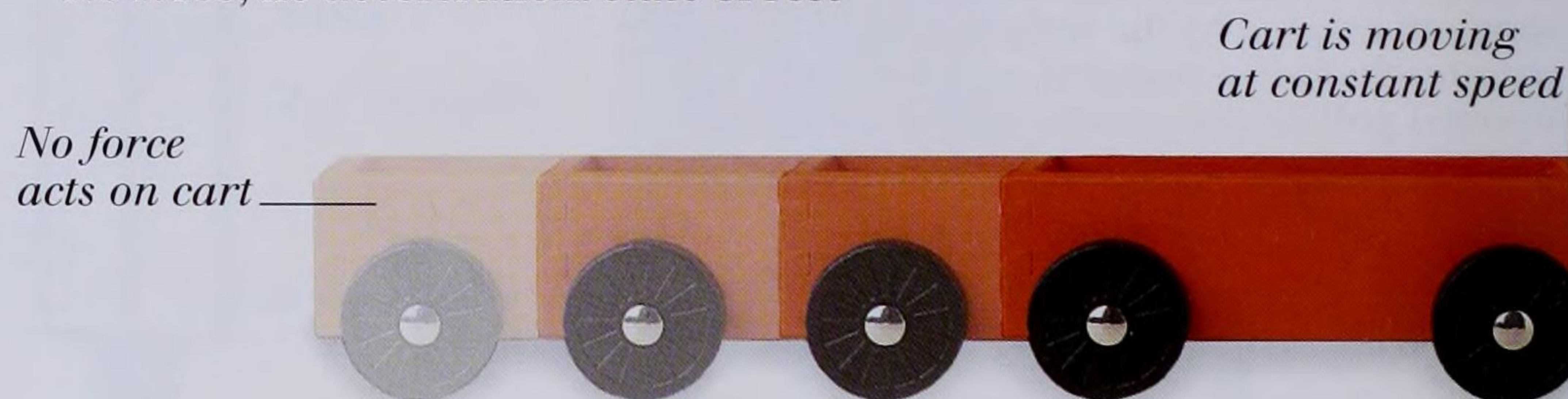
## NEWTON'S LAWS

### NEWTON'S FIRST LAW

When no force acts on an object, it will remain in a state of rest or continue its uniform motion in a straight line.



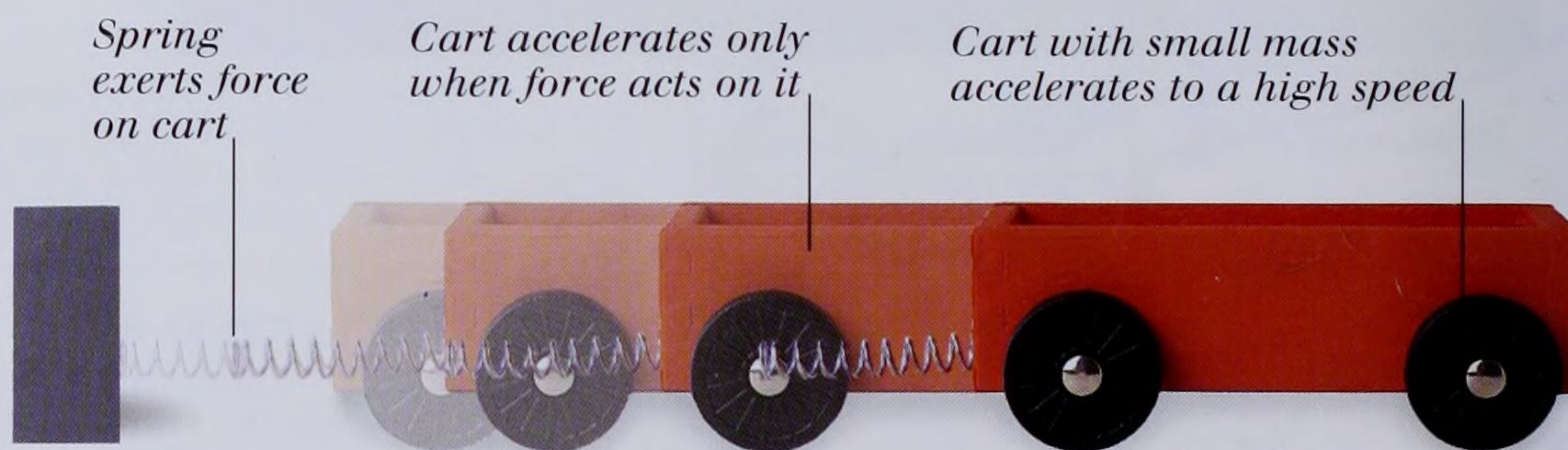
No force, no acceleration: state of rest



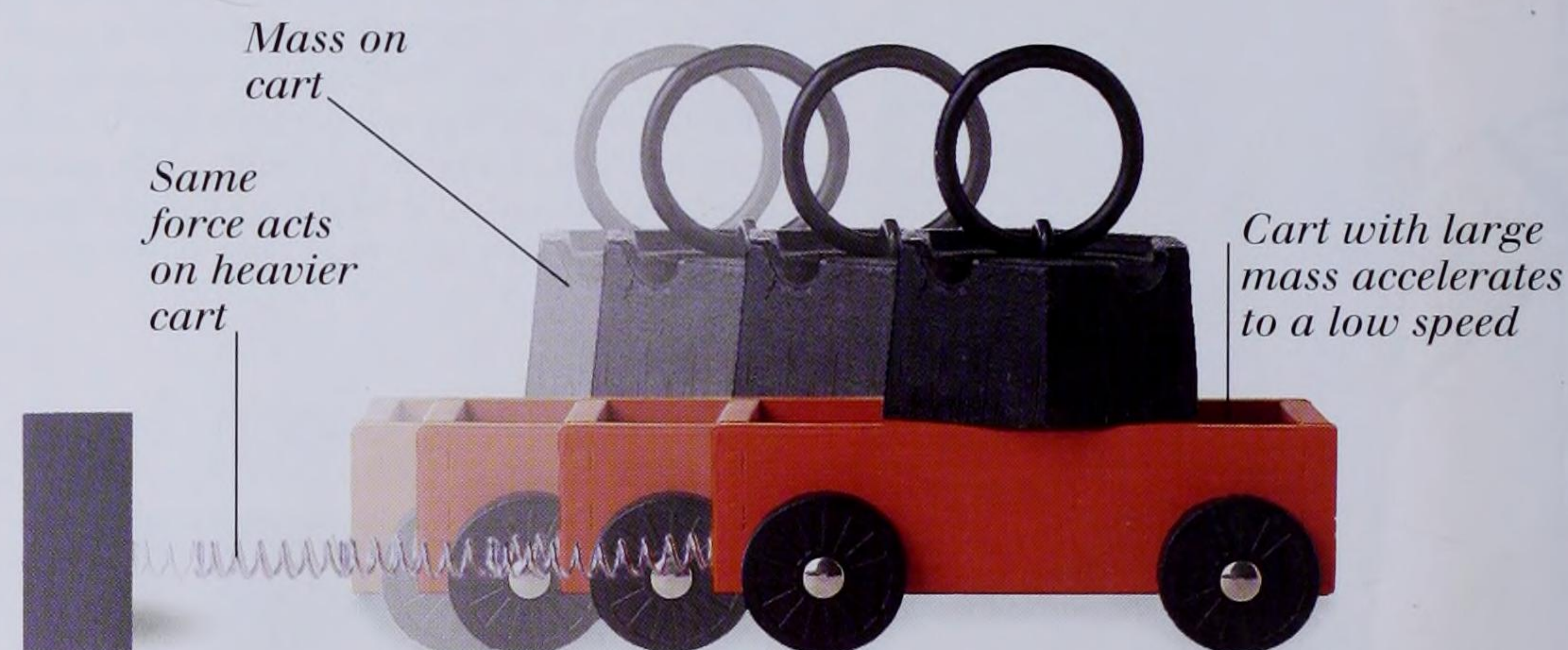
No force, no acceleration: uniform motion

### NEWTON'S SECOND LAW

When a force acts on an object, the motion of the object will change. This change in motion is called acceleration and is equal to the size of the force divided by the mass of the object on which it acts (see p. 54).



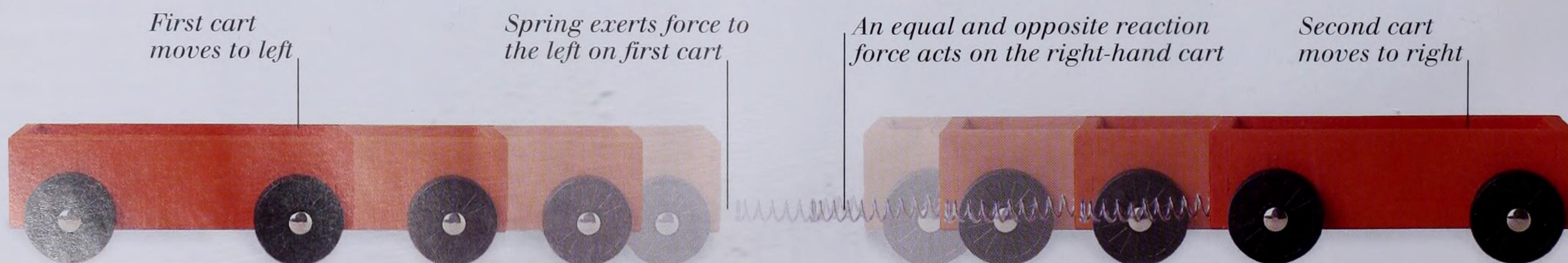
Force acts on small mass: large acceleration



Same force acts on large mass: small acceleration

### NEWTON'S THIRD LAW

If one object exerts a force on another, an equal and opposite force, called the reaction force, is applied by the second to the first.



Action and reaction

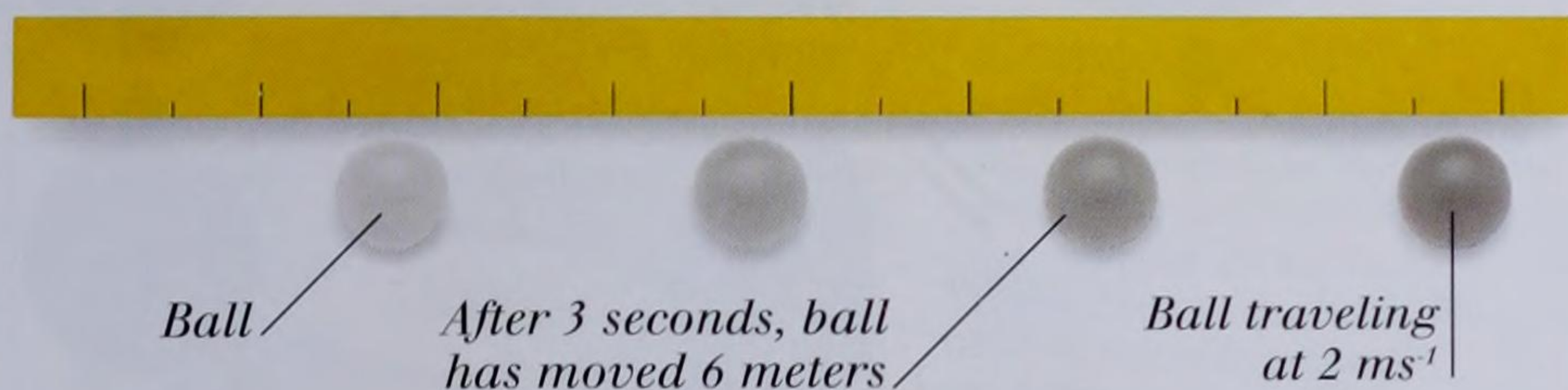
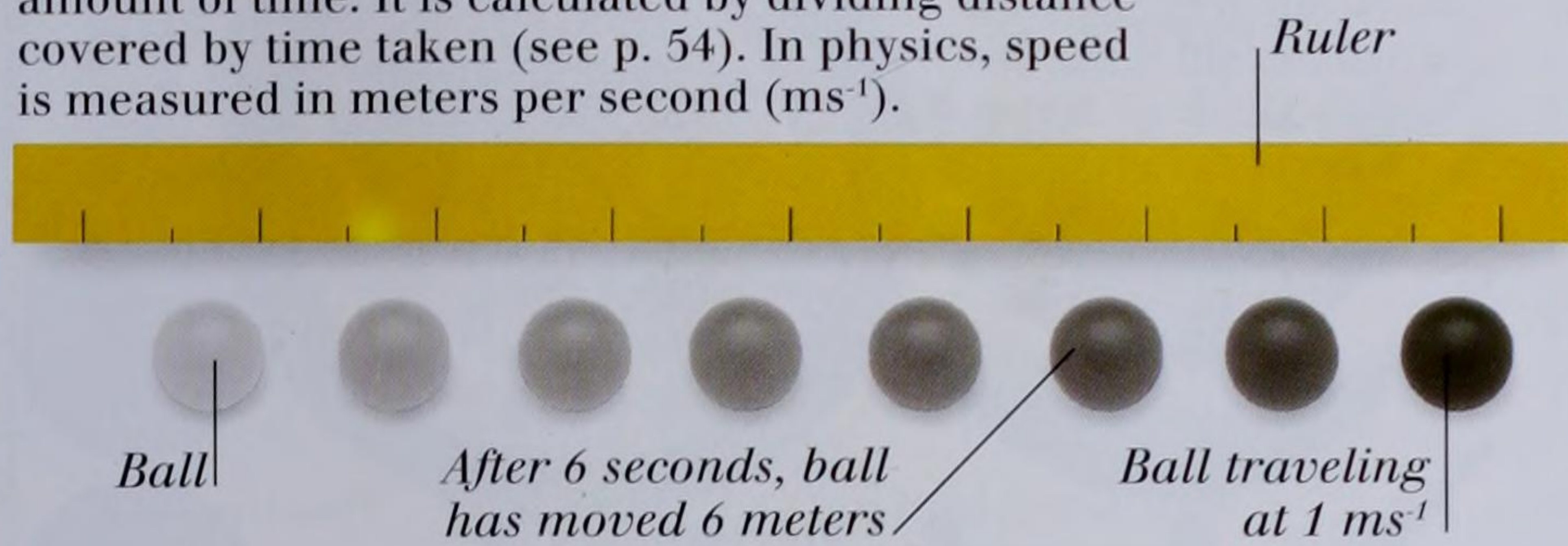


## FORCE AND MOTION

In the images below, each row of balls is a record of the motion of one ball, photographed once each second beside a ruler. This shows how far the ball moved during that second and each subsequent second, giving a visual representation of speed and acceleration.

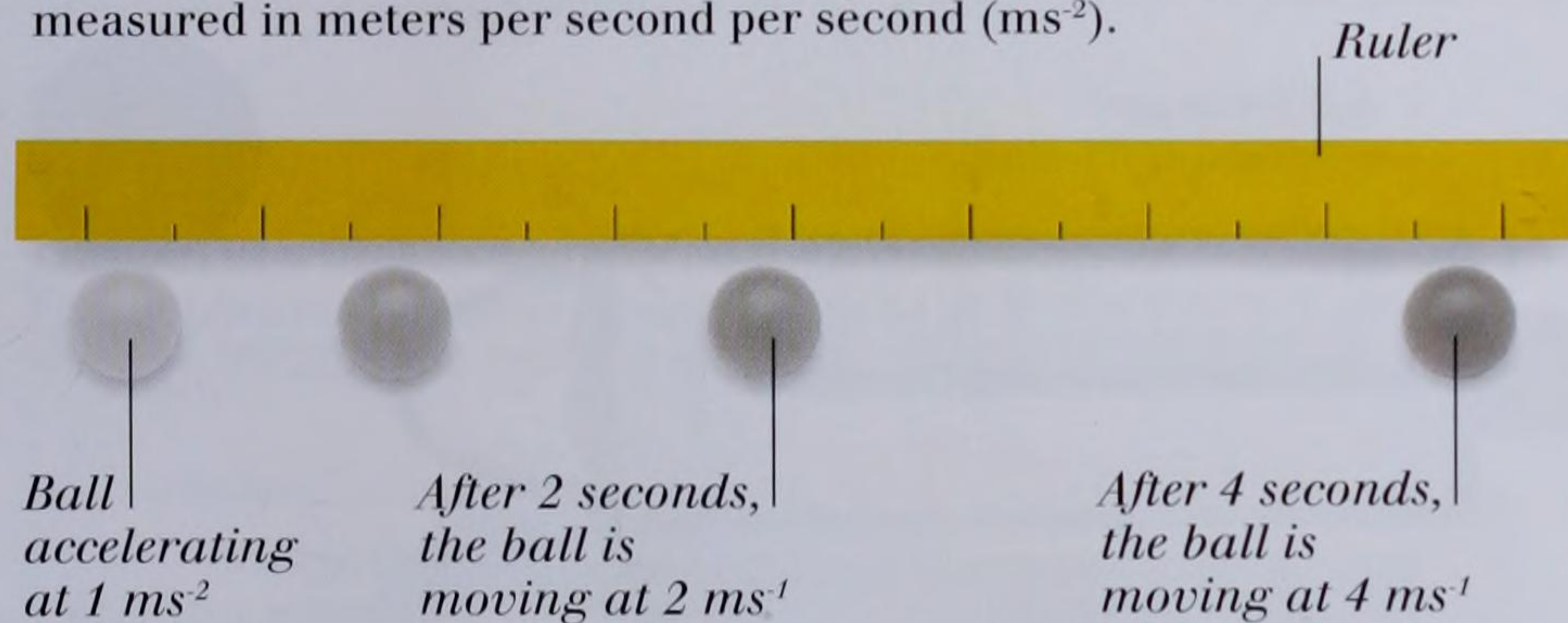
### SPEED

Speed is the distance an object travels in a set amount of time. It is calculated by dividing distance covered by time taken (see p. 54). In physics, speed is measured in meters per second ( $\text{ms}^{-1}$ ).



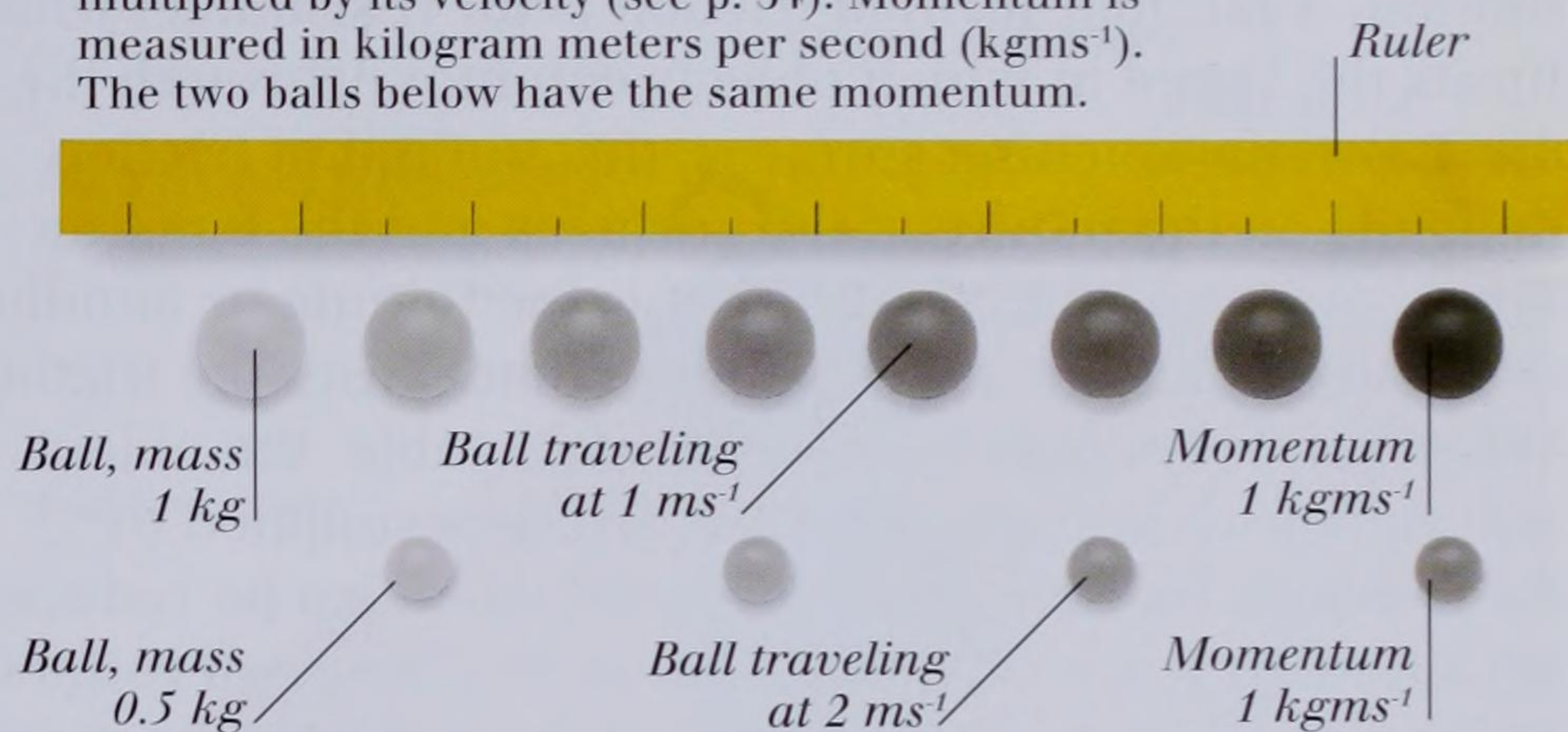
### ACCELERATION

Acceleration is the rate that the speed of an object changes. It is calculated by dividing the change in speed by the time it took for that change (see p. 54). It is measured in meters per second per second ( $\text{ms}^{-2}$ ).

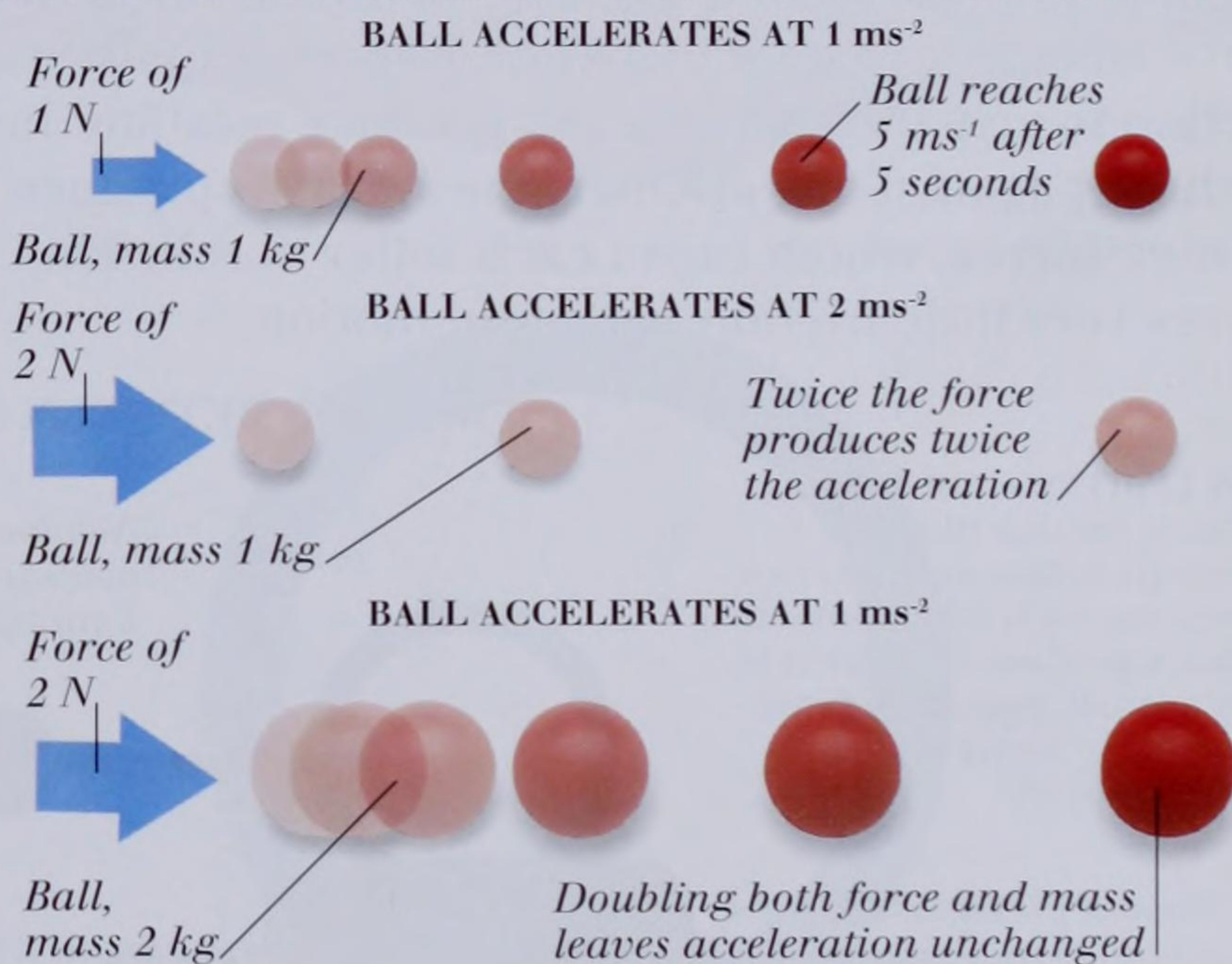


### MOMENTUM

The momentum of an object is equal to its mass multiplied by its velocity (see p. 54). Momentum is measured in kilogram meters per second ( $\text{kgms}^{-1}$ ). The two balls below have the same momentum.

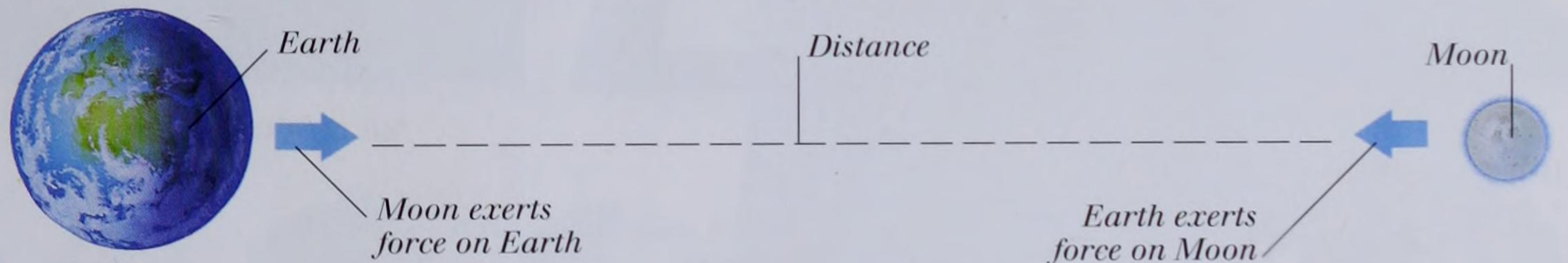


### NEWTON'S SECOND LAW APPLIED TO ACCELERATION

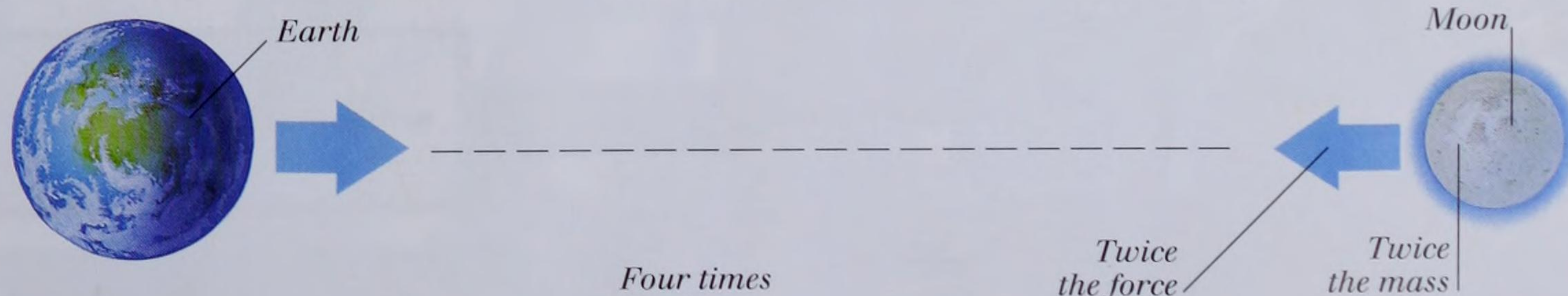


## GRAVITATIONAL FORCE

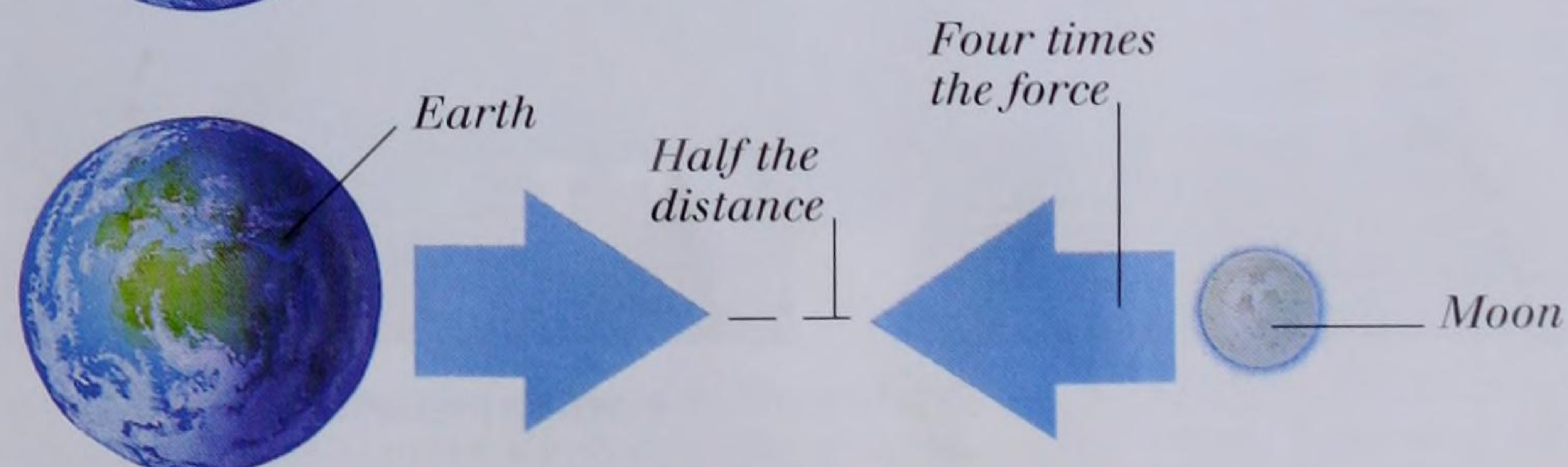
Gravitation, or gravity, is a force that acts on all matter. The force between any two objects depends upon their masses and the distance between them (see p. 54).



If the Moon had twice the mass that it does, the force between the Earth and Moon would be twice as large.



If the Moon were half the distance from the Earth, the gravitational force would be four times as large. This is because the force depends upon the square of the distance.



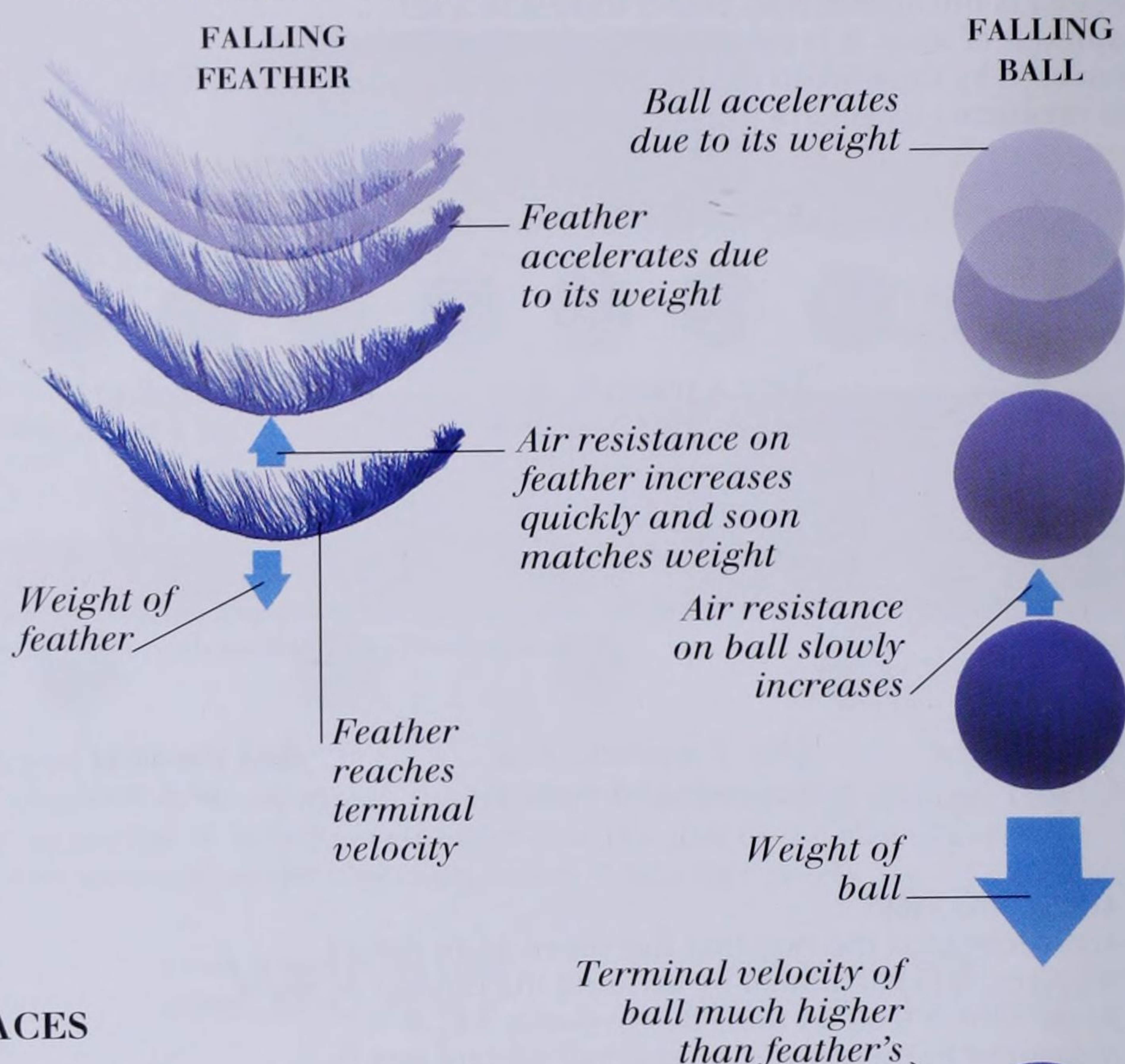


# Friction

**FRICTION IS A FORCE THAT SLOWS DOWN** or prevents motion. A familiar form of friction is air resistance, which limits the speed at which objects can move through the air. Between touching surfaces, the amount of friction depends on the nature of the surfaces and the force or forces pushing them together. It is the joining or bonding of the **atoms** at each of the surfaces that causes the friction. When you try to pull an object along a table, the object will not move until the **limiting friction** supplied by these bonds has been overcome. Friction can be reduced in two main ways: by lubrication or by the use of rollers. Lubrication involves the presence of a **fluid** between two surfaces; fluid keeps the surfaces apart, allowing them to move smoothly past one another. Rollers actually use friction to grip the surfaces and produce rotation. Instead of sliding against one another, the surfaces produce turning forces, which cause each roller to roll. This leaves very little friction to oppose motion.

## AIR RESISTANCE

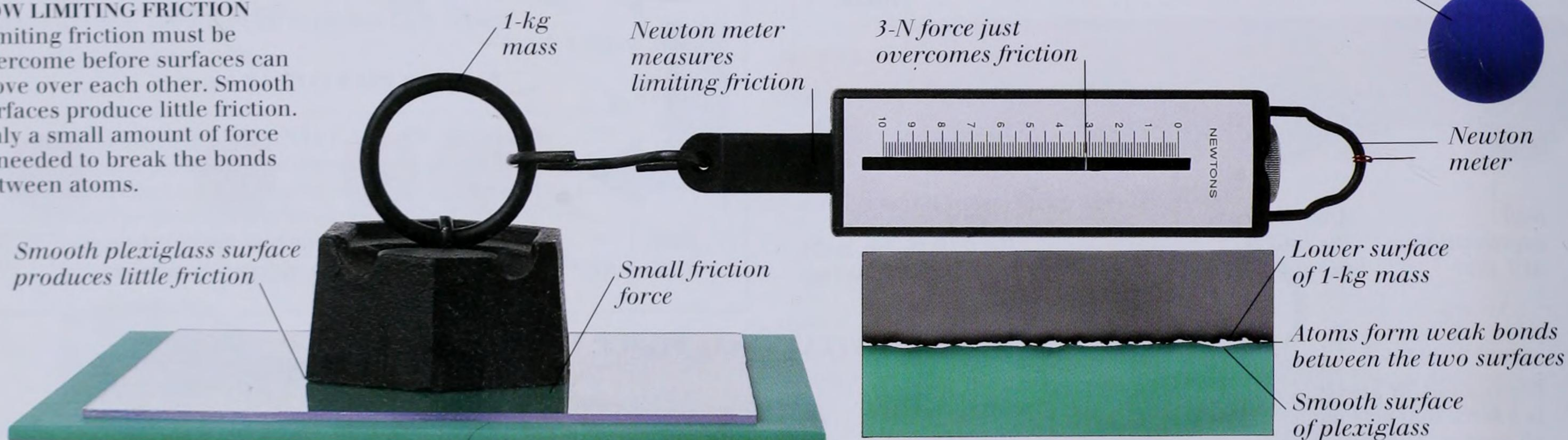
Air resistance is a type of friction that occurs when an object moves through the air. The faster an object moves, the greater the air resistance. Falling objects accelerate to a speed called **terminal velocity**, at which the air resistance exactly balances the object's weight. At this speed, there is no **resultant force** and so no further acceleration can occur.



## FRICTION BETWEEN SURFACES

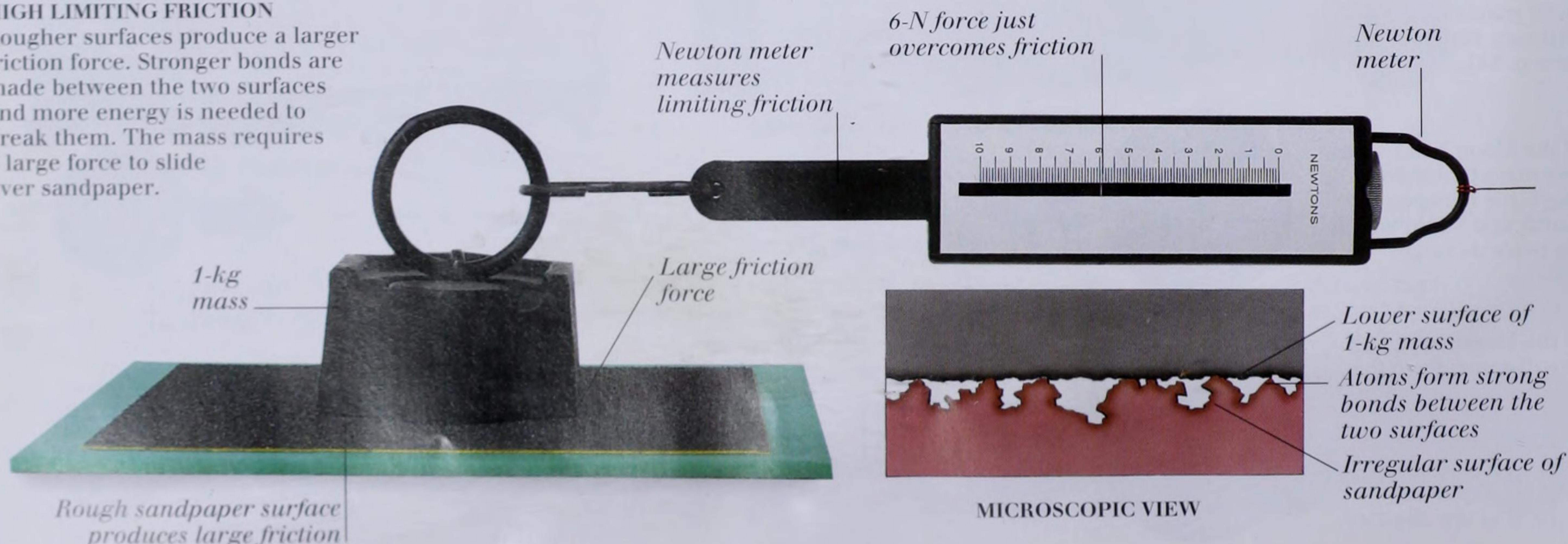
### LOW LIMITING FRICTION

Limiting friction must be overcome before surfaces can move over each other. Smooth surfaces produce little friction. Only a small amount of force is needed to break the bonds between atoms.



### HIGH LIMITING FRICTION

Rougher surfaces produce a larger friction force. Stronger bonds are made between the two surfaces and more energy is needed to break them. The mass requires a large force to slide over sandpaper.



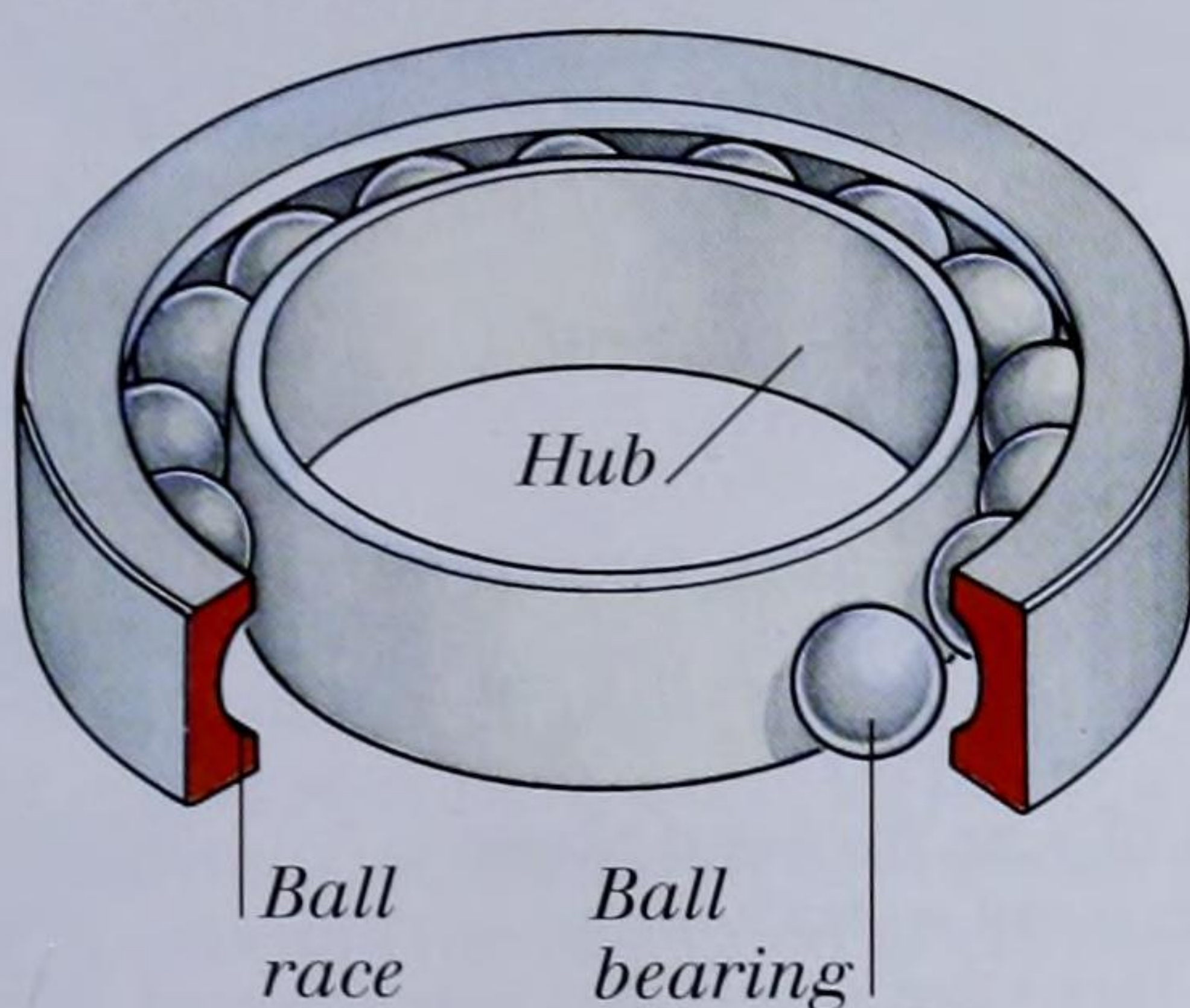


## MOTORCYCLE BRAKE



Friction is put to good use in the disk brakes of a motorcycle. The friction force between disk and brake pad slows down the rotation of the wheel, reducing the vehicle's speed. In doing so, it converts the **kinetic energy** of the vehicle into heat (see p. 7).

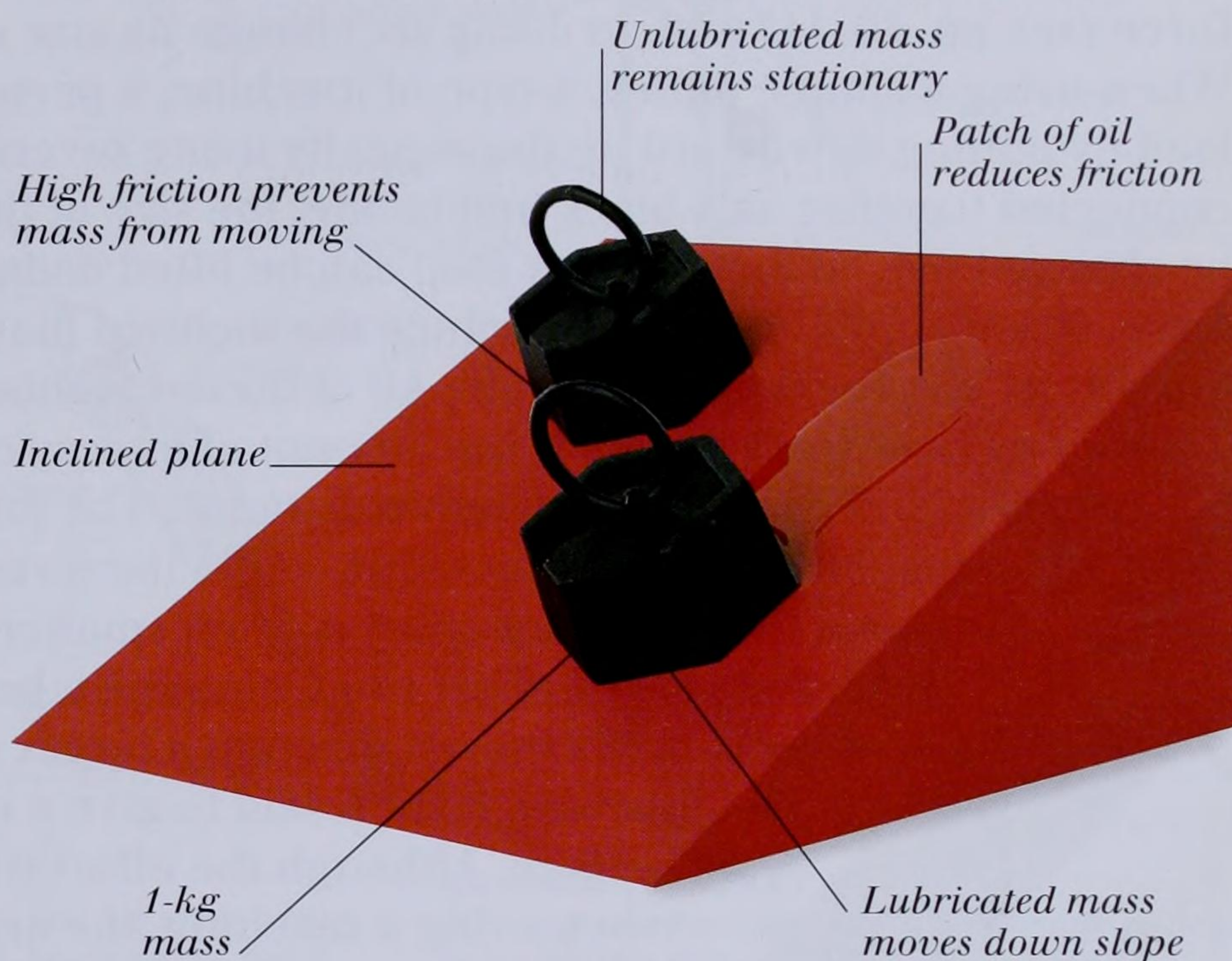
## BALL BEARINGS



Bearings are a type of roller used to reduce friction between moving machine parts such as a wheel and its axle. As a wheel turns on its axle, the balls roll around inside the bearing, drastically reducing the friction between wheel and axle.

## LUBRICATION

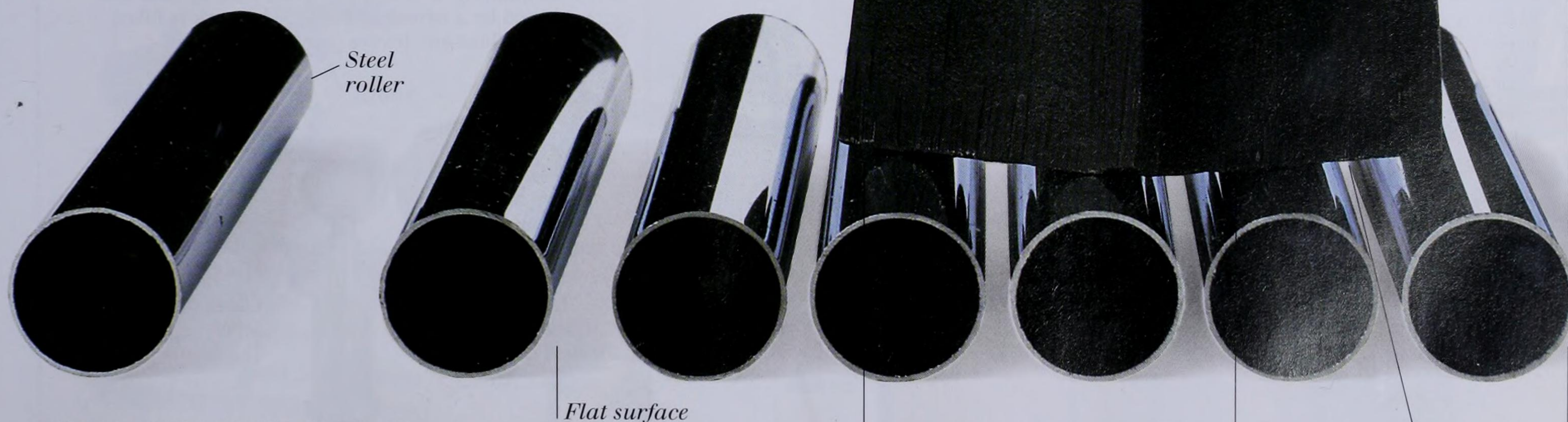
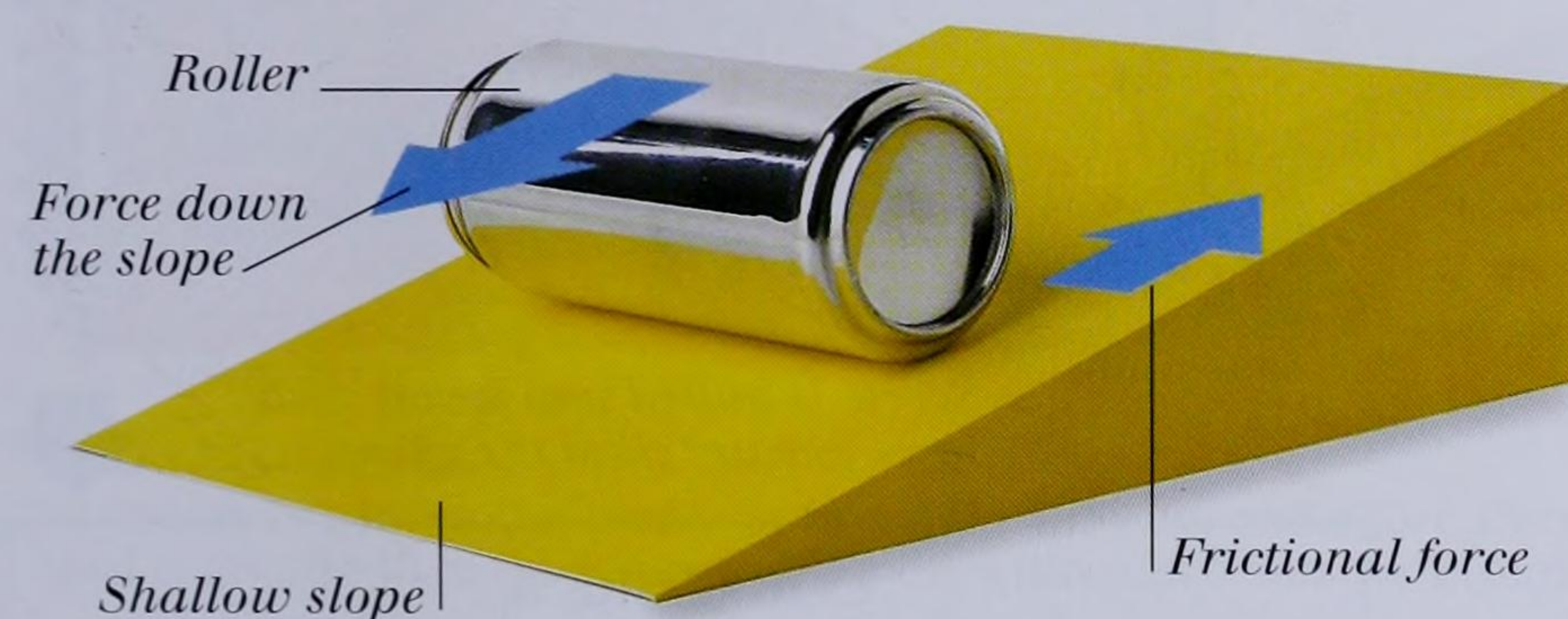
The presence of oil or another fluid between two surfaces keeps the surfaces apart. Because fluids (liquids or gases) flow, they allow movement between surfaces. Here, a lubricated kilogram mass slides down a slope, while an unlubricated one is prevented from moving by friction.



## ROLLERS

### THE ACTION OF A ROLLER ON A SLOPE

Friction causes the roller to grip the slope so that it turns. If there were no friction, the roller would simply slide down the slope.



### USING ROLLERS TO AVOID FRICTION

Rollers placed between two surfaces keep the surfaces apart. The rollers allow the underside of the kilogram mass to move freely over the ground. An object placed on rollers will move smoothly if pushed or pulled.

Friction forces between surfaces create a turning force that turns the rollers

Mass moves smoothly over surface

Underside of 1-kg mass



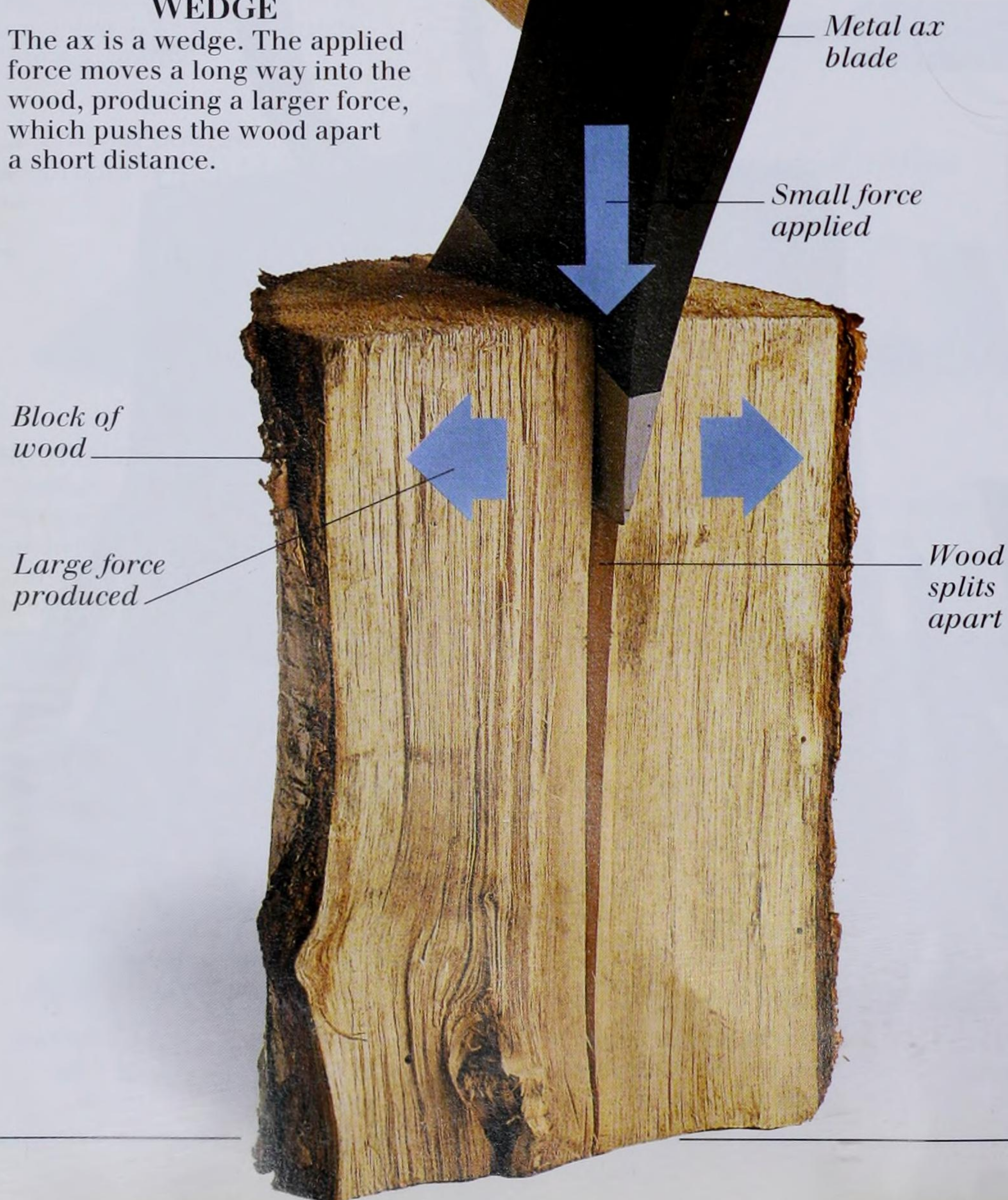
# Simple machines

IN PHYSICS, A MACHINE IS ANY DEVICE that can be used to transmit a **force** (see pp. 10–11) and, in doing so, change its size or direction. When using a simple pulley, a type of machine, a person can lift a load by pulling downward on the rope. By using several pulleys connected together as a block and tackle, the size of the force can be changed too, so that a heavy load can be lifted using a small force. Other simple machines include the inclined plane, the lever, the screw, and the wheel and axle. All of these machines illustrate the concept of **work**. Work is the amount of energy expended when a force is moved through a distance. The force applied to a machine is called the **effort**, while the force it overcomes is called the **load**. The effort is often smaller than the load, for a small effort can overcome a heavy load if the effort is moved through a larger distance.

The machine is then said to give a mechanical advantage. Although the effort will be smaller when using a machine, the amount of work done, or energy used, will be equal to or greater than that without the machine.

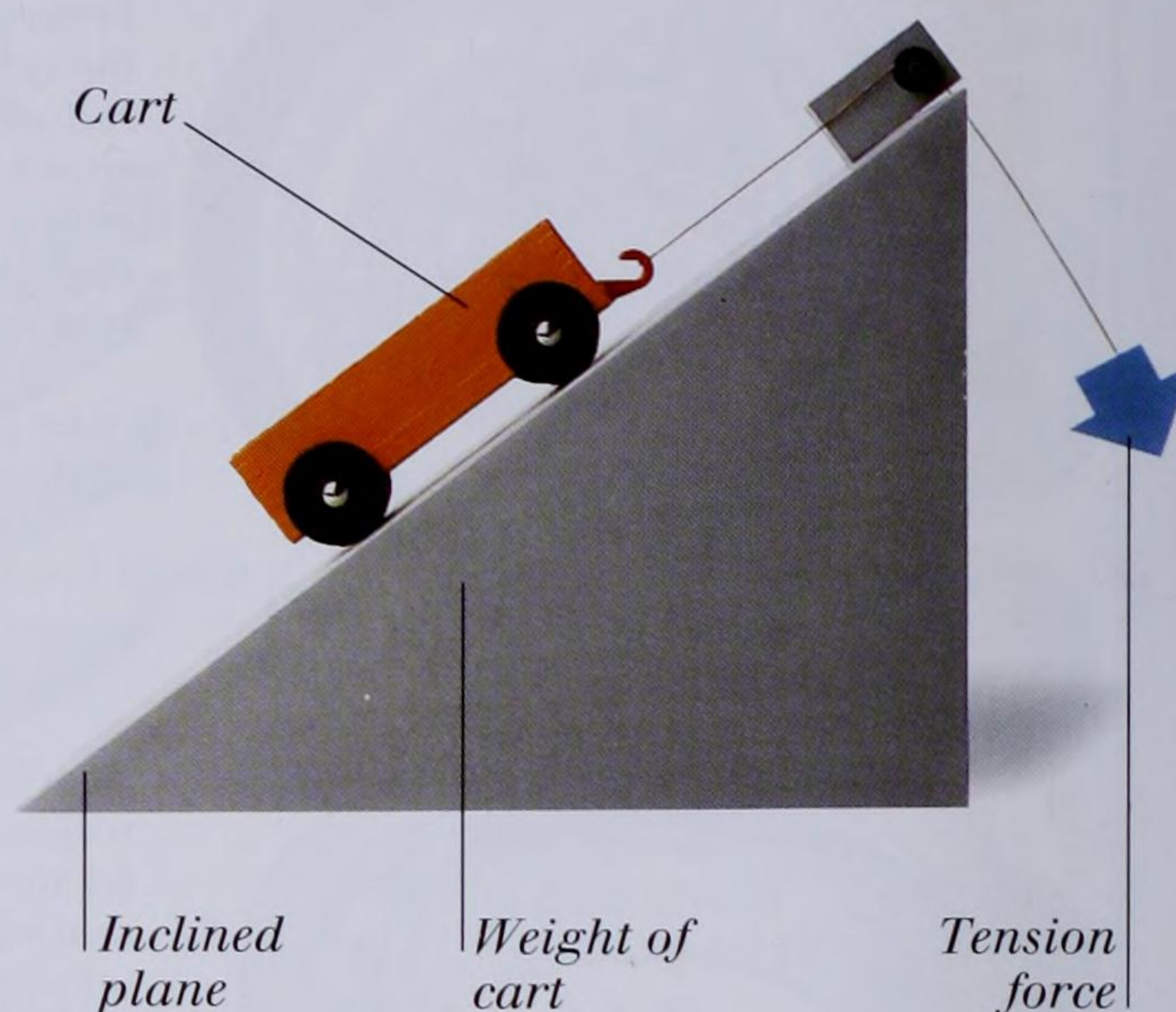
## WEDGE

The ax is a wedge. The applied force moves a long way into the wood, producing a larger force, which pushes the wood apart a short distance.



## AN INCLINED PLANE

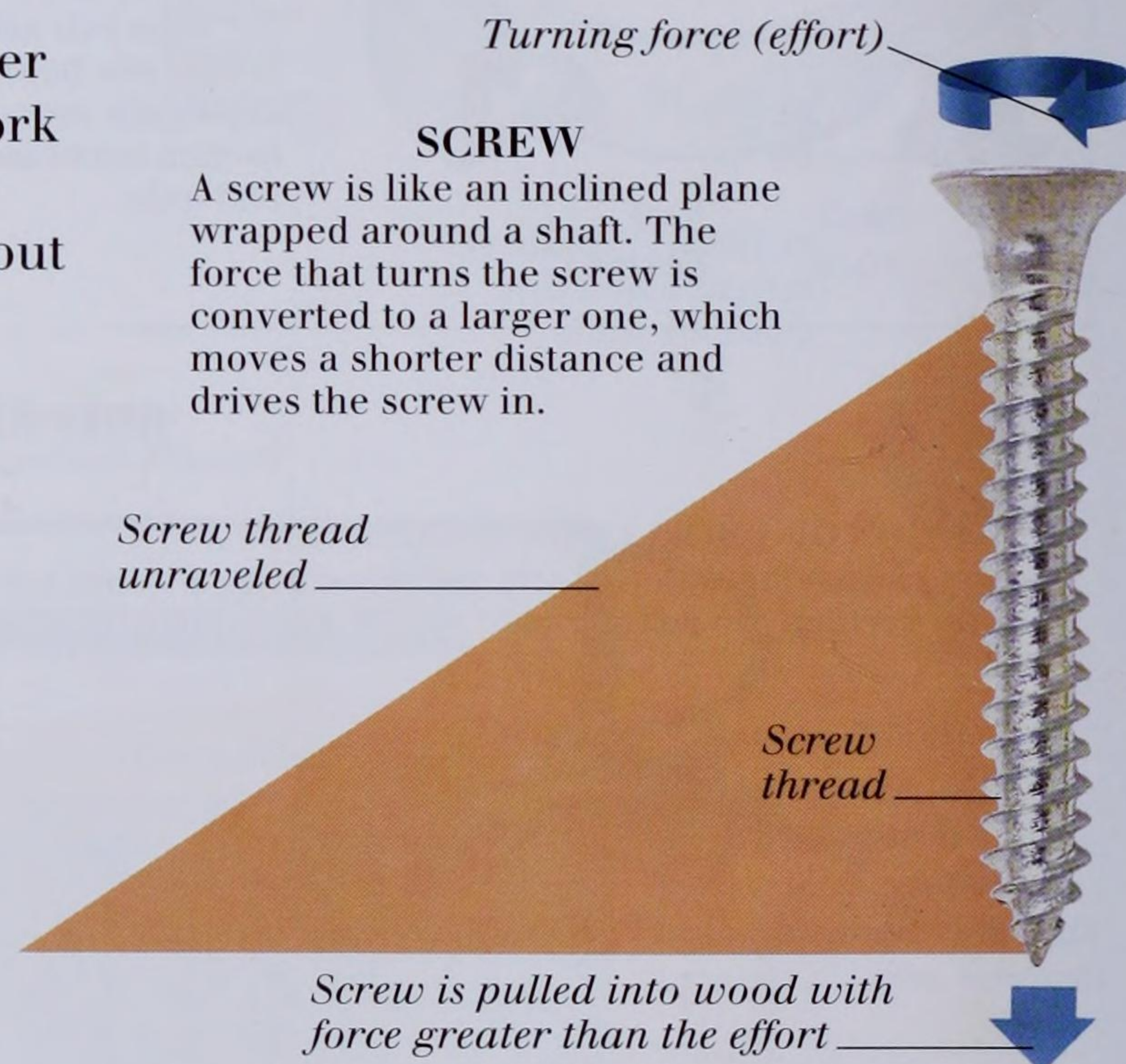
The force needed to drag an object up a slope is less than that needed to lift it vertically. However, the distance moved by the object is greater when pulled up the slope than if it were lifted vertically.



Turning force (effort)

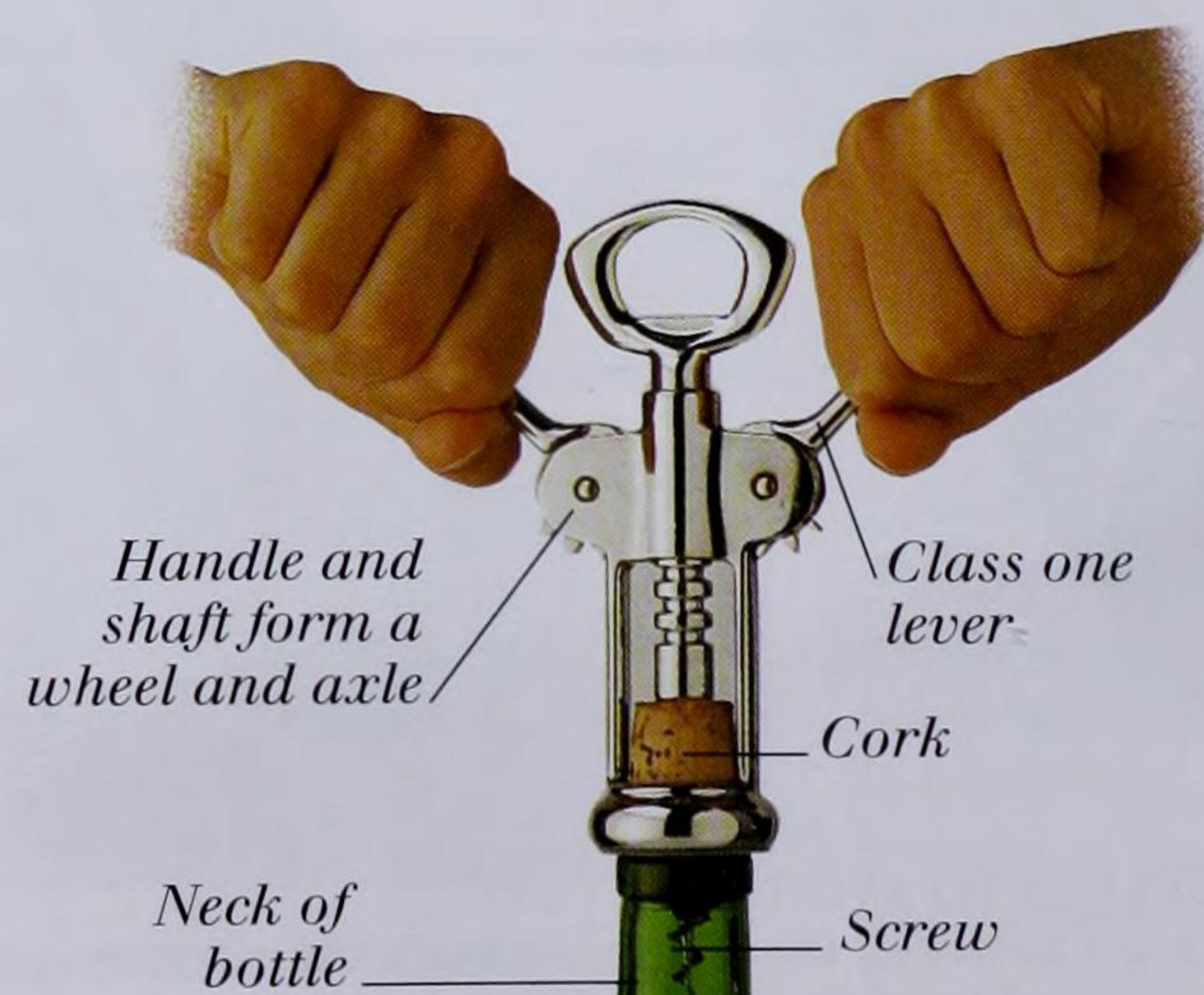
## SCREW

A screw is like an inclined plane wrapped around a shaft. The force that turns the screw is converted to a larger one, which moves a shorter distance and drives the screw in.



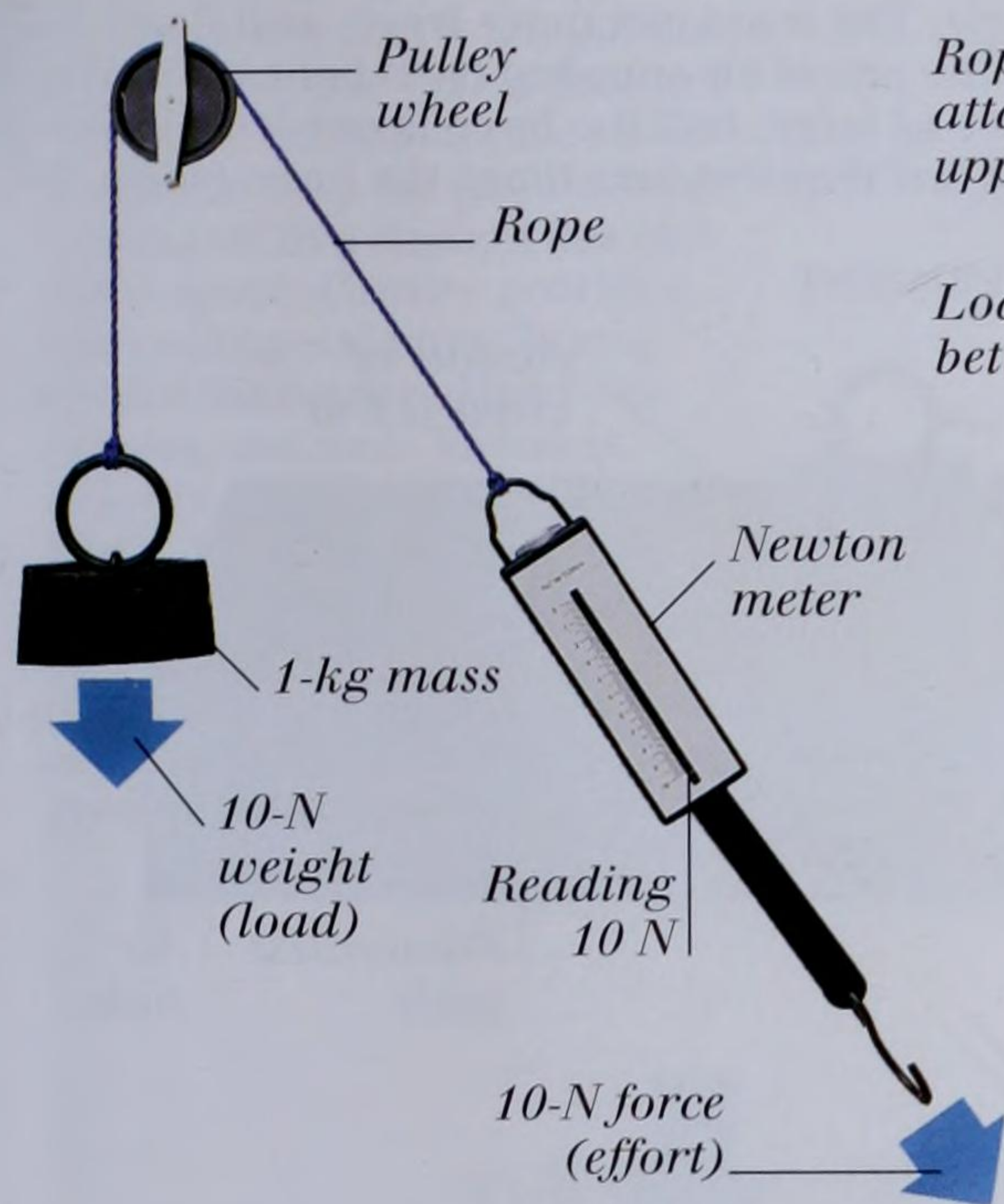
## CORKSCREW

The corkscrew is a clever combination of several different machines. The screw pulls its way into the cork, turned by a wheel and axle. The cork is lifted by a pair of class one levers (see opposite).



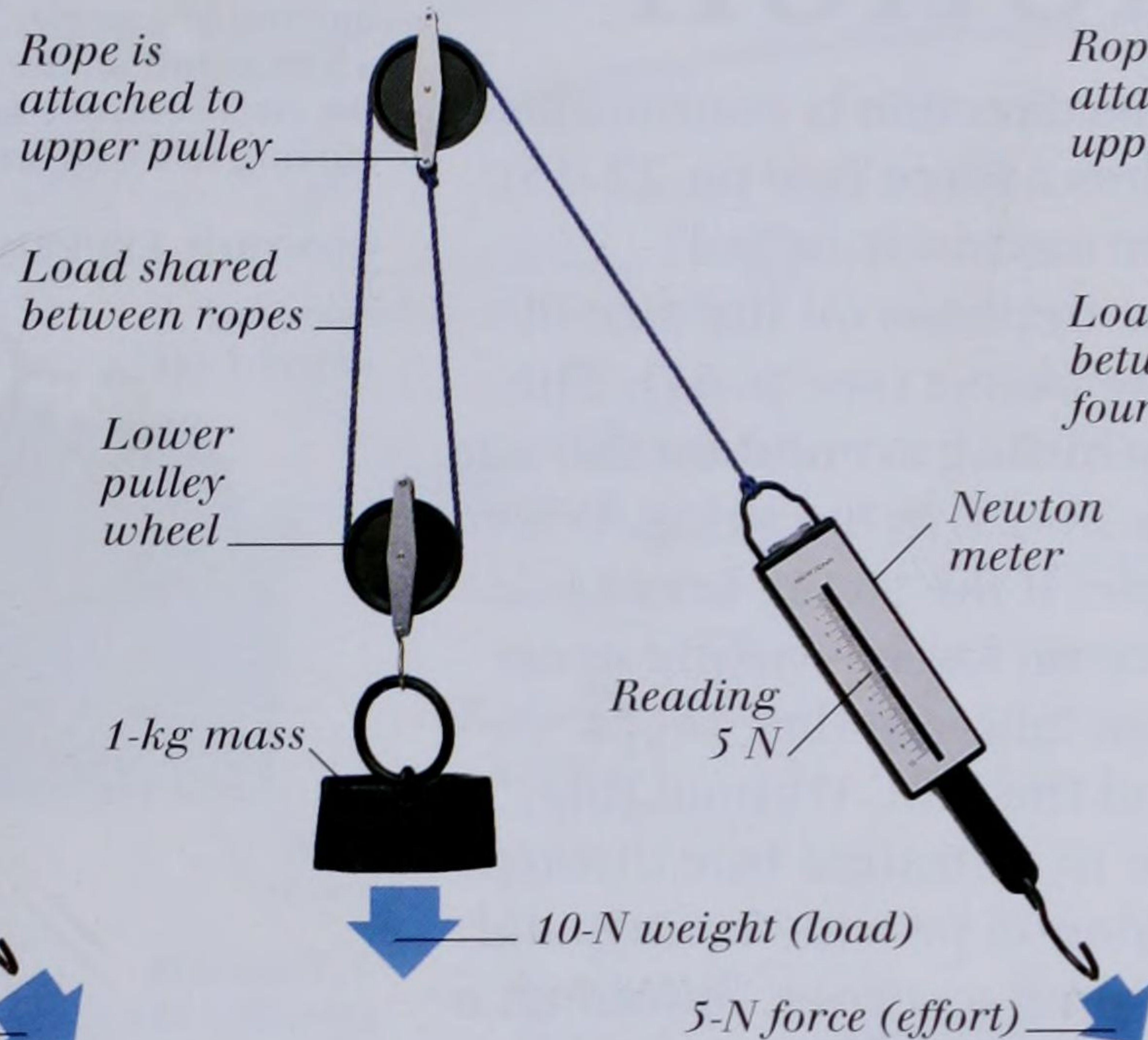


## PULLEYS



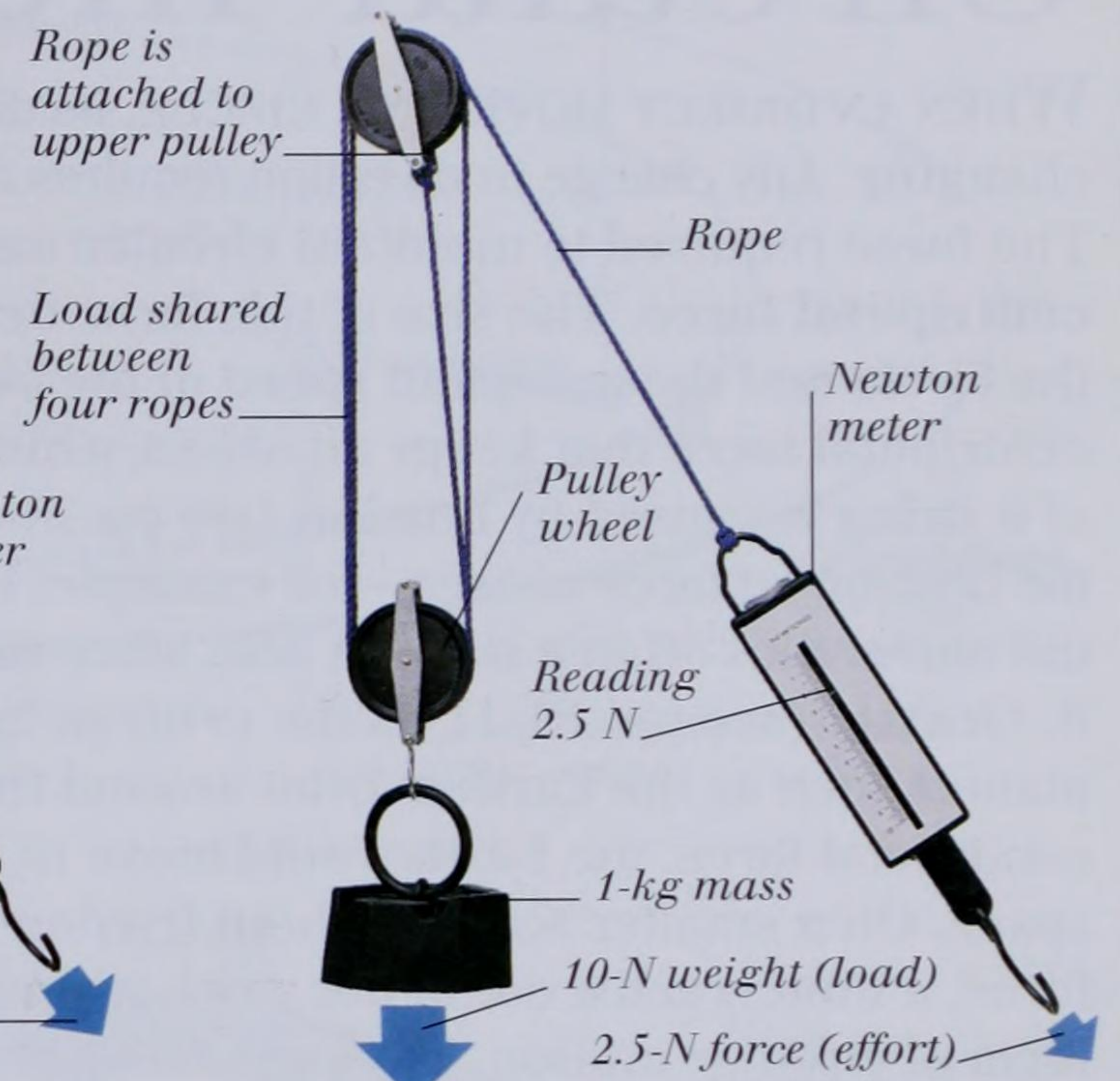
### SIMPLE PULLEY

A simple pulley changes the direction of a force but not its size. Here a one-kg mass, weighing ten newtons, is lifted by a ten-newton force. The mass and the other end of the rope move through the same distance.



### DOUBLE PULLEY

A double pulley will lift a one-kg mass with only a five-newton effort, because the force in the rope doubles up as the rope does. However, pulling the rope by one meter only raises the mass by half a meter.



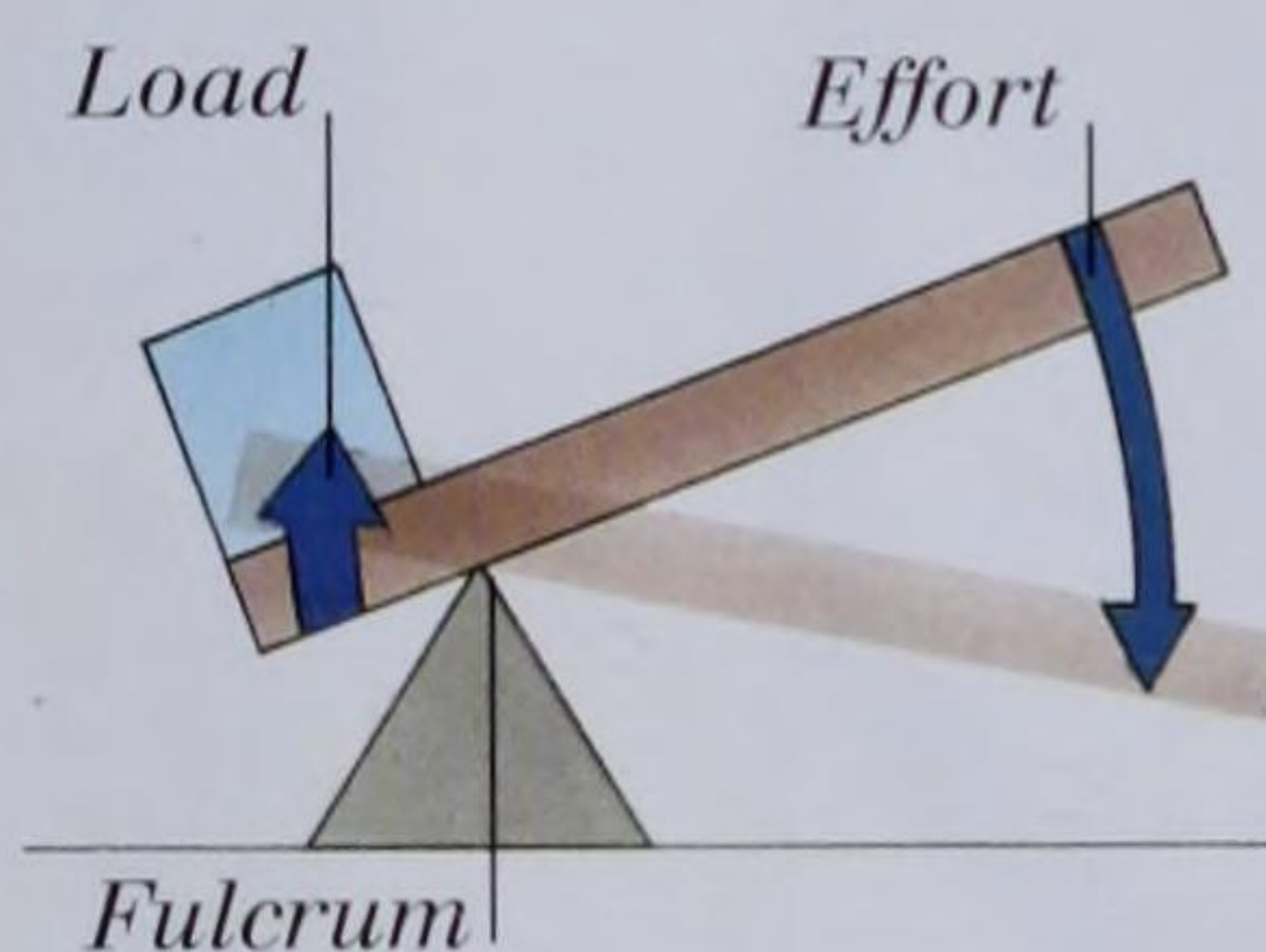
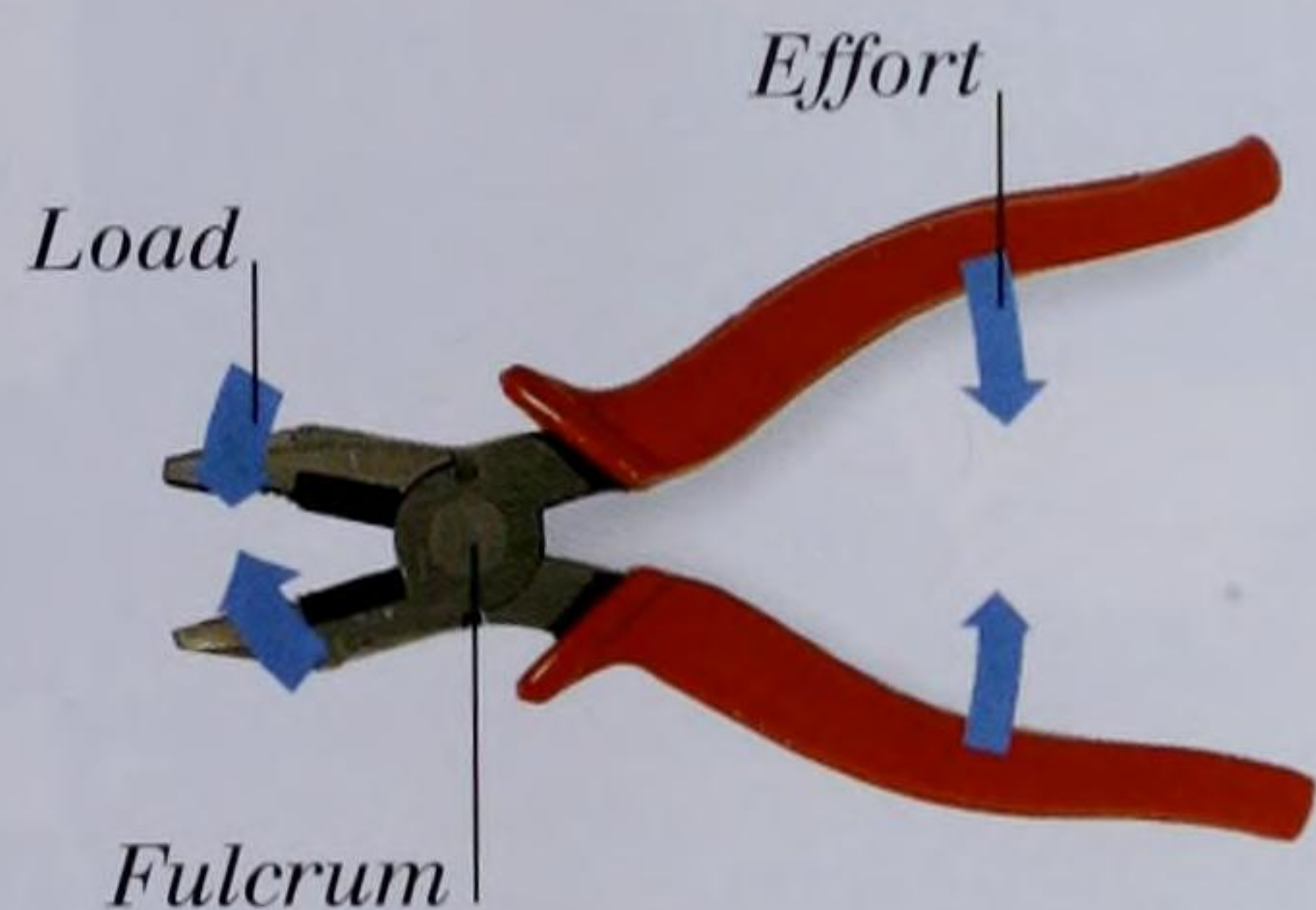
### QUADRUPLE PULLEY

Lifting a one-kg mass with a quadruple pulley, in which the rope goes over four pulley wheels, feels almost effortless. However, pulling the rope by one meter lifts the mass by only one quarter of a meter.

## THREE CLASSES OF LEVER

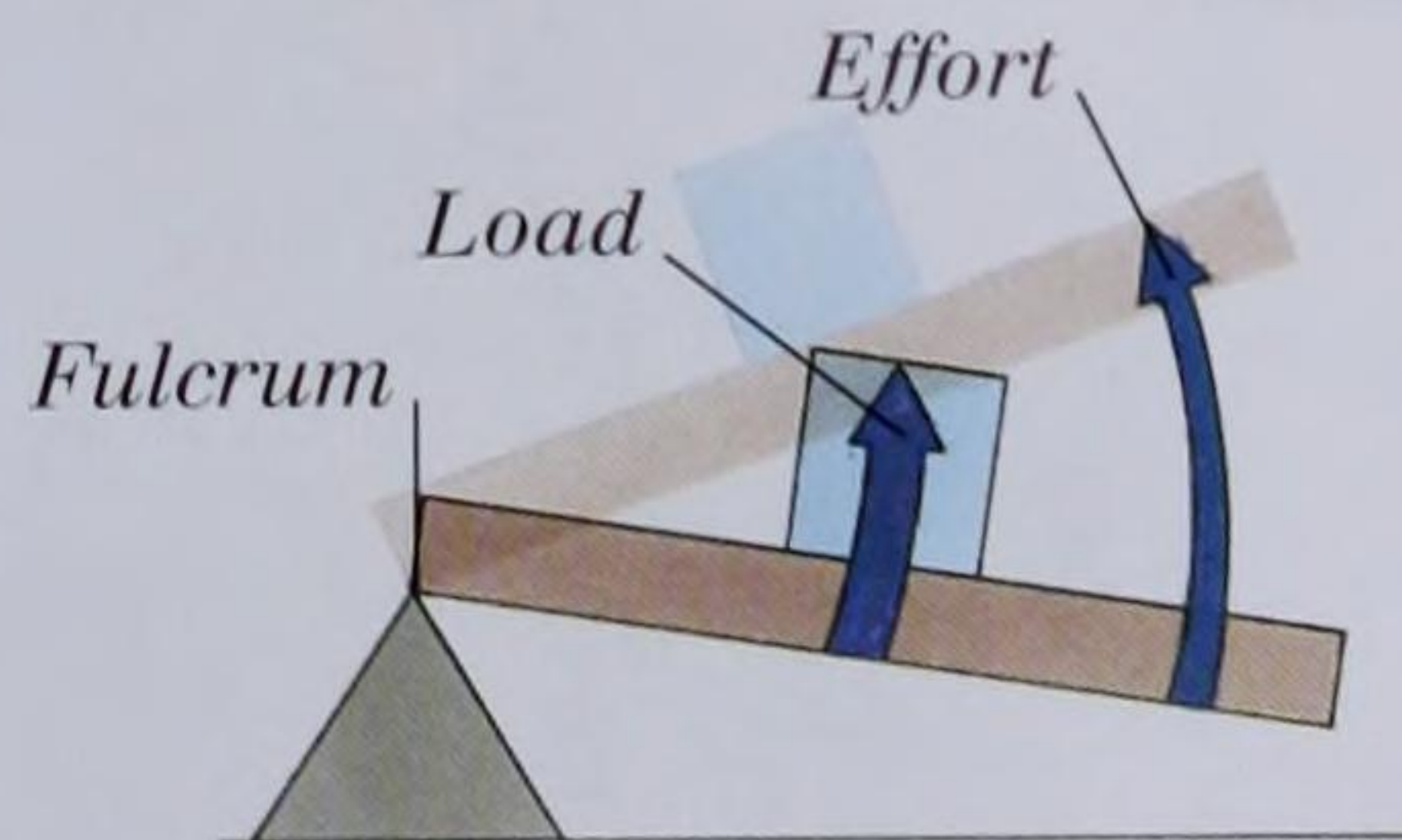
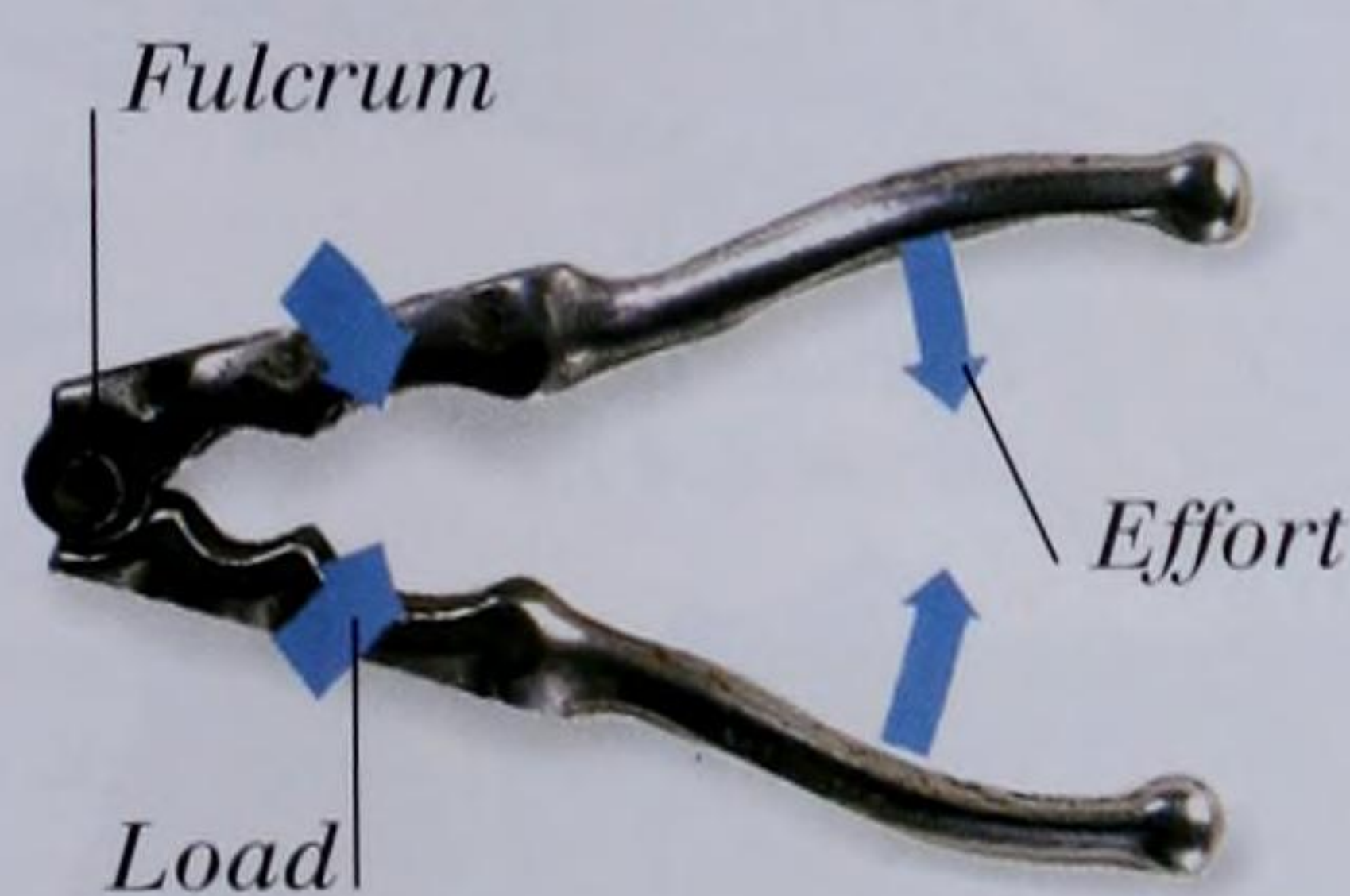
### CLASS ONE LEVER

In a class one lever, the **fulcrum** (pivot point) is between the effort and the load. The load is larger than the effort, but it moves through a smaller distance.



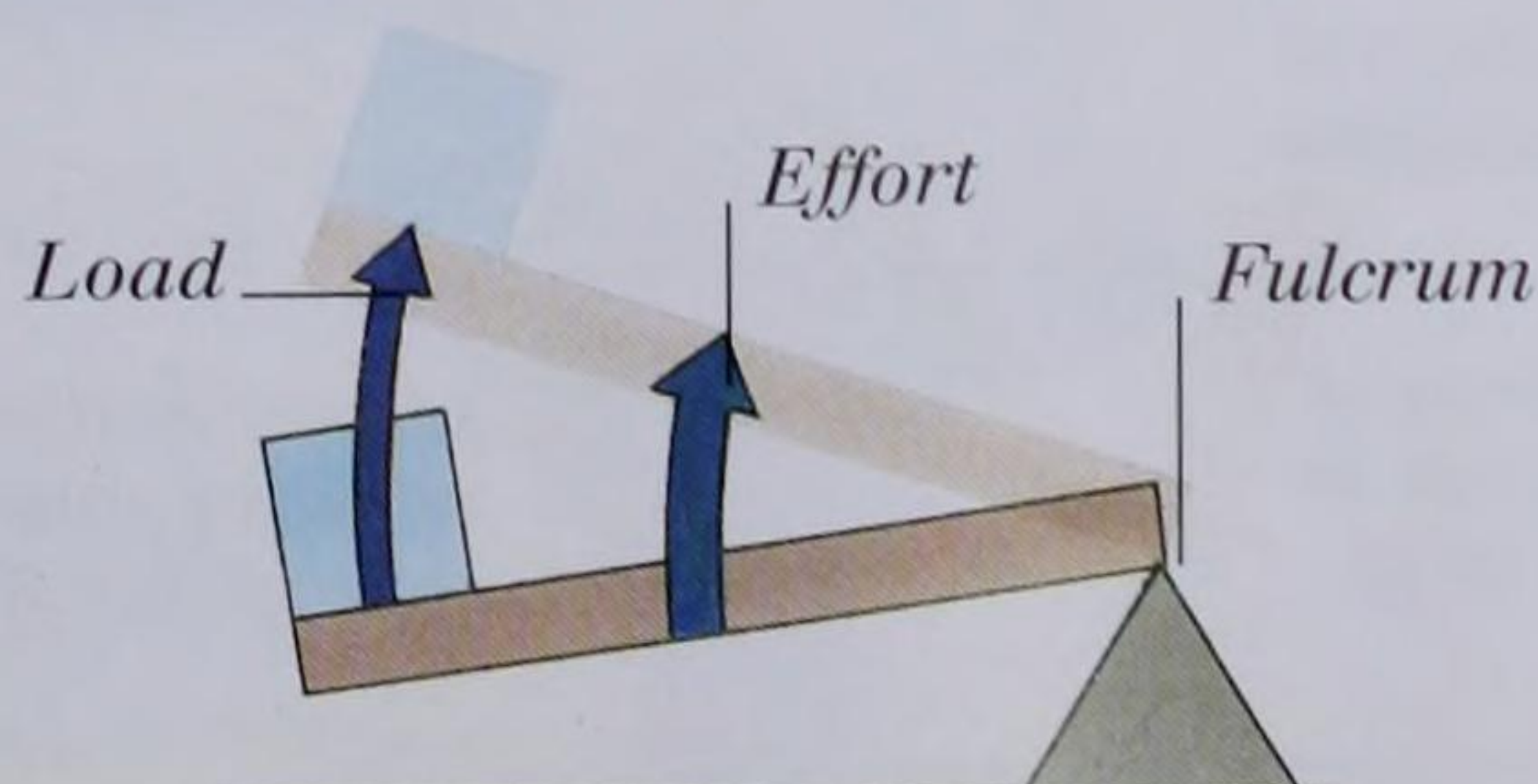
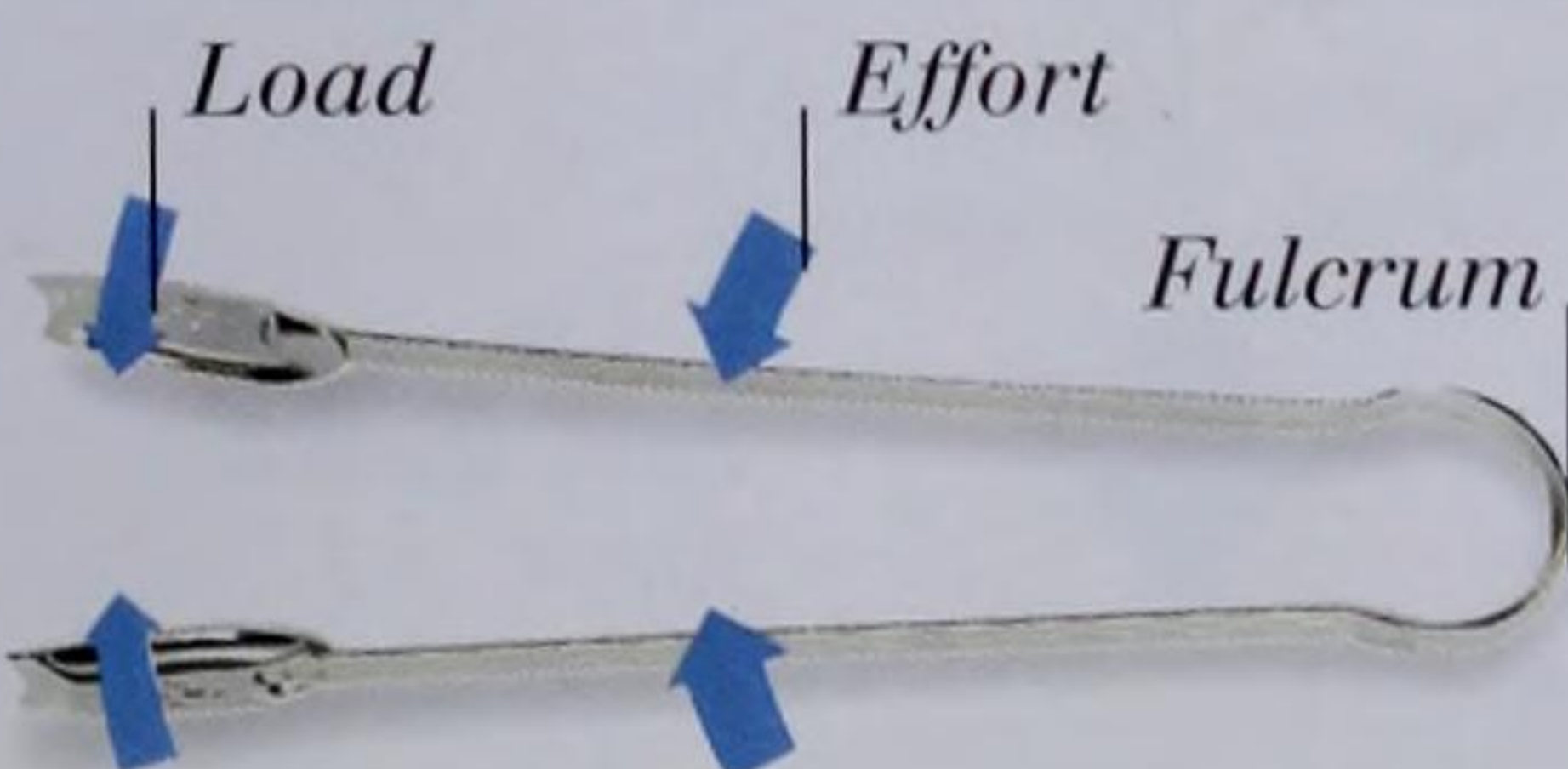
### CLASS TWO LEVER

In a class two lever, the load is between the fulcrum and effort. Here again, the load is greater than the effort and it moves through a smaller distance.



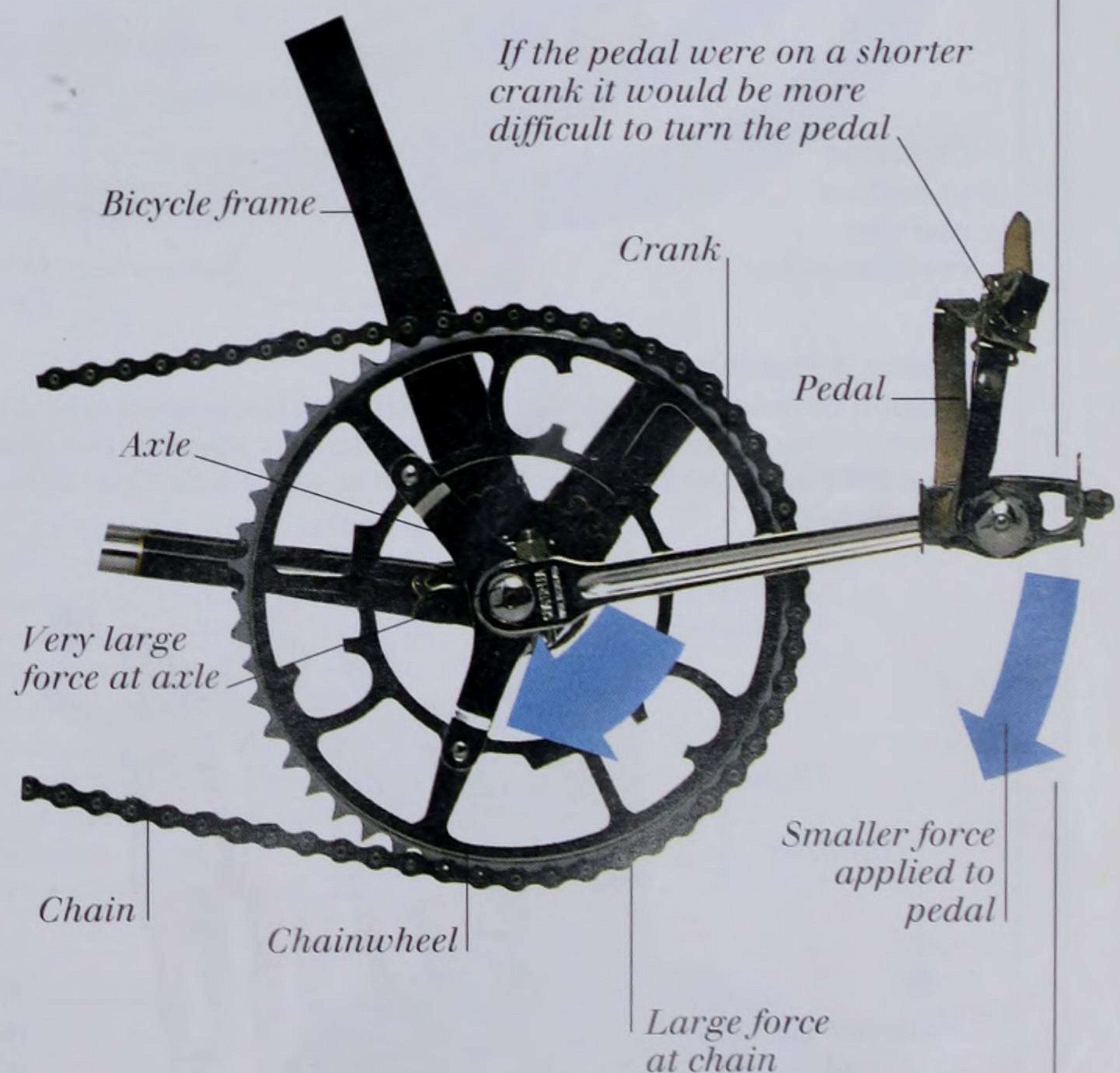
### CLASS THREE LEVER

In a class three lever, the effort is between the fulcrum and the load. In this case, the load is less than the effort but it moves through a greater distance.



## WHEEL AND AXLE

As the pedal and chainwheel of a bicycle turn through one revolution, the pedal moves farther than the links of the chain. For this reason, the force applied to the chain is greater than the force applied to the pedal. The steering wheel of a car is another example of a wheel and axle.





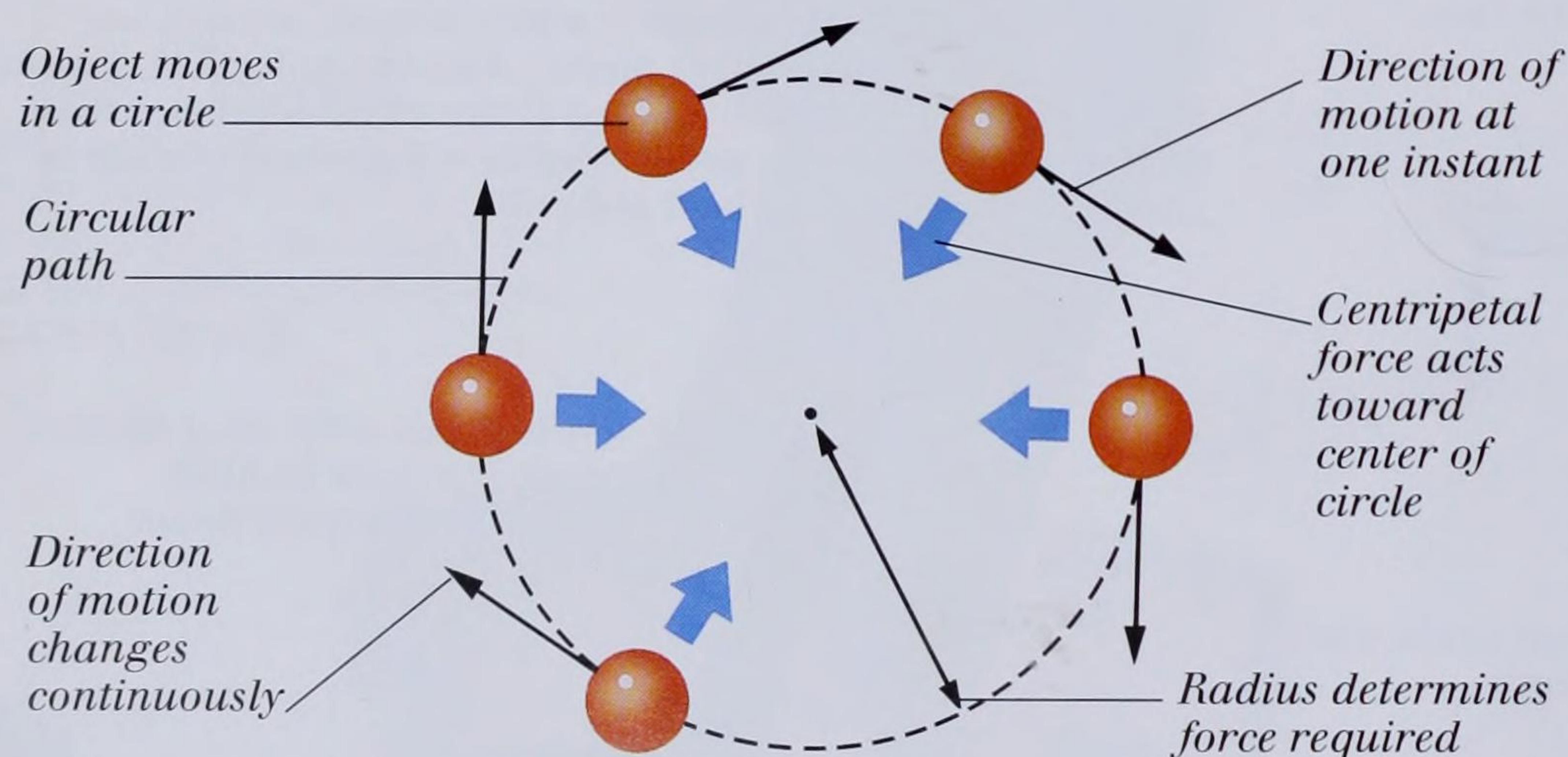
# Circular motion

WHEN AN OBJECT MOVES IN A CIRCLE, its direction is continuously changing. Any change in direction requires a **force** (see pp. 12–13). The force required to maintain circular motion is called **centripetal force**. The size of this force depends on the size of the circle and the mass and speed of the object (see p. 54). The centripetal force that keeps an object whirling around on the end of a string is caused by **tension** (see pp. 24–25) in the string. When the centripetal force ceases—for example, if the string breaks—the object flies off in a straight line, since no force is acting upon it. **Gravity** (see pp. 10–11) is the centripetal force that keeps planets such as the Earth in orbit around the Sun. Without this centripetal force, the Earth would move in a straight line through space. On a smaller scale, without friction to provide centripetal force, a motorcyclist could not steer around a corner. Spinning, a form of circular motion, gives **gyroscopes** stability.

## MOTION IN A CIRCLE

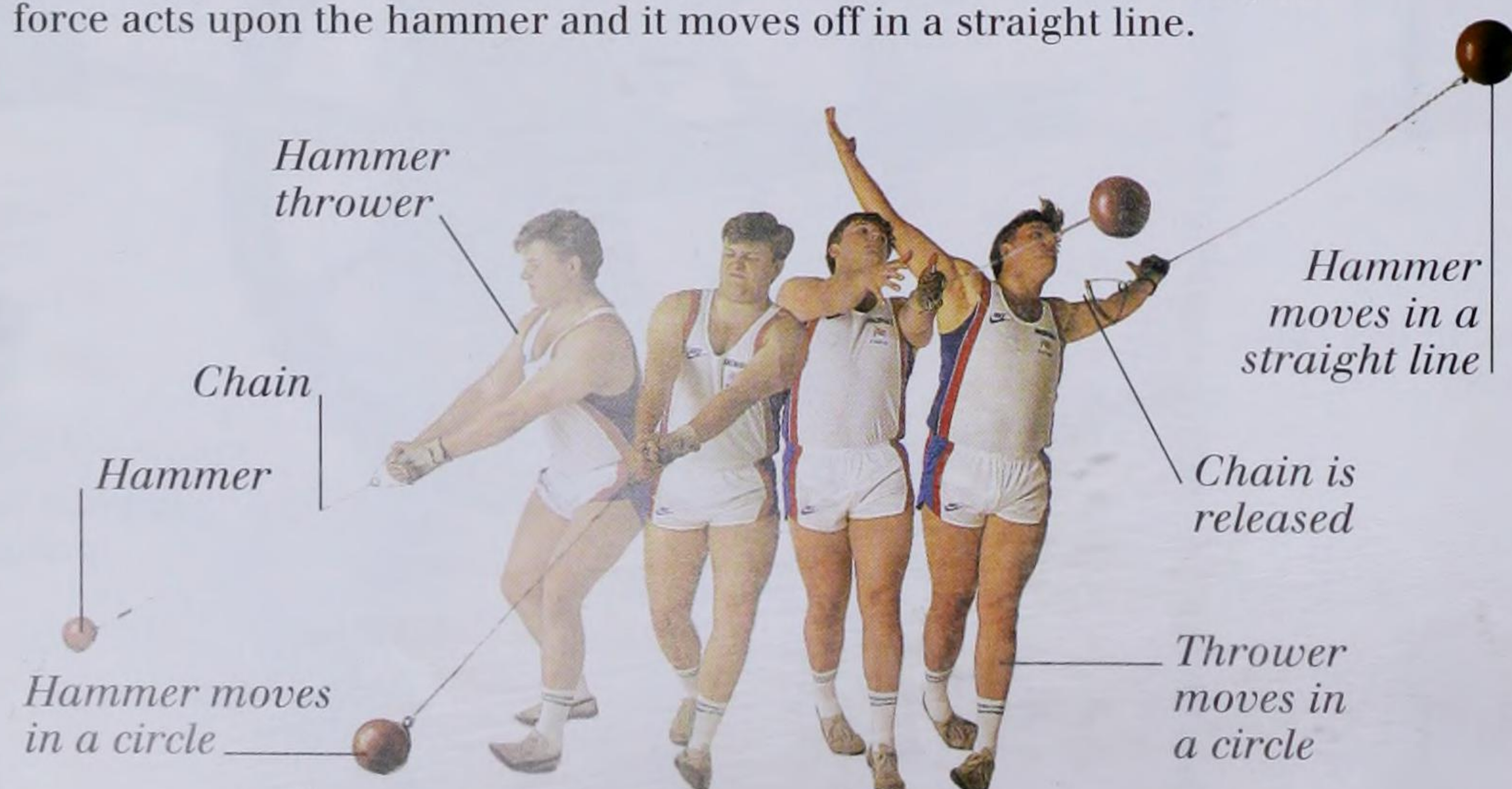
### ASPECTS OF CIRCULAR MOTION

The force that continuously changes the direction of an object moving in a circle is called centripetal force. It is directed toward the center of the circle. The smaller the radius of the circle, the larger the force needed.



### HAMMER THROWER

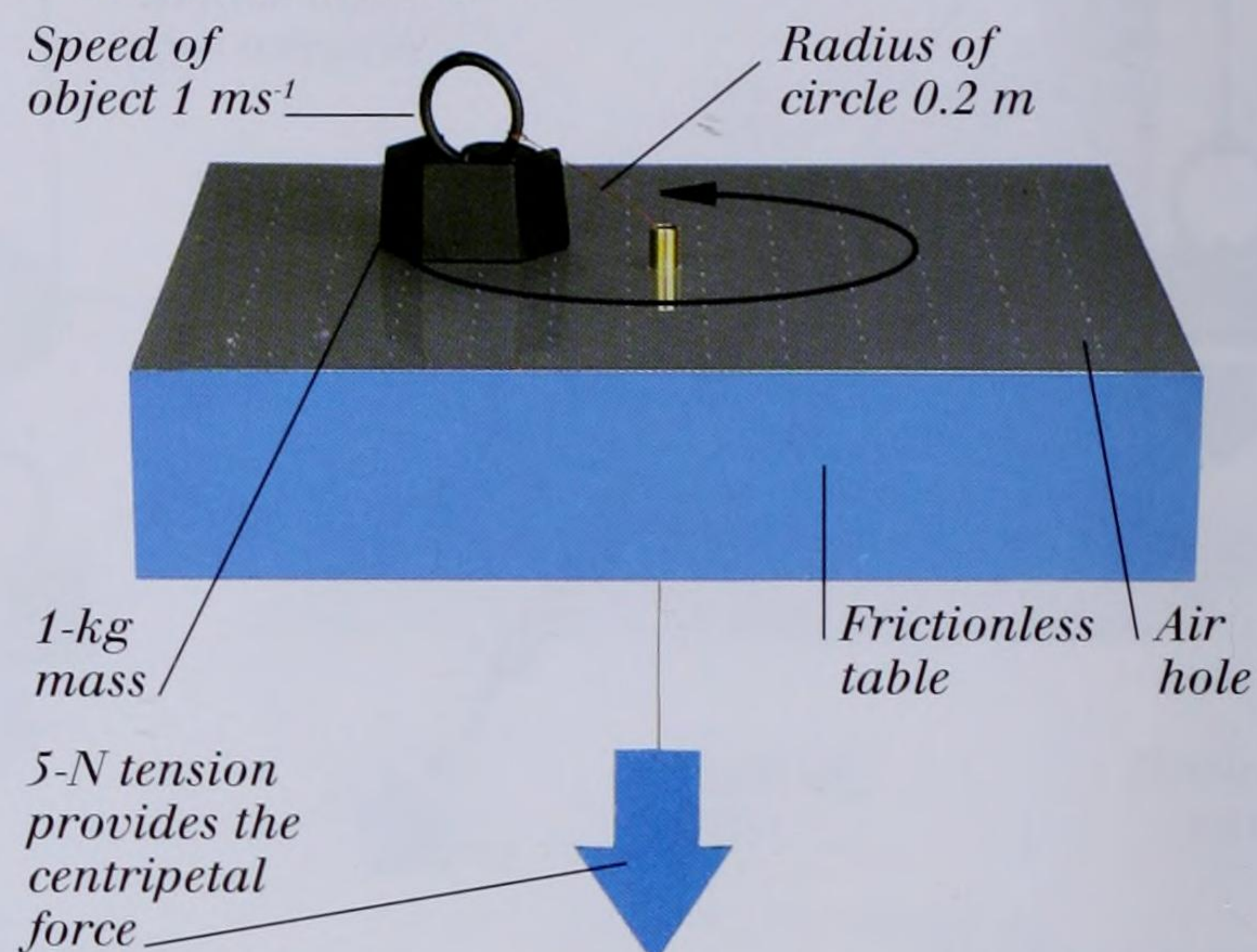
Tension in muscles provides the centripetal force needed to whirl a hammer around in a circle. When the thrower releases the chain, no force acts upon the hammer and it moves off in a straight line.



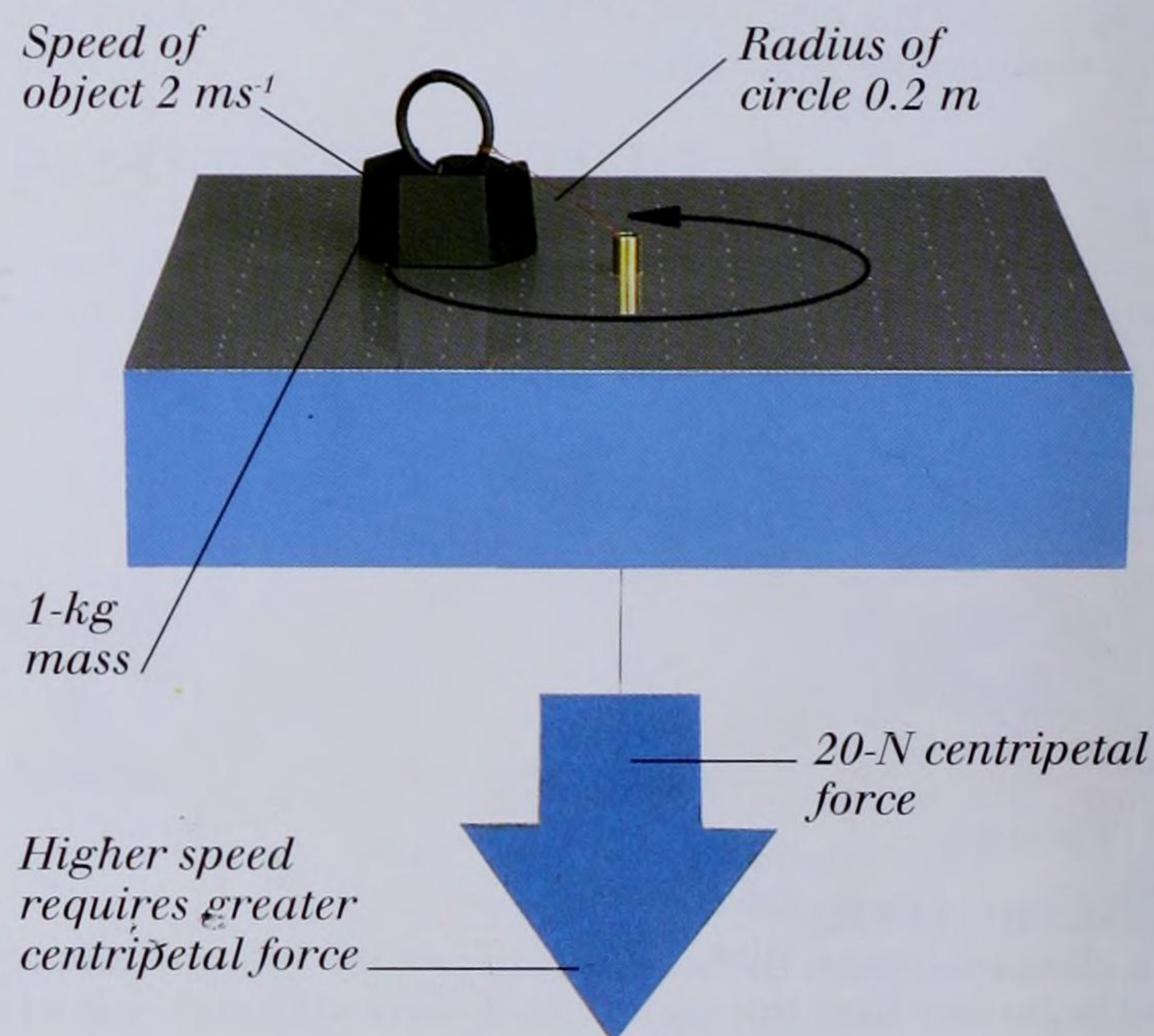
## CENTRIPETAL FORCE

In the experiment below, centripetal force is provided by tension in a length of string, which keeps a 1-kg mass moving in a circle. The mass can move freely as it floats like a hovercraft on the jets of air supplied from beneath it. When the circle is twice as large, half the force is needed. However, moving twice as fast requires four times the force (see p. 54).

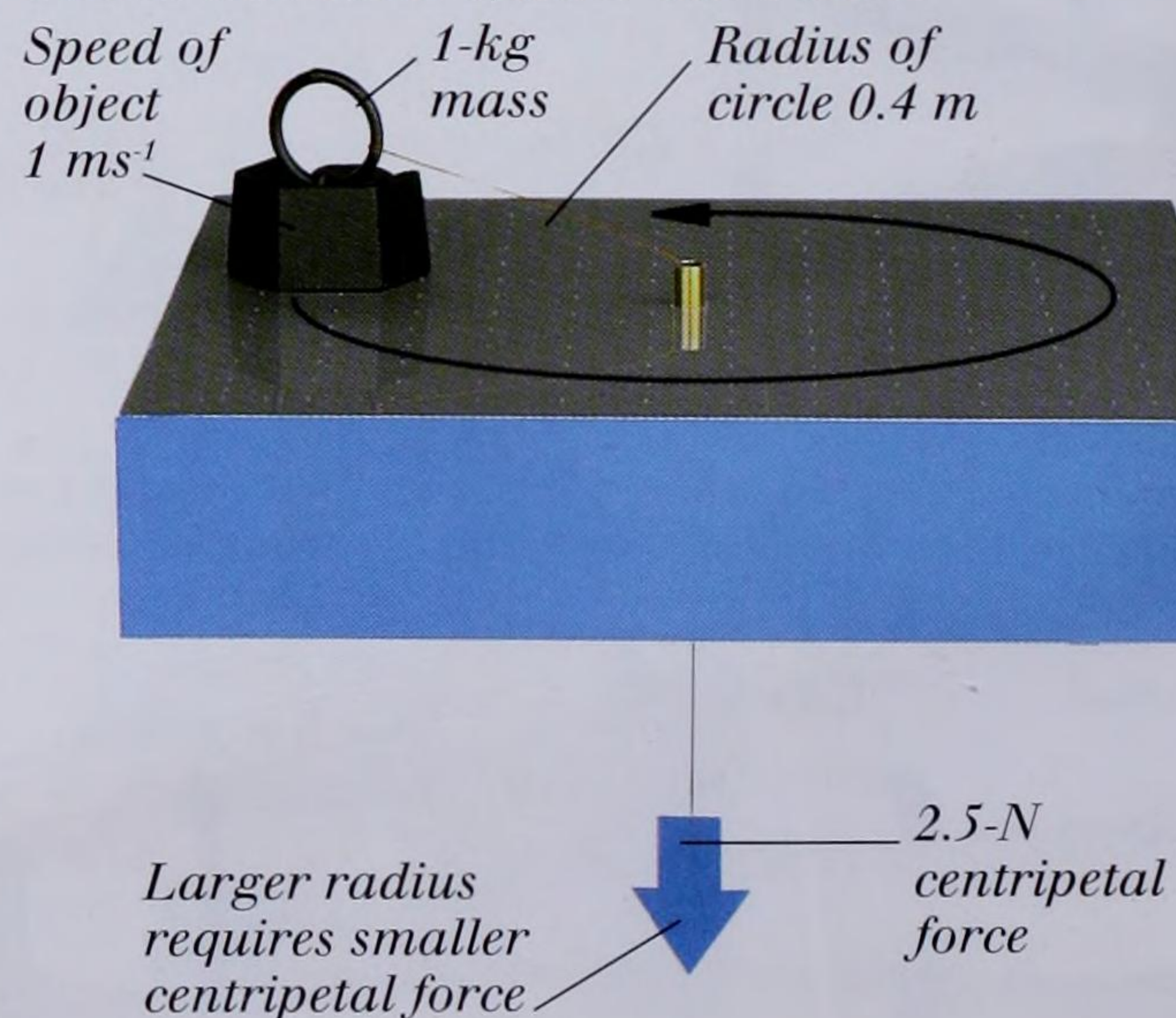
### CONTROL EXPERIMENT



### TWICE THE SPEED, FOUR TIMES THE FORCE



### TWICE THE RADIUS, HALF THE FORCE

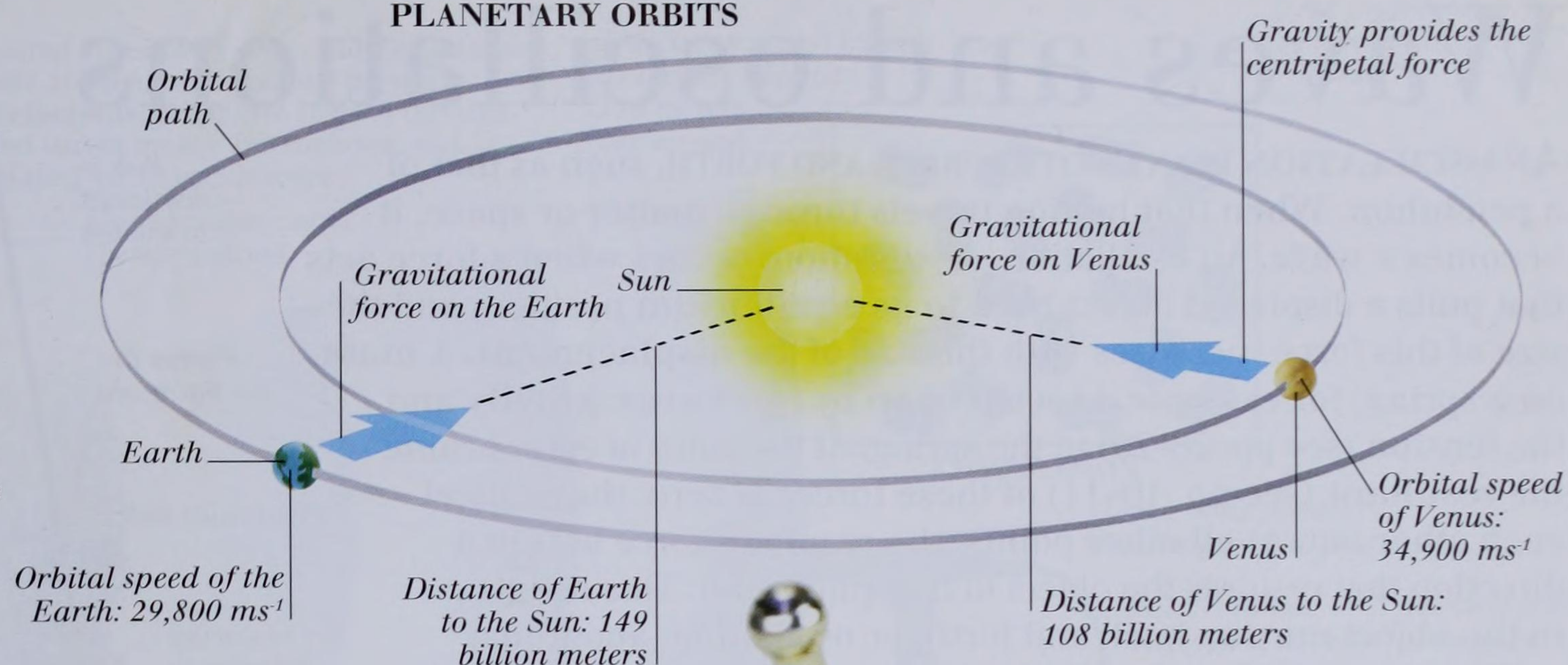




## PLANETARY ORBITS

### GRAVITATIONAL FORCES

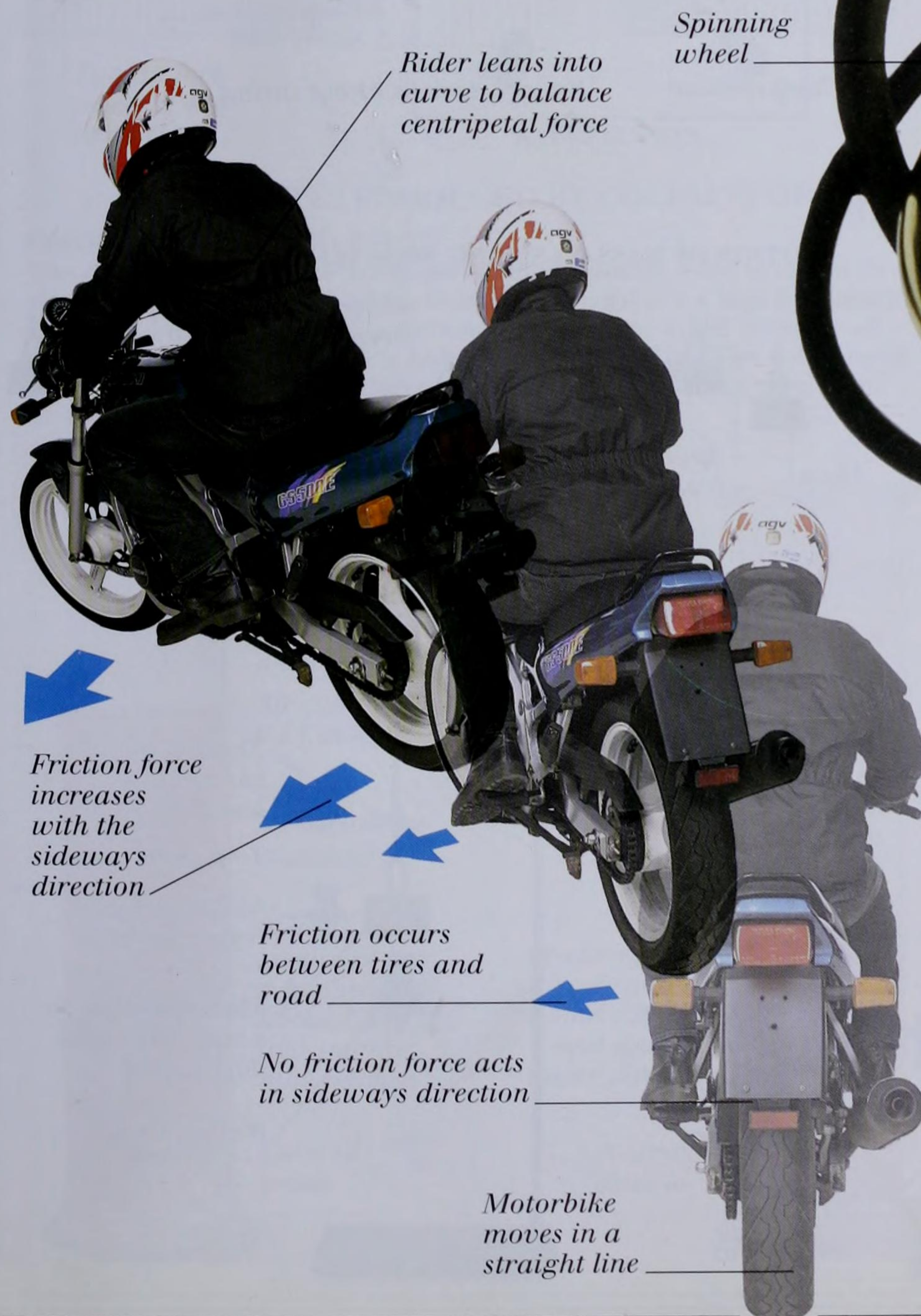
The orbit of a planet around the Sun is an **ellipse** (like a flattened circle). Centripetal force is needed to keep the planets from moving off in a straight line into outer space. Gravity provides this centripetal force. It acts toward the center of the Solar System, the Sun. Venus is roughly the same mass as the Earth, but travels much faster. This is possible because Venus is closer to the Sun, so the force of gravity, and therefore the centripetal force, is much greater (see p. 54).



## TURNING A CORNER

### FRICTION

One of the forces acting on a motorcycle as it turns a corner is the centripetal force caused by the friction between the tires and the road. Without this friction, for example on an icy surface, a motorcycle would simply continue in a straight line.



## GYROSCOPE



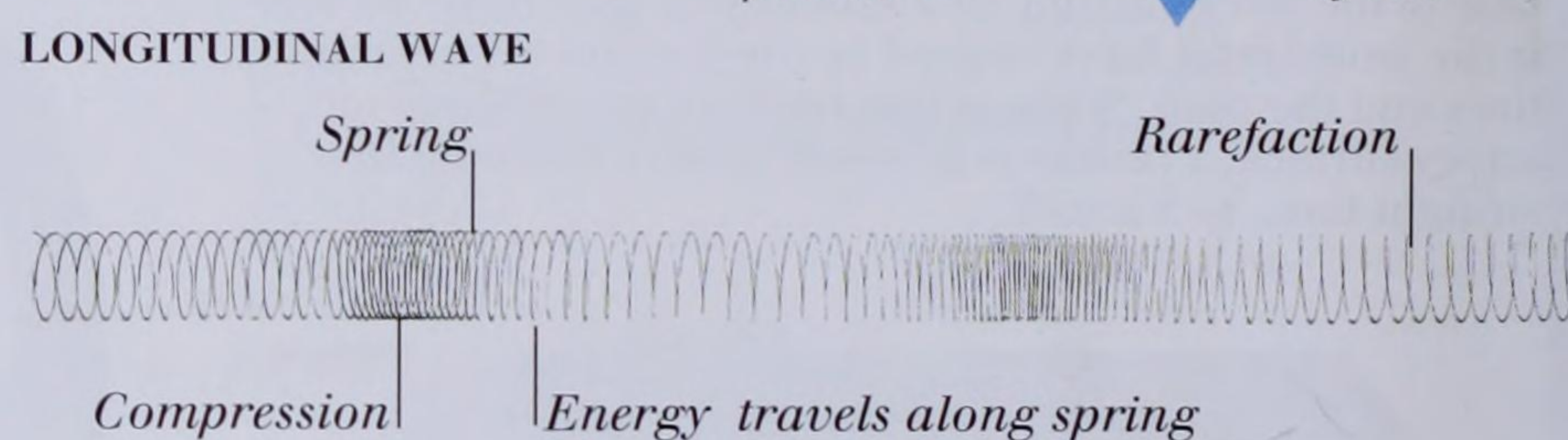
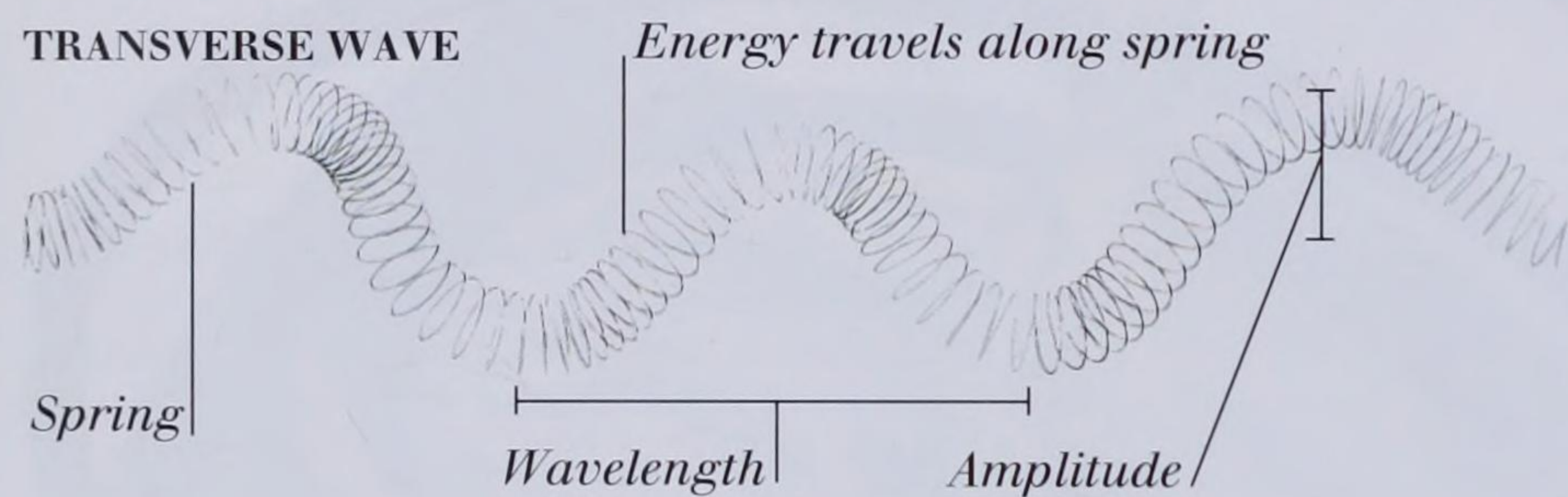
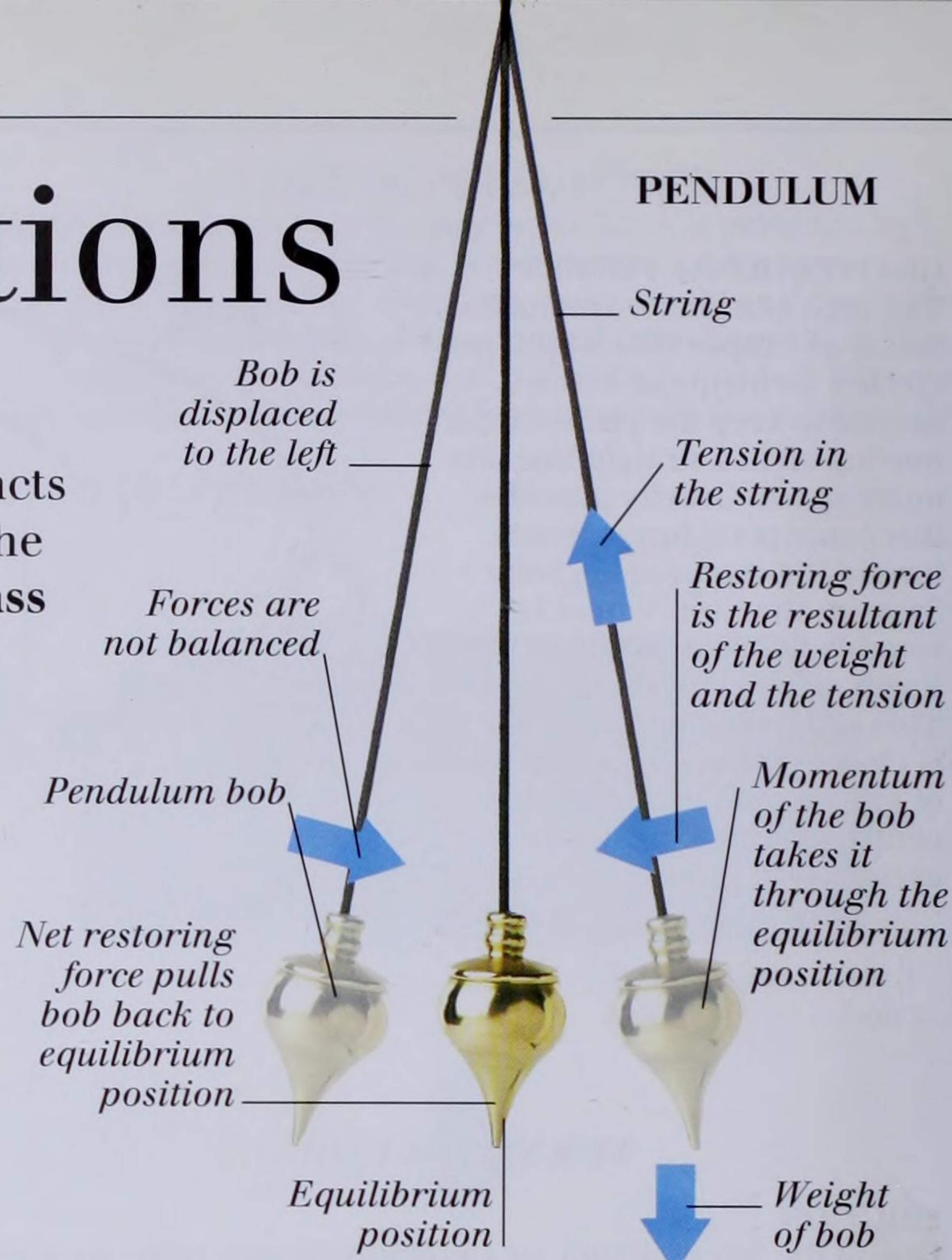
### ANGULAR MOMENTUM

Any spinning object, like a wheel or a top, will behave like a gyroscope. Once spinning, a gyroscope possesses **angular momentum**. This gives the gyroscope stability. The force of gravity acting on the gyroscope will not topple it. As gravity tries to tilt the axis, its axis moves at right angles to gravity's force. This causes a motion called **precession**, in which the axis traces a small circle.



# Waves and oscillations

AN OSCILLATION IS ANY MOTION BACK AND FORTH, such as that of a pendulum. When that motion travels through matter or space, it becomes a wave. An oscillation, or vibration, occurs when a force acts that pulls a displaced object back to its **equilibrium** position, and the size of this force increases with the size of the **displacement**. A mass on a spring, for example, is acted upon by two forces: **gravity** and the **tension** (see pp. 28–29) in the spring. At the point of equilibrium, the **resultant** (see pp. 10–11) of these forces is zero: they cancel each other out. At all other points, the resultant force acts in a direction that restores the object to its equilibrium. This results in the object moving back and forth, or oscillating, about that position. Vibration is very common, and results in the phenomenon of sound. In air, the vibrations that cause sound are transmitted as a wave between air molecules; many other substances transmit sound in a similar way.

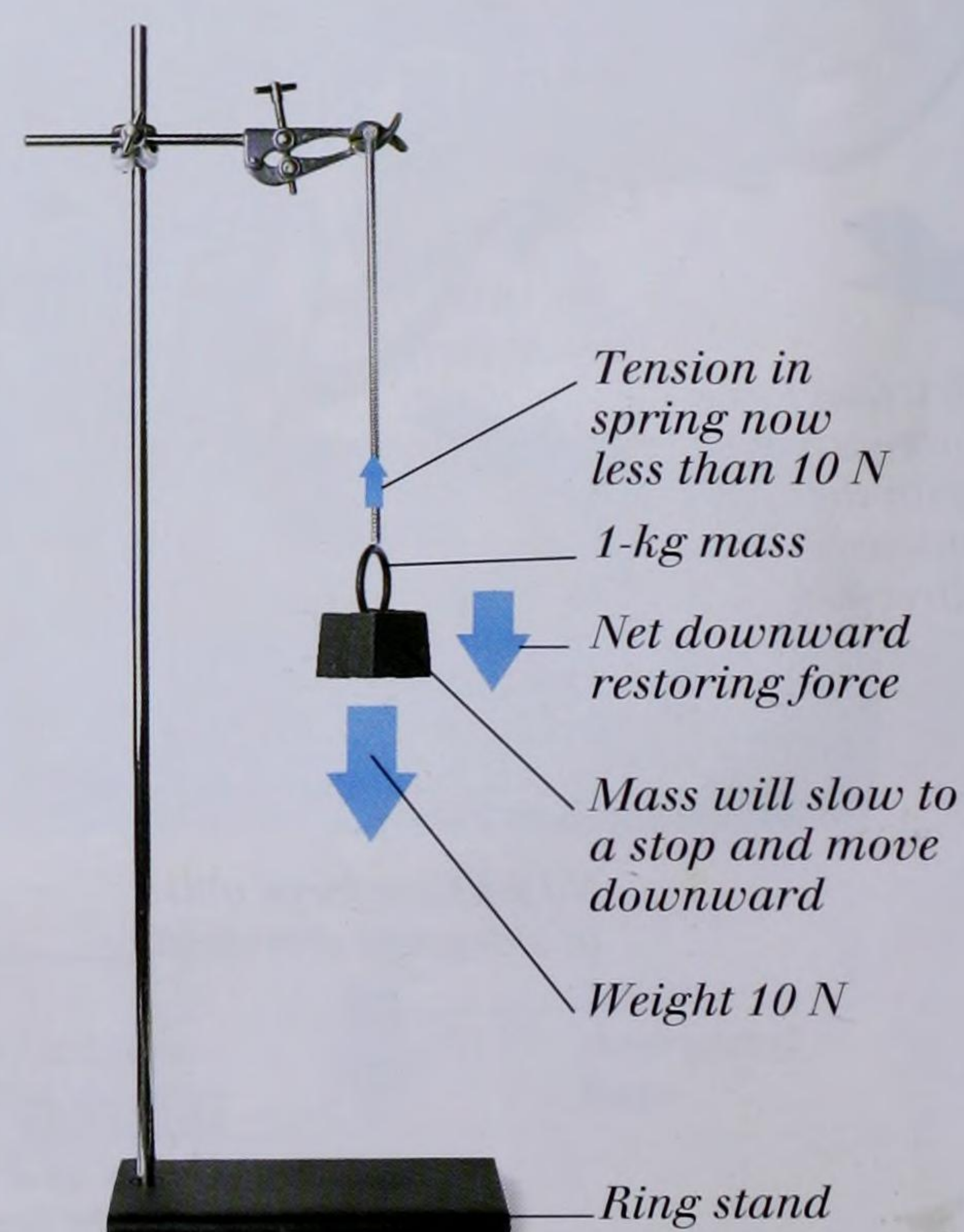
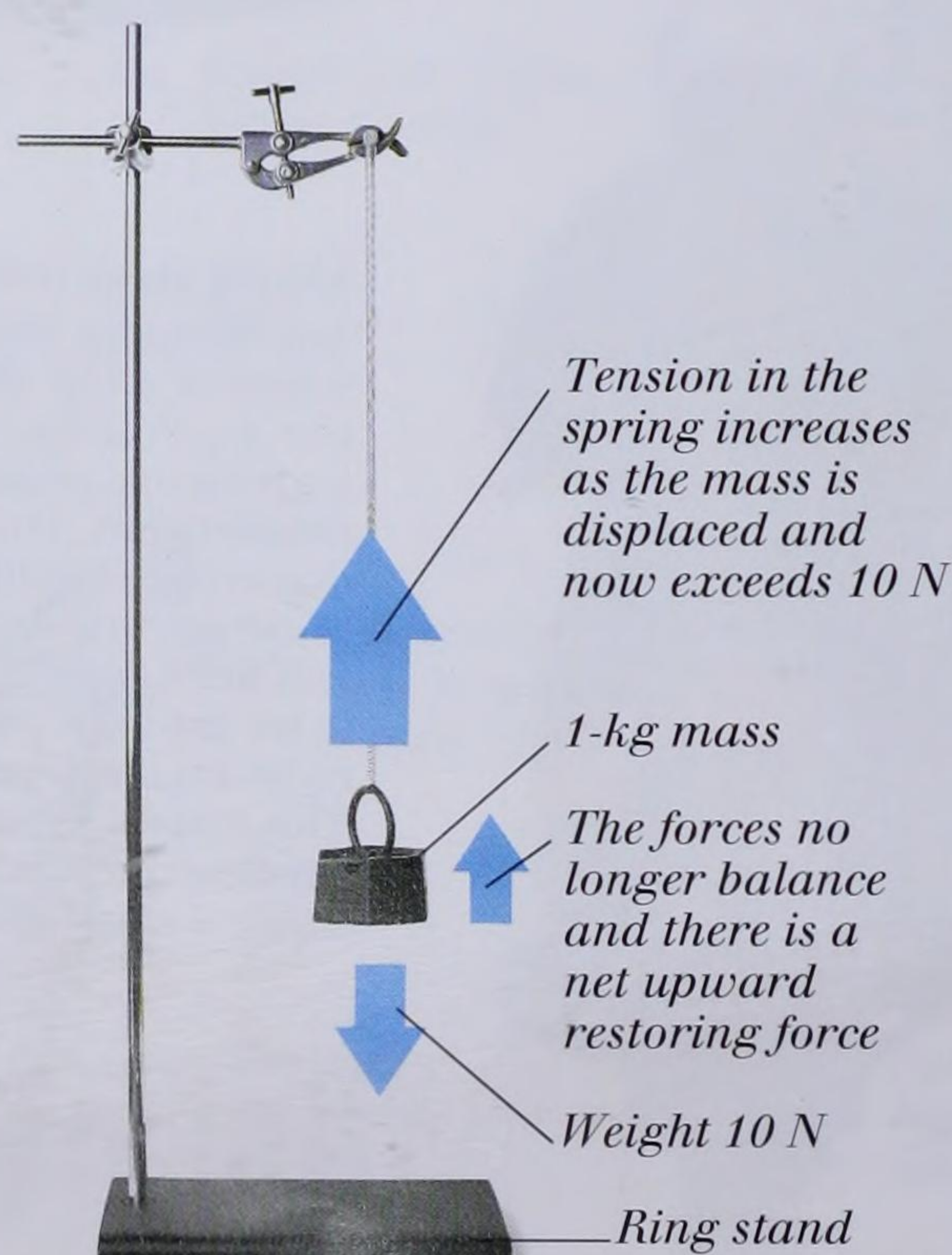
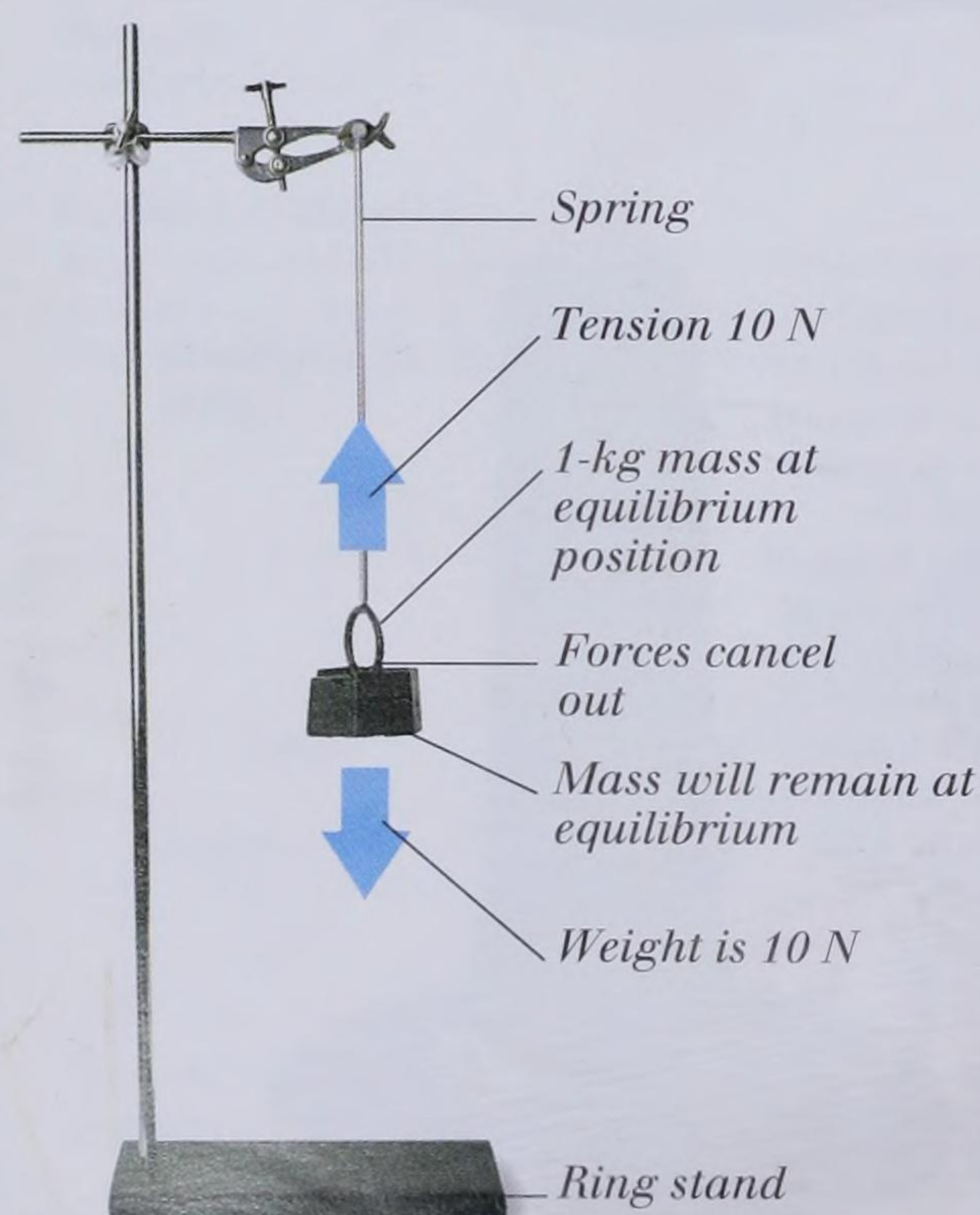
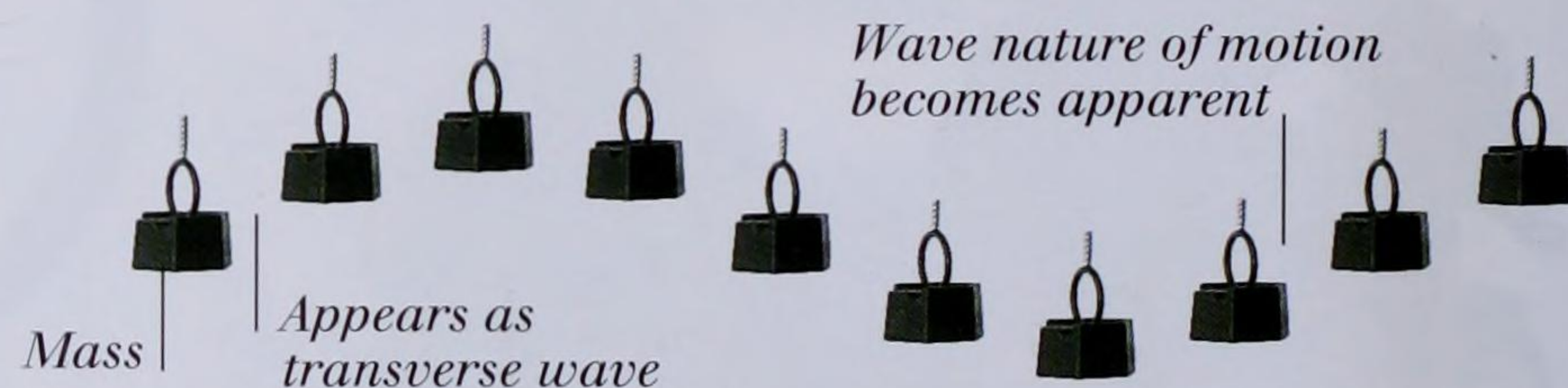


## OSCILLATION

### MOTION OF MASS ON SPRING

The first mass shown (below left) is in equilibrium. The two forces acting on it—its **weight** and the tension in the spring—exactly cancel each other out. The mass is given an initial downward push. Once the mass is displaced downward (below center), the tension in the spring exceeds the weight. The resultant upward force accelerates the mass back up toward its original position, by which time it has momentum, carrying it farther upward. When the weight exceeds the tension in the spring (below right), the mass is pulled down again. This cycle repeats.

### MOTION OF MASS ON SPRING, MASS SEEN IN ISOLATION

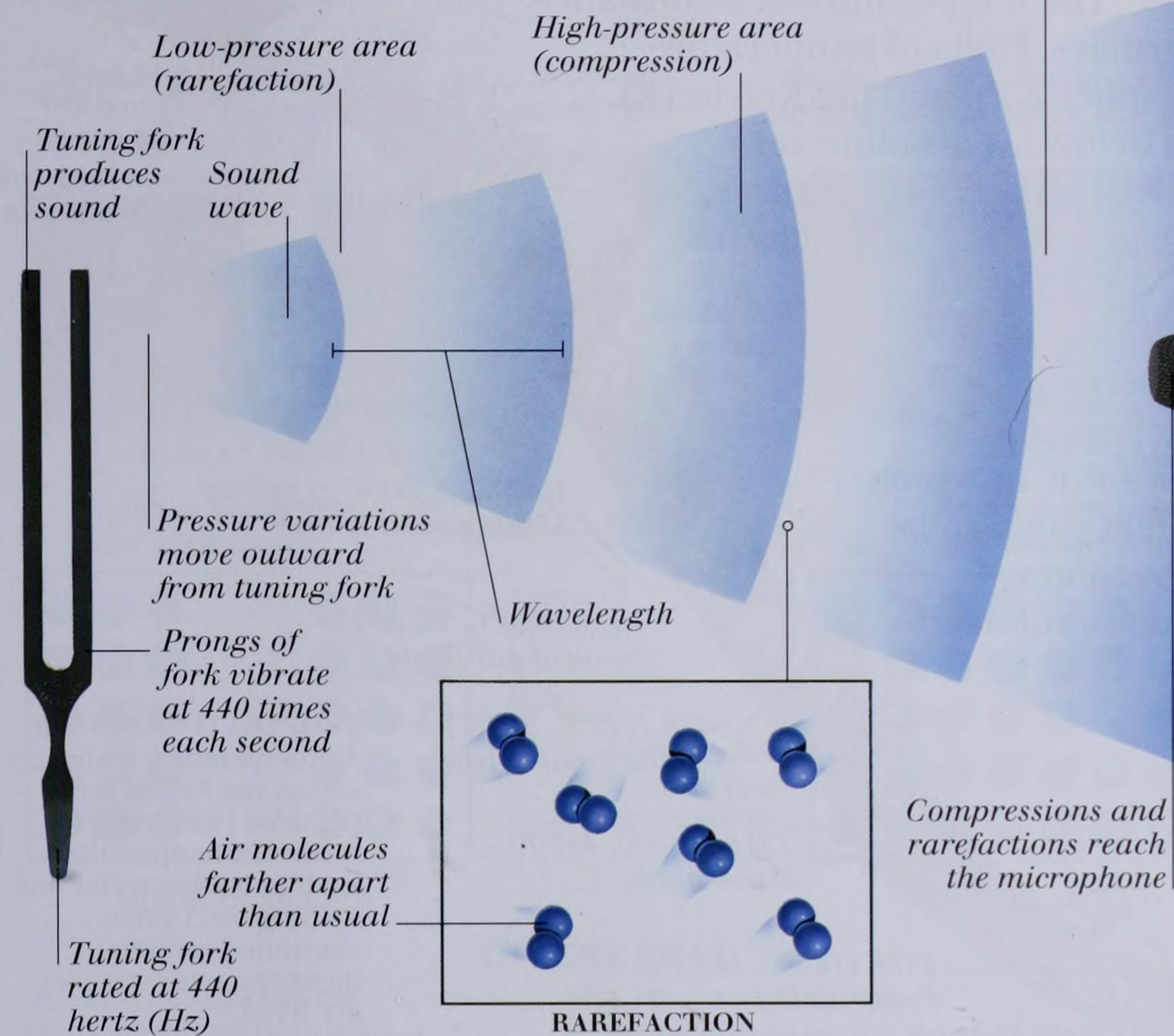




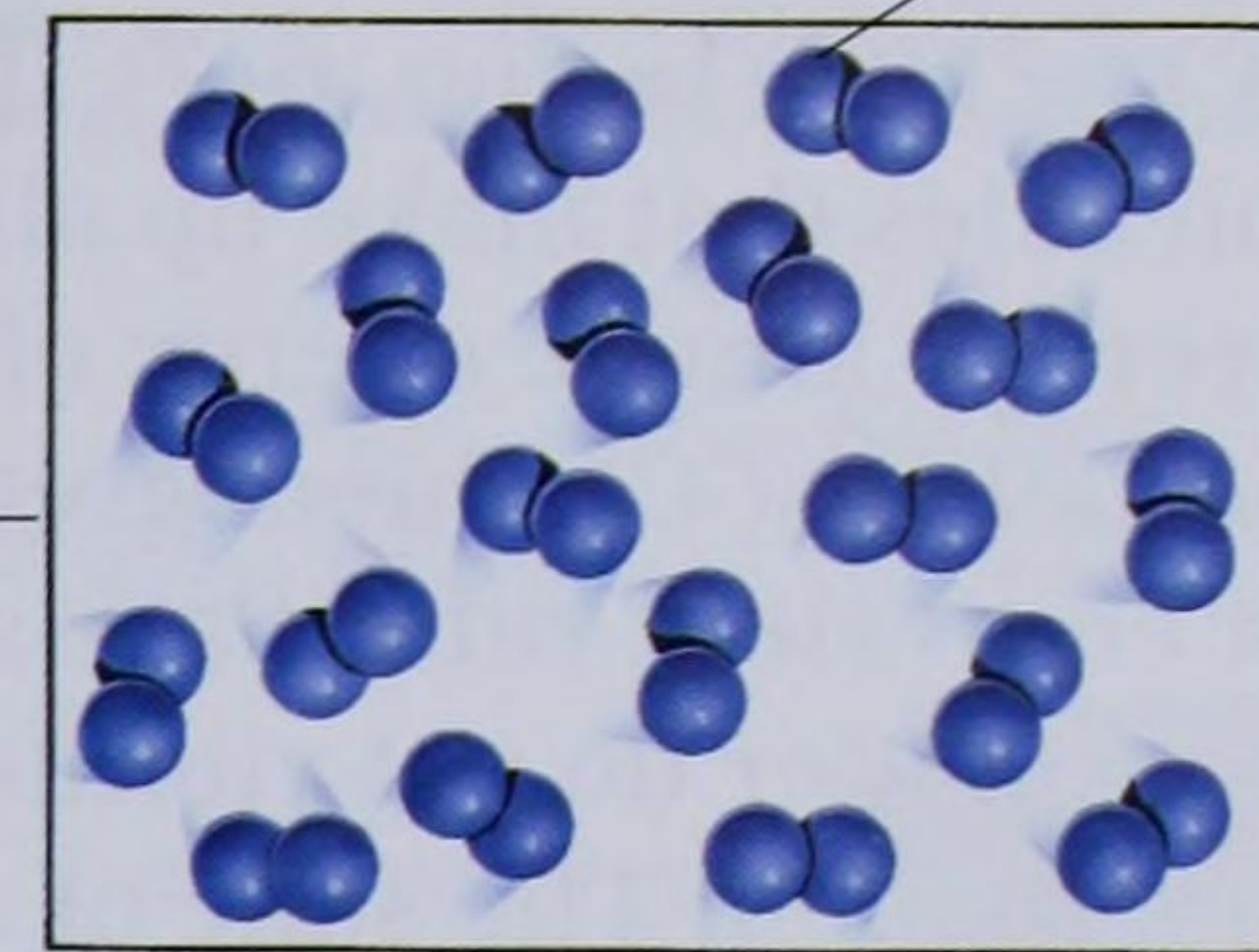
## SOUND AS VIBRATION OF THE AIR

### PROPAGATION OF SOUND

A vibrating object, such as the tuning fork shown here, causes variations in pressure in the surrounding air. Areas of high and low **pressure**, known as **compressions** and **rarefactions**, propagate (move) through the air as sound waves. The sound waves meet a microphone, and create electrical oscillations displayed on an oscilloscope.

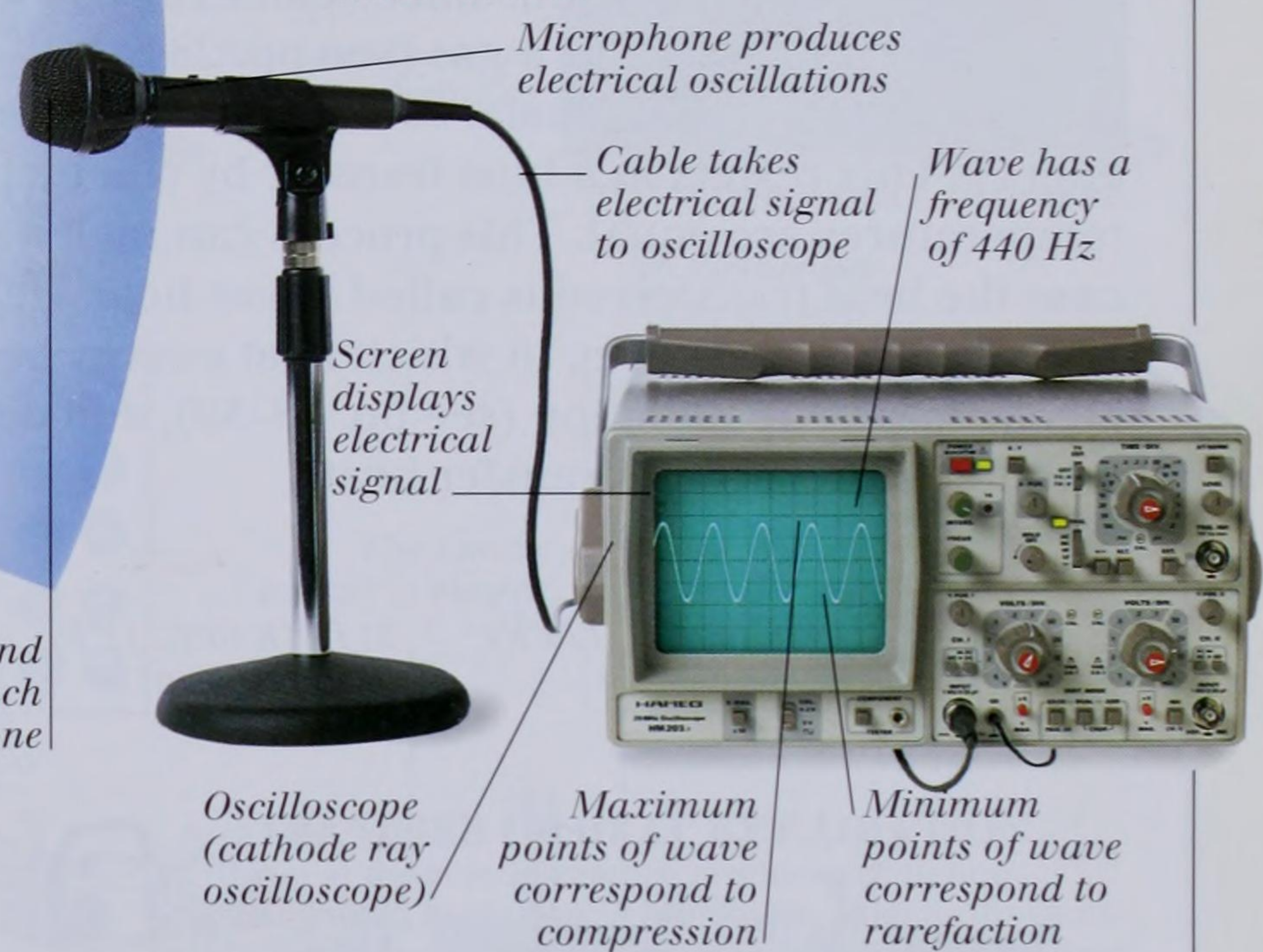


The compression travels as a wave at about 330 meters per second



COMPRESSION

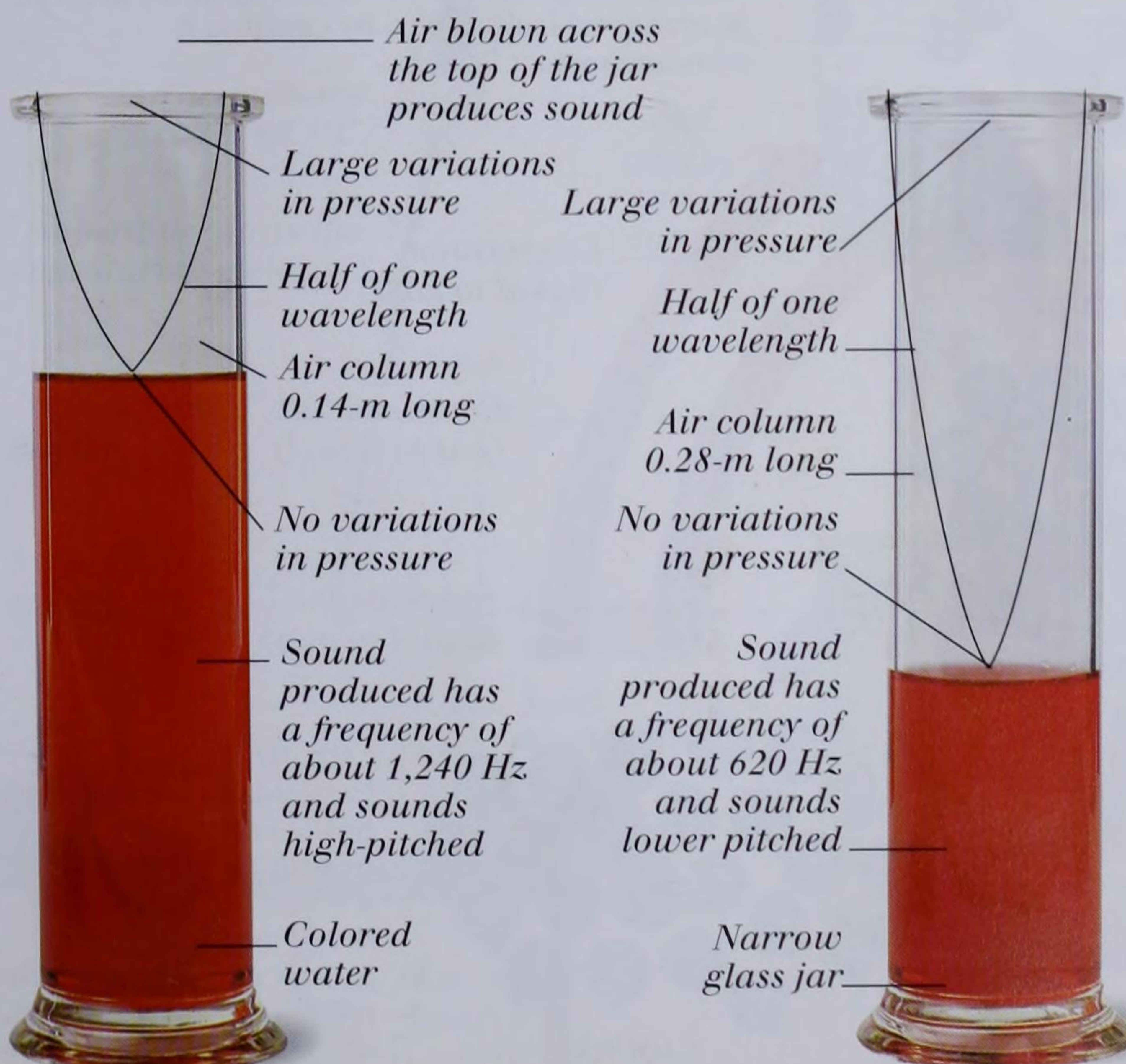
Air molecules closer than usual



### NOTES PRODUCED BY COLUMNS OF AIR

#### FREQUENCY AND WAVELENGTH

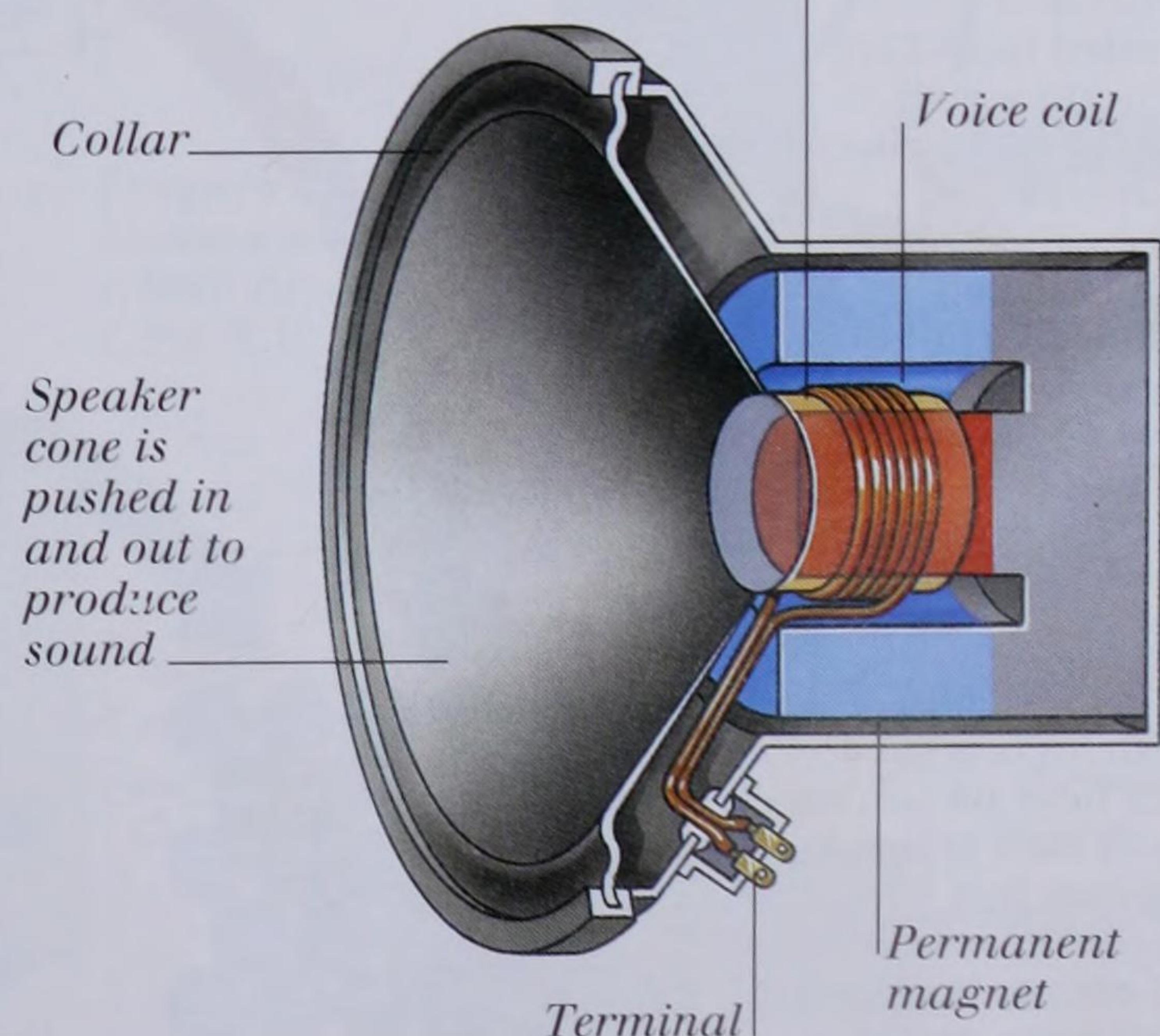
The distance between each compression of a sound wave is called its **wavelength**. Sound waves with a short wavelength have a high **frequency** and sound high-pitched. The frequency of a note is the number of vibrations each second, and is measured in hertz (Hz). The columns of air in these jars produce different notes when air is blown over them.



### LOUDSPEAKER

A changing electrical signal is fed to the voice coil of a loudspeaker, which lies within the **magnetic field** of a **permanent magnet**. The signal in the coil causes it to behave like an **electromagnet** (see pp. 34-35), making it push against the field of the permanent magnet. The speaker cone is then pushed in and out by the coil in time with the signal.

If the signal is from a sound recording the original sound will be reproduced



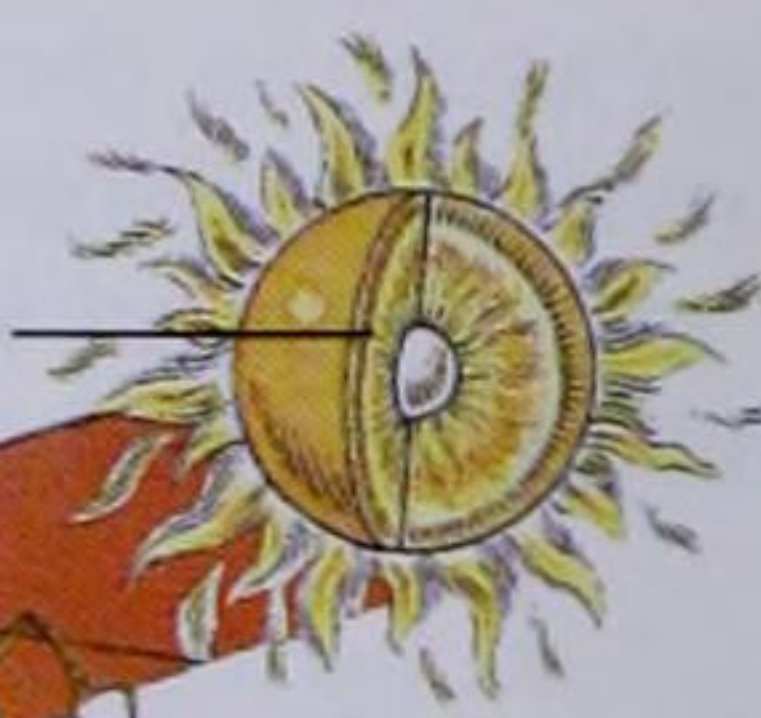


# Heat and temperature

HEAT IS A FORM OF ENERGY (see pp. 6–7). This energy is the kinetic energy of the atoms and molecules that make up all matter. The temperature of a substance is related to the average kinetic energy of its particles. Units of temperature include the degree Celsius ( $^{\circ}\text{C}$ ), the degree Fahrenheit ( $^{\circ}\text{F}$ ), and the Kelvin ( $\text{K}$ ). Some examples of equivalent values are shown below. At absolute zero (zero  $\text{K}$ ), particles of matter do not vibrate, but at all other temperatures, particles have some motion. The state of a substance is determined by its temperature and most substances can exist as a solid (see pp. 24–25), a liquid (see pp. 26–27), or a gas (see pp. 28–29). If two substances at different temperatures make contact, their particles will share their energy. This results in a heat transfer by conduction, until the temperatures are equal. This process can melt a solid, in which case the heat transferred is called latent heat. Heat can also be transferred by radiation, in which heat energy becomes electromagnetic radiation (see pp. 38–39), and does not need a material medium to transfer heat.

## RANGE OF TEMPERATURES

About 14 million  $\text{K}$   
(14 million  $^{\circ}\text{C}$ ,  
25 million  $^{\circ}\text{F}$ ):  
Center of the Sun



30,000 $\text{K}$   
(30,000 $^{\circ}\text{C}$ ,  
54,000 $^{\circ}\text{F}$ ):  
Average bolt  
of lightning

5,800 $\text{K}$  (5,530 $^{\circ}\text{C}$ , 10,000 $^{\circ}\text{F}$ ):  
Surface of the Sun

3,300 $\text{K}$  (3,027 $^{\circ}\text{C}$ ; 5,480 $^{\circ}\text{F}$ ):  
Metals can be welded

1,808 $\text{K}$  (1,535 $^{\circ}\text{C}$ , 2,795 $^{\circ}\text{F}$ ):  
Melting point of iron

933 $\text{K}$  (660 $^{\circ}\text{C}$ , 1,220 $^{\circ}\text{F}$ ):  
Natural gas flame

600 $\text{K}$  (327 $^{\circ}\text{C}$ , 620 $^{\circ}\text{F}$ ):  
Melting point of lead

523 $\text{K}$  (250 $^{\circ}\text{C}$ , 482 $^{\circ}\text{F}$ ):  
Wood burns

457 $\text{K}$  (184 $^{\circ}\text{C}$ , 363 $^{\circ}\text{F}$ ):  
Paper ignites

373.15 $\text{K}$  (100 $^{\circ}\text{C}$ , 212 $^{\circ}\text{F}$ ):  
Boiling point of water

331 $\text{K}$  (58 $^{\circ}\text{C}$ , 136 $^{\circ}\text{F}$ ):  
Earth's highest  
temperature

184 $\text{K}$  (-89 $^{\circ}\text{C}$ , -128 $^{\circ}\text{F}$ ):  
Earth's lowest  
temperature

73 $\text{K}$  (-200 $^{\circ}\text{C}$ ,  
-328 $^{\circ}\text{F}$ ): Air  
liquifies

0 $\text{K}$  (-273.15 $^{\circ}\text{C}$ ,  
-459.67 $^{\circ}\text{F}$ ):  
Absolute zero

234 $\text{K}$  (-39 $^{\circ}\text{C}$ ,  
-38.2 $^{\circ}\text{F}$ ):  
Freezing point  
of mercury

273.15 $\text{K}$  (0 $^{\circ}\text{C}$ , 32 $^{\circ}\text{F}$ ):  
Freezing point of  
water

## TEMPERATURE SCALES

All temperature scales except the Kelvin scale ( $\text{K}$ ) need two or more reference temperatures, such as boiling water and melting ice. Under controlled conditions, these two temperatures are fixed.

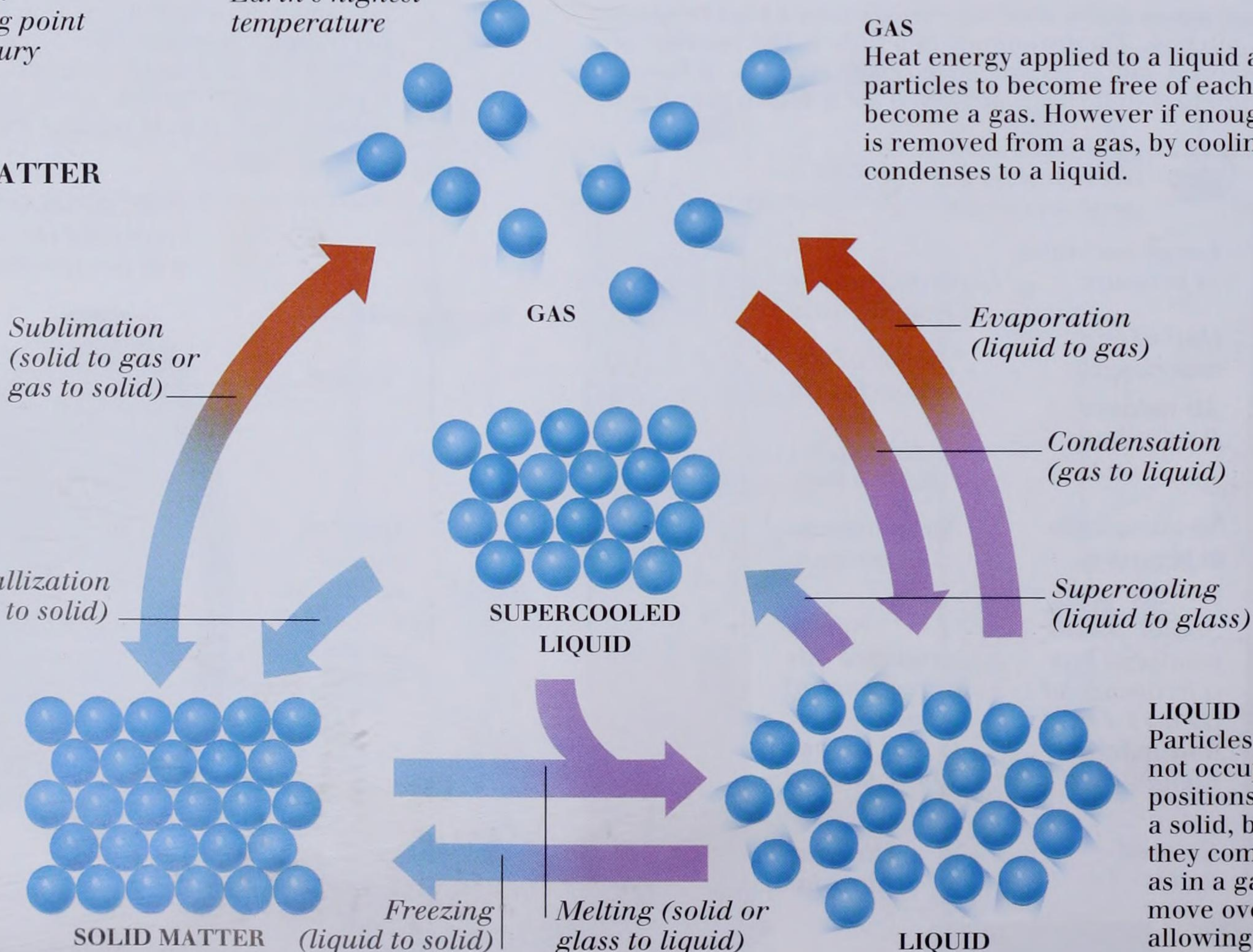
## STATES OF MATTER

### SUPERCOOLED LIQUID

The particles of a **supercooled liquid** are in fixed positions, like those of a solid, but they are disordered and cannot be called a true solid. Supercooled liquids flow like liquids, but very slowly, and they have no definite melting point.

### SOLID

The particles of a solid normally have no motion relative to each other, as they are only free to vibrate about a fixed position. An input of energy breaks the bonds between particles, and the solid melts.



### GAS

Heat energy applied to a liquid allows particles to become free of each other and become a gas. However if enough energy is removed from a gas, by cooling, it condenses to a liquid.

### LIQUID

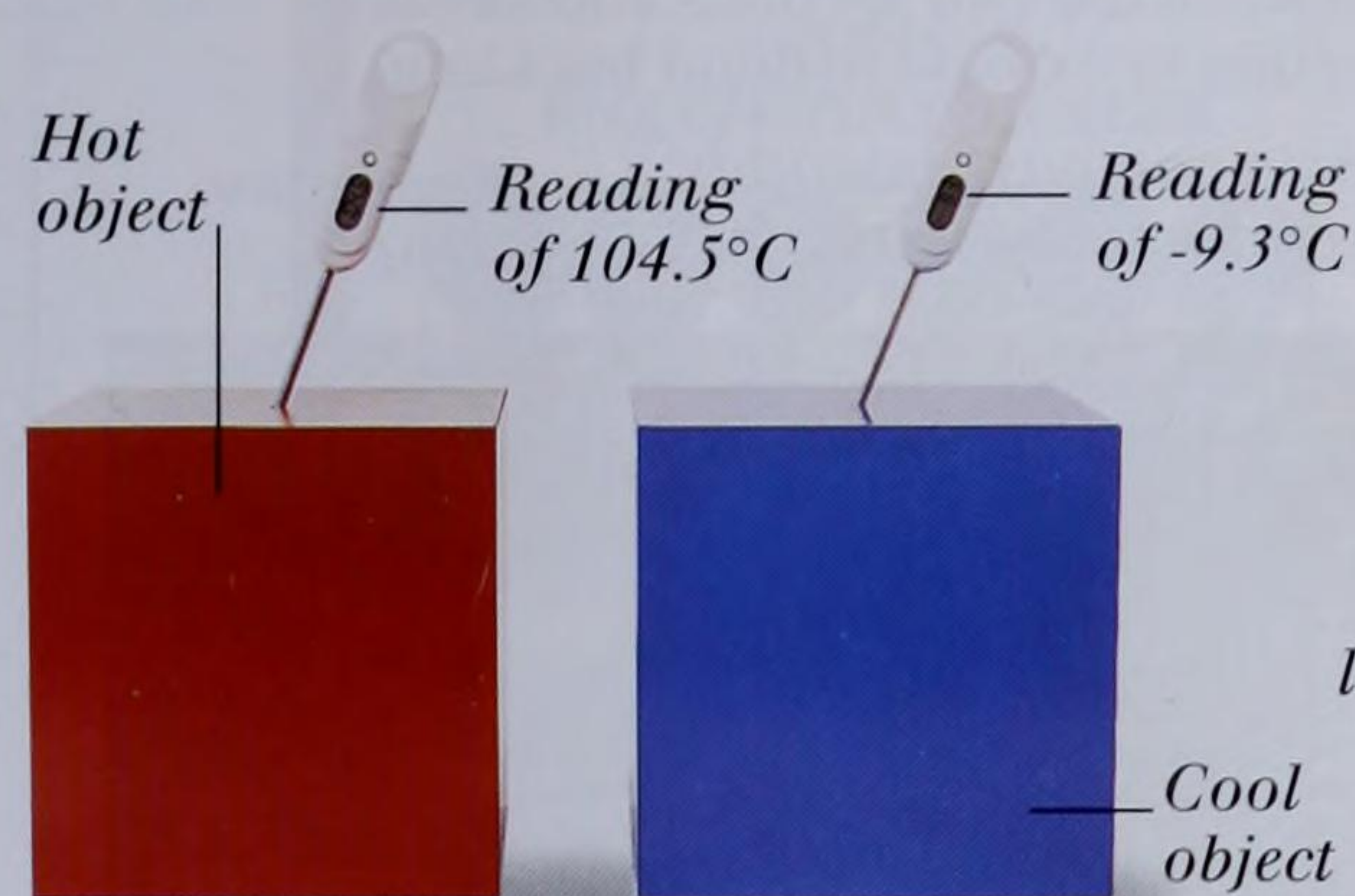
Particles in a liquid do not occupy fixed positions like those in a solid, but neither are they completely free, as in a gas. The particles move over one another, allowing a liquid to flow.



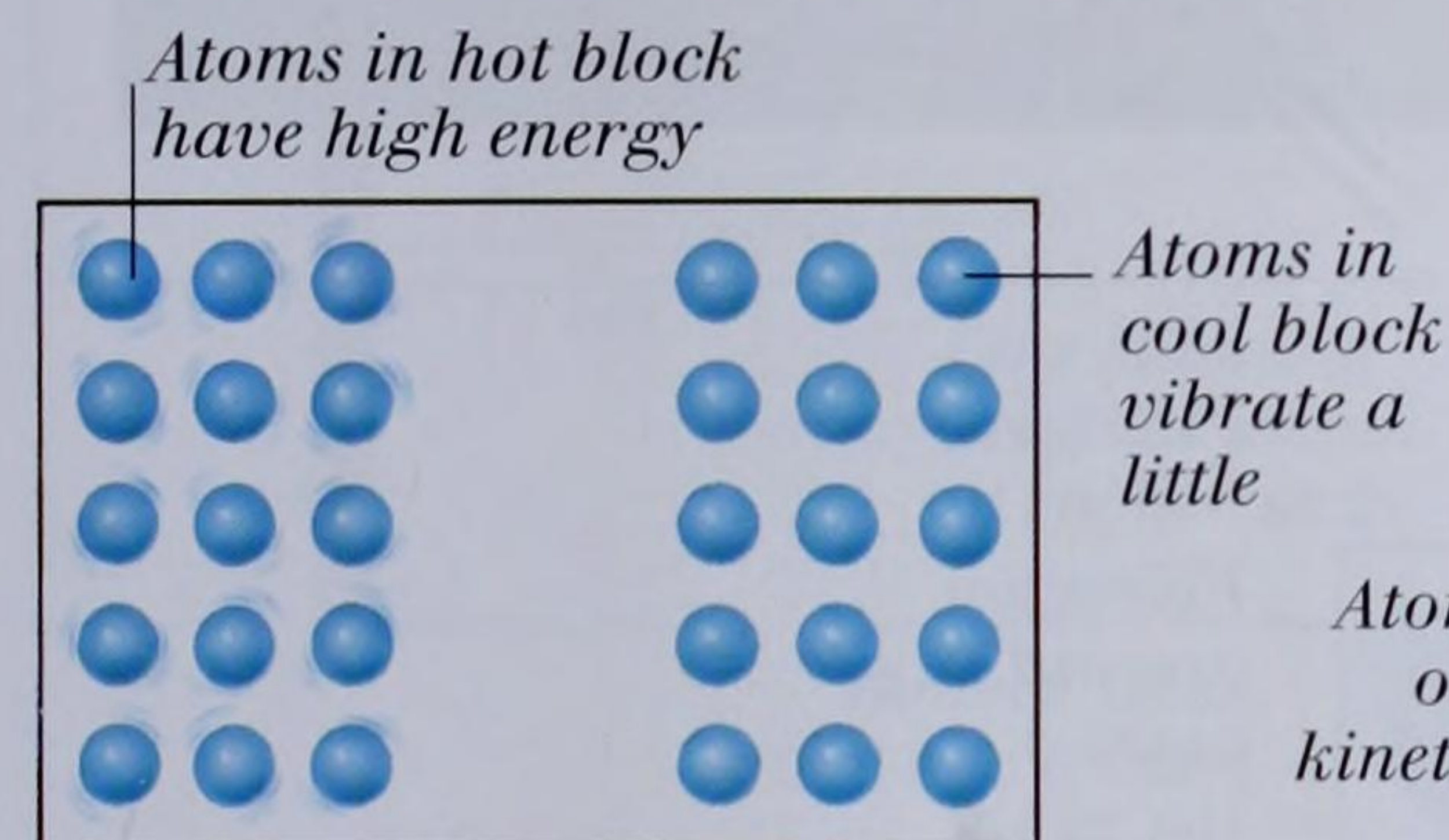
## EQUALIZATION OF TEMPERATURES

### OBJECTS AT DIFFERENT TEMPERATURES

The particles of objects at different temperatures have different kinetic energies. The colors of the blocks below are an indication of their temperature.



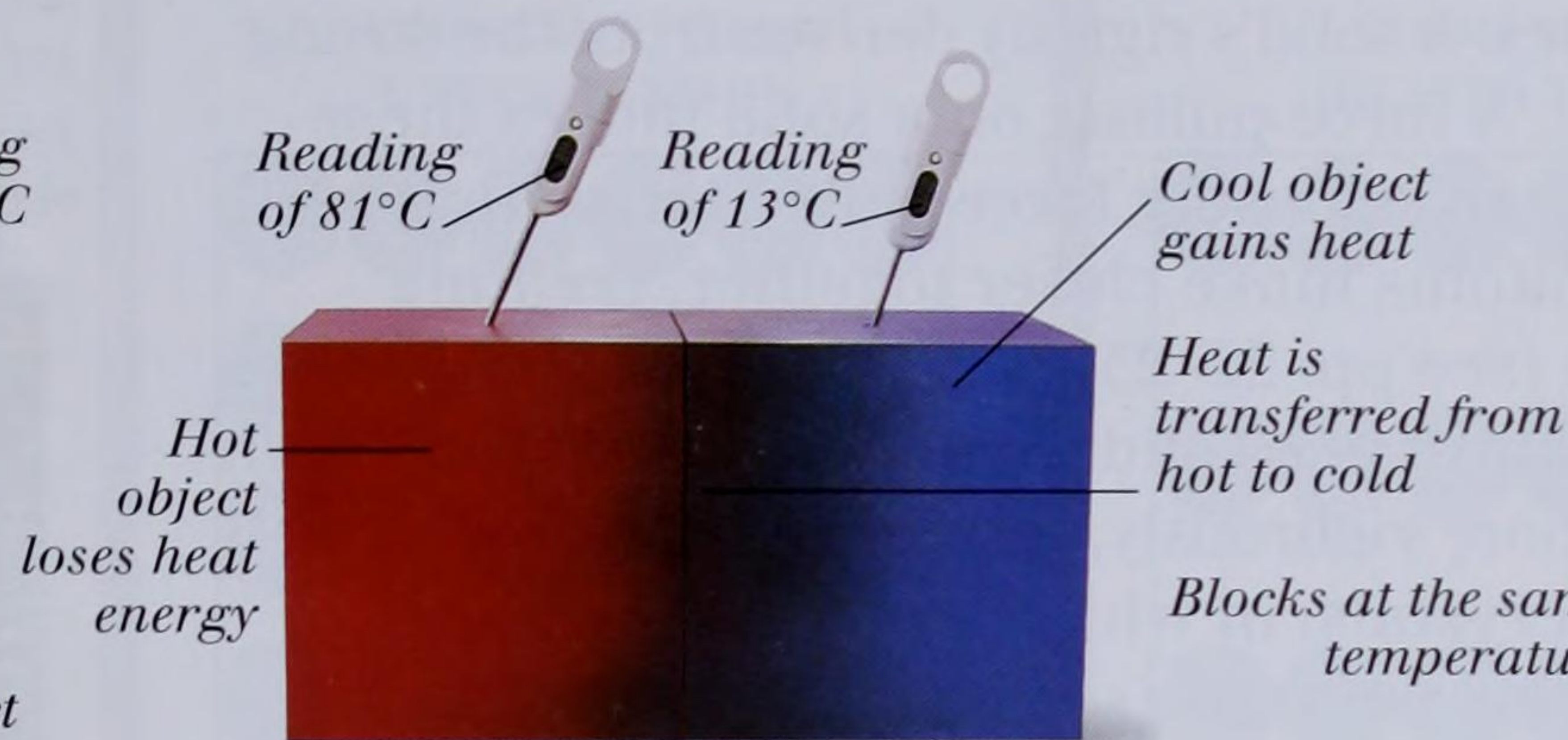
BLOCKS SEPARATED



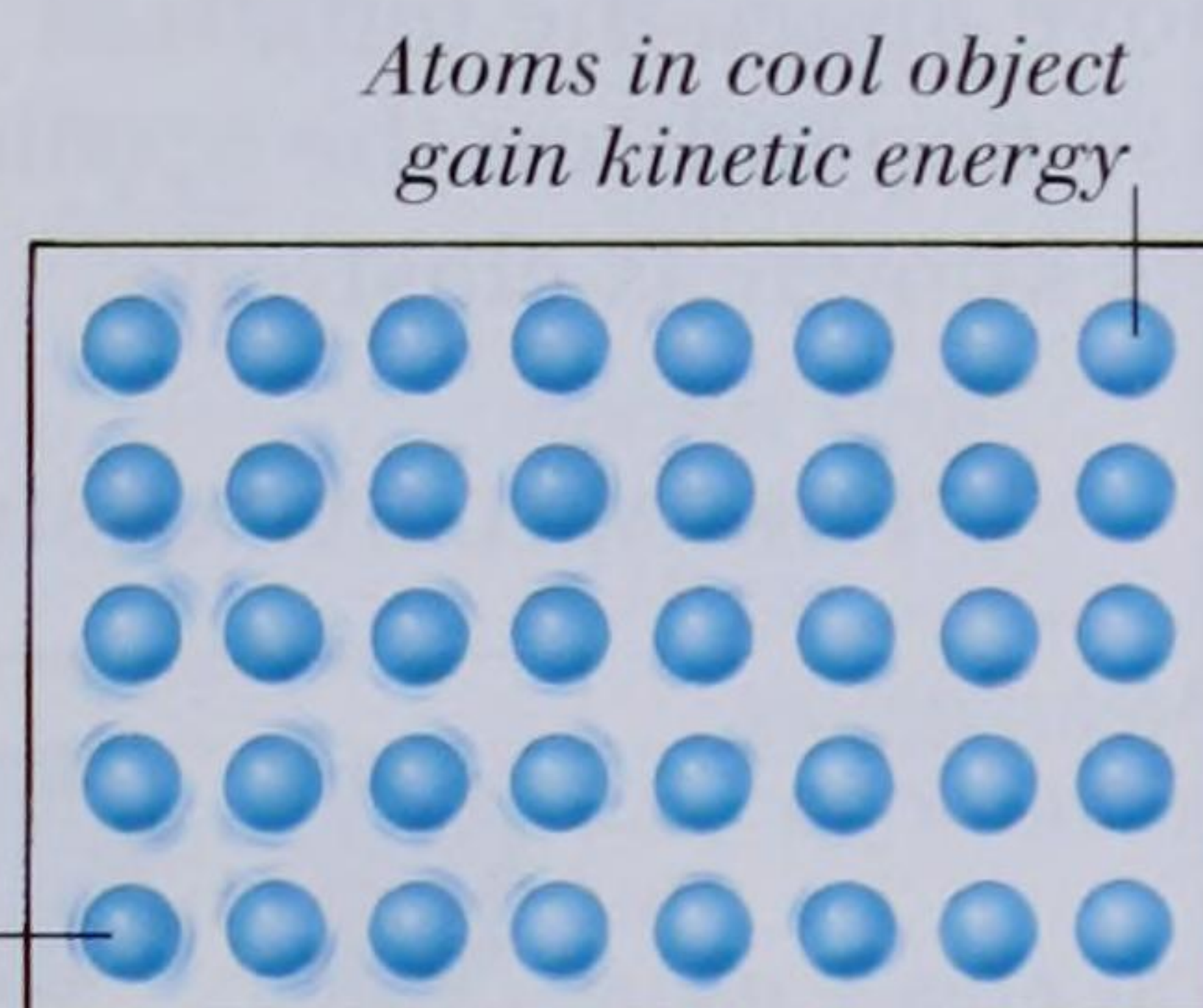
MOLECULAR VIEW

### TRANSFER OF HEAT

When two objects at different temperatures are brought into contact, a transfer of kinetic energy—in the form of heat—takes place. Here, the hot and cold blocks are touching.

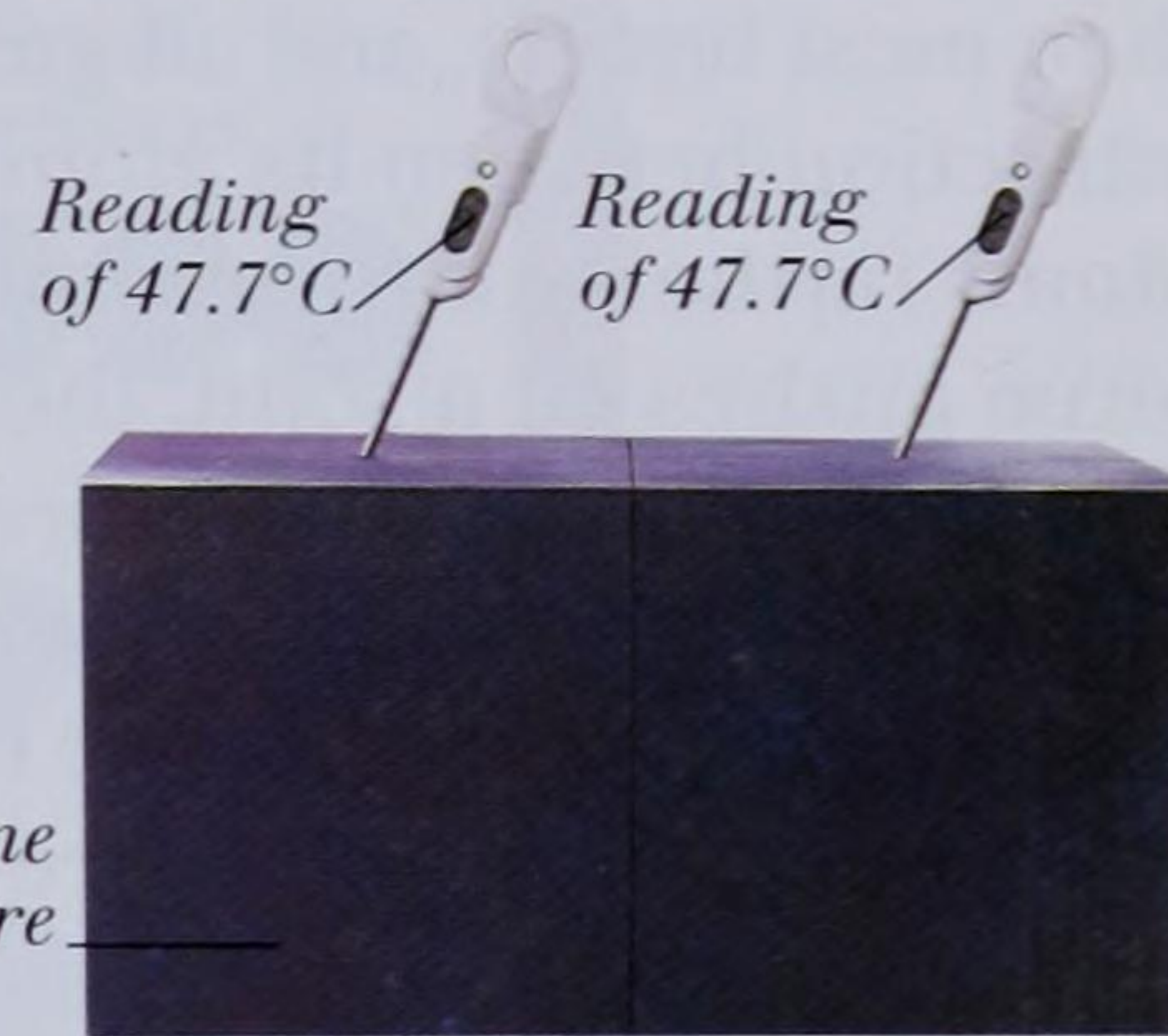


BLOCKS IN CONTACT

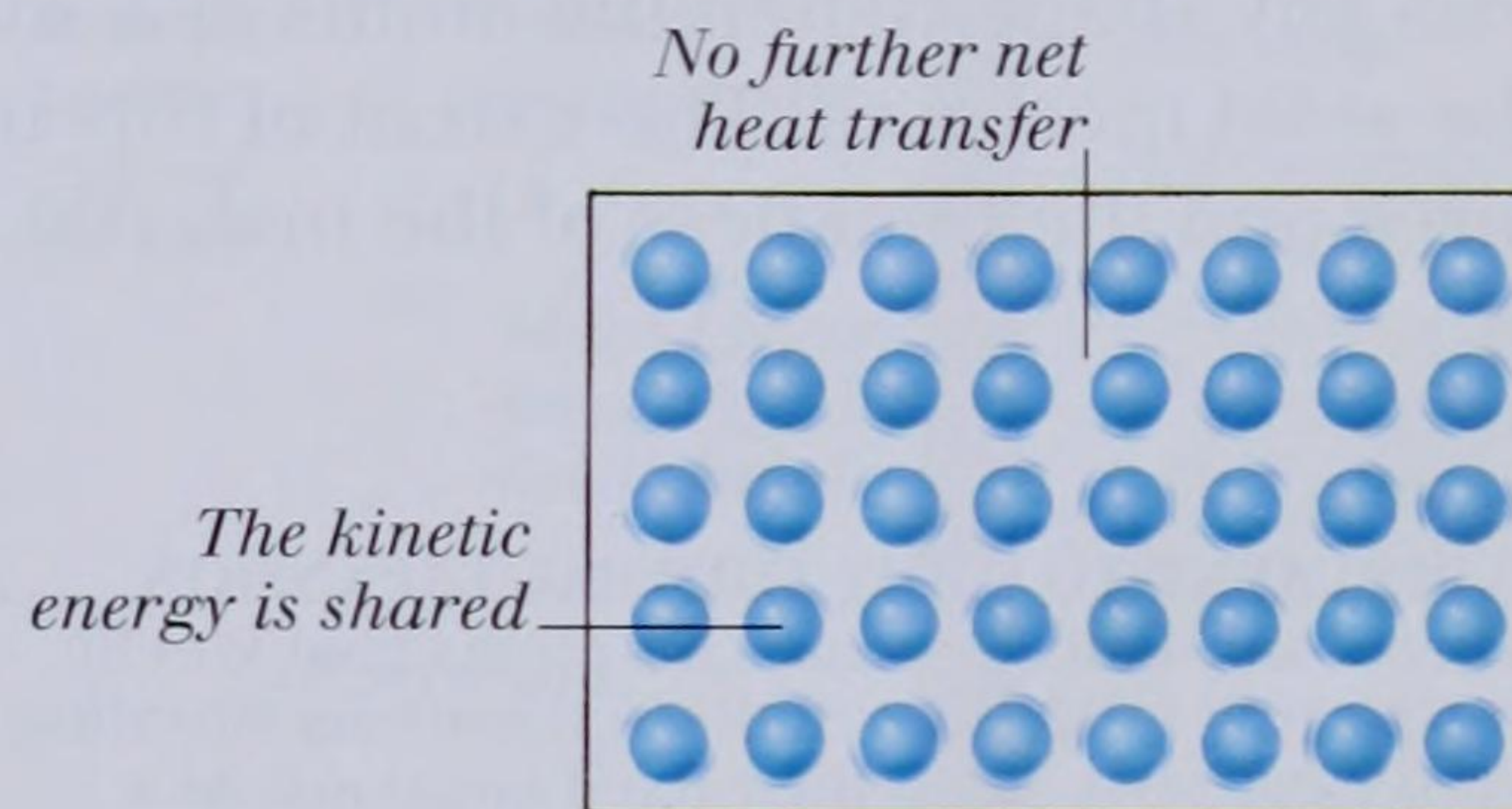


MOLECULAR VIEW

Eventually, the average kinetic energies of particles in two touching objects become equal. The temperatures of the two objects are then said to be equal, as shown by the blocks below.



NO FURTHER HEAT TRANSFER



MOLECULAR VIEW

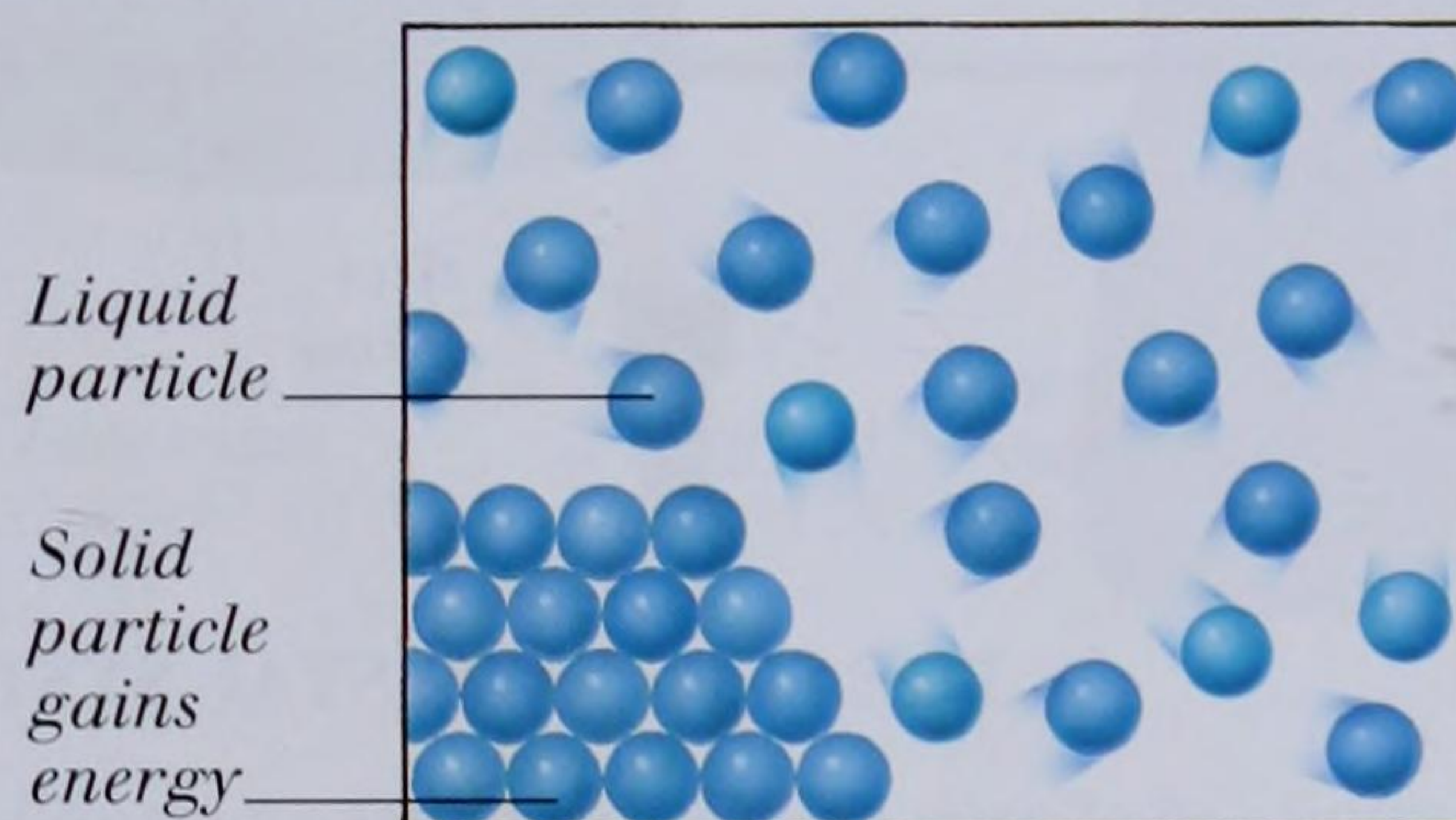
## LATENT HEAT

### HEATING A SUBSTANCE

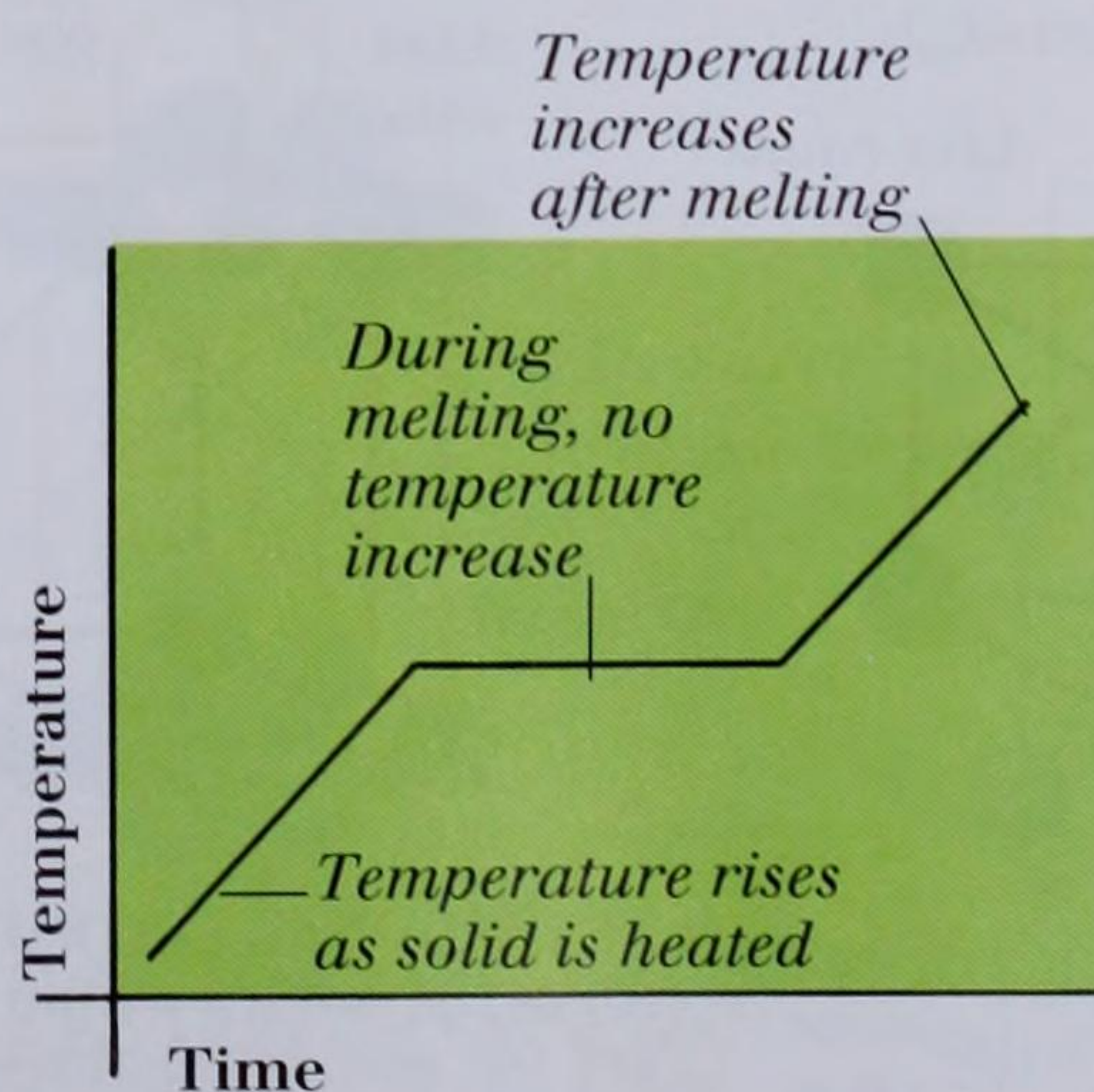
Heat transferred from a hot flame to a cooler substance can cause the substance to melt. The temperature of the substance (here, naphthalene) rises with the transfer of more energy, until it reaches the **melting point**.

### MELTING A SUBSTANCE

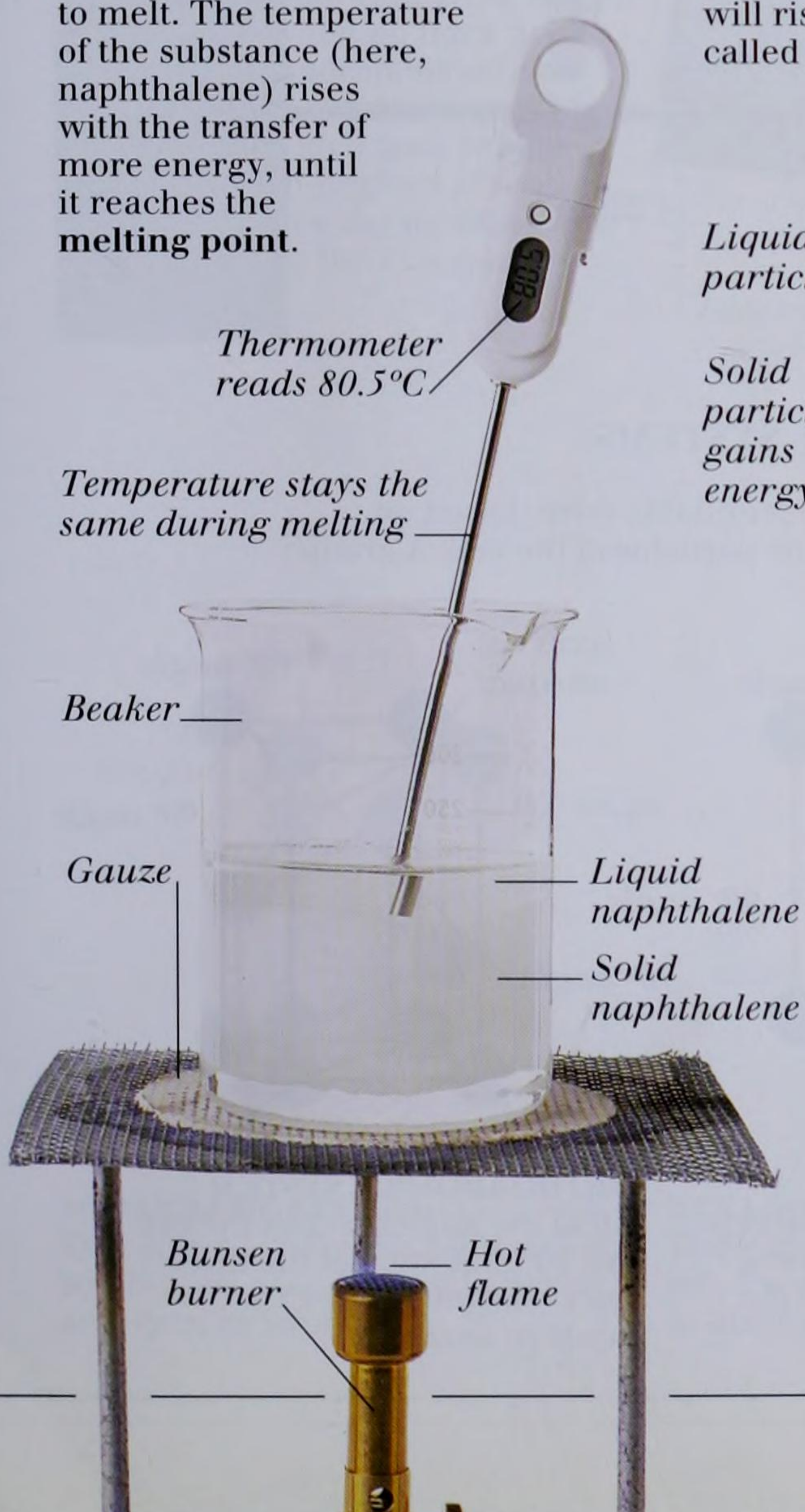
At the melting point, the supplied energy must break the attraction between all the particles, melting all the solid, before the temperature will rise again. This extra supplied energy is called **latent heat**.



MELTING



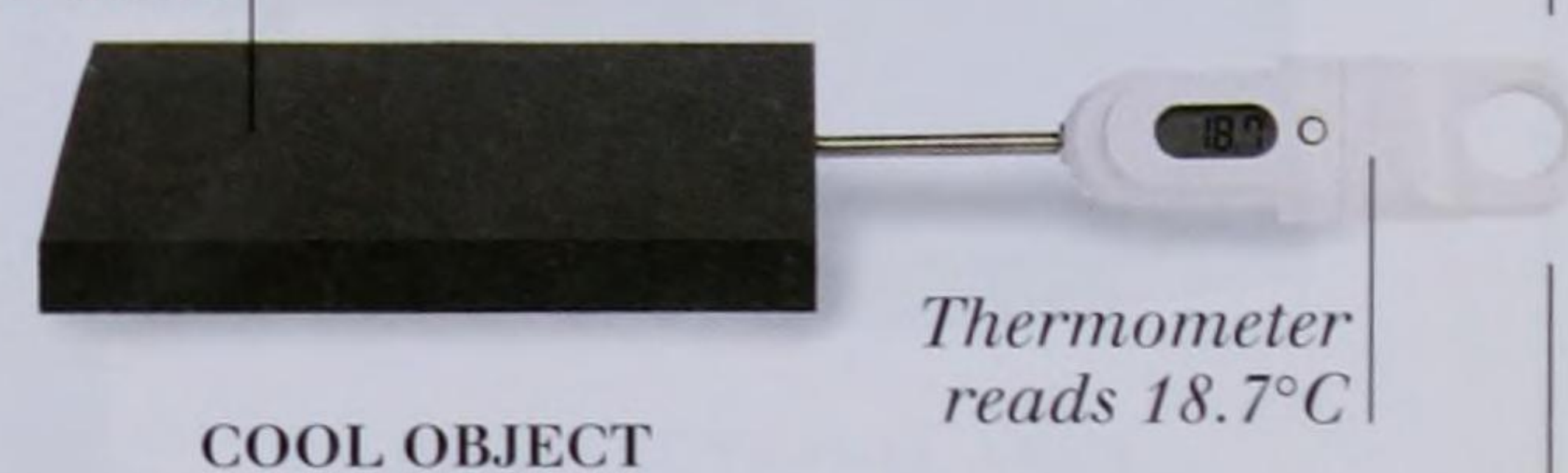
GRAPH TO SHOW MELTING



## TRANSFER OF HEAT BY RADIATION

An object at room temperature produces radiation—called **infrared radiation**. A hot object, such as the lamp below, produces a lot of infrared. This radiation can heat up other objects. The hot object cools as it loses energy as radiation.

Metal block at room temperature



RADIATION



# Solids

THE ATOMS OF A SOLID ARE CLOSELY PACKED, giving it a greater **density** than most liquids, and all gases. A solid's rigidity derives from the strong attraction between its atoms. A force pulling on a solid moves these atoms farther apart, creating an opposing force called **tension**. If a force pushes on a solid, the atoms move closer together, creating **compression**. Temperature (see pp. 22–23) can also affect the nature of a solid. When the temperature of a solid increases, its particles gain **kinetic energy** and vibrate more vigorously, resulting in **thermal expansion**. Most solids are **crystals**, in which atoms are arranged in one of seven regular, repeating patterns (see below). **Amorphous solids**, such as glass, are not composed of crystals and can be molded into any shape. When the atoms of a solid move apart, the length of the solid increases. The extent of this increase depends on the applied force and the thickness of the material, and is known as **elasticity**.

## STEEL RAILS

The expansion of a solid with an increase in temperature (see below) would cause rails to buckle badly in hot weather. To prevent this, rails are made in sections. The gap between the two sections allows each section to expand without buckling.

Train can pass smoothly over diagonal joint

Expansion joint



## THERMAL EXPANSION

### EXPERIMENT TO SHOW THERMAL EXPANSION

When a substance is heated, its atoms gain kinetic energy. In a solid, this results in the atoms vibrating more vigorously about their fixed positions. As a result, solids expand when heated. Below, a thin steel rod is heated by a gas flame, and the resulting expansion is measured using a **micrometer**.

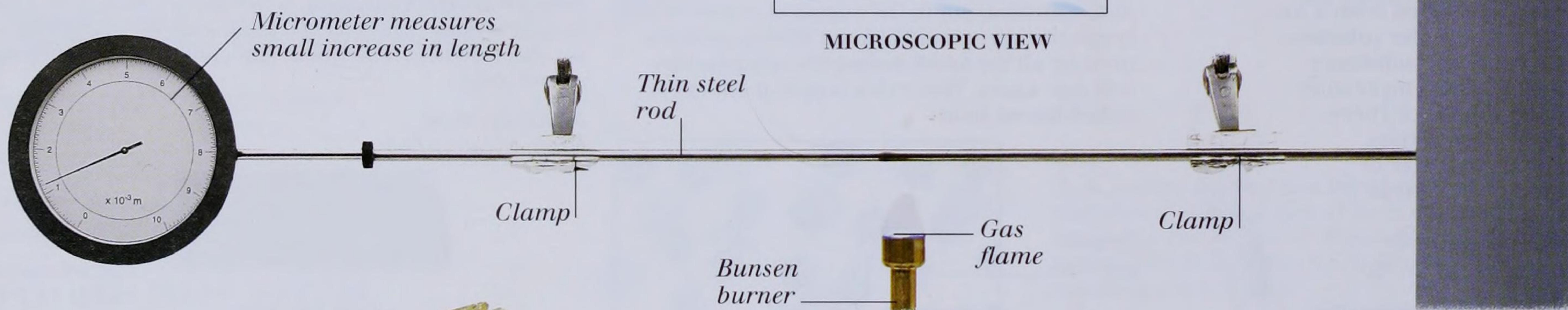
Metal atoms gain energy

The higher the temperature, the greater the vibration

Vibration around fixed point

Steel rod pushes against rigid block

### MICROSCOPIC VIEW



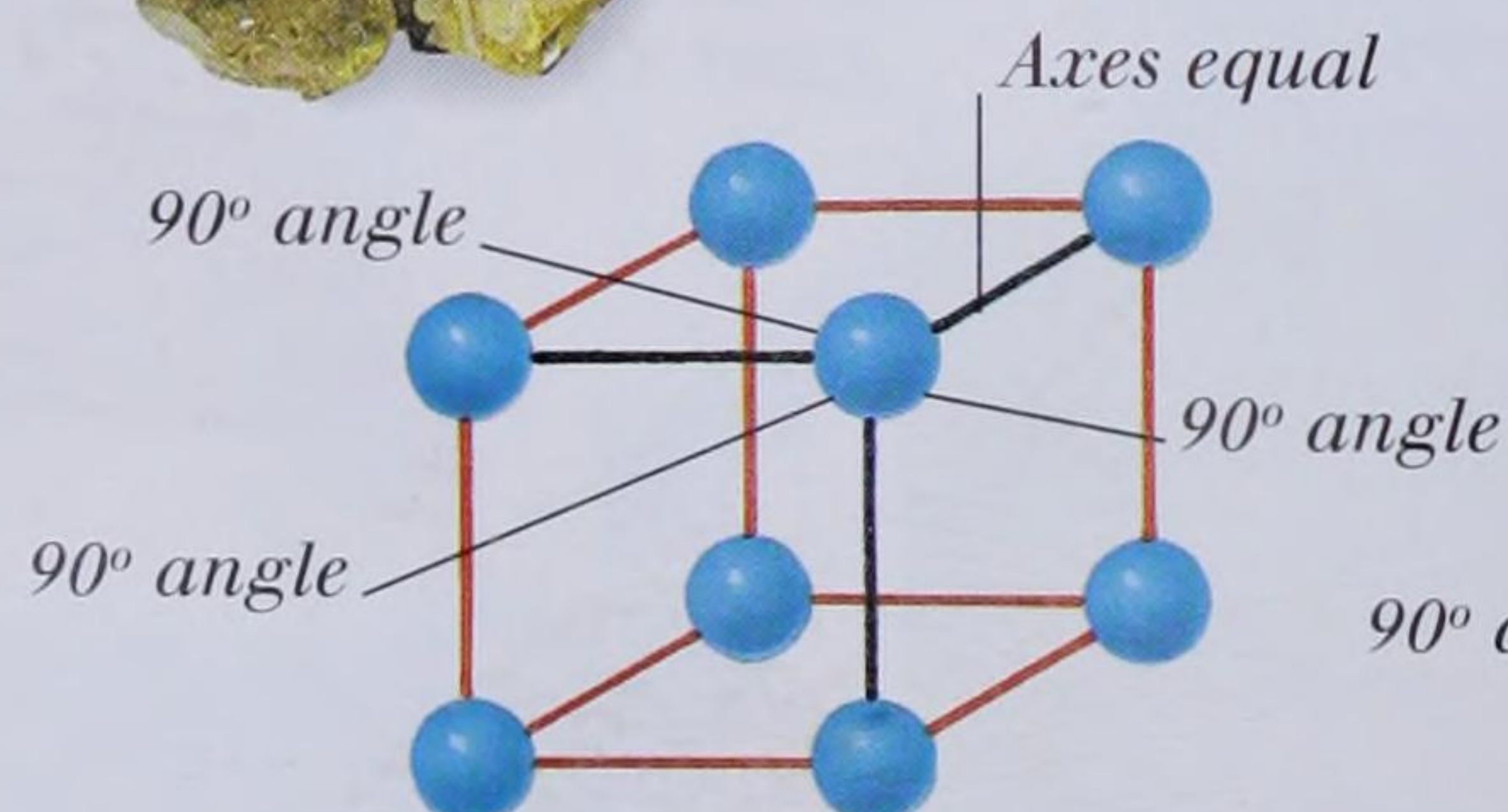
### EXTERNAL FEATURES

The seven crystal systems are based on the external shapes of crystals, but they also correspond to the arrangement of atoms within. The basic arrangement that is repeated in the crystal is called the **unit cell**.



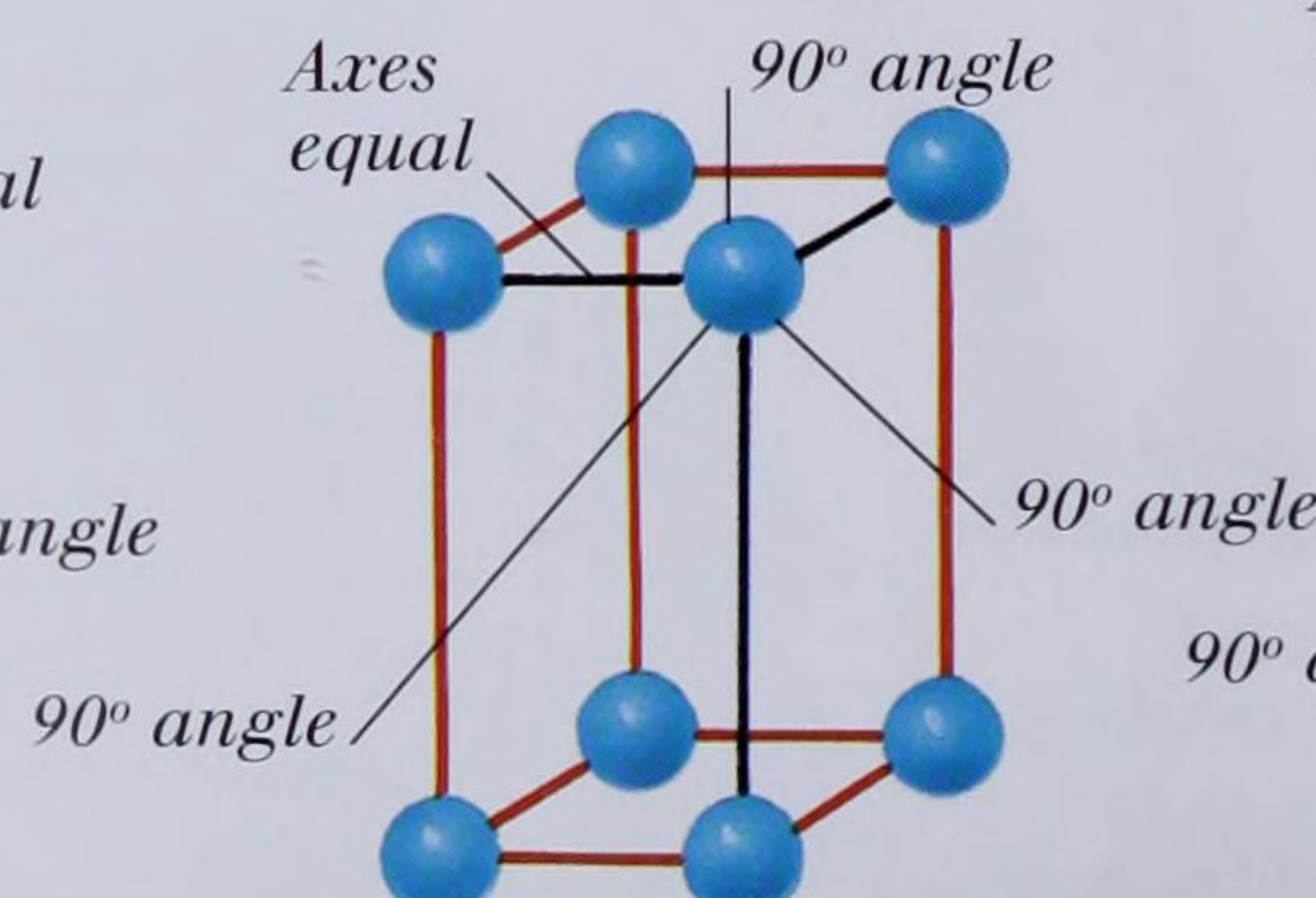
## THE SEVEN CRYSTAL SYSTEMS

The unit cell of each crystal system has an identifiable form, based on hypothetical axes composed by joining up the particles of the cell. A group of unit cells form a **crystal lattice**.



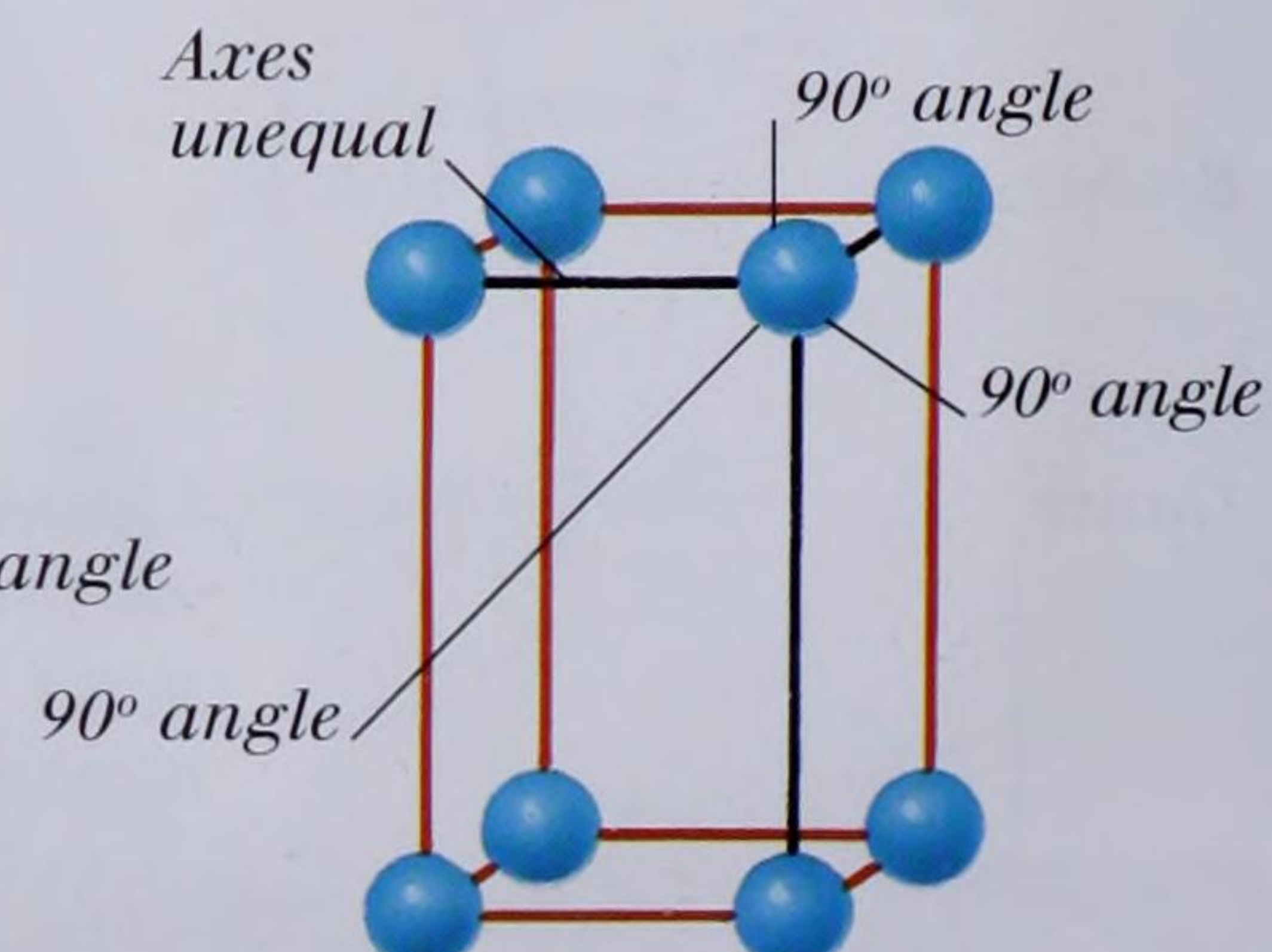
### CUBIC SYSTEM

Atoms in a cubic system are equally spaced, and the angle between each axis of the repeating cell is always 90°.



### TETRAGONAL SYSTEM

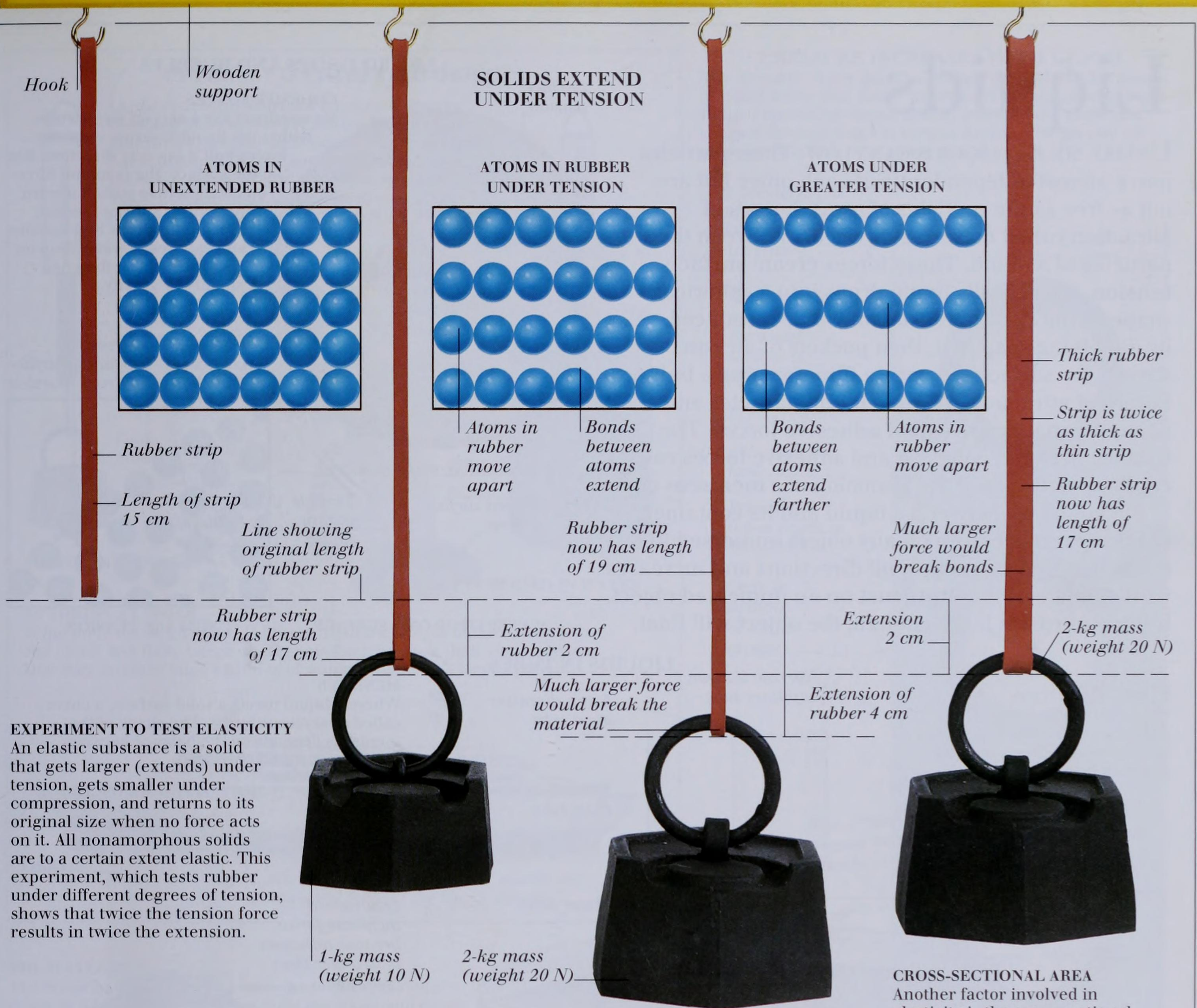
All of the angles within the cell are 90° and, of the three axes (shown in black), two are the same length.



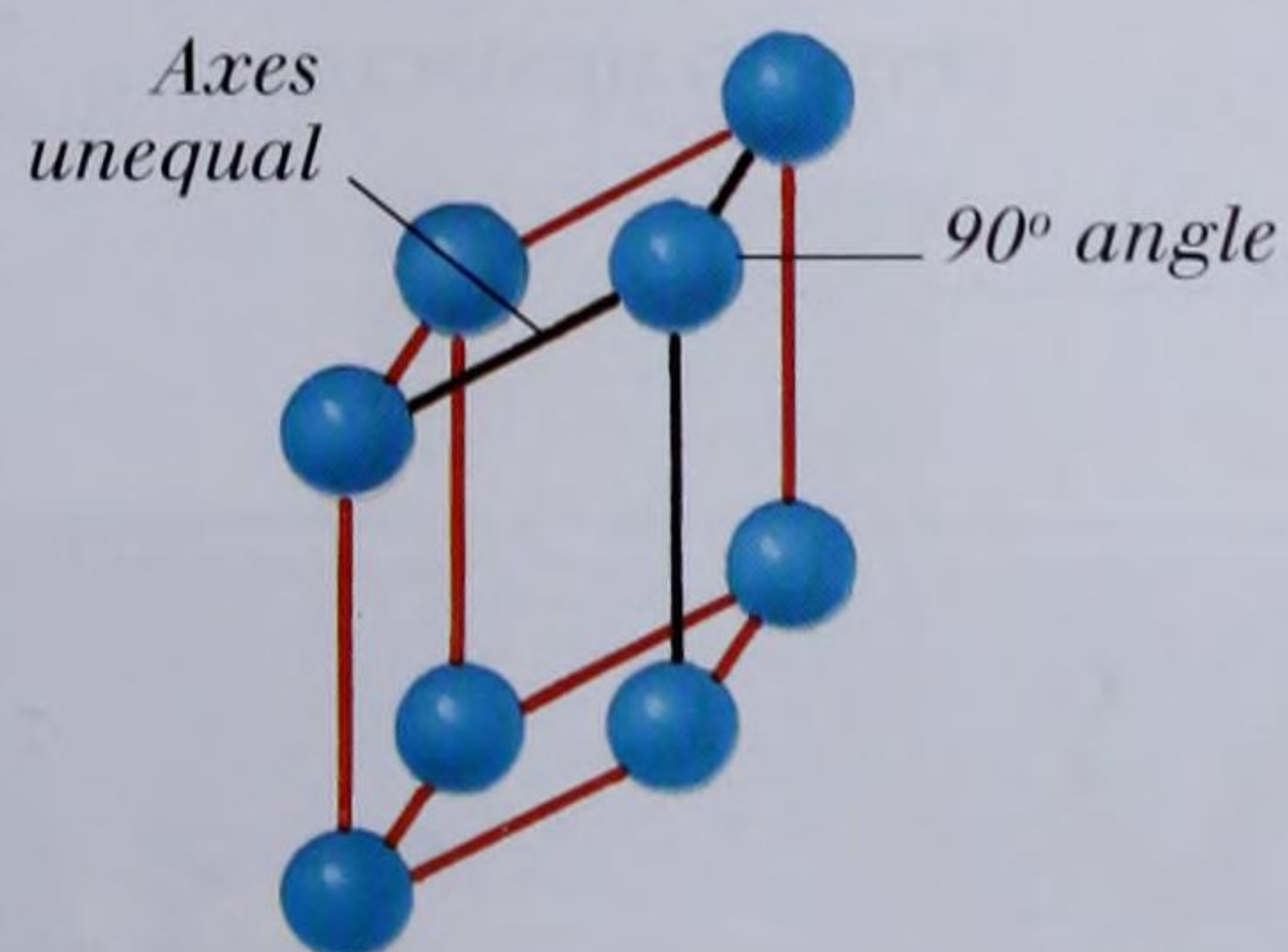
### ORTHORHOMBIC SYSTEM

All of the angles within the cell are 90°, but none of the three axes (shown in black) is equal in length.

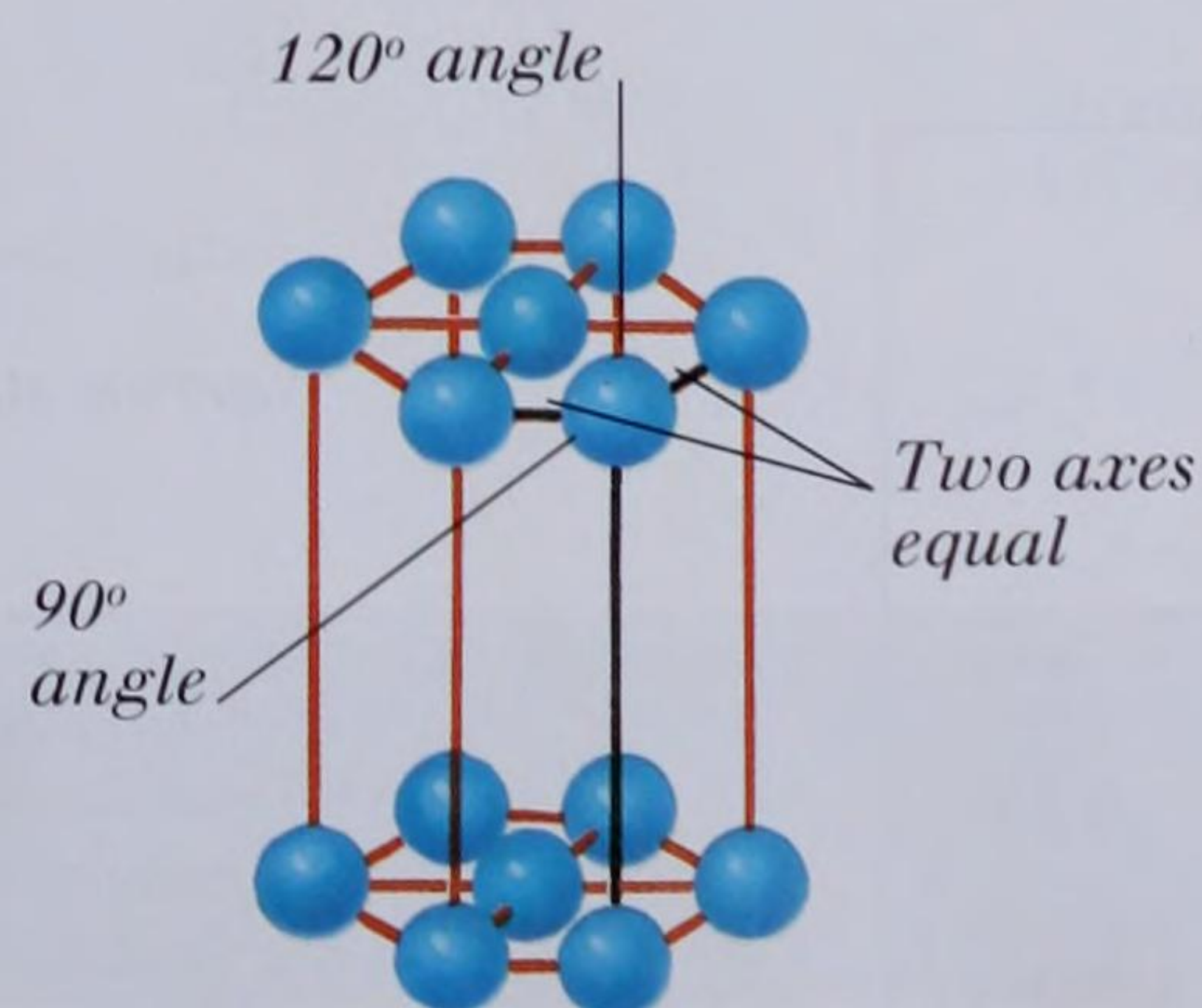




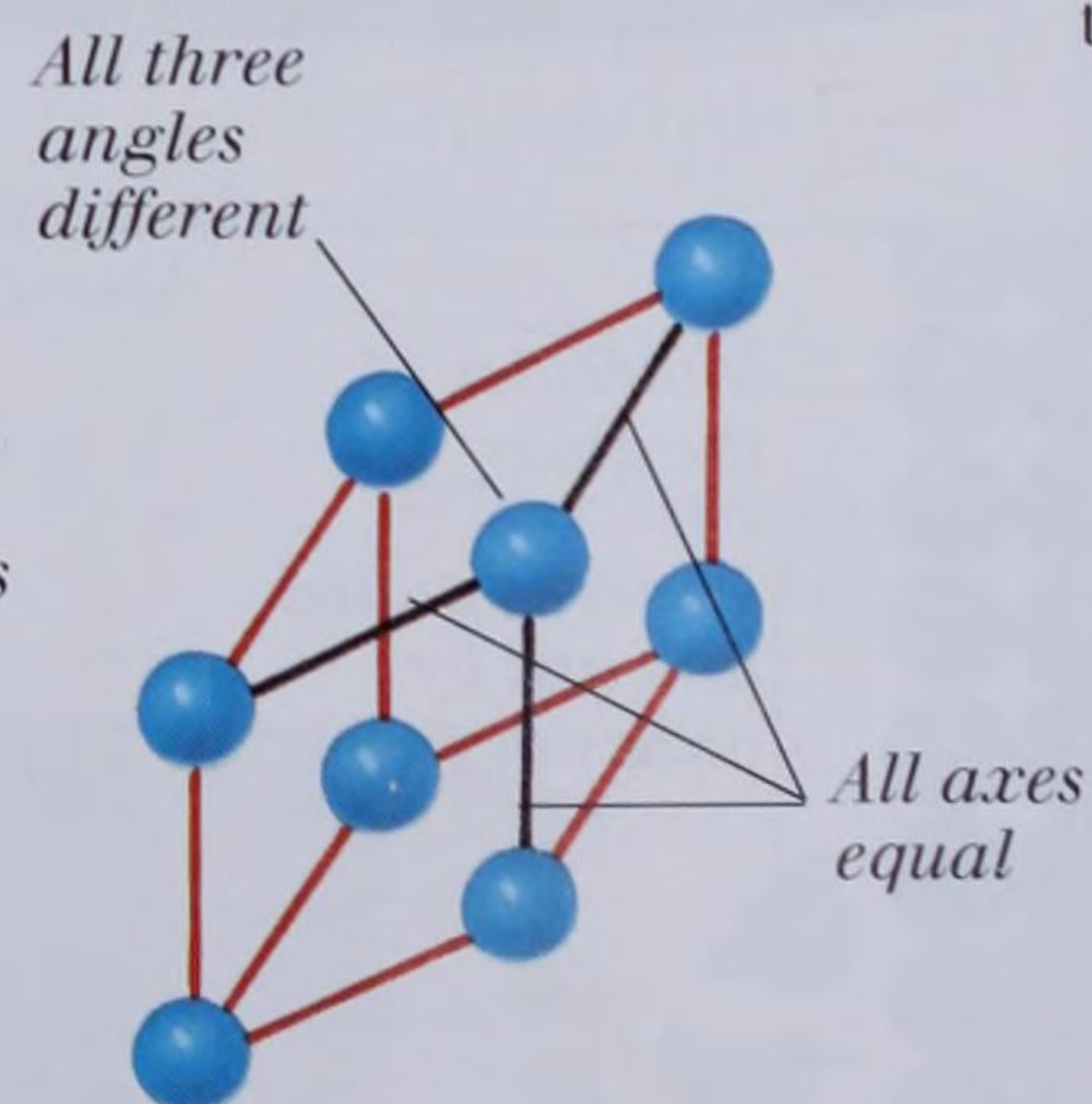
**CROSS-SECTIONAL AREA**  
Another factor involved in elasticity is the cross-sectional area of the material involved. The thick rubber strip (above) extends less under the same tension than the thinner one (above left).



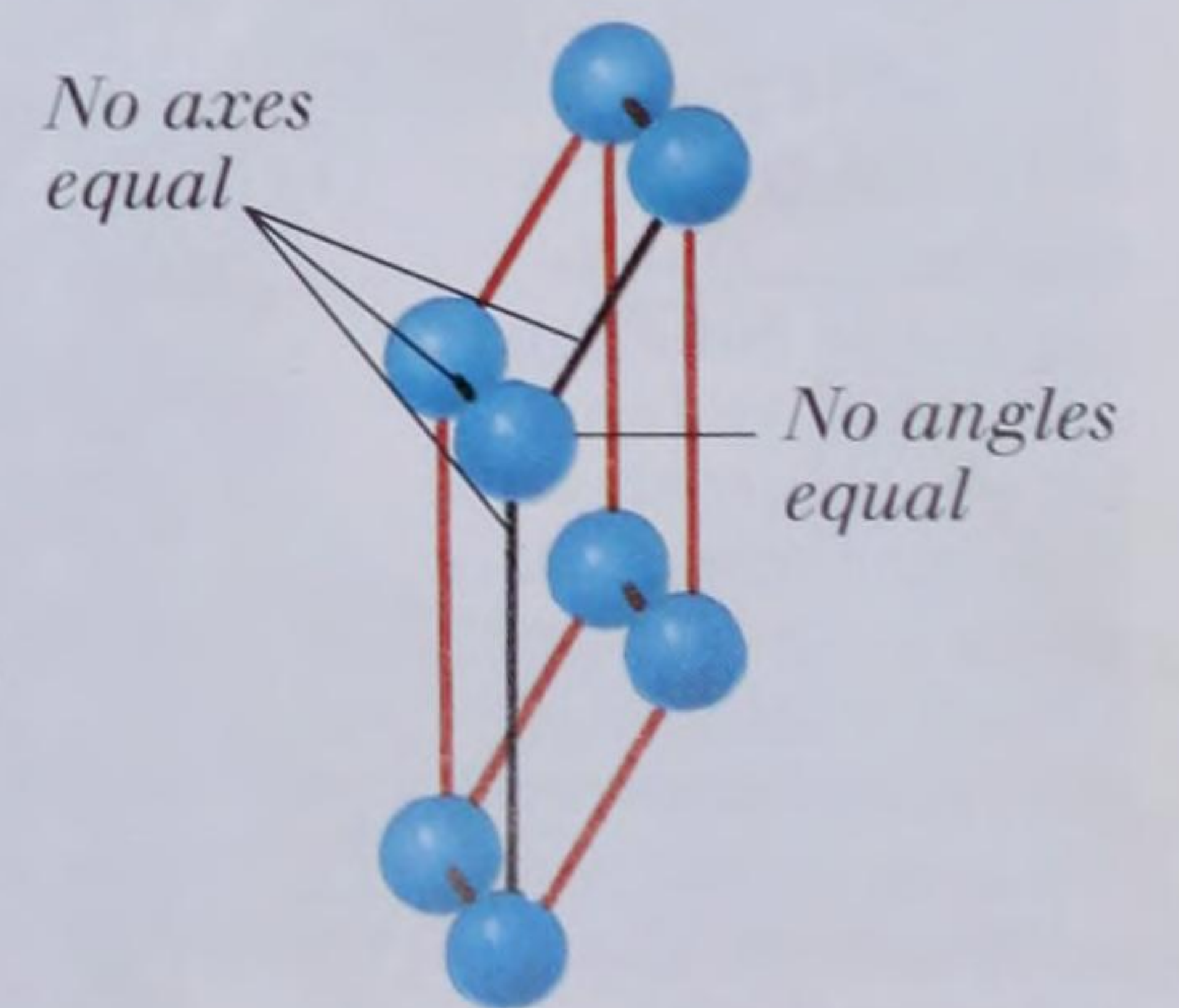
**MONOCLINIC SYSTEM**  
Two of the axes of the cell meet at 90°. No two axes (shown in black) are equal in length.



**HEXAGONAL SYSTEM**  
The edges form angles of 120° and 90°. Two of the three axes (shown in black) are equal in length.



**TRIGONAL SYSTEM**  
No two edges meet at 90°. All of the edges are equal in length.



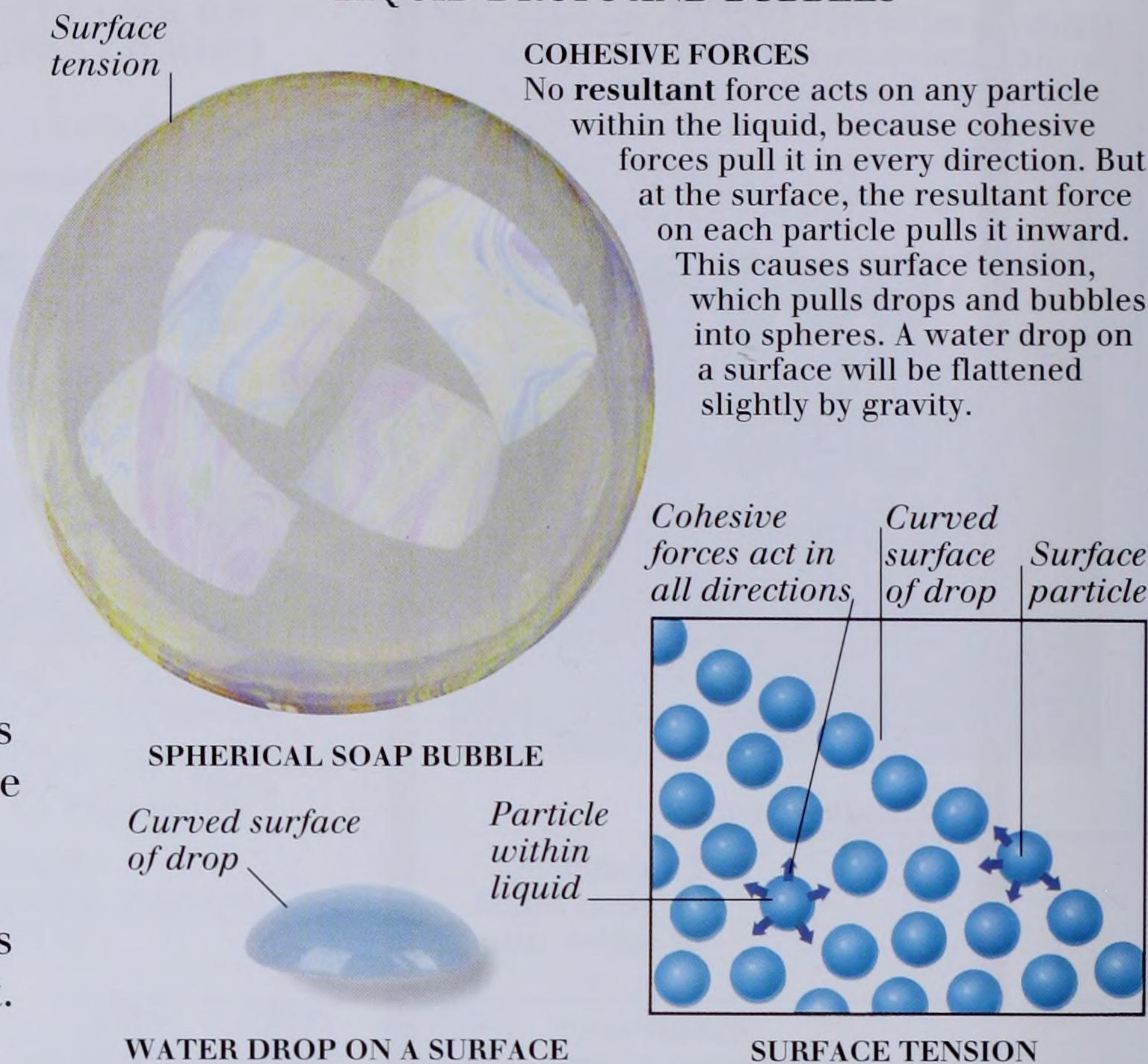
**TRICLINIC SYSTEM**  
No two edges meet at 90°. No two axes (shown in black) are equal in length.



# Liquids

UNLIKE SOLIDS, LIQUIDS CAN FLOW. Their particles move almost independently of each other but are not as free as the particles of a gas. Forces of attraction called **cohesive** forces act between the particles of a liquid. These forces create **surface tension**, which pulls liquid drops into a spherical shape. If the surface tension of water is reduced by dissolving soap in it, then pockets of air can stretch the surface into a thin film, forming a bubble. Forces of attraction between liquid particles and adjoining matter are called **adhesive** forces. The balance between cohesive and adhesive forces causes **capillary action**, and the formation of a **meniscus** curve at the boundary between a liquid and its container. Liquids exert pressure on any object immersed in them; the pressure acts in all directions and increases with depth, creating **upthrust** on an immersed object. If the upthrust is large enough, the object will float.

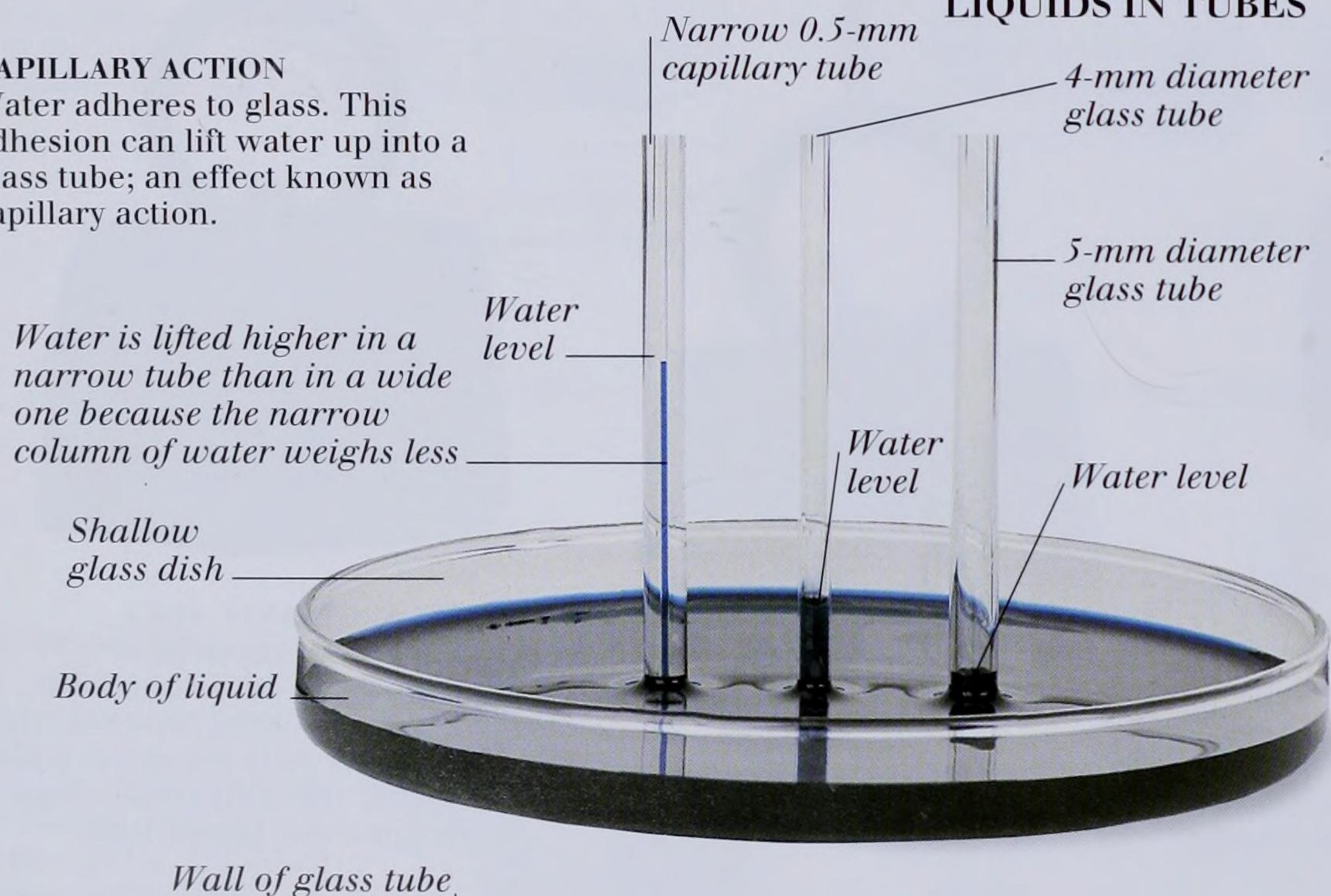
## LIQUID DROPS AND BUBBLES



### CAPILLARY ACTION

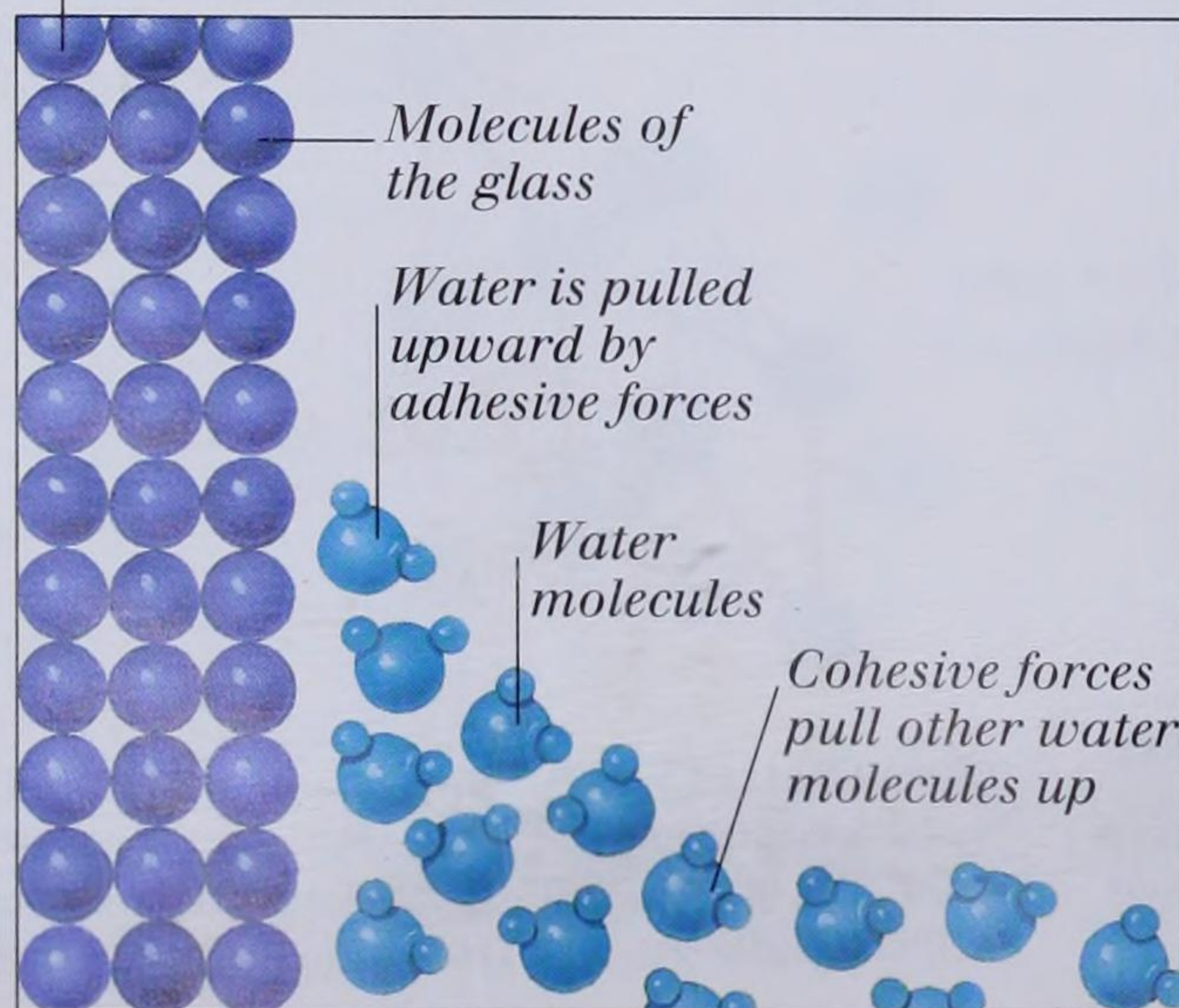
Water adheres to glass. This adhesion can lift water up into a glass tube; an effect known as capillary action.

## LIQUIDS IN TUBES



### MOLECULAR VIEW

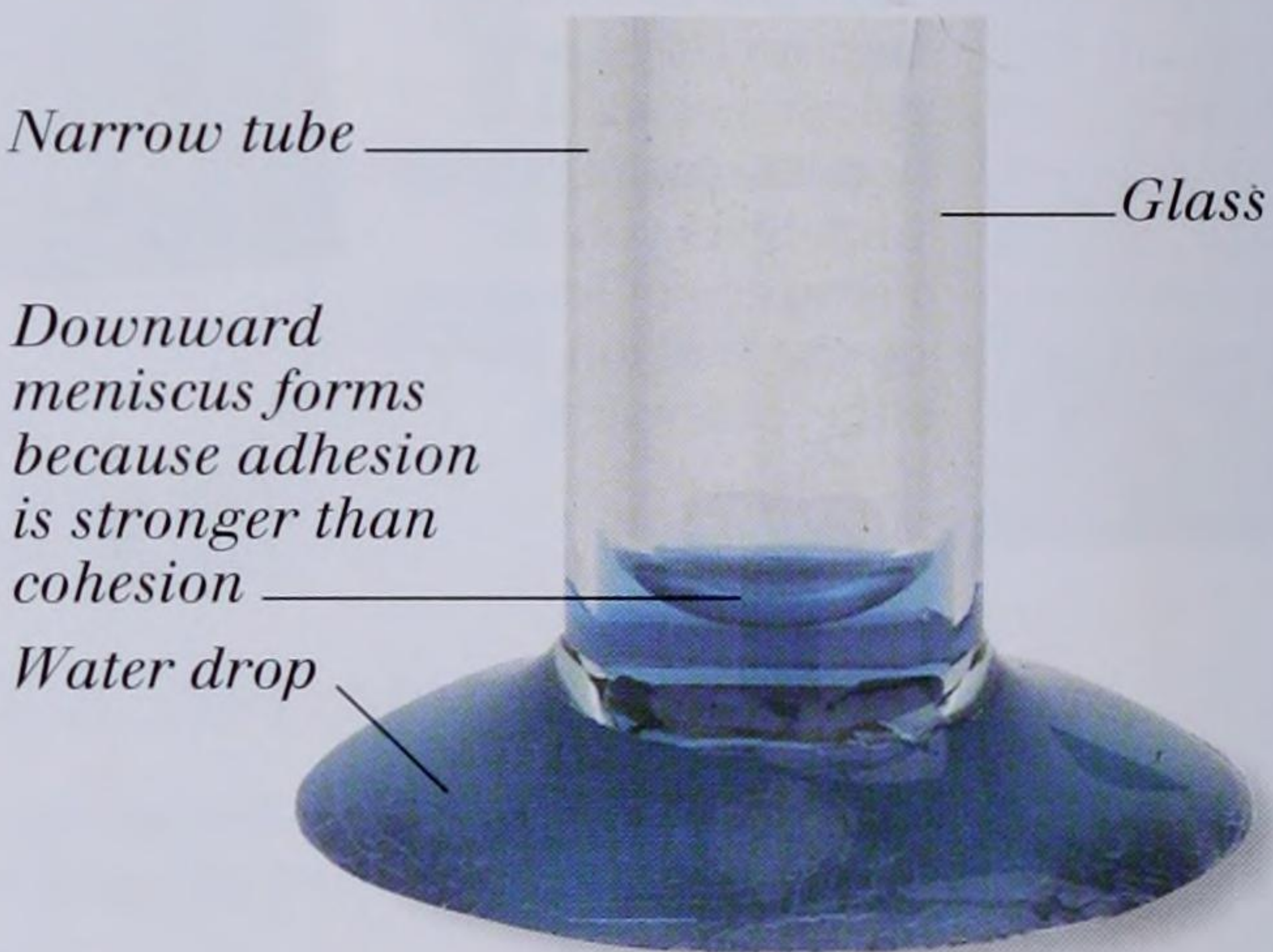
Capillary action is caused by adhesive and cohesive forces between particles of glass and water. Here, water molecules adhere to glass and the adhesive force lifts the edge of the water up the glass. The cohesive forces between water molecules means that this lifted edge also raises water molecules lying farther out from the edge of the glass.



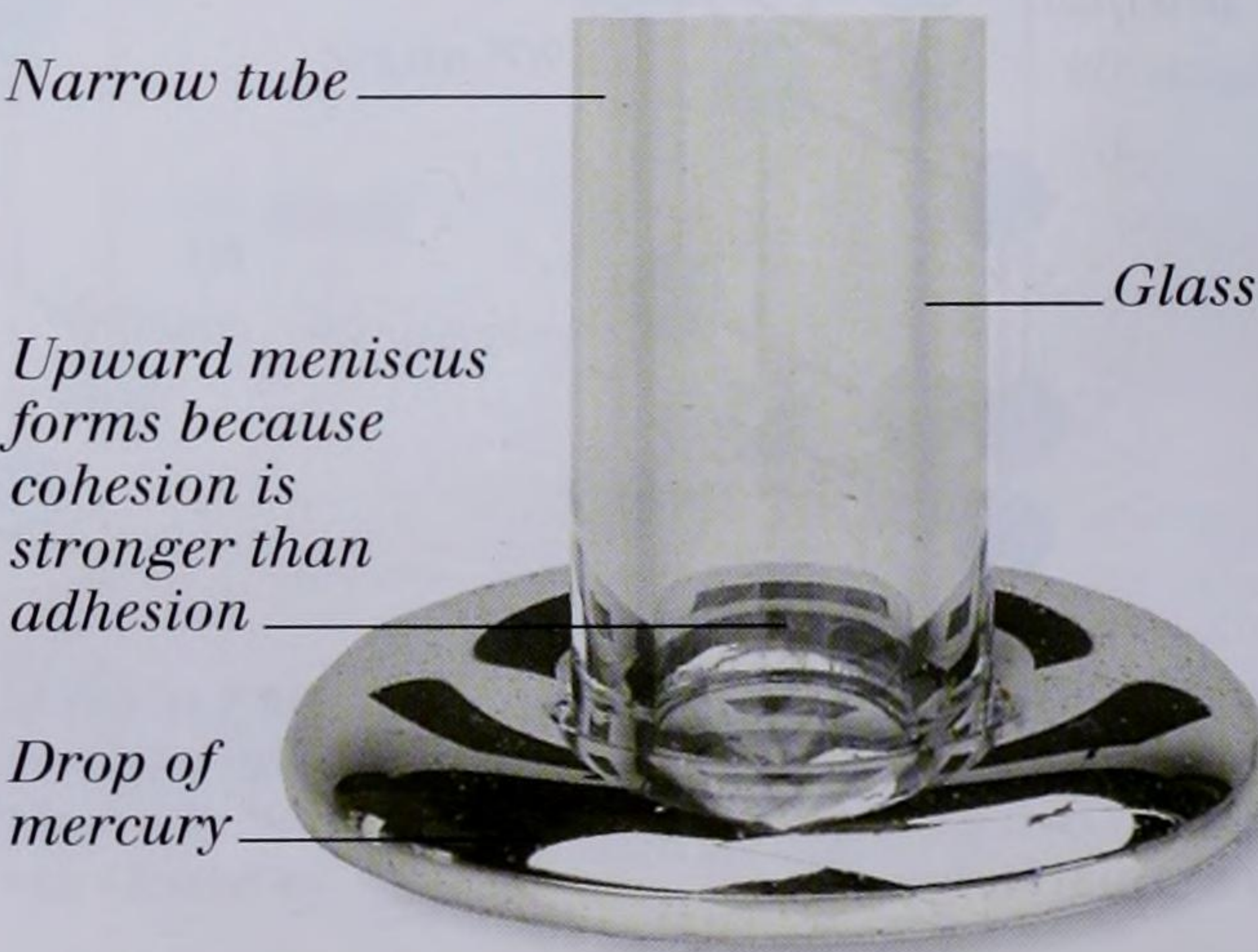
### MENISCUS

Where a liquid meets a solid surface, a curve called a meniscus forms. The shape of the meniscus depends on the balance between cohesive and adhesive forces.

#### DOWNWARD MENISCUS



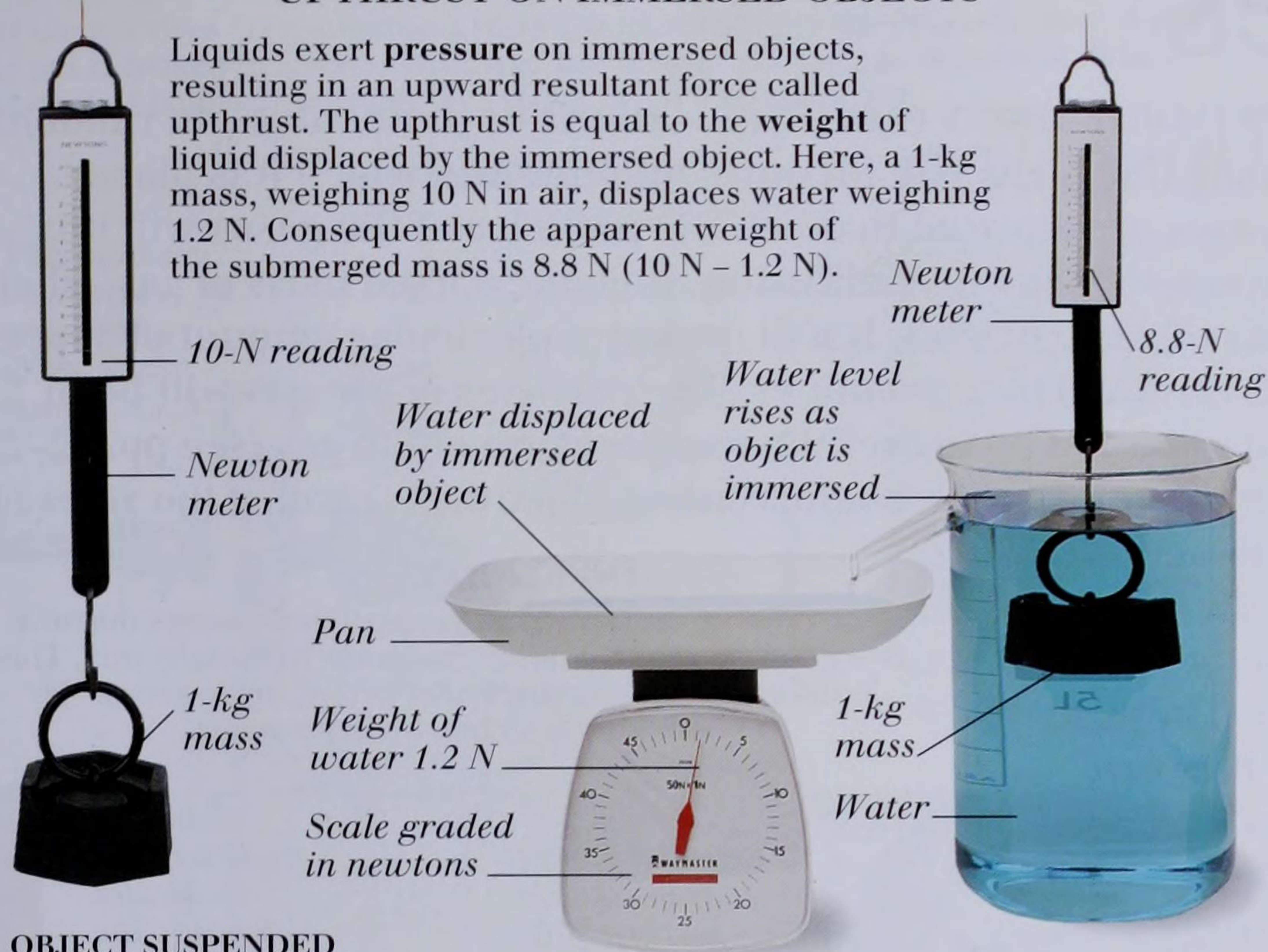
#### UPWARD MENISCUS





## UPTHRUST ON IMMERSED OBJECTS

Liquids exert **pressure** on immersed objects, resulting in an upward resultant force called **upthrust**. The upthrust is equal to the **weight** of liquid displaced by the immersed object. Here, a 1-kg mass, weighing 10 N in air, displaces water weighing 1.2 N. Consequently the apparent weight of the submerged mass is 8.8 N ( $10\text{ N} - 1.2\text{ N}$ ).



OBJECT SUSPENDED  
IN AIR

OBJECT IMMERSED IN WATER

### UPTHRUST AT WORK

If the upthrust on an object is greater than the weight of the object, then the object will float. Large metal ships float, because their shape means that they displace huge amounts of water, producing a large upthrust.



Metal hull

Upthrust force from the water equals the downward force of the ship's weight

### THE WATER JETS

The water in the jets coming from the tank breaks into drops as it falls. Surface tension pulls the water into drops as the jet weakens and cohesive forces keep the drops in a near spherical shape. When the drops fall into the tray, they form a pool. Unlike solids, liquids can flow, so under the influence of **gravity** the surface of this pool becomes flat and horizontal.

Shallow tray

Flat and horizontal surface

Colored water

Stream is almost horizontal

Water escapes very quickly

## PRESSURE INCREASES WITH DEPTH

The pressure at any point in a liquid depends on the weight of liquid above that point. So pressure increases with depth. In the experiment shown below, water from a large tank escapes through holes at various depths. The greater the pressure, the faster the water escapes.

The pressure of a liquid is measured in newtons per square meter ( $\text{Nm}^{-2}$ )

Atmospheric pressure above the water's surface is  $100,000\text{ Nm}^{-2}$

Pressure gauge

Pressure at 0.1-m depth is  $101,000\text{ Nm}^{-2}$

Only a dribble of water escapes

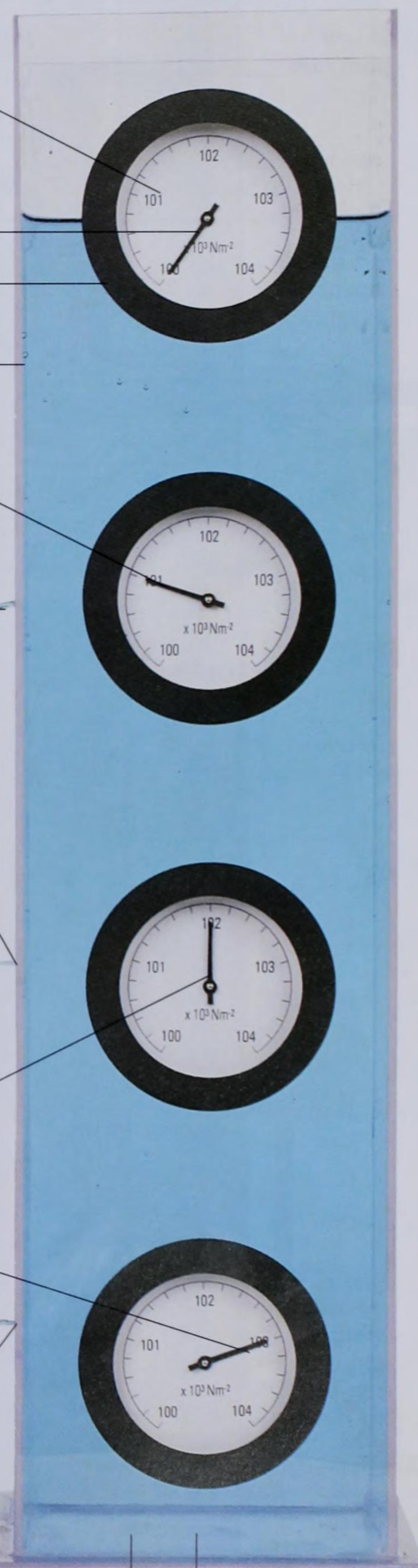
Water escapes quickly

Pressure at 0.2-m depth is  $102,000\text{ Nm}^{-2}$

Pressure at 0.3-m depth is  $103,000\text{ Nm}^{-2}$

Colored water

Water pressure greatest at the base of the tank





# Gases

A GAS COMPRISES INDEPENDENT PARTICLES—**atoms** or **molecules**—in random motion. This means that a gas will fill any container into which it is placed. If two different gases are allowed to meet, the particles of the gases will mix together. This process is known as **diffusion**. Imagine a fixed mass of gas—that is, a fixed number of gas particles. It will occupy a particular amount of space, or **volume**, often confined by a container. The particles of the gas will be in constant, random motion. The higher the **temperature** of the gas (see pp. 22–23), the faster the particles move. The bombardment of particles against the sides of the container produces **pressure**

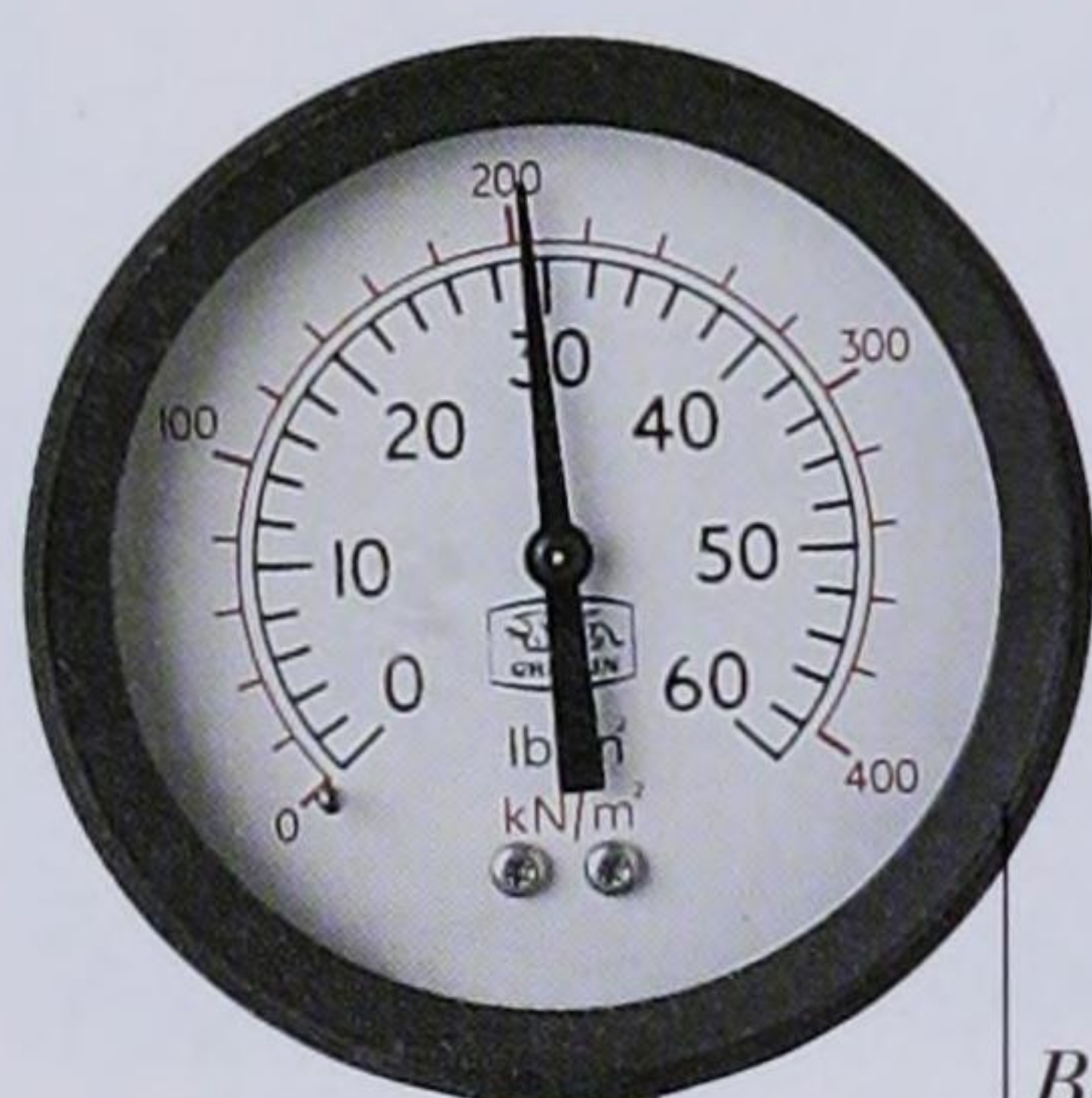
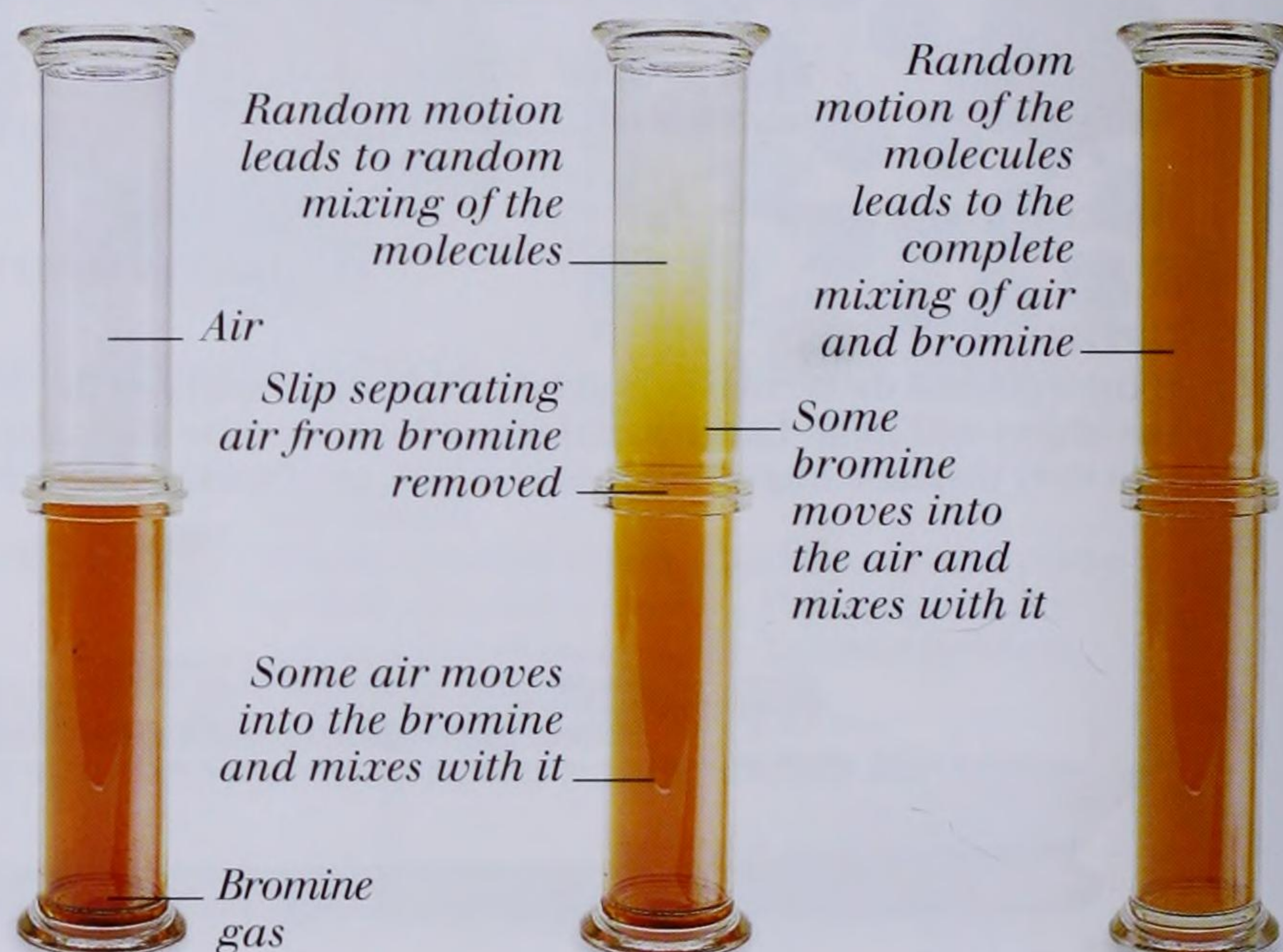
(see pp. 10–11). Three simple laws describe the predictable behavior of gases. They are Boyle's Law, the Pressure Law, and Charles' Law. Each of the gas laws describes a relationship between the pressure, volume, and temperature of a gas.

## BOYLE'S LAW

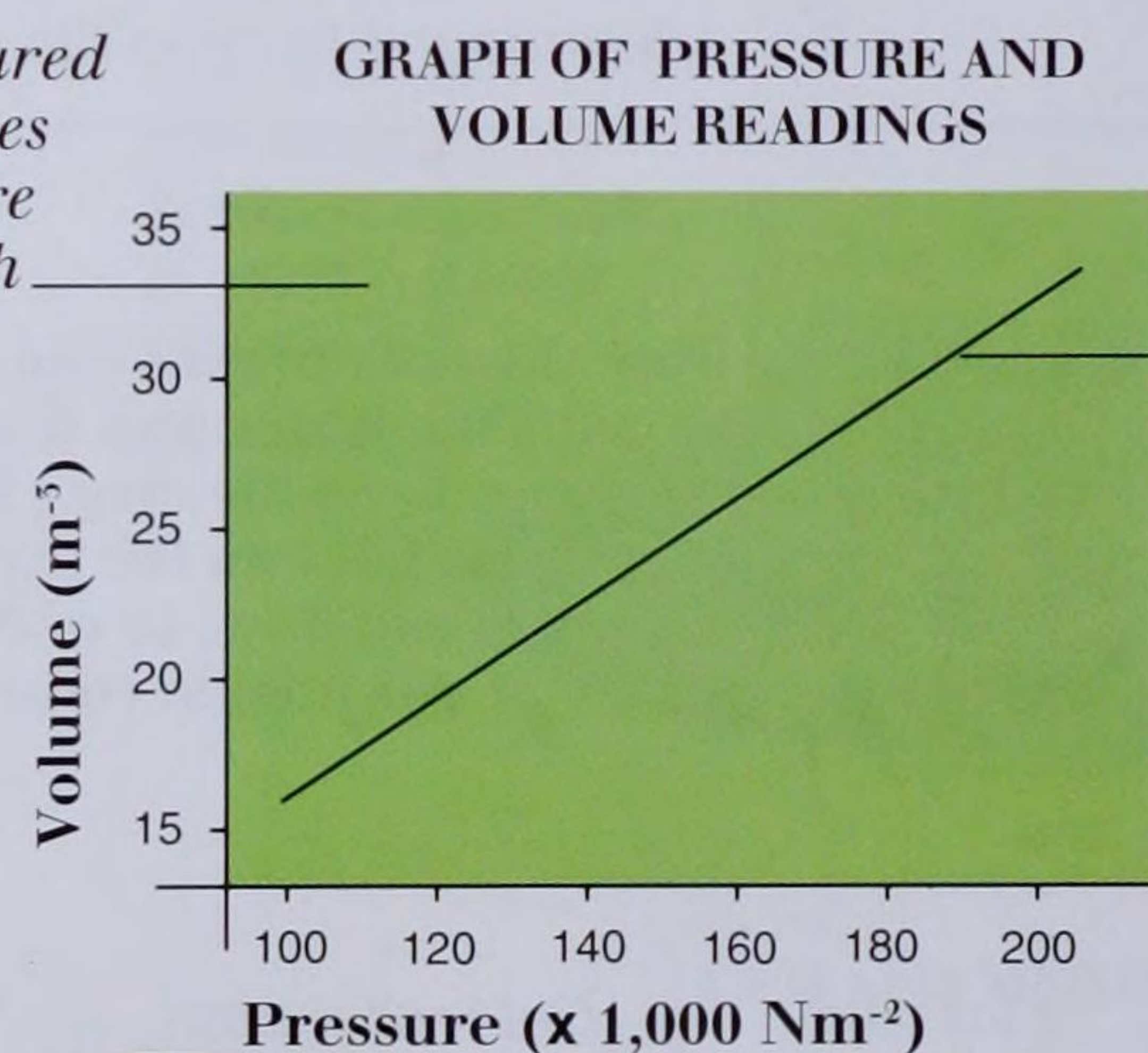
The volume of a mass of gas at a fixed temperature will change in relation to the pressure. If the pressure on a gas increases, its volume will decrease. The apparatus on the left is used to illustrate Boyle's Law. A foot pump is used to push a column of oil up a sealed tube, reducing the volume occupied by the gas in the top part of the tube.

## DIFFUSION

The random movement of gas particles ensures that any two gases sharing the same container will totally mix. This is diffusion. In the experiment below, the lower gas jar contains bromine, the top one air.



Pressure is measured at various volumes and the results are shown as a graph



Connecting pipe

Clear plexiglass shield

Closed glass tube

Thick tube wall withstands pressure

Volume of trapped air

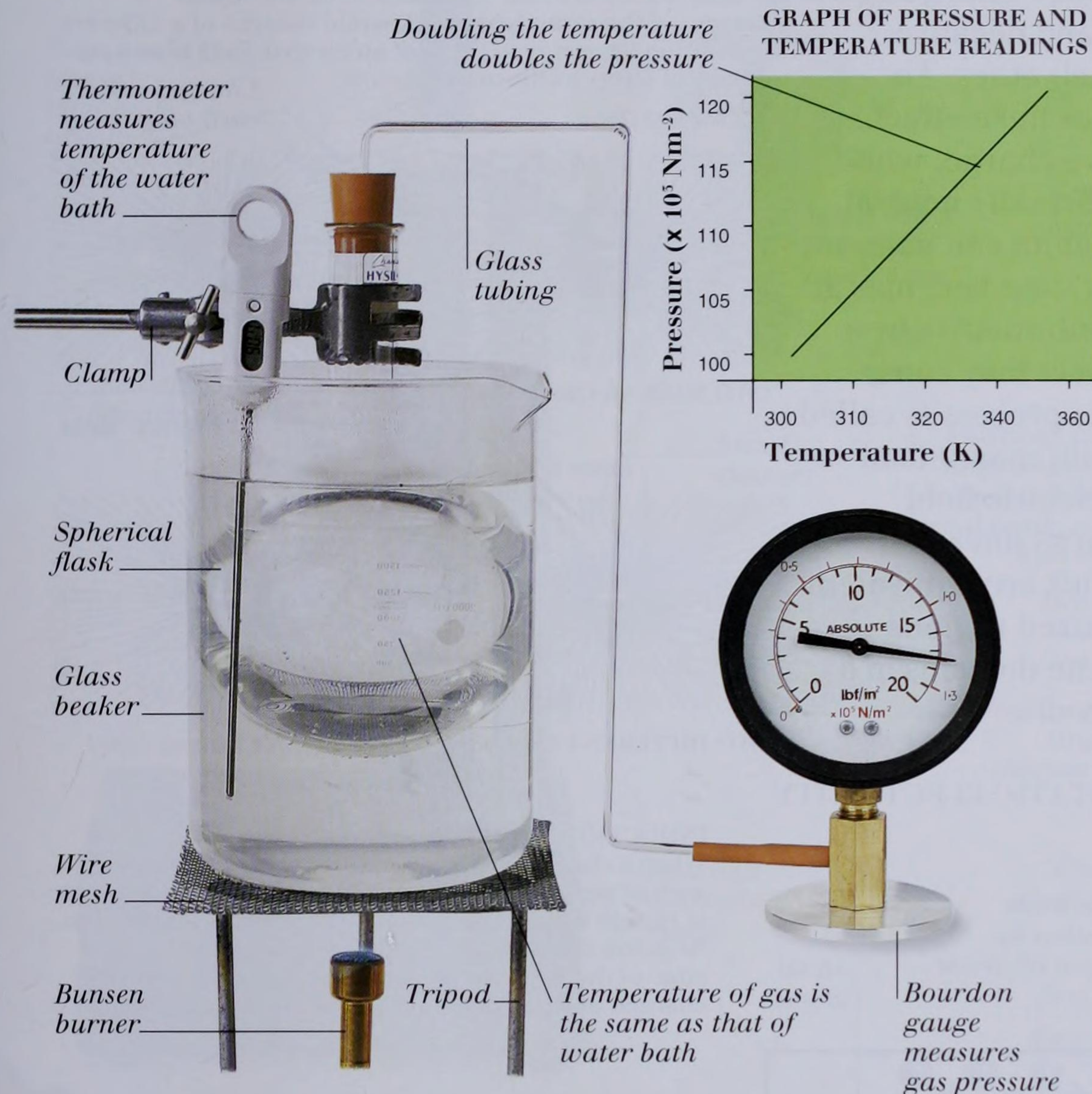
Column of oil

After each pressure change, apparatus is allowed to revert to room temperature



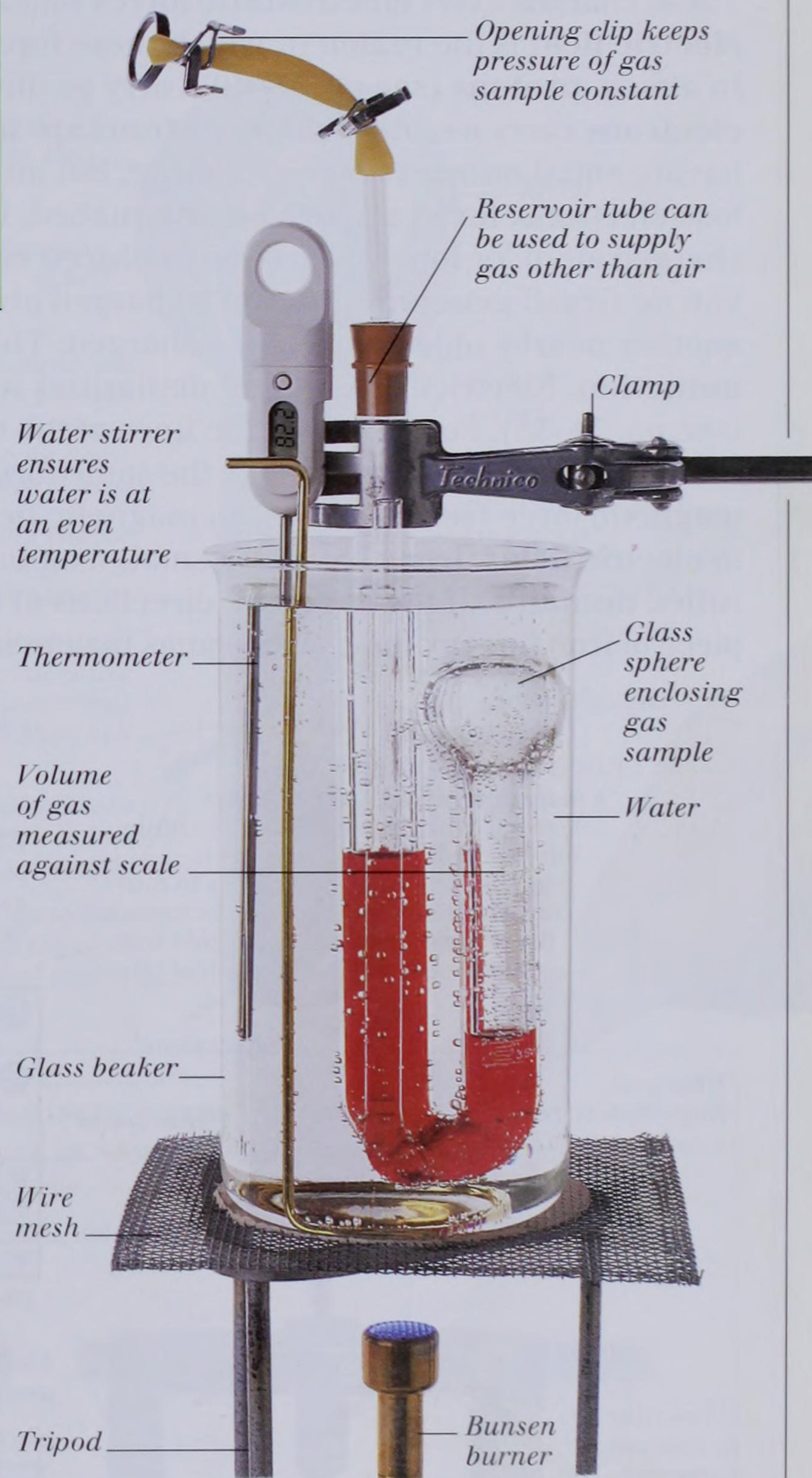
## PRESSURE LAW

The pressure exerted by a gas at constant volume increases as the temperature of the gas rises. The apparatus shown is used to verify the Pressure Law. A mass of gas is heated in a water bath, and the pressure of the gas measured. When plotted as points on a graph the results lie on a straight line.



## CHARLES' LAW

The volume of a mass of gas at a fixed pressure depends on its temperature. The higher the temperature, the greater the volume. The apparatus shown is used to illustrate Charles' Law. The volume of a gas sample in the glass bulb is noted at various temperatures. A graph shows the results.

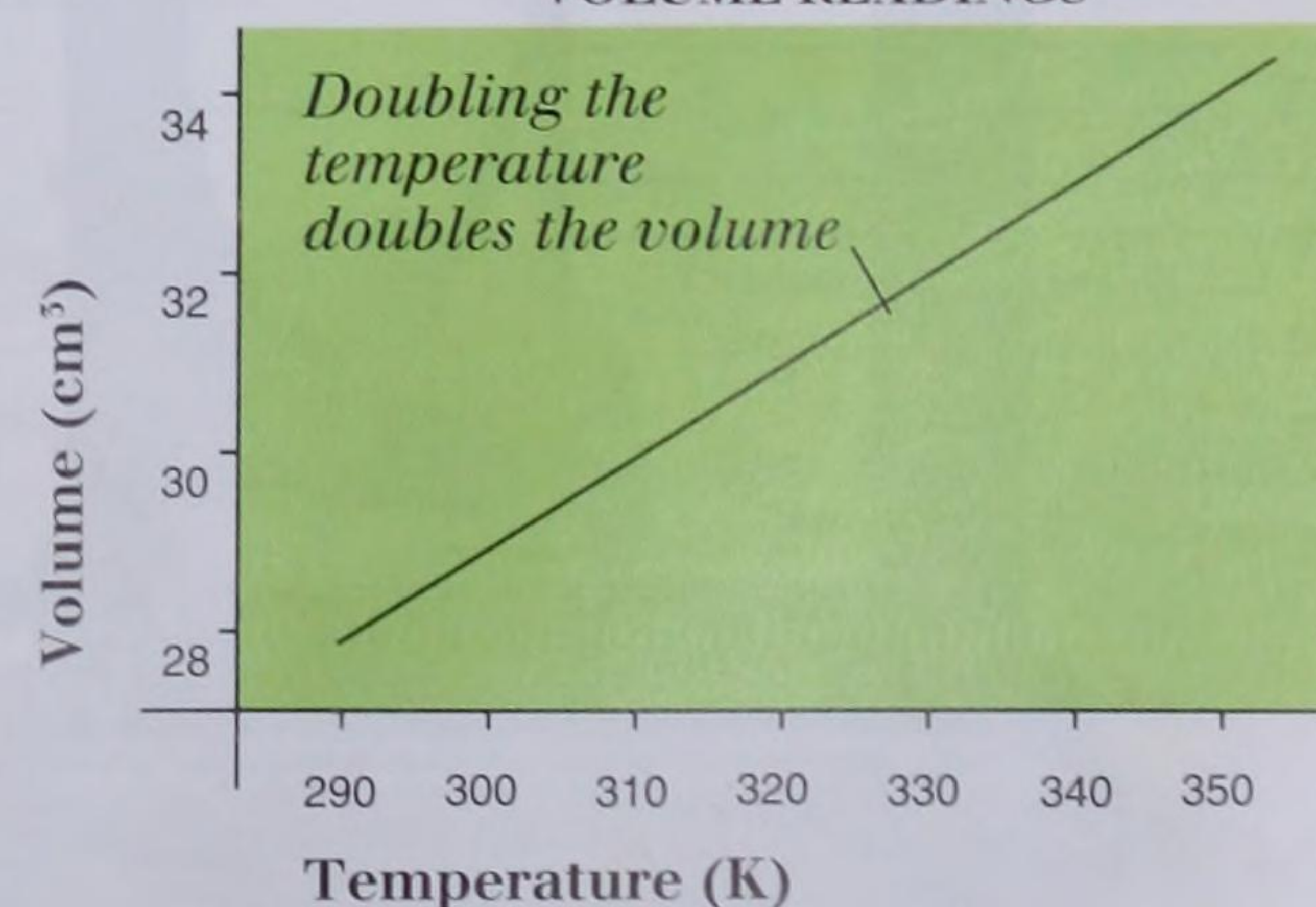


## HOT AIR BALLOON - CHARLES' LAW IN ACTION

The air in the envelope of a hot air balloon is heated by a gas burner. As its temperature rises, the gas expands in accordance with Charles' Law. The envelope is open at the bottom, so some hot air escapes. Because air has mass (and therefore **weight**), the balloon weighs less once some air has escaped, although its volume is still large. The pressure of the air outside the envelope produces an **upthrust**, which, if enough air has been lost from the envelope, will be great enough to lift the balloon.



## GRAPH OF TEMPERATURE AND VOLUME READINGS



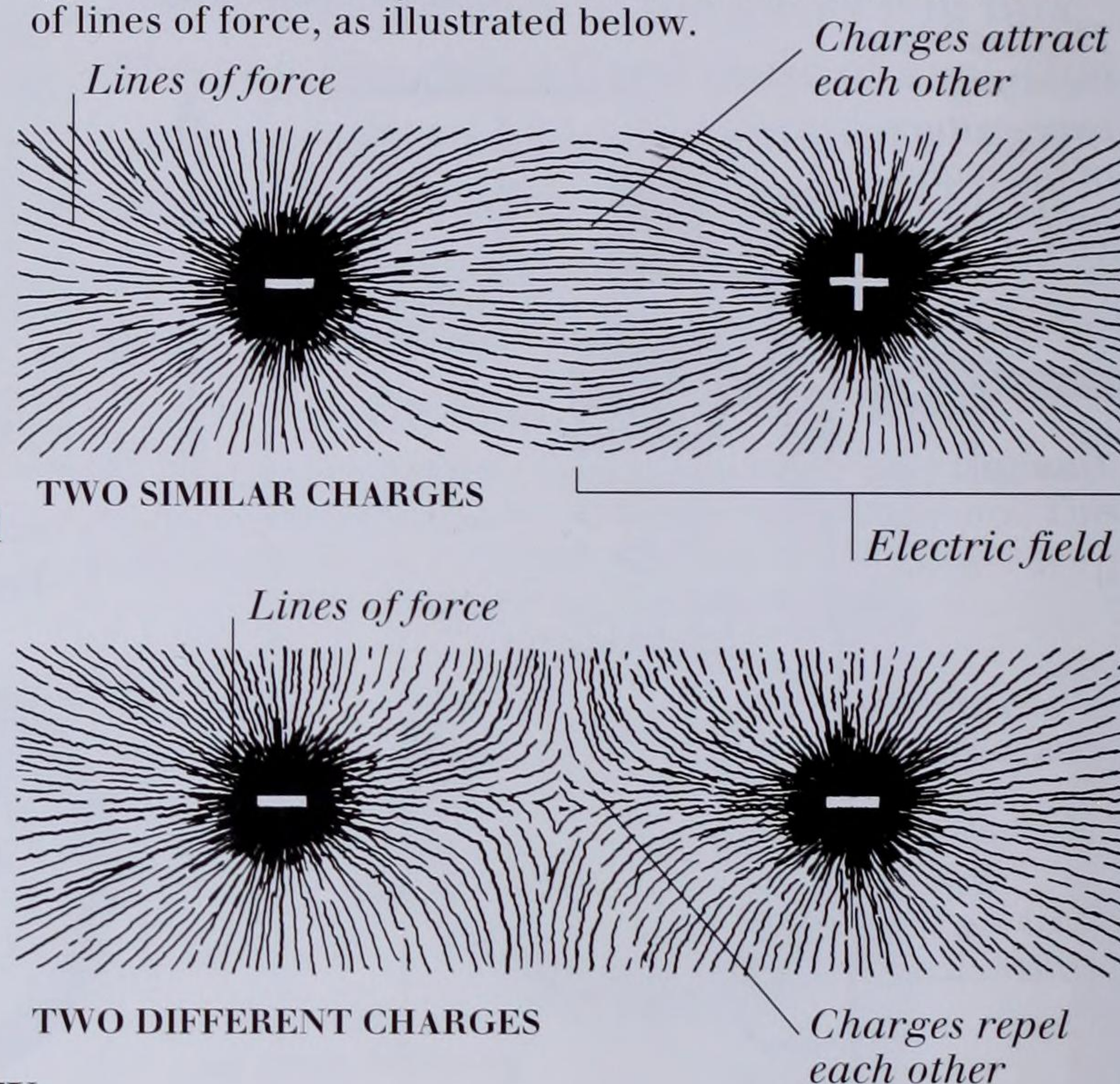


# Electricity and magnetism

ALL ELECTRICAL EFFECTS ARE CAUSED by **electric charges**. There are two types of electric charge, positive and negative. These charges exert **electrostatic** forces on each other. An **electric field** is the region in which these forces have effect. In atoms, **protons** (see pp. 48–49) carry positive charge, while **electrons** carry negative charge. **Atoms** are normally neutral, having equal numbers of each charge, but an atom can gain or lose electrons, for example by being rubbed. It then becomes a charged atom, or **ion**. Ions can be produced continuously by a Van de Graaff generator. Ions in a charged object may cause another nearby object to become charged. This process is called **induction**. Electricity has many similarities with magnetism (see pp. 34–35). For example, the lines of the electric field between charges (see right) take the same form as lines of magnetic force (see opposite), so magnetic fields are equivalent to electric fields. Iron consists of small magnetized regions called **domains**. If the magnetic directions of the domains in a piece of iron line up, the iron becomes magnetized.

## ELECTRIC FIELDS AND FORCES

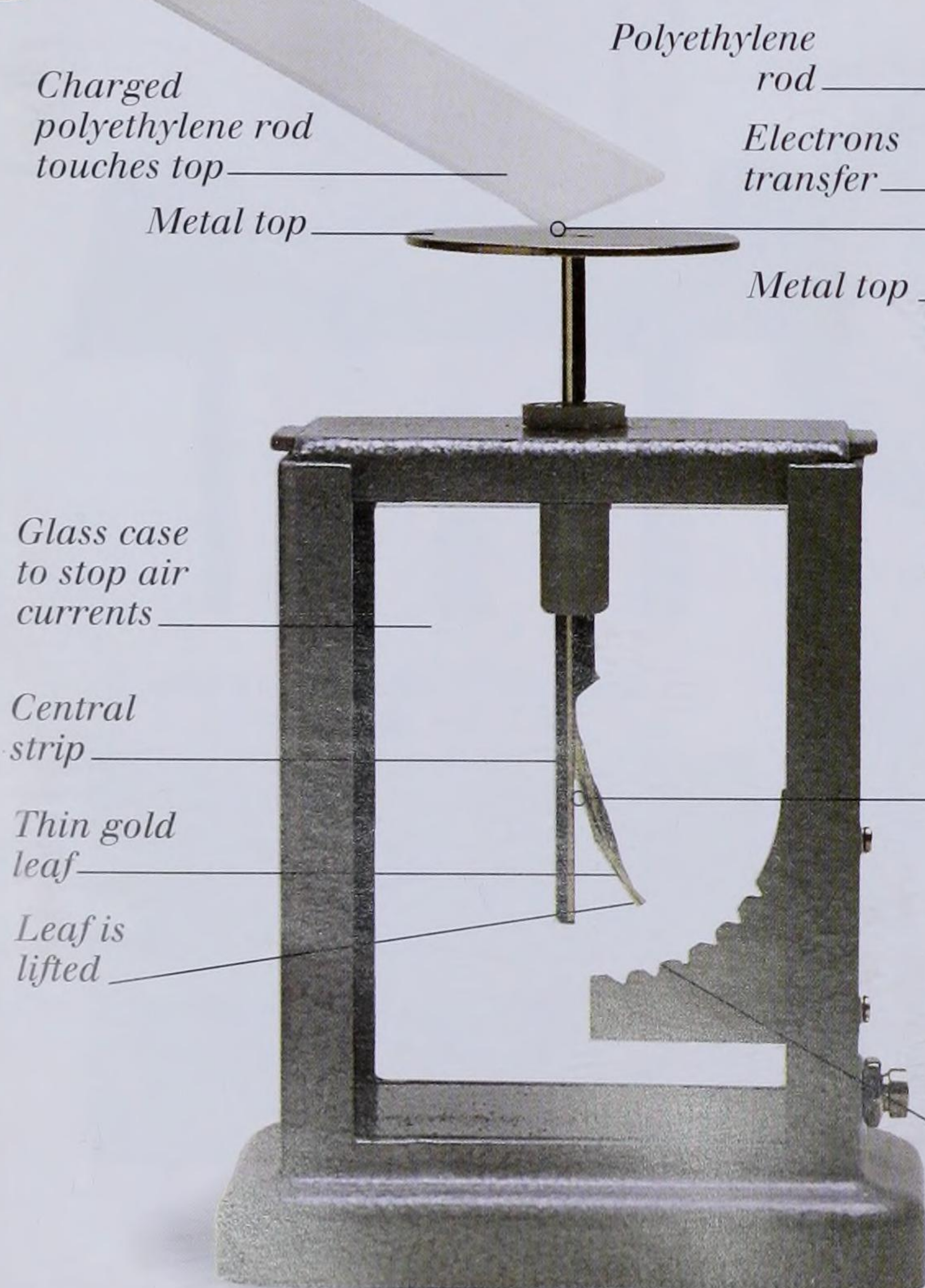
Charges of the same type repel, while charges of a different type attract. One way to think of an electric field is as a set of lines of force, as illustrated below.



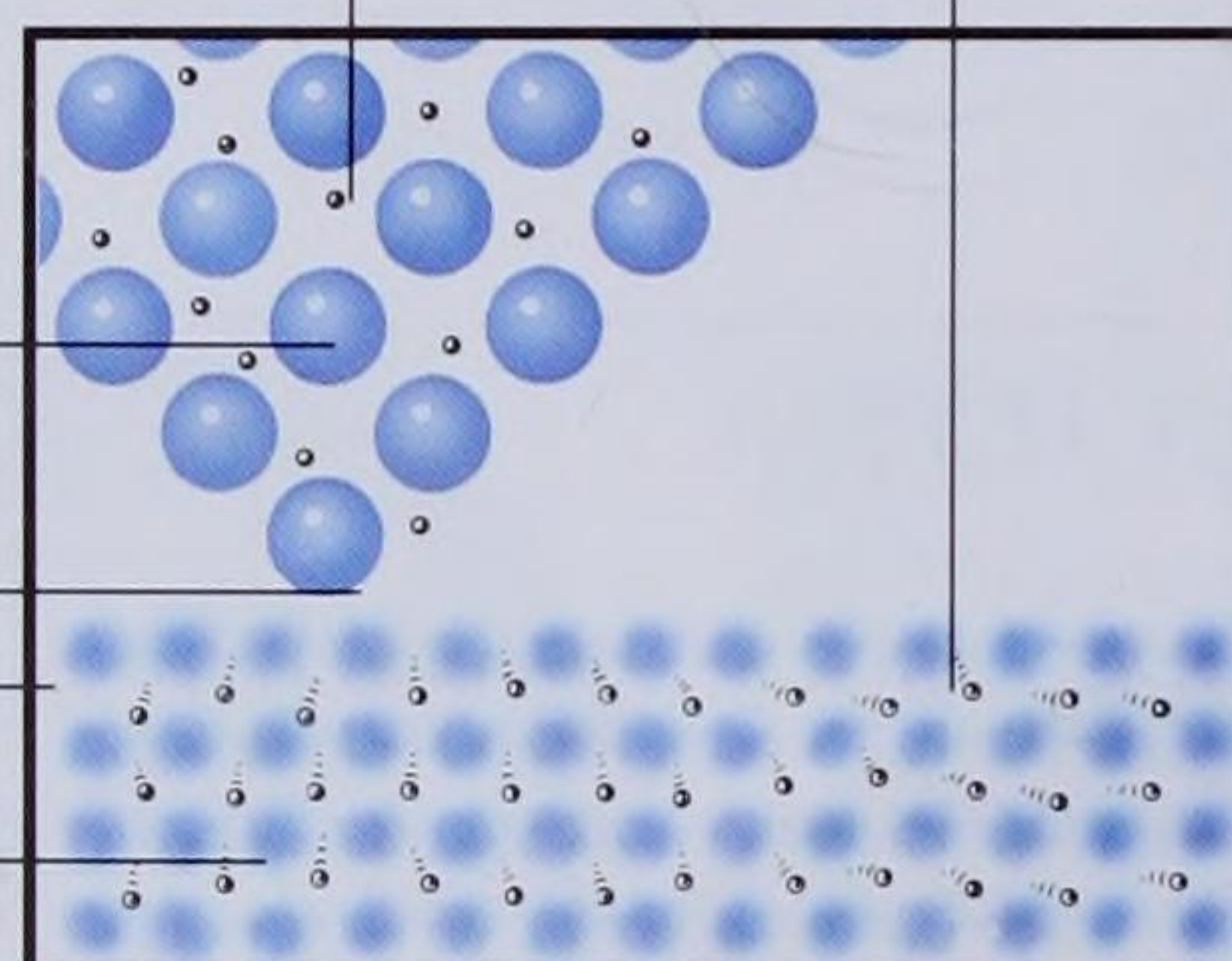
## STATIC ELECTRICITY

### GOLD LEAF ELECTROSCOPE

A polyethylene rod can gain extra electrons when it is rubbed. Touching the charged rod to the top of an electroscope causes electrons to move into the electroscope. The electrons in the central strip and in the gold leaf repel each other, and the leaf lifts.

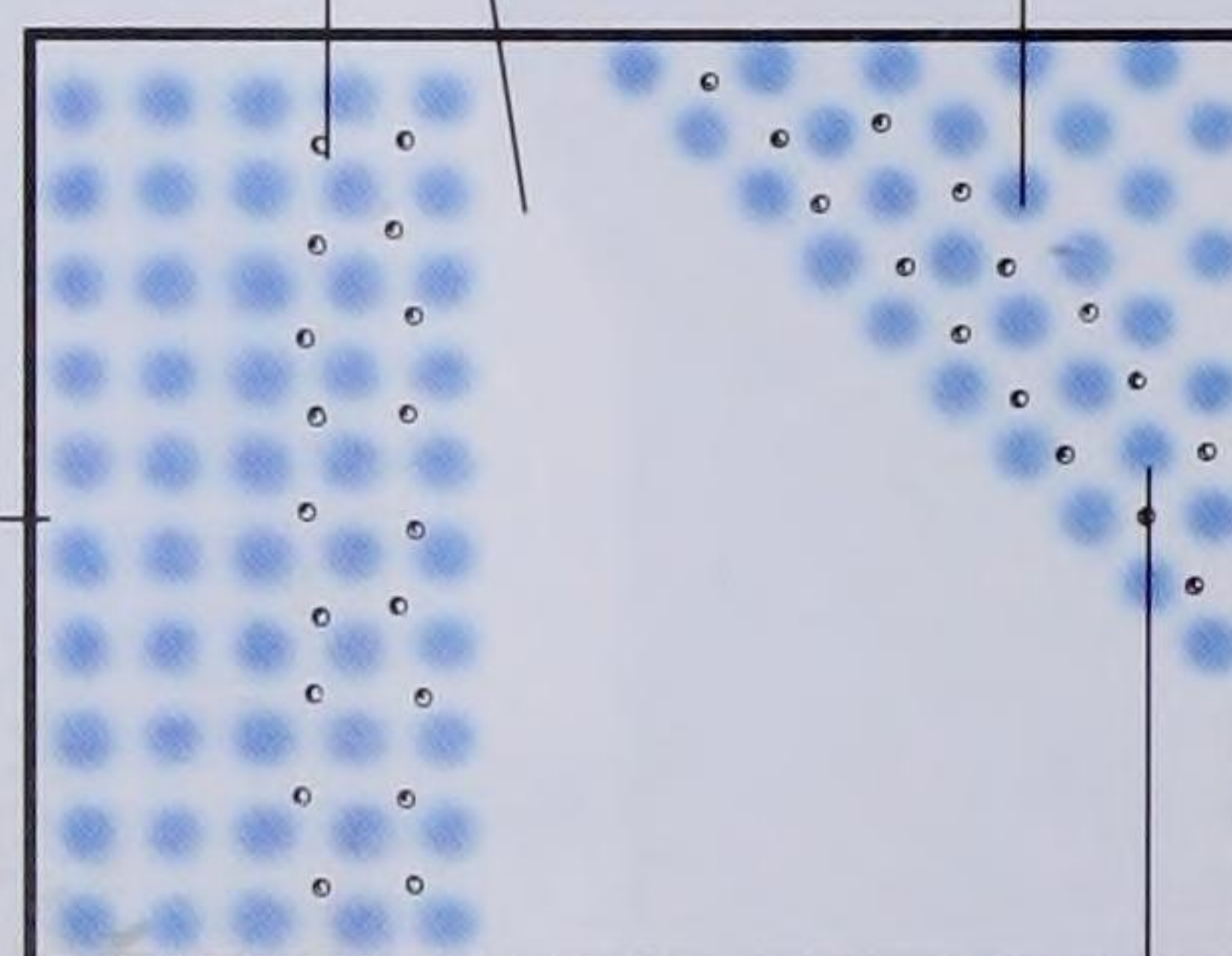


*Electrons pushed by extra electrons in rod*



### TRANSFER OF CHARGE

*Electrons push apart*

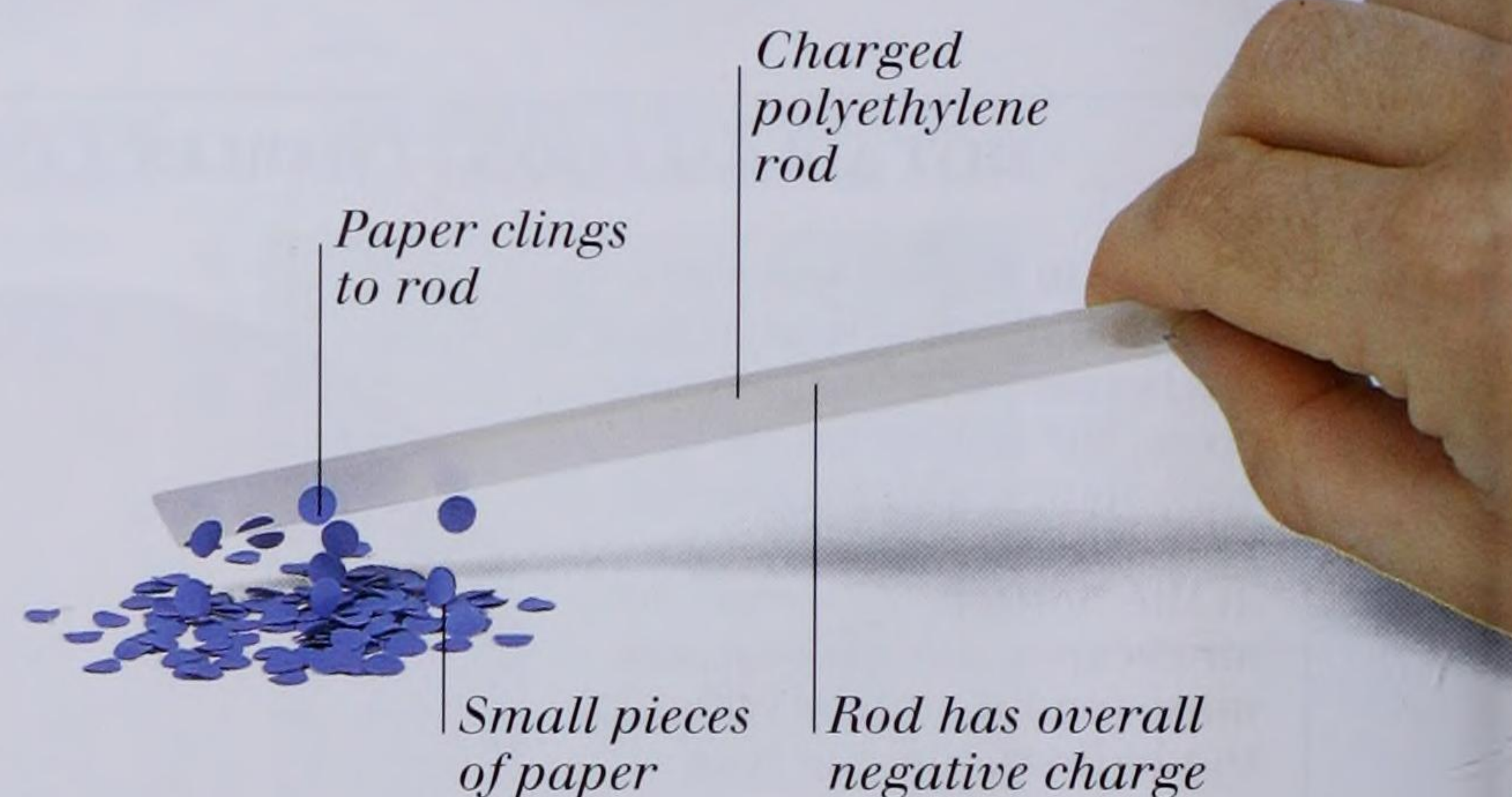


### REPULSIVE FORCE

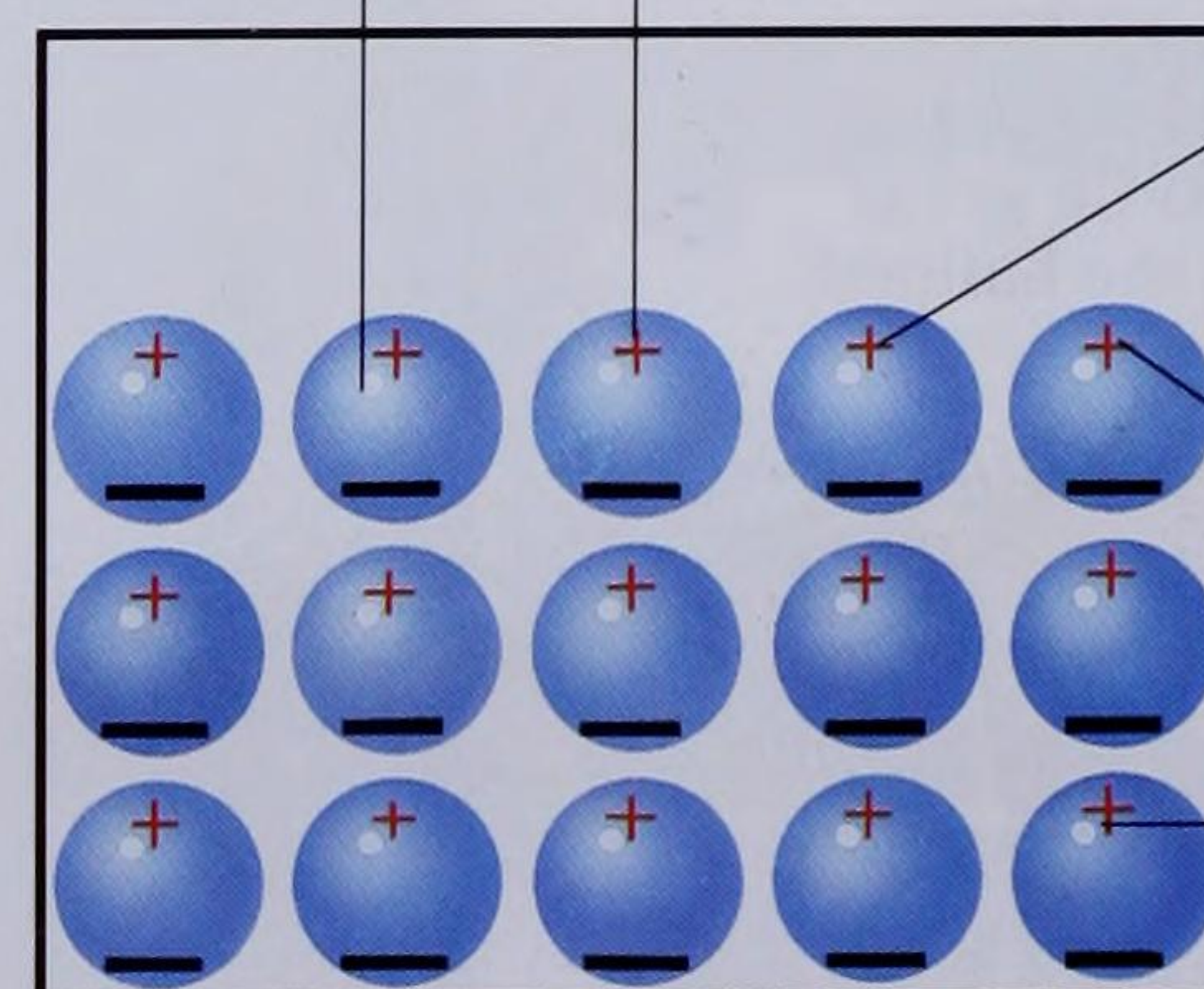
*Gold leaf*

### INDUCTION

When a charged object is brought near to other materials, such as paper, electrostatic forces cause a displacement of charge within that material. This is called induction. Negative charges in the paper are displaced so the edge of the paper nearest the rod becomes positively charged and clings to the negatively charged rod.



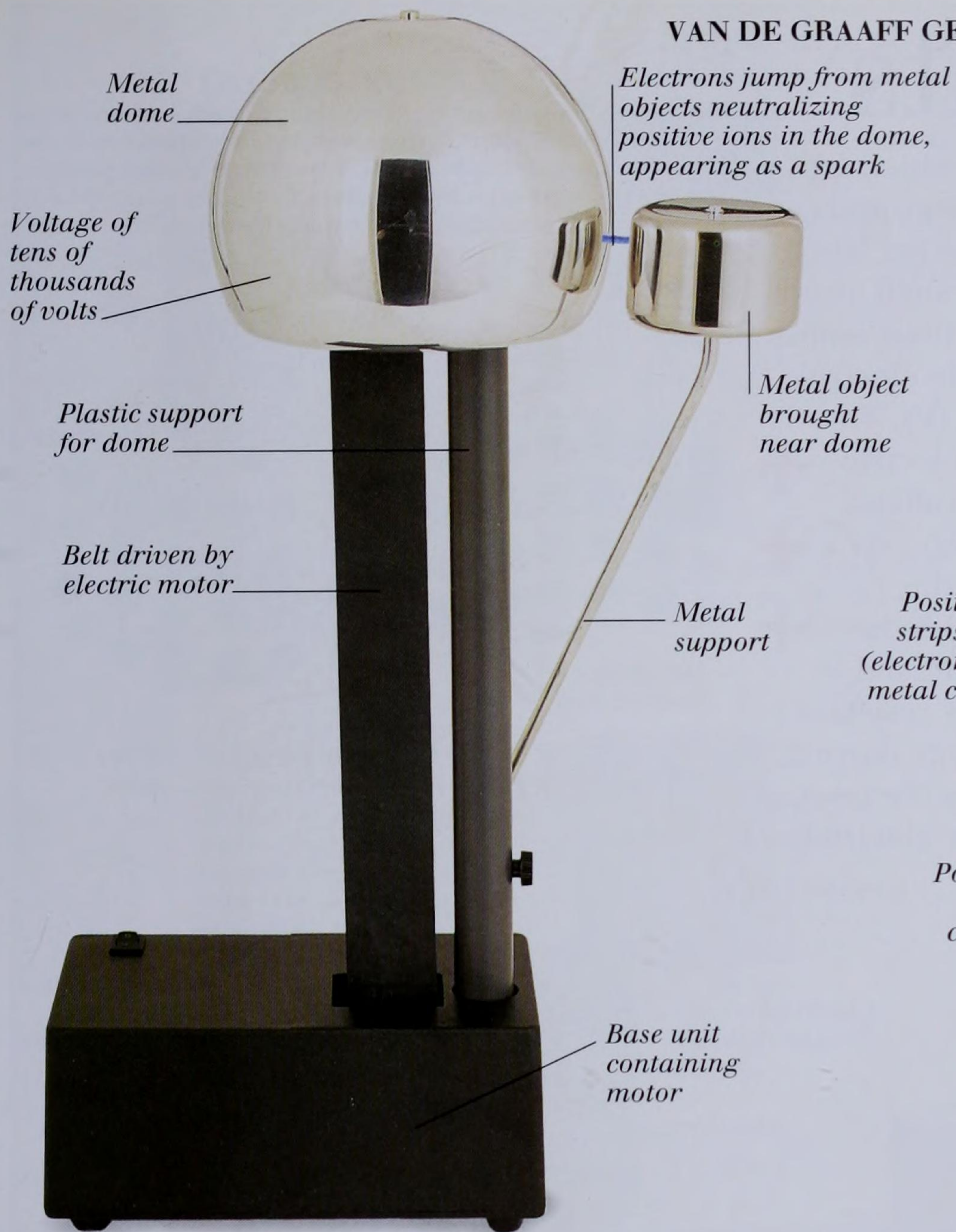
*Molecule in paper*



### INDUCTION IN PAPER

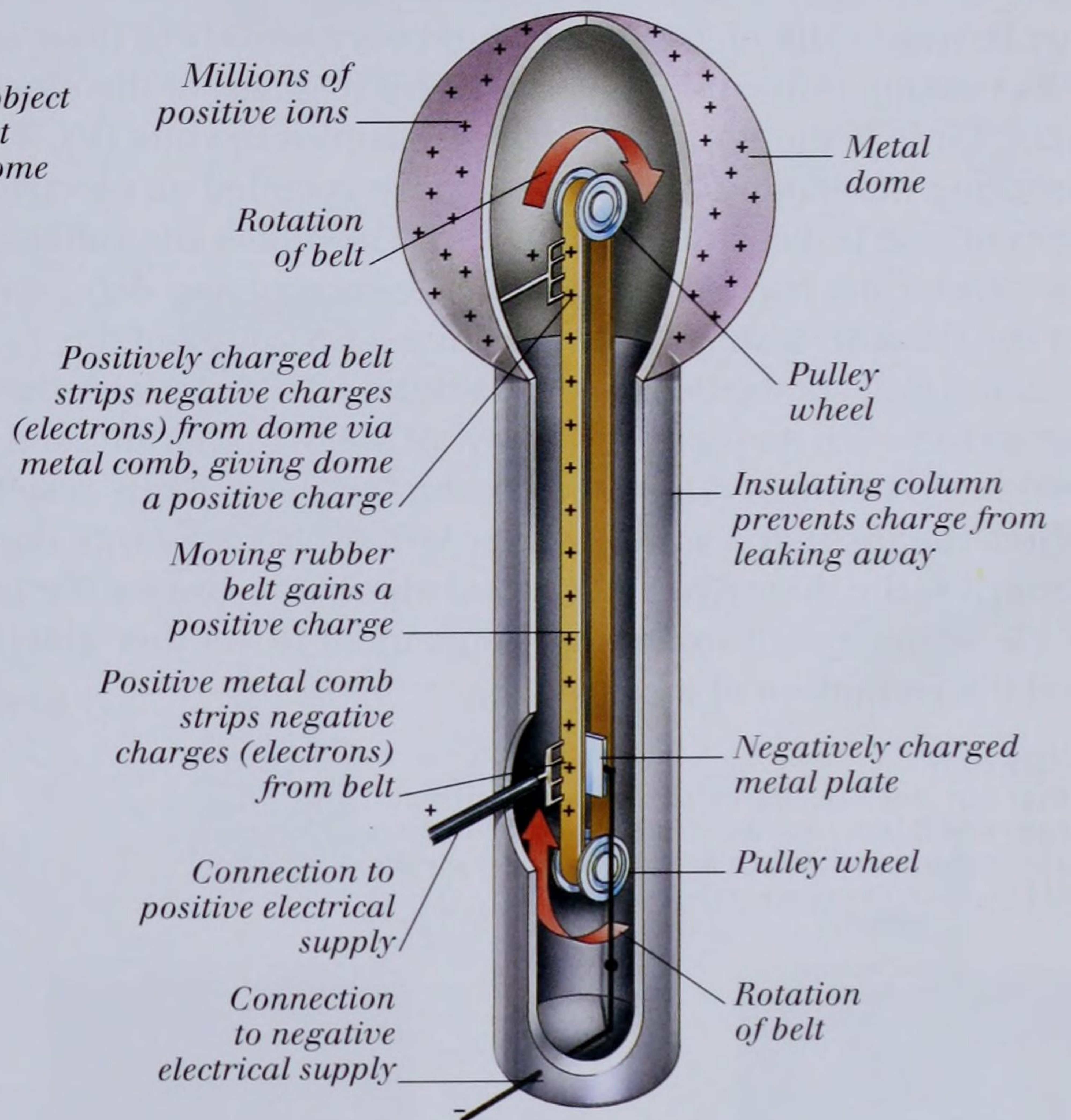


## VAN DE GRAAFF GENERATOR



### GENERATION OF IONS

A Van de Graaff generator separates electrons from the atoms of a moving belt. The positive ions created are carried upward by the belt, and take electrons from atoms of a metal dome. The electric field around the dome becomes very strong.



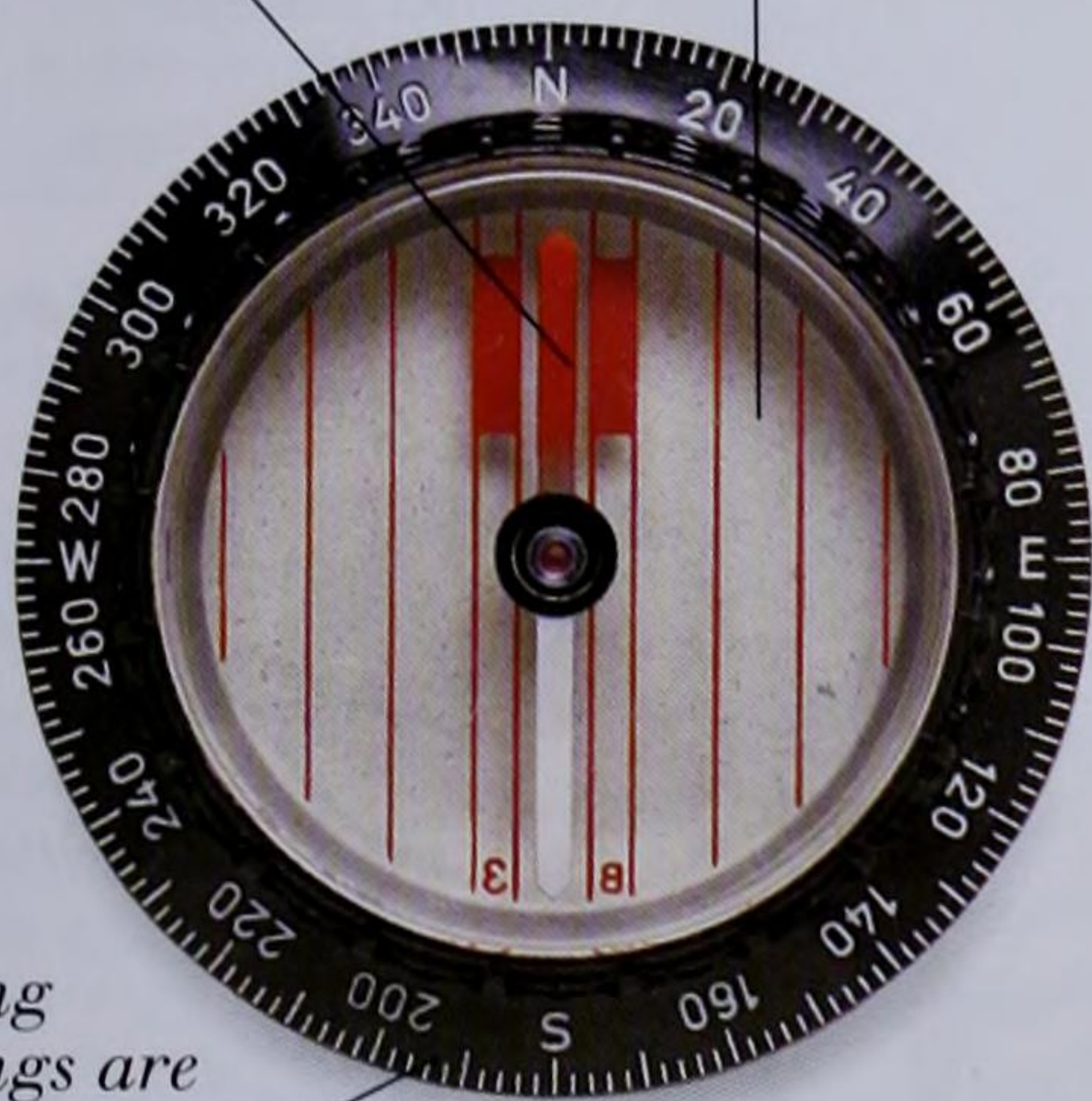
## MAGNETISM

### MAGNETIC COMPASS

Walkers and sailors use magnetic compasses to find their way. The needle of a compass lines up with the Earth's magnetic field, and always points North-South. The Earth's magnetism is thought to be caused by currents in its molten iron core.

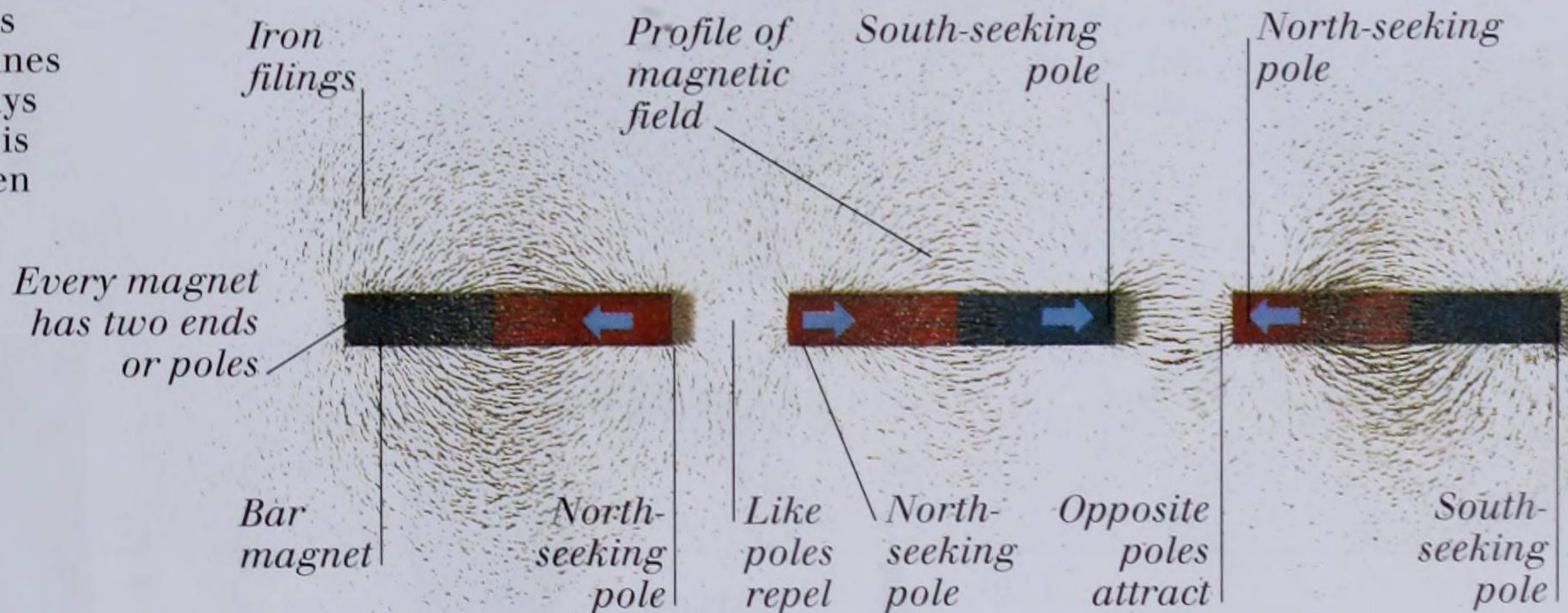
*Needle is a small magnet that is free to turn*

*Needle is suspended in fluid*



*Bearing readings are taken from this scale*

### MAGNETIC FIELDS AND FORCES

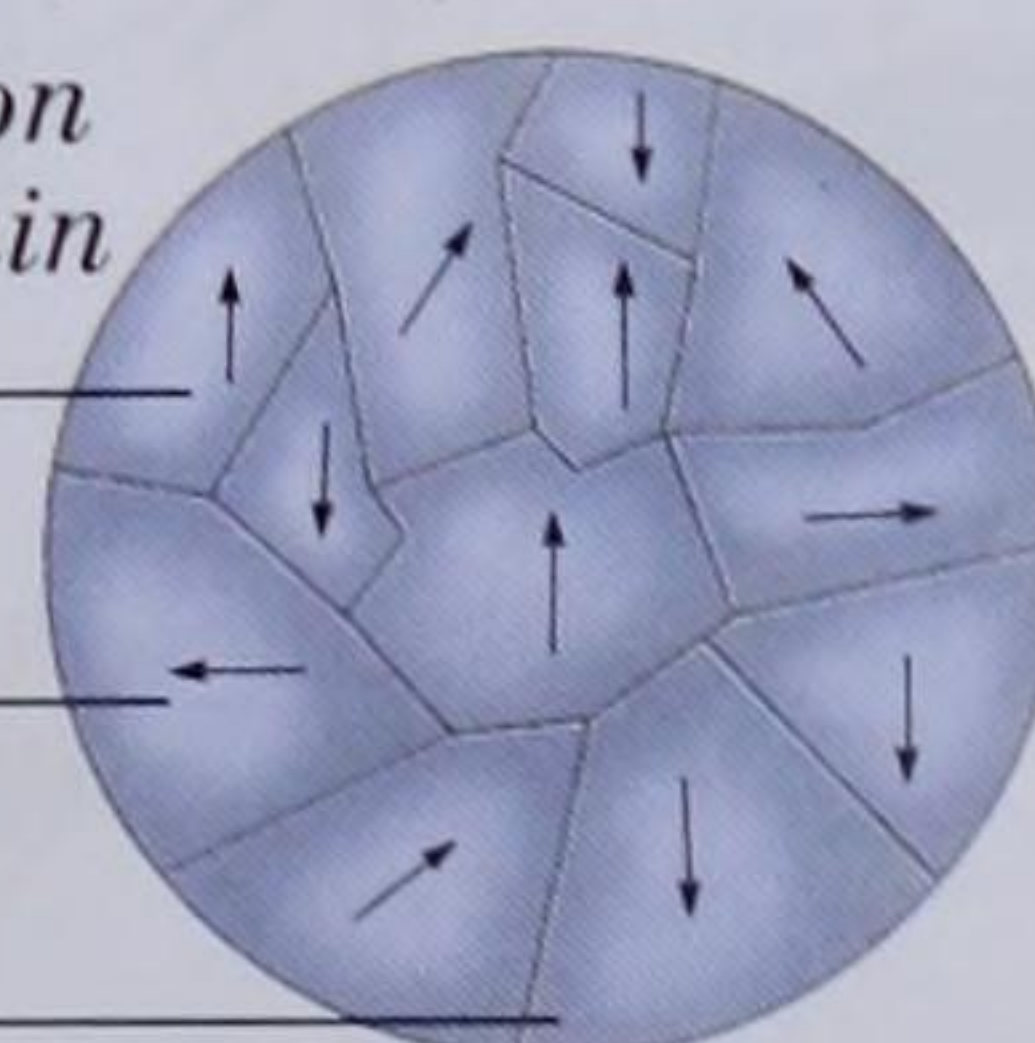


### MAGNET DOMAINS

*Direction of magnetization within domain is random*

*Domain*

*Domain boundary*



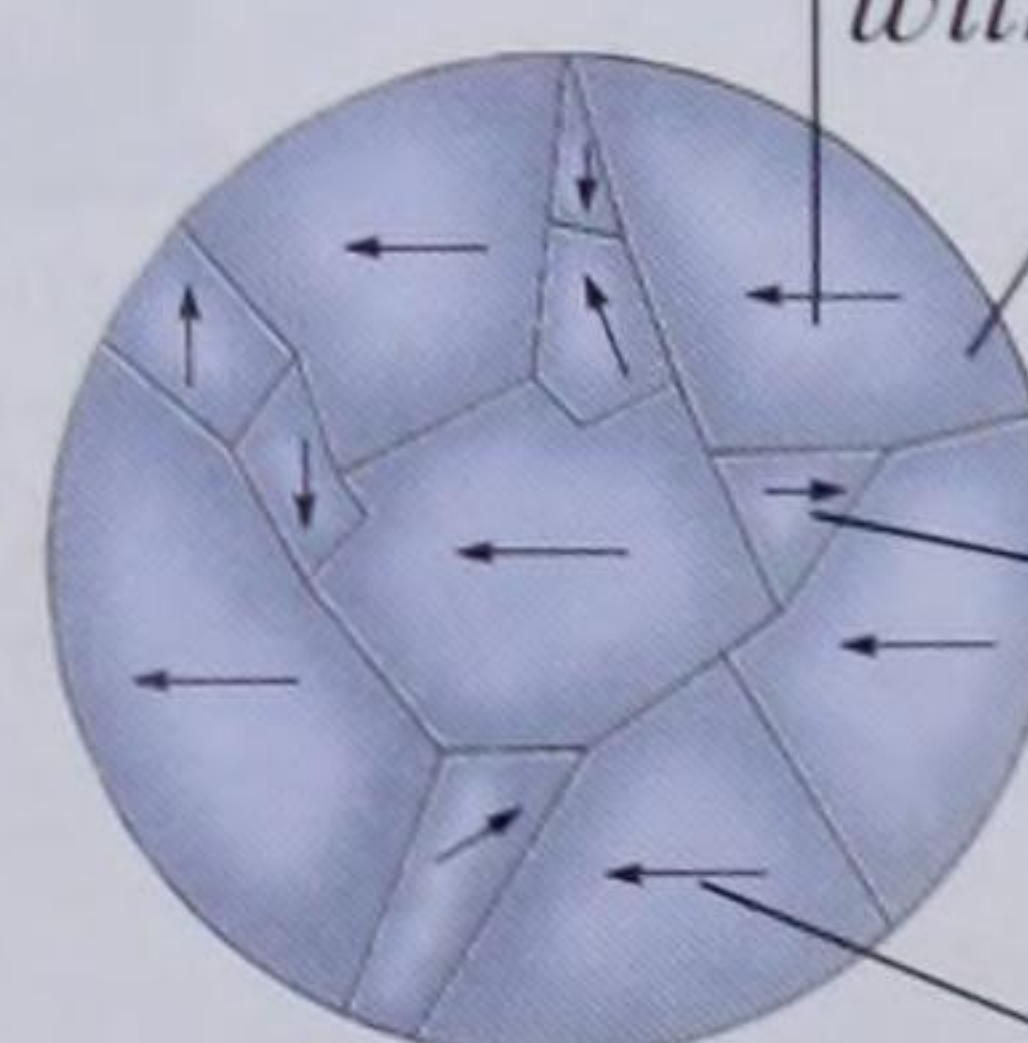
UNMAGNETIZED IRON

*Direction of magnetization within domain has aligned*

*Domain aligned with magnetization has grown*

*Domain not aligned with magnetization has shrunk*

*Direction of overall magnetization*



MAGNETIZED IRON



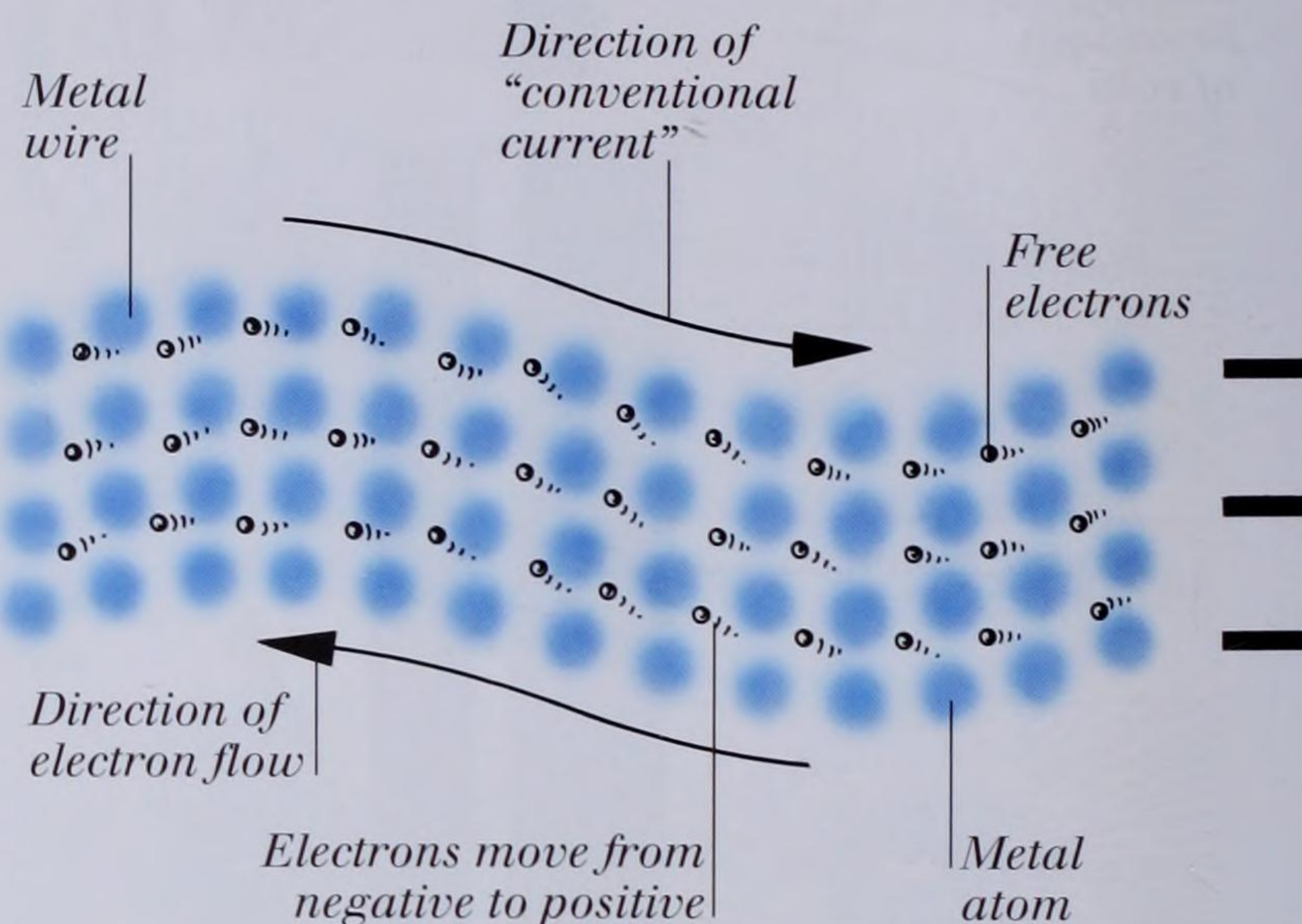
# Electric circuits

AN ELECTRIC CIRCUIT IS SIMPLY THE COURSE along which an **electric current** flows. **Electrons** carry negative charge and can be moved around a circuit by **electrostatic forces** (see pp. 30–31). A circuit usually consists of a **conductive** material, such as a metal, where the electrons are held very loosely to their atoms, thus making movement possible. The strength of the electrostatic force is the **voltage**, and is measured in volts (V). The resulting movement of electric charge is called an electric current, and is measured in amps (A). The higher the voltage, the greater the current will be. But the current also depends on the thickness, length, temperature, and nature of the material that conducts it. The **resistance** of a material is the extent to which it opposes the flow of electric current, and is measured in ohms ( $\Omega$ ). Good conductors have a low resistance, which means that a small voltage will produce a large current. In batteries, the dissolving of a metal **electrode** causes the freeing of electrons, resulting in their movement to another electrode and the formation of a current.



## ELECTRIC CURRENT

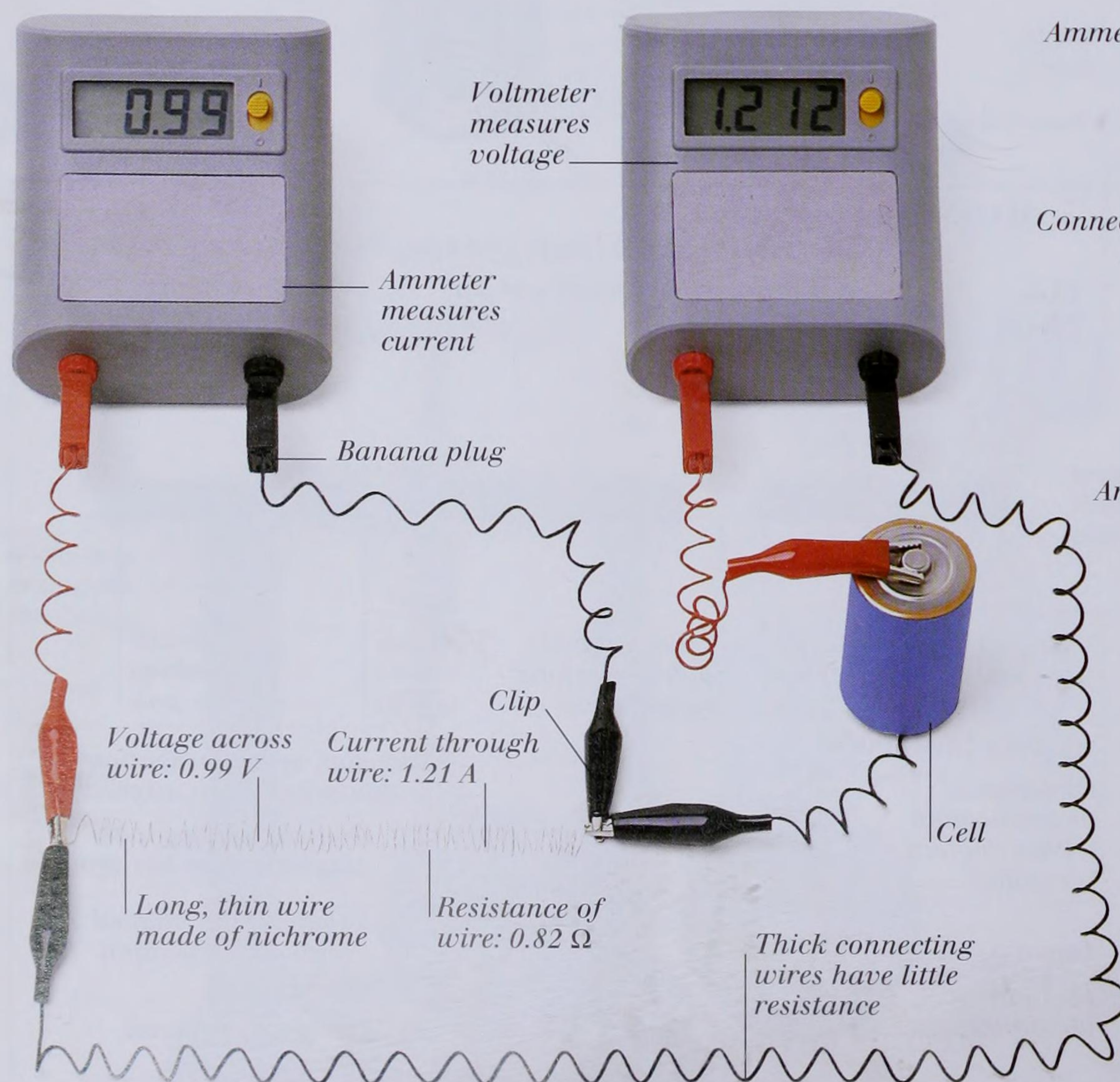
Regions of positive or negative charge, such as those at the terminals of a battery, force electrons through a conductor. The electrons move from negative charge toward positive. Originally, current was thought to flow from positive to negative. This is called “conventional current.”



## RESISTANCE

### OHM'S LAW

A thin wire has a resistance to the flow of current. The longer and thinner the wire, the higher the resistance. An object's resistance can be worked out by dividing the voltage by the current (see p. 54).



Current flowing through resistor: 0.18 A

Ammeter

### 22 $\Omega$ RESISTANCE

Electrical components called resistors allow current in circuits to be controlled. The current flowing around a circuit can be worked out using Ohm's Law.

Connecting wire

Negative terminal

Positive terminal

4.5 V battery

22  $\Omega$  resistor

Ammeter

### 47 $\Omega$ RESISTANCE

The larger the resistor, the smaller the current. The smaller the resistor, the larger the current.

Current flowing through resistor: 0.09 A

Negative terminal

Positive terminal

4.5 V battery

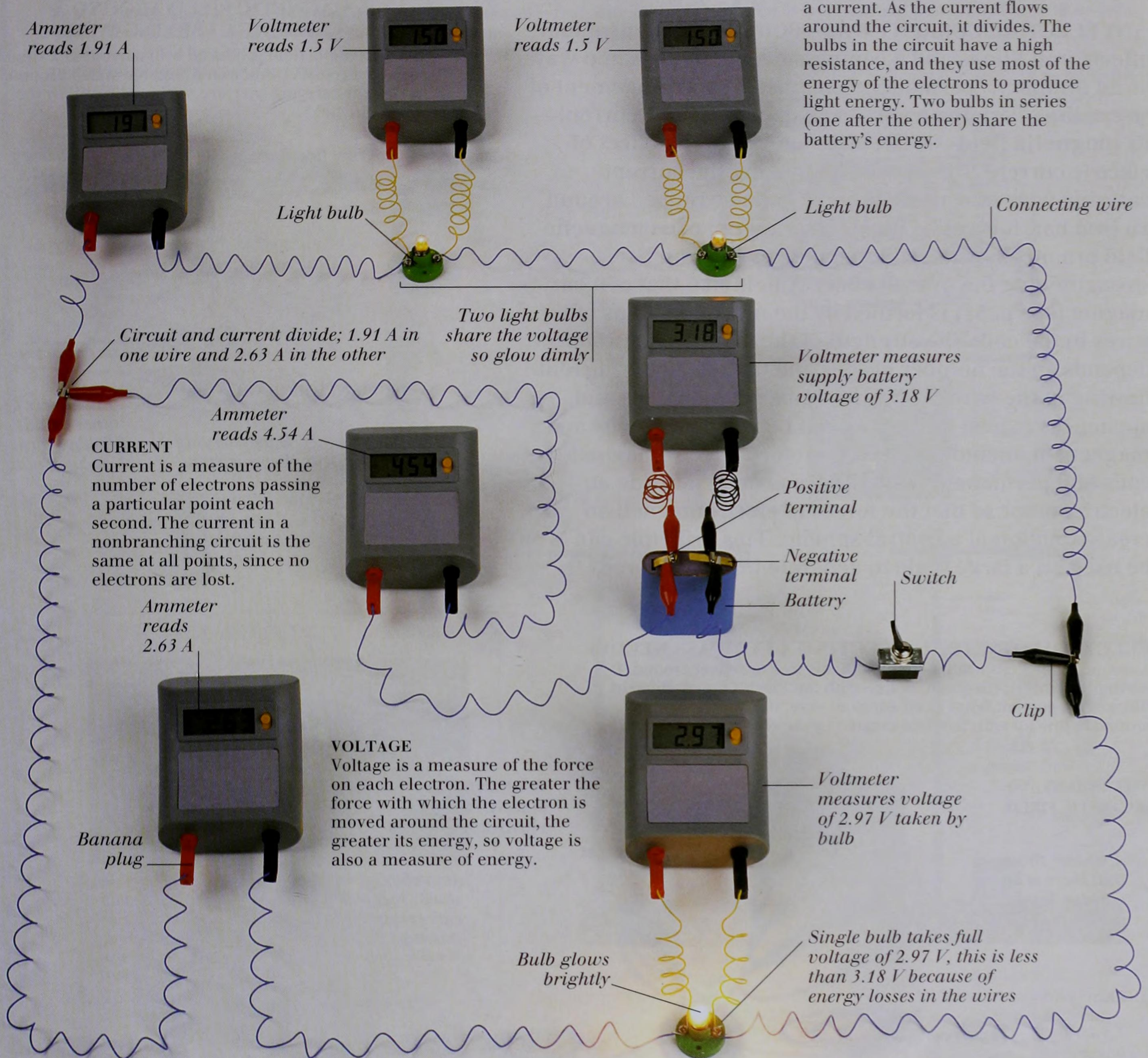
47  $\Omega$  resistor



## WORKING ELECTRIC CIRCUIT

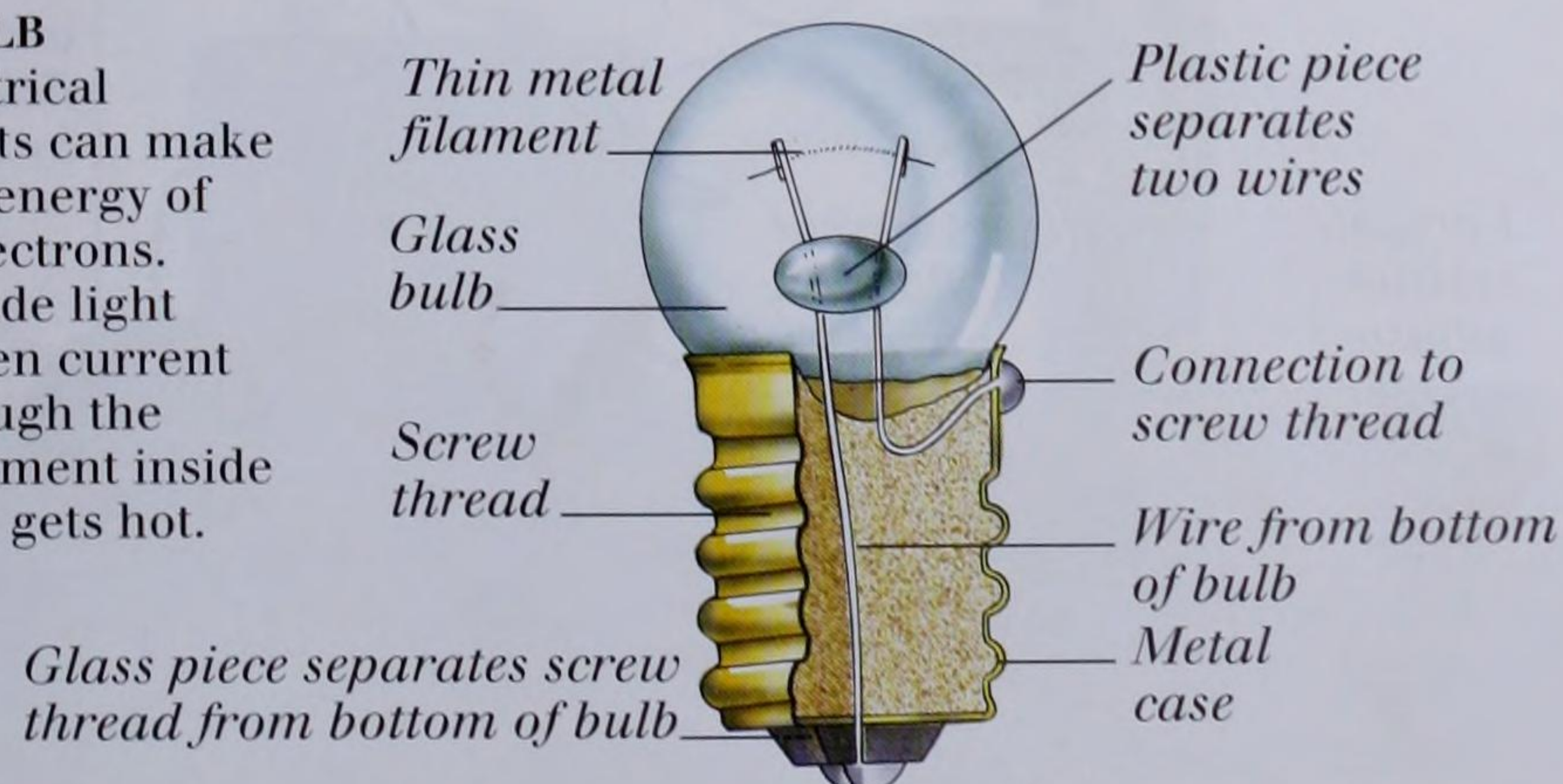
### BULBS IN A CIRCUIT

In this circuit, a 4.5V battery creates a current. As the current flows around the circuit, it divides. The bulbs in the circuit have a high resistance, and they use most of the energy of the electrons to produce light energy. Two bulbs in series (one after the other) share the battery's energy.



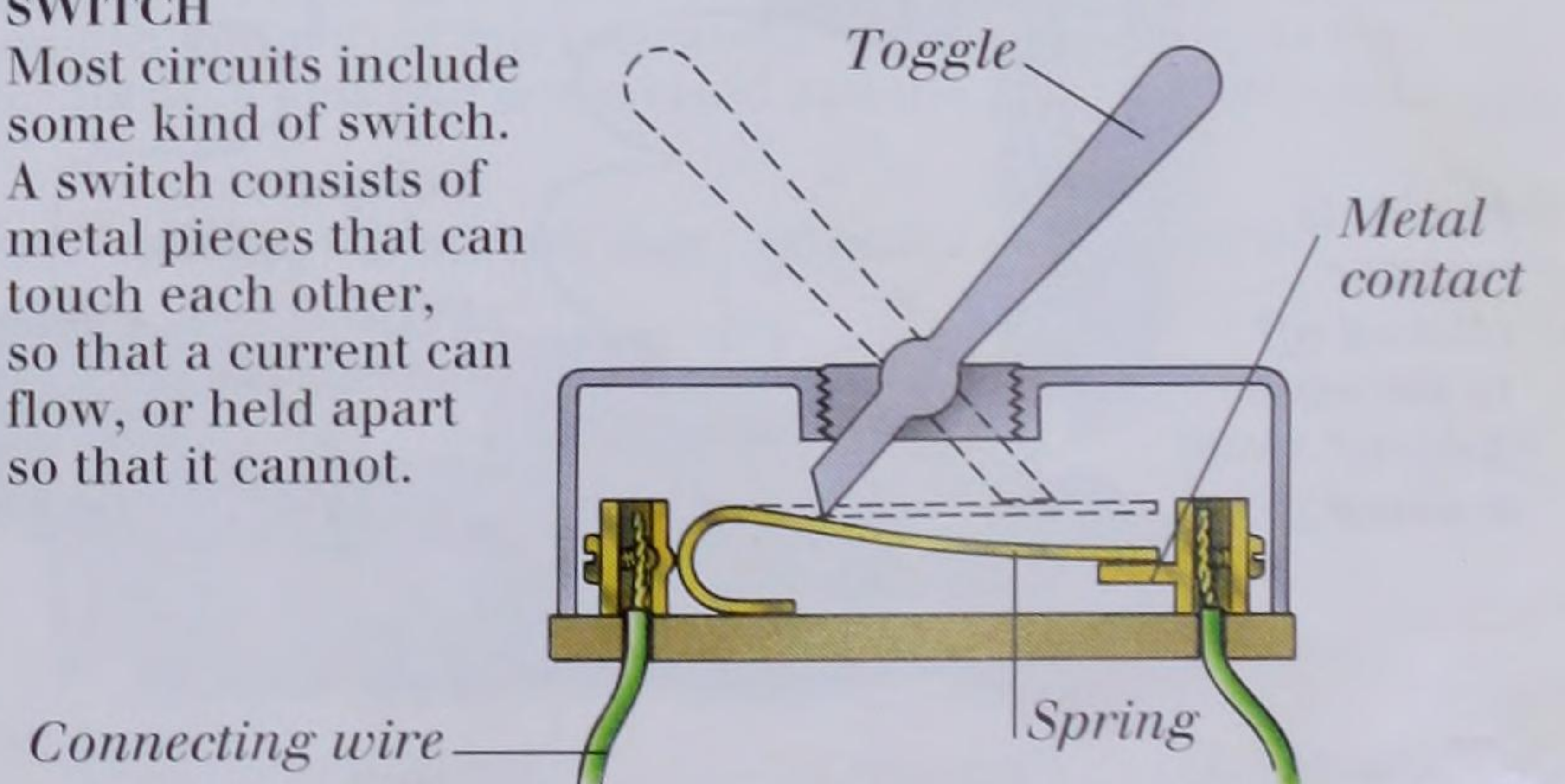
### LIGHT BULB

Many electrical components can make use of the energy of moving electrons. They include light bulbs. When current flows through the bulb, a filament inside glows as it gets hot.



### SWITCH

Most circuits include some kind of switch. A switch consists of metal pieces that can touch each other, so that a current can flow, or held apart so that it cannot.





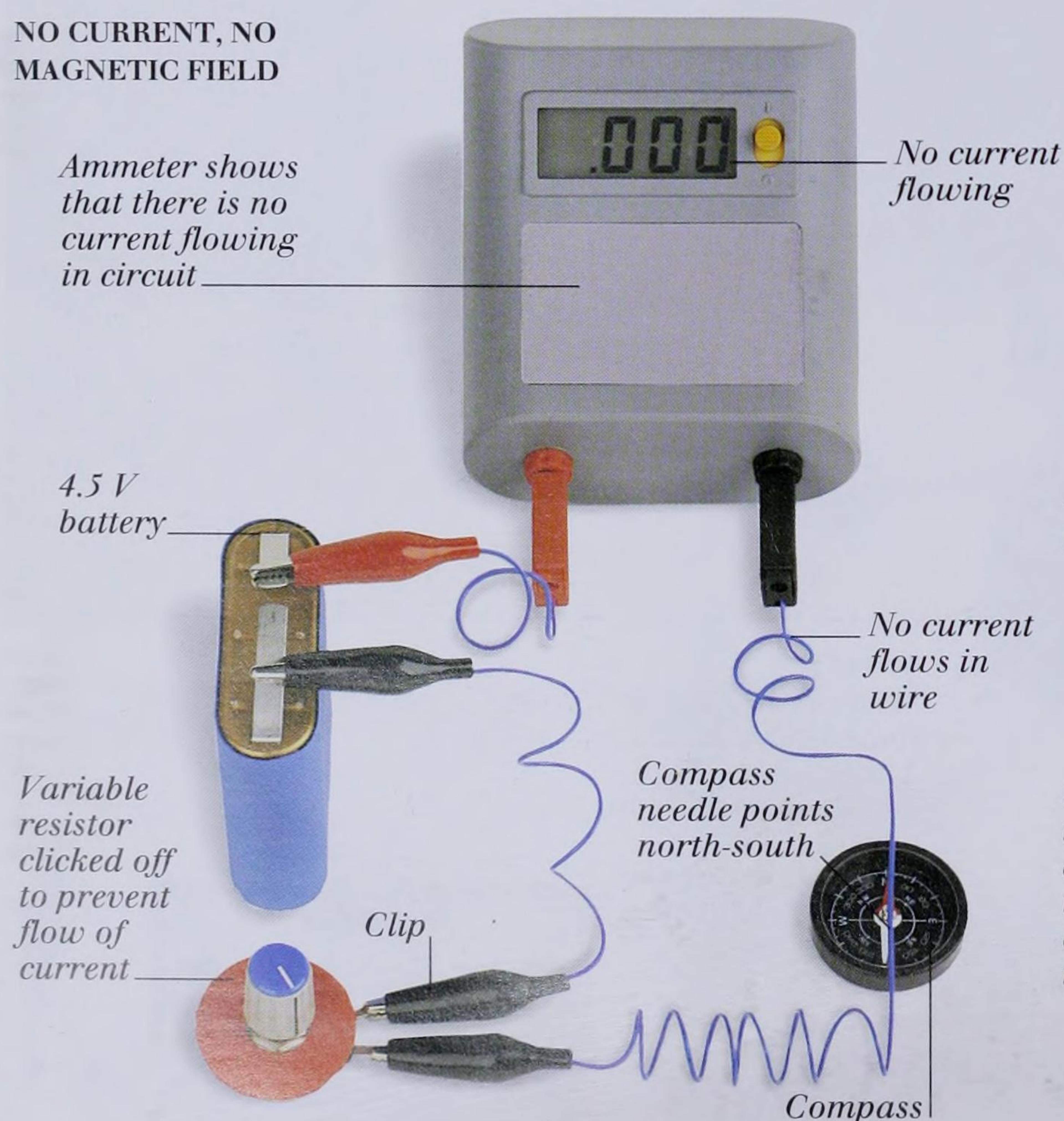
# Electromagnetism

ANY ELECTRIC CURRENT WILL PRODUCE magnetism that affects iron filings and a compass needle in the same way as an ordinary, “permanent” magnet. The arrangement of “force lines” around a wire carrying an electric current—its **magnetic field**—is circular. The magnetic effect of electric current is increased by making the current-carrying wire into a coil. When a coil is wrapped around an iron bar, it is called an **electromagnet**. The magnetic field produced by the coil magnetizes the iron bar, strengthening the overall effect. A field like that of a bar magnet (see p. 31) is formed by the magnetic fields of the wires in the coil. The strength of the magnetism produced depends on the number of coils and the size of the current flowing in the wires. A huge number of machines and appliances exploit the connection between electricity and magnetism, including electric motors. Electromagnetic coils and permanent magnets are arranged inside an electric motor so that the forces of electromagnetism create rotation of a central spindle. This principle can be used on a large scale to generate immense forces.

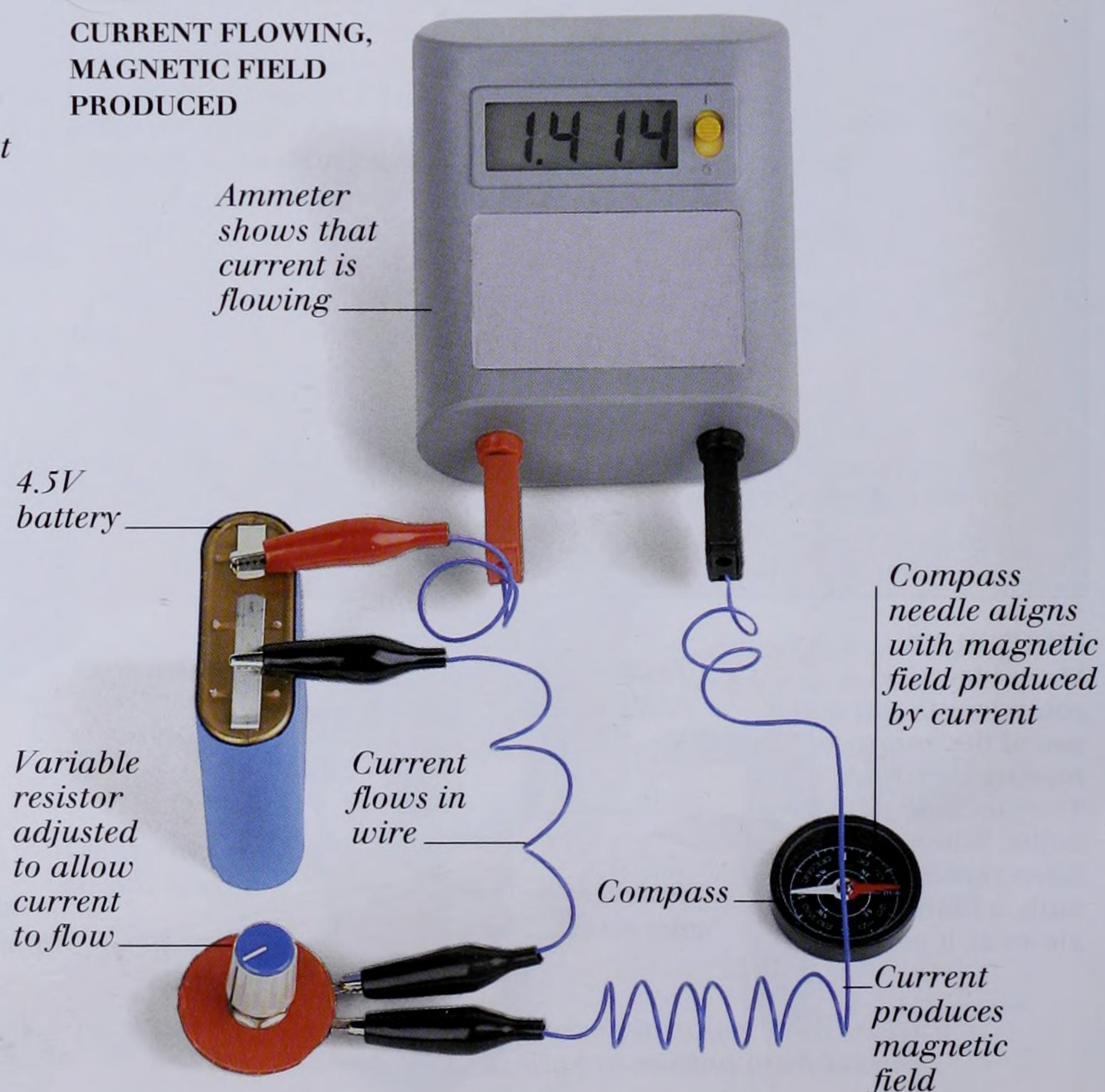
## ELECTROMAGNETISM AFFECTING A COMPASS NEEDLE

A compass needle is a small magnet that is free to swivel around. It normally points north-south, in line with the Earth’s magnetic field. But when a current flows in an adjacent wire, the needle swings around to line up with the field created by the current.

### NO CURRENT, NO MAGNETIC FIELD

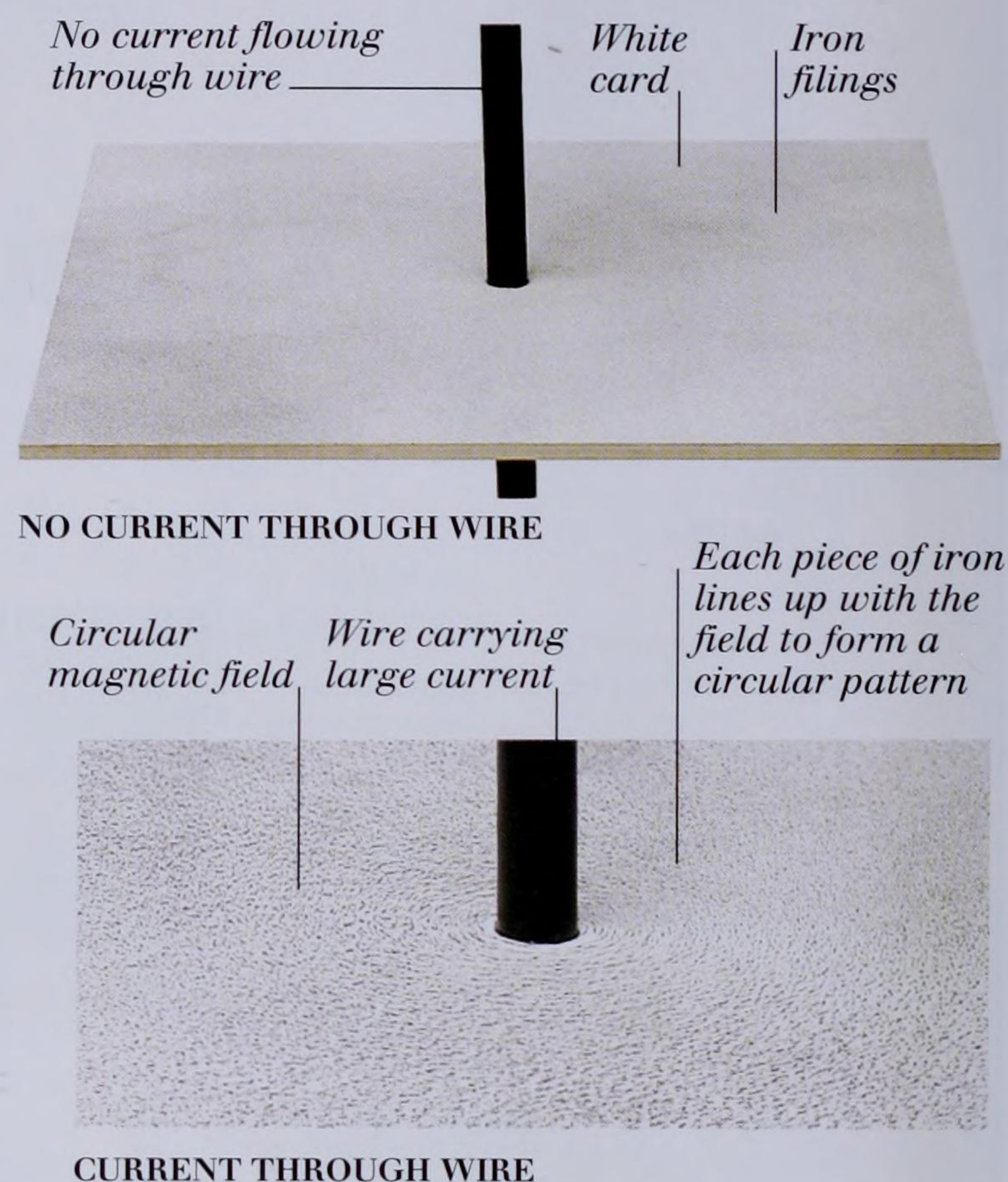


### CURRENT FLOWING, MAGNETIC FIELD PRODUCED



## MAGNETIC FIELD AROUND A CURRENT-CARRYING WIRE

The magnetic field produced by a current in a single wire is circular. Here, iron filings sprinkled around a current-carrying wire are made to line up by the magnetic field.

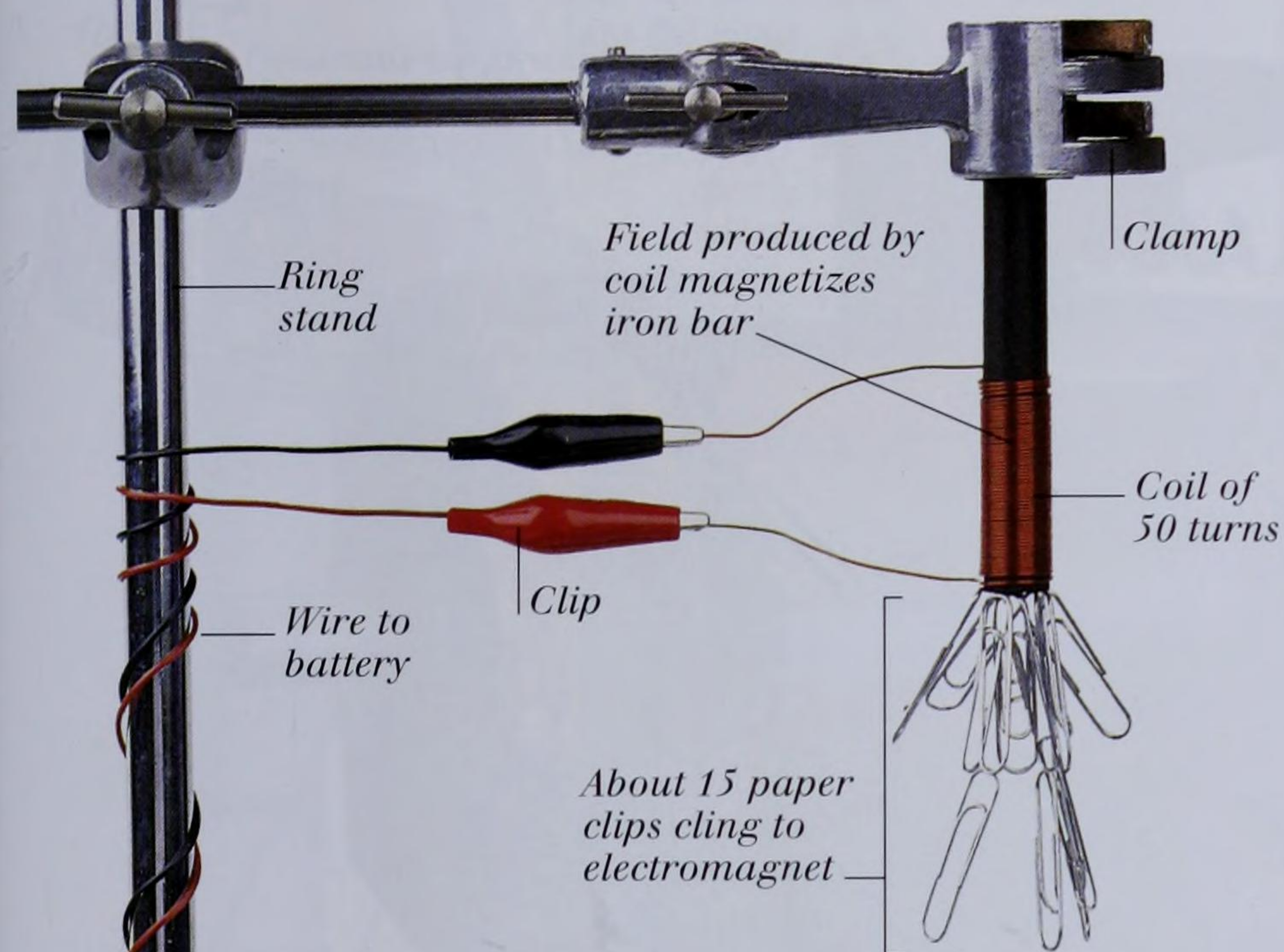




## ELECTROMAGNETS

### THE STRENGTH OF AN ELECTROMAGNET

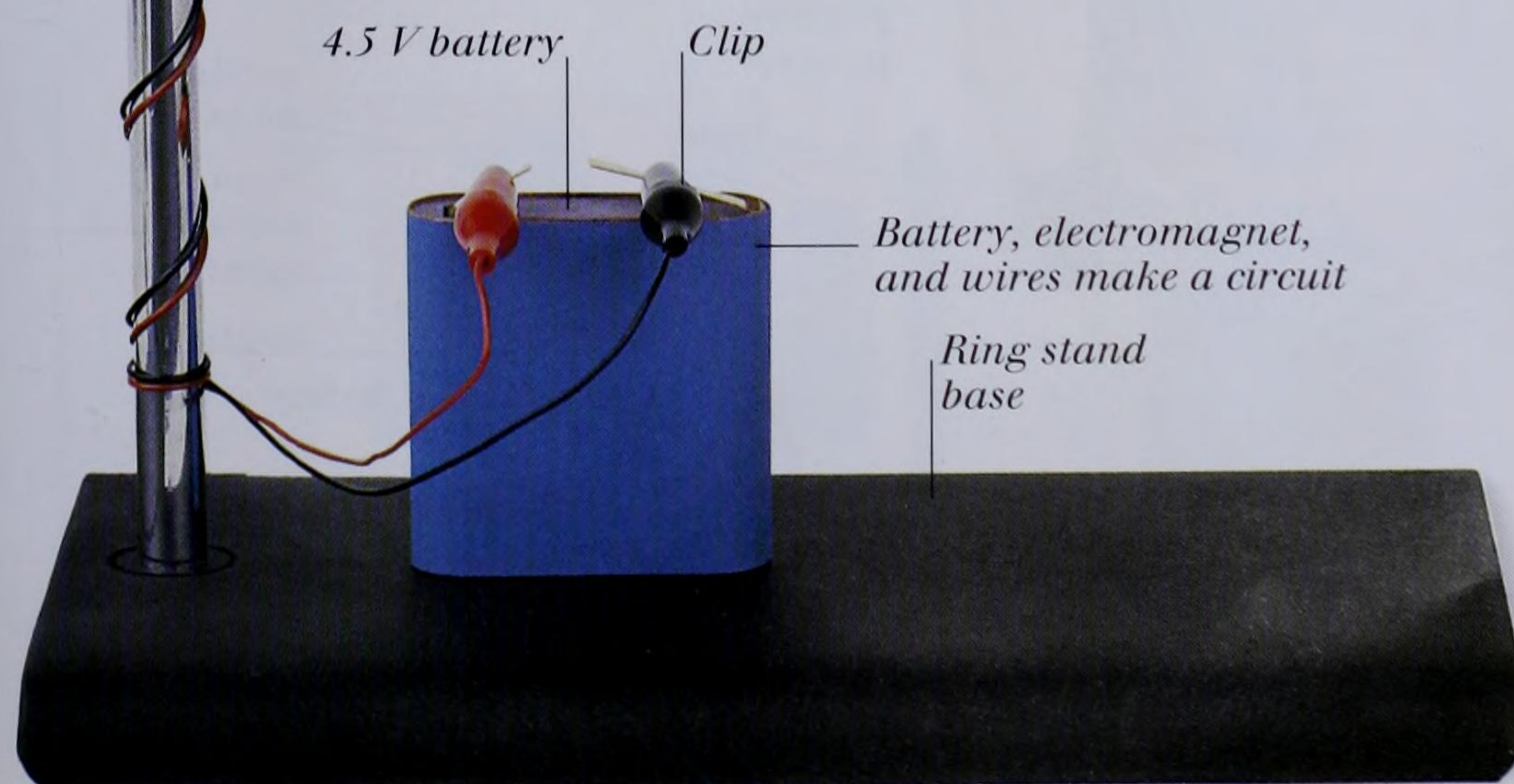
An electromagnet is a coil of wire wrapped around an iron bar. It behaves like a permanent magnet, except that it can be turned off. Here, the size of the magnetic force produced by an electromagnet is measured by the number of paper clips it can lift. The strength of an electromagnet depends on the number of turns in the coil and the current flowing through the wire.



### SCRAPYARD ELECTROMAGNET

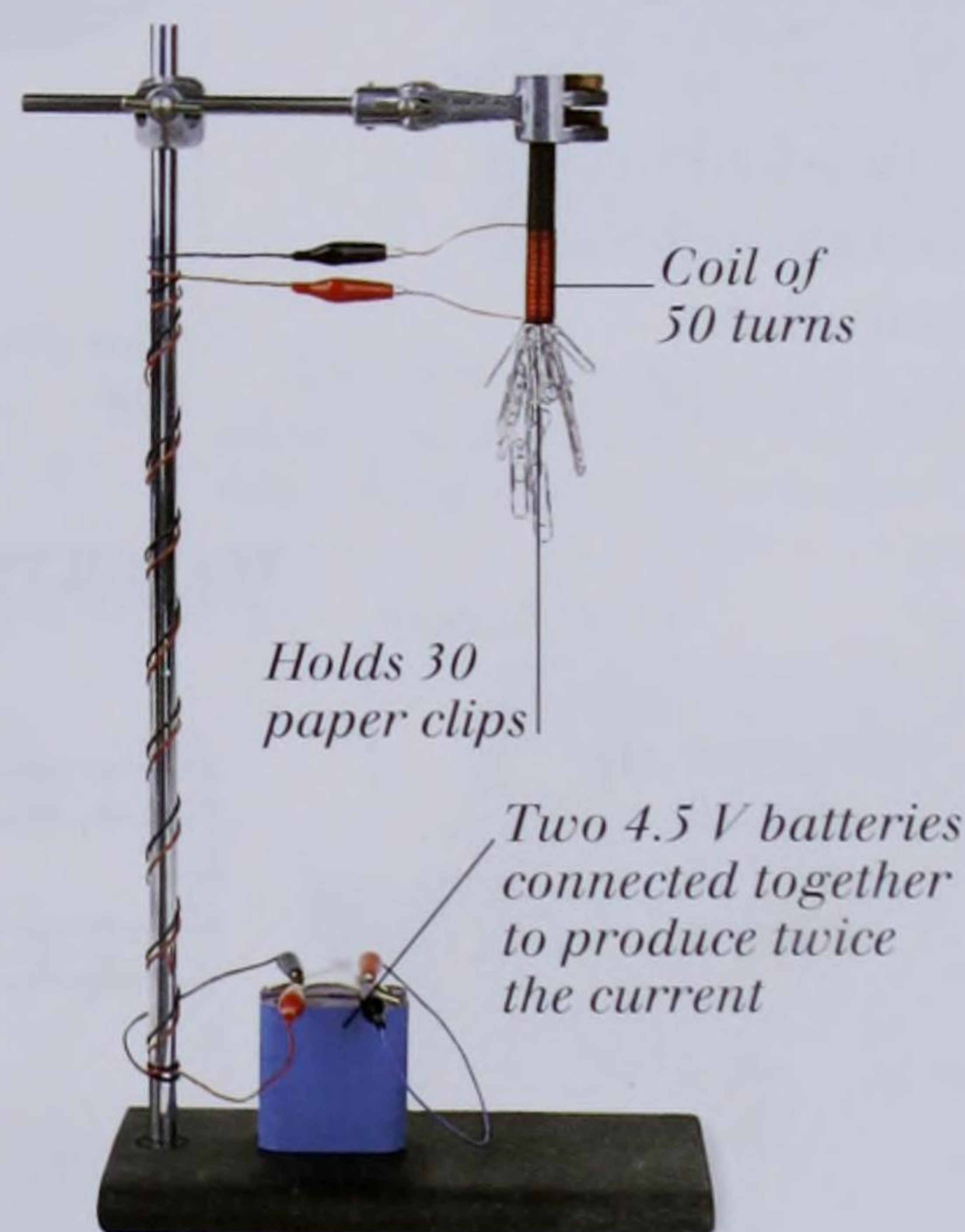
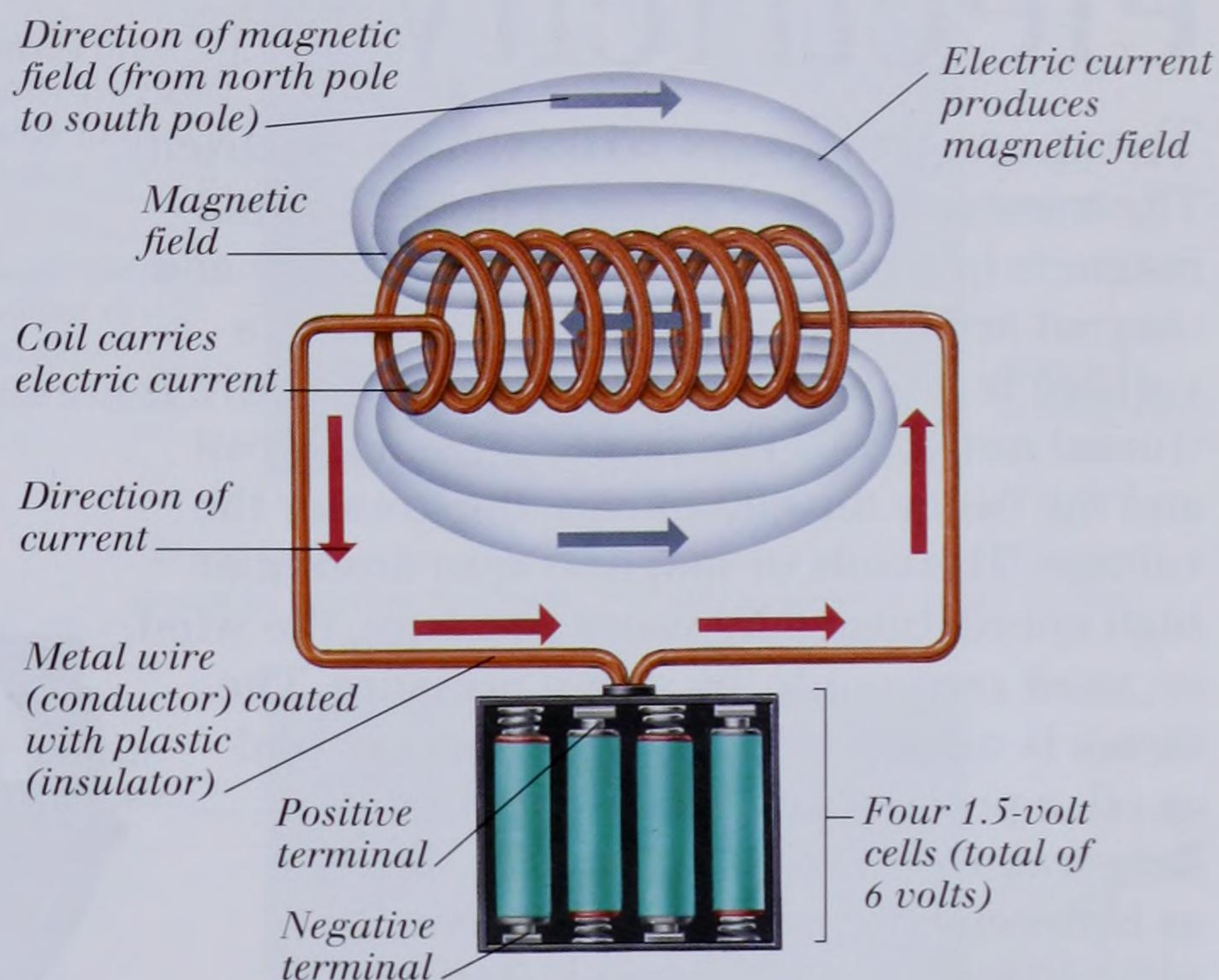


An electromagnetic crane picks up scrap metal using a powerful electromagnet. The electromagnet is switched on, scrap metal containing iron clings to it, and can be moved around. The metal is dropped by switching the magnet off.

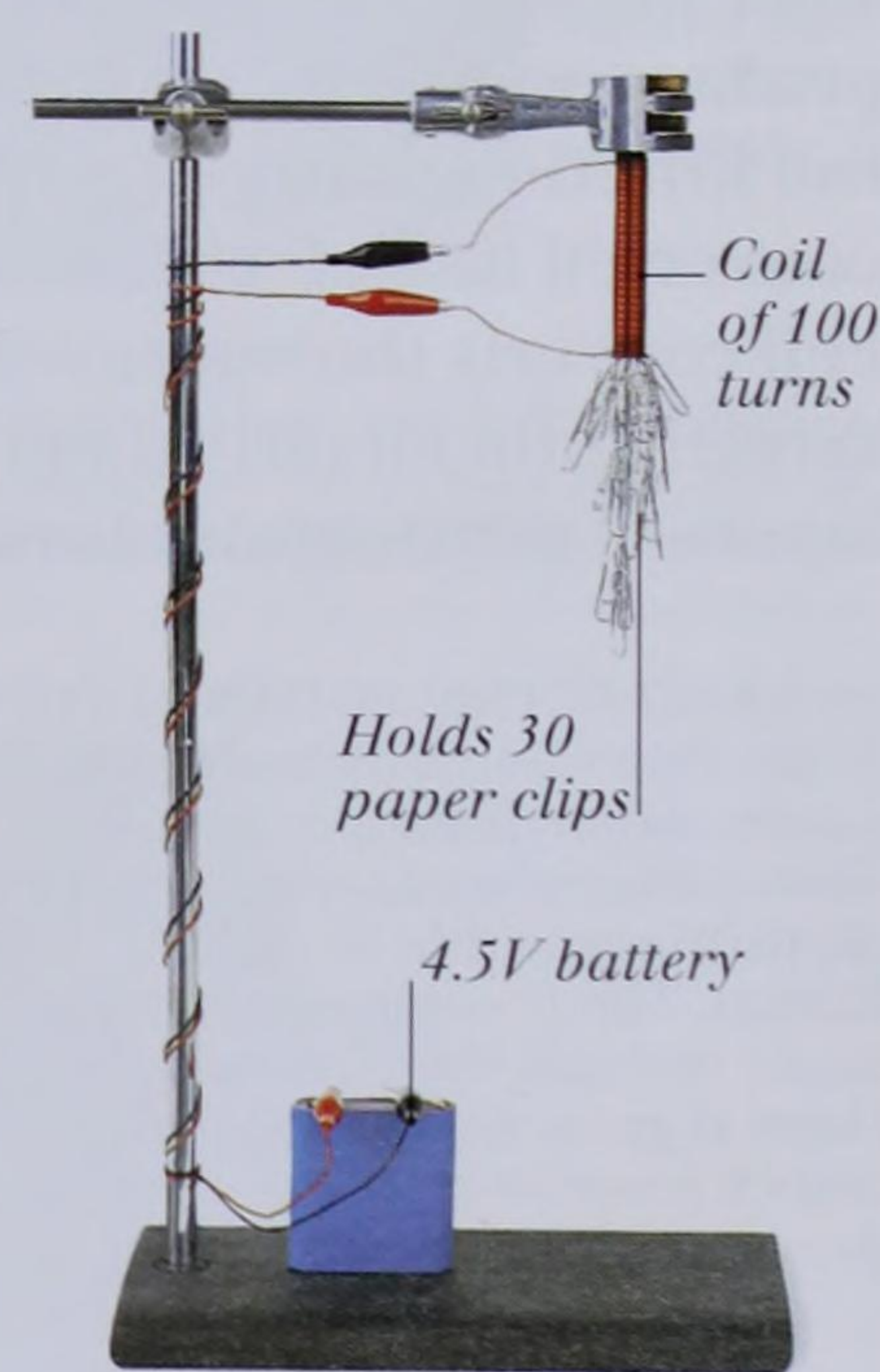


### A SOLENOID

The magnetic field around a coil of current-carrying wire resembles that around an ordinary bar magnet. The fields of each individual wire add up to give the overall pattern. A coil like this, with no iron bar at its core, is called a **solenoid**.



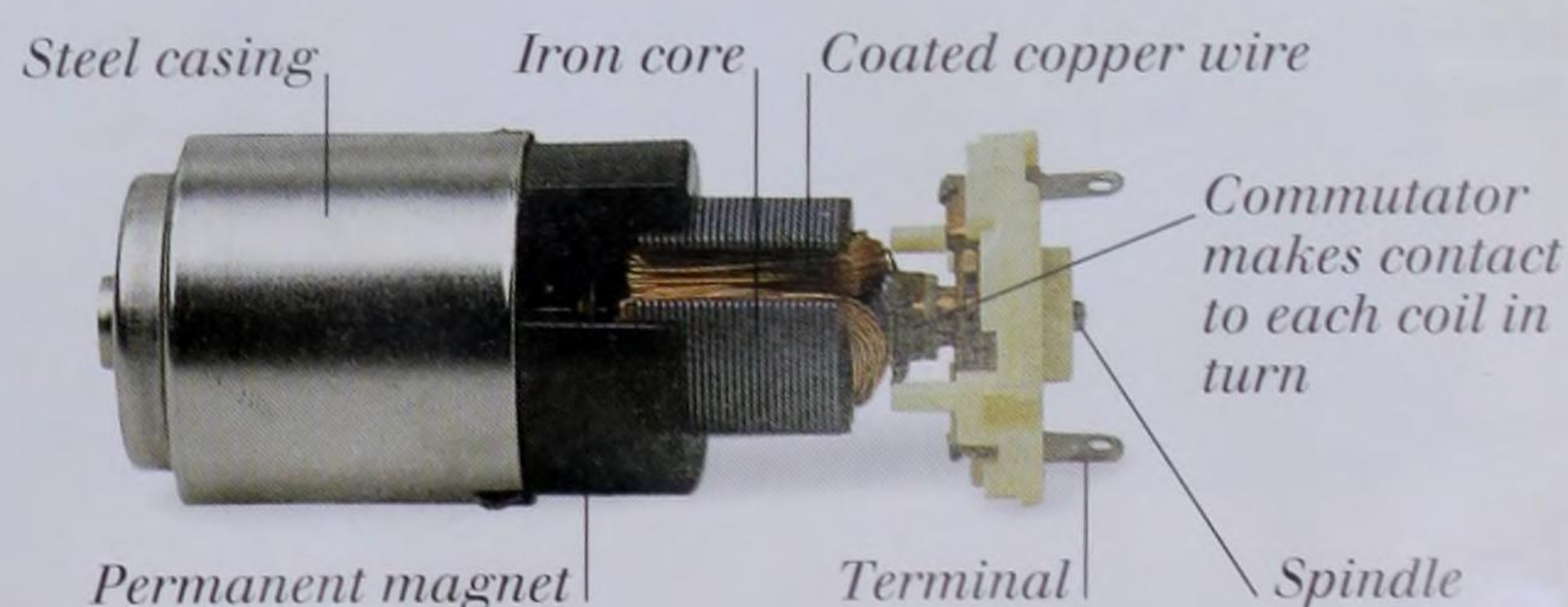
EFFECT OF DOUBLING CURRENT



EFFECT OF DOUBLING NUMBER OF TURNS ON COIL

### ELECTRIC MOTORS

Inside the motor, an electric current is sent through a series of wire coils one by one, providing a magnetic field around each coil, one after the other. The magnetism of the coils interacts with the magnetic fields of permanent magnets placed around them. The push and pull of this interaction turns the motor. As the rotor turns, a new coil is activated and the motion continues.



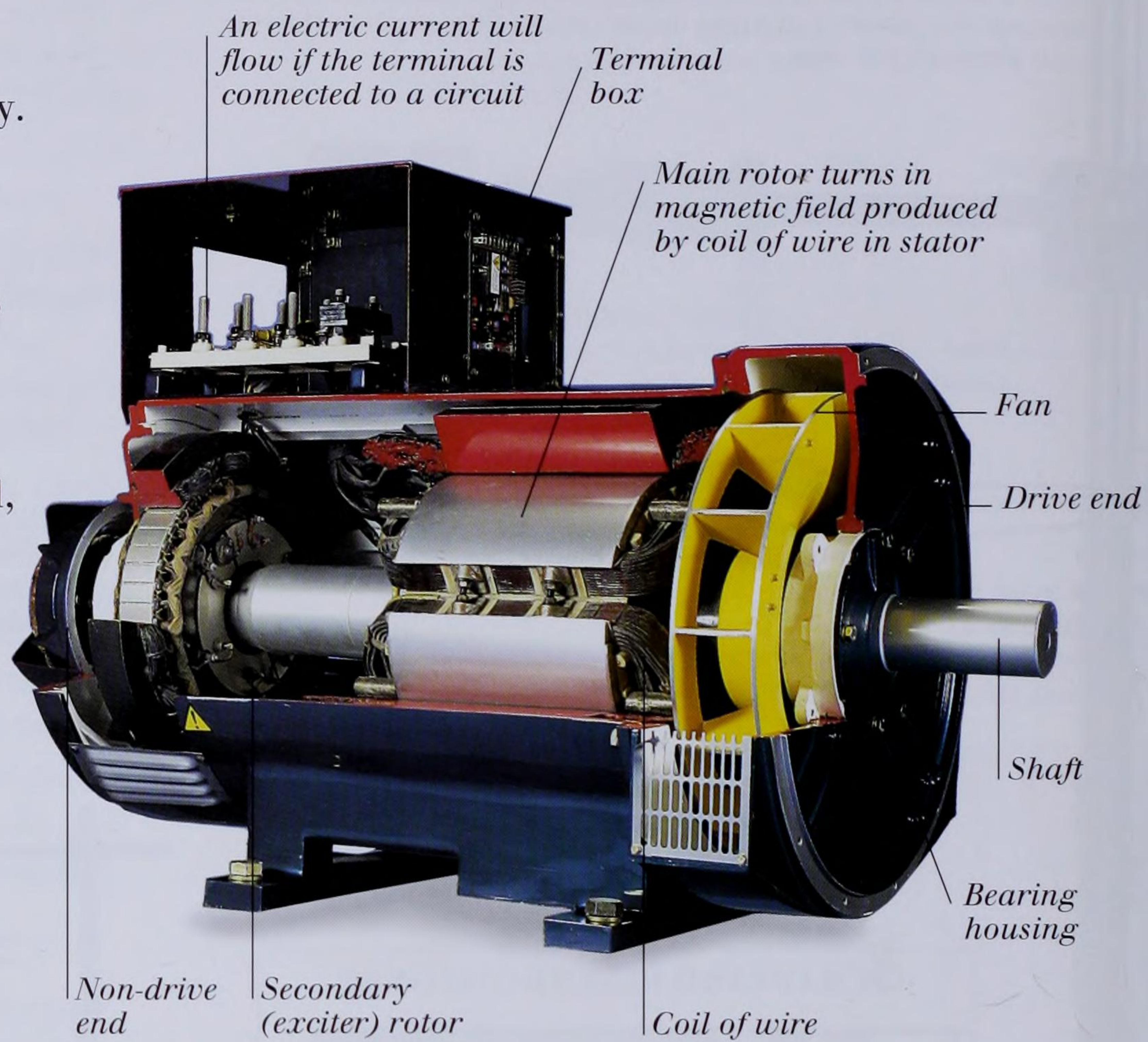


# Generating electricity

THERE ARE MANY WAYS TO GENERATE electricity. The most common is to use coils of wire and magnets in a **generator**. Whenever a wire and magnet are moved relative to each other, a **voltage** is produced. In a generator, the wire is wound into a coil. The more turns in the coil and the faster the coil moves, the greater the voltage. The coils or magnets spin around at high speed, turned by water pressure, the wind, or, most commonly, by steam pressure. The steam is usually generated by burning coal or oil, a process that creates pollution. Renewable sources of electricity—such as hydroelectric power, wind power, solar energy, and geothermal power—produce only heat as pollution. In a generator, the **kinetic energy** of a spinning object is converted into electrical energy. A solar cell converts the energy of sunlight directly into electrical energy, using layers of **semiconductors**.

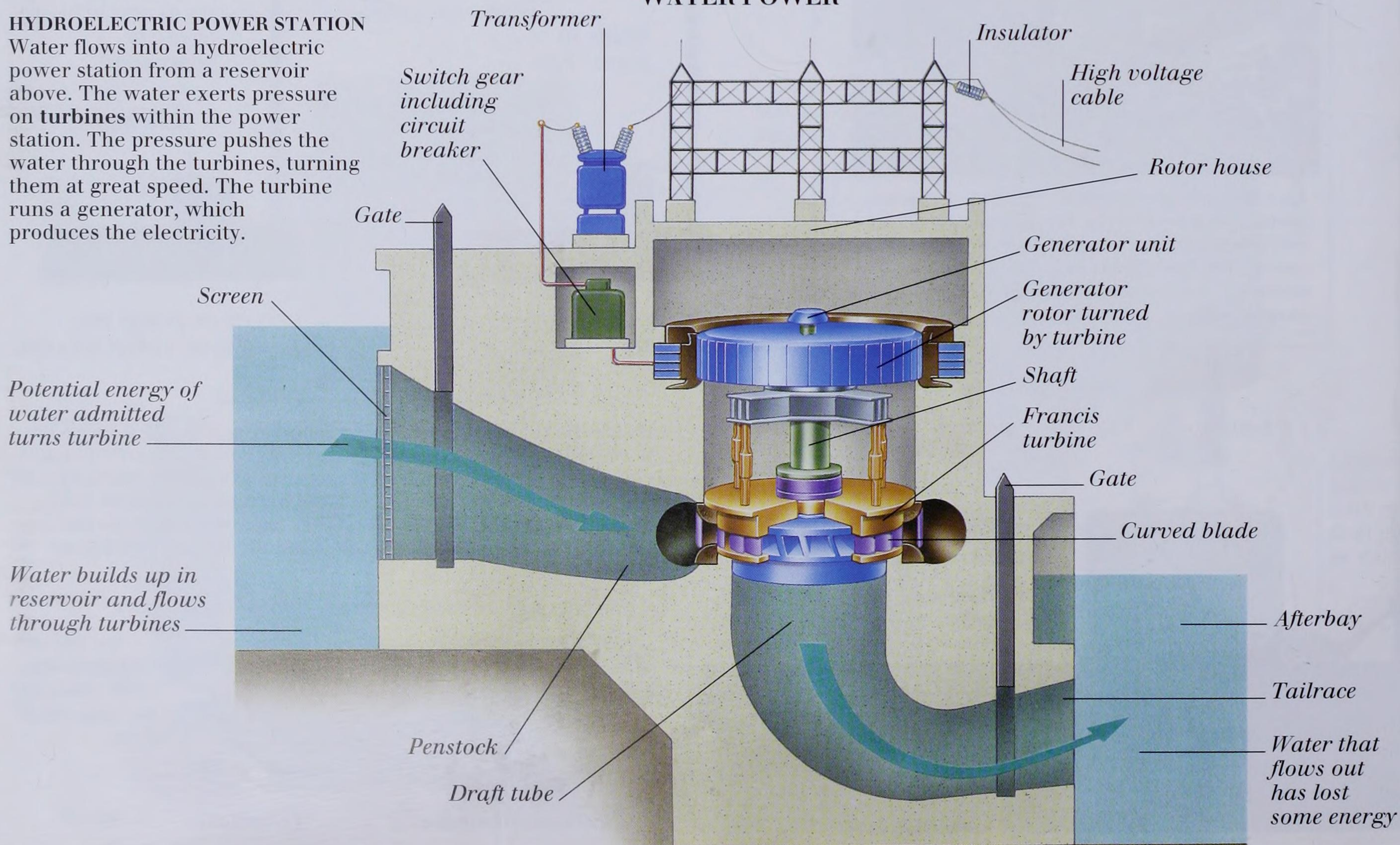
## GENERATOR

Inside a generator, you will find coils of wire and magnets (or electromagnets). In the generator shown, electromagnets spin rapidly inside stationary coils of wire. A voltage is then produced in the coils.



## HYDROELECTRIC POWER STATION

Water flows into a hydroelectric power station from a reservoir above. The water exerts pressure on **turbines** within the power station. The pressure pushes the water through the turbines, turning them at great speed. The turbine runs a generator, which produces the electricity.

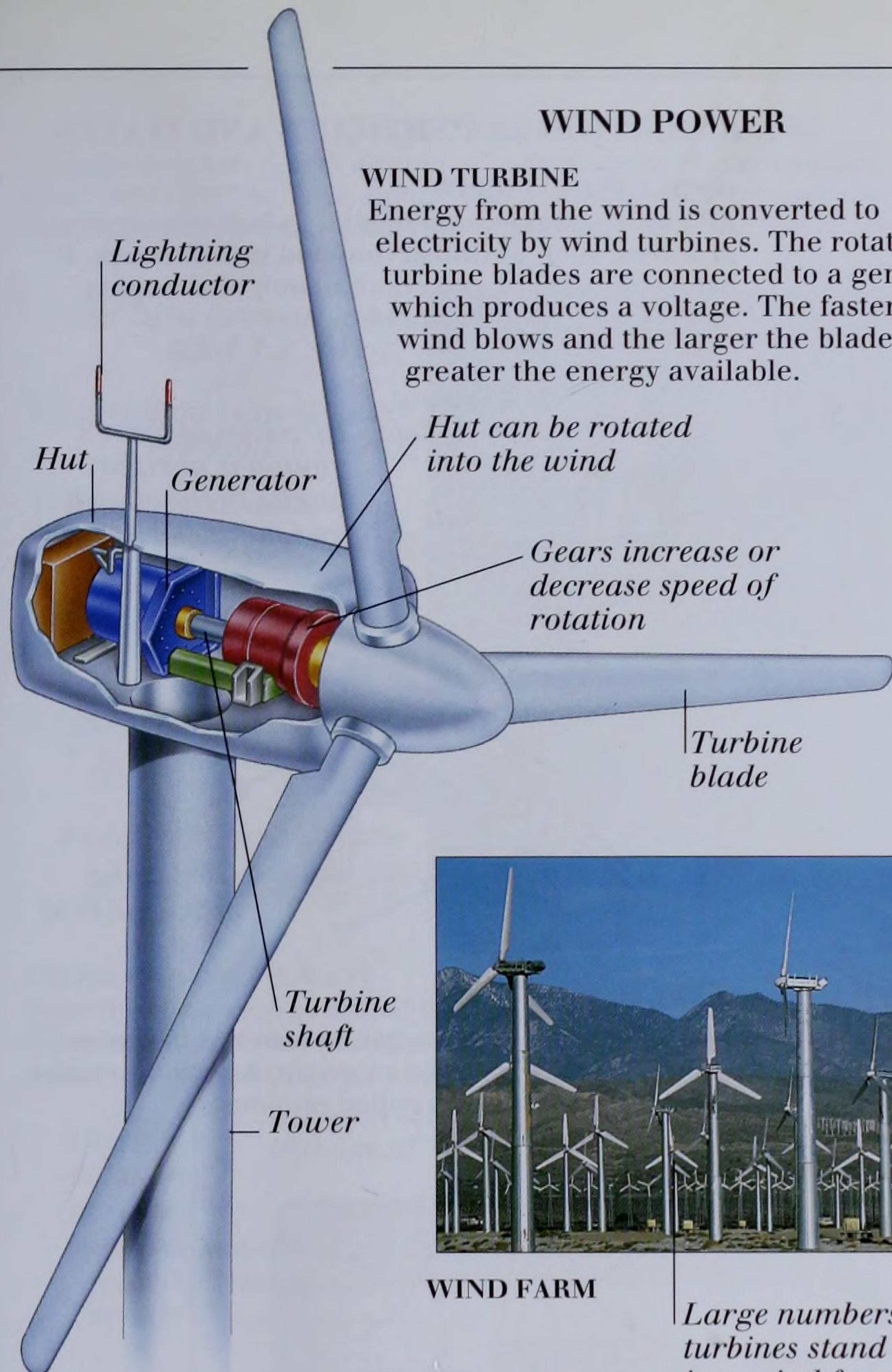




## WIND POWER

### WIND TURBINE

Energy from the wind is converted to electricity by wind turbines. The rotating turbine blades are connected to a generator, which produces a voltage. The faster the wind blows and the larger the blades, the greater the energy available.



WIND FARM

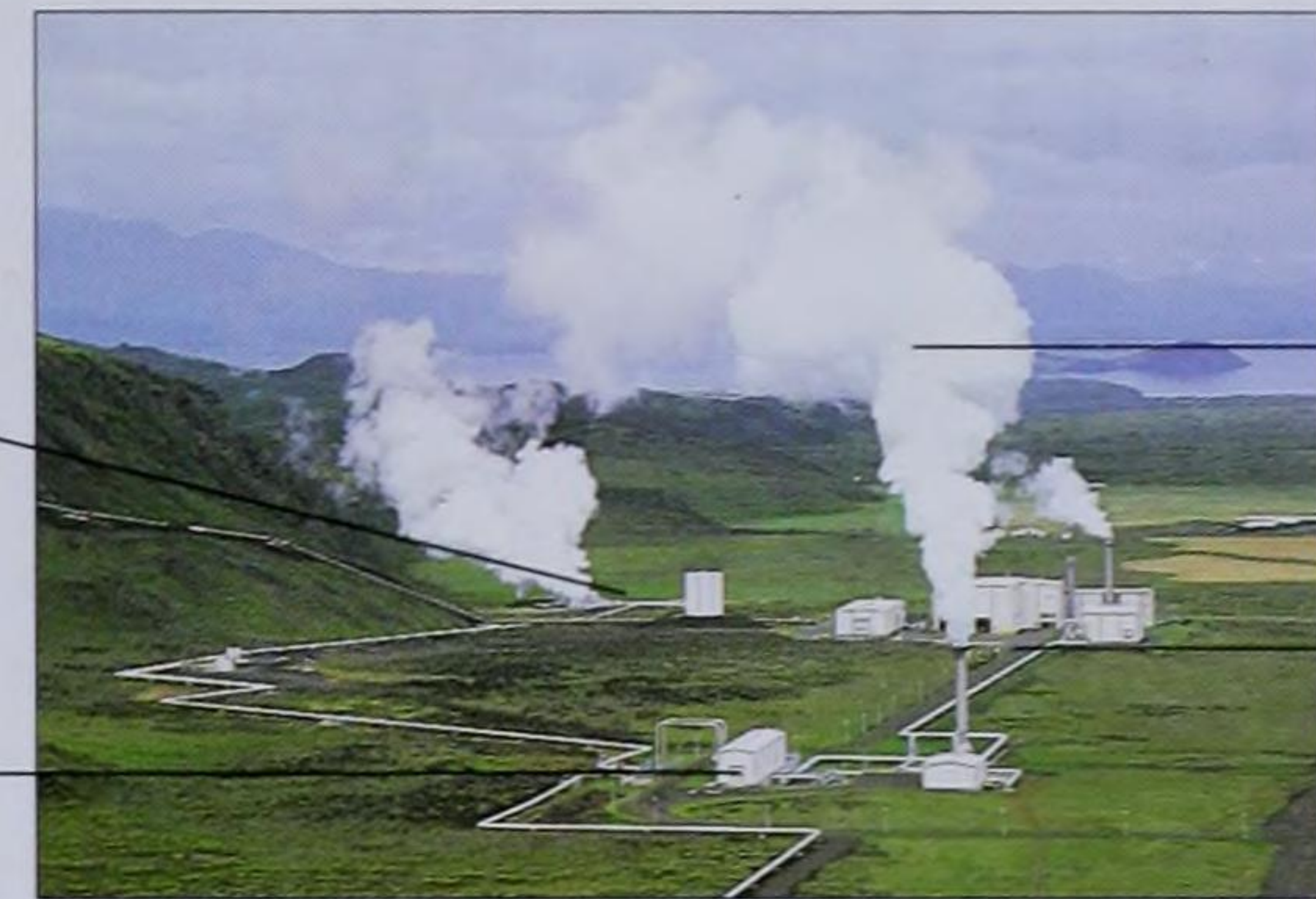
Large numbers of turbines stand together in a wind farm

## OTHER SOURCES

Two further examples of renewable sources are tidal power and geothermal power. The tides are a result of the gravitational pull of the Moon. Geothermal heat is produced by the disintegration of radioactive atoms in the Earth's core.

Excess hot water carried away to heat homes

Steam turns turbine to produce electricity



Steam emerges

Water pumped underground becomes very hot

### GEOTHERMAL POWER

Water pumped underground is turned into high-pressure steam by geothermal heat. The steam returns to the surface under pressure and turns turbines.

Tide water



Barrier

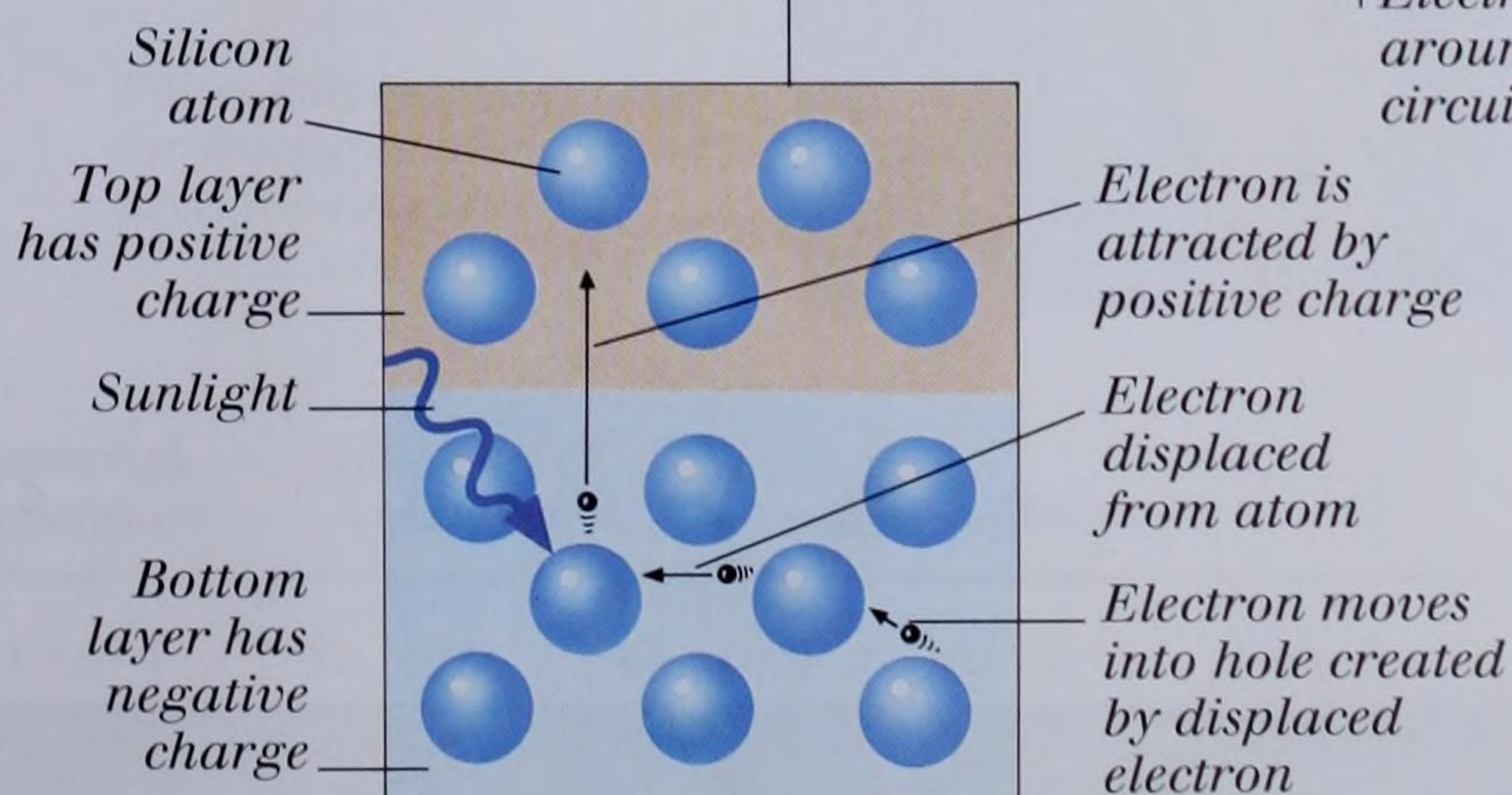
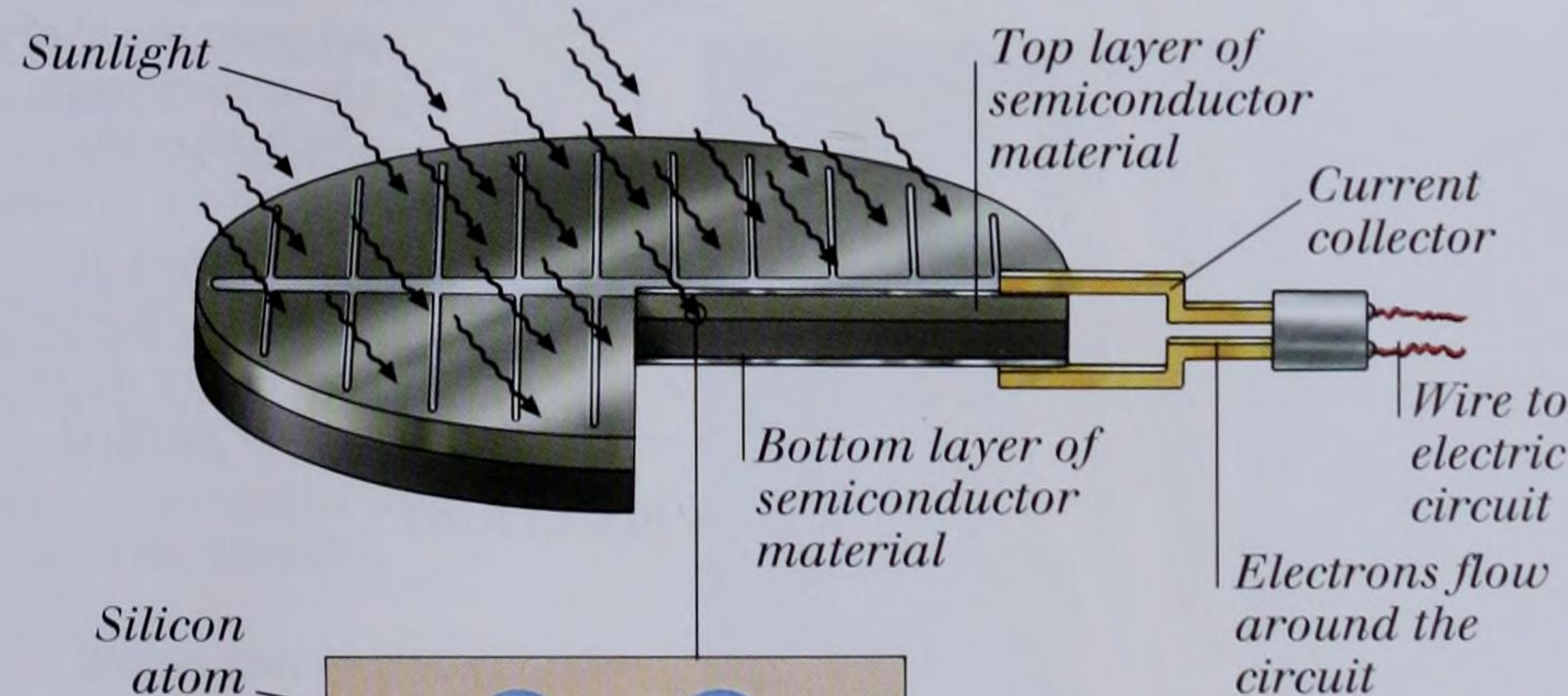
Turbines in barrier turn to produce electricity

### TIDAL POWER STATION

Seawater is held back by a barrage as it rises and falls. When there is a difference in height between the water on either side of the barrage, the water escapes through tunnels, turning turbines.

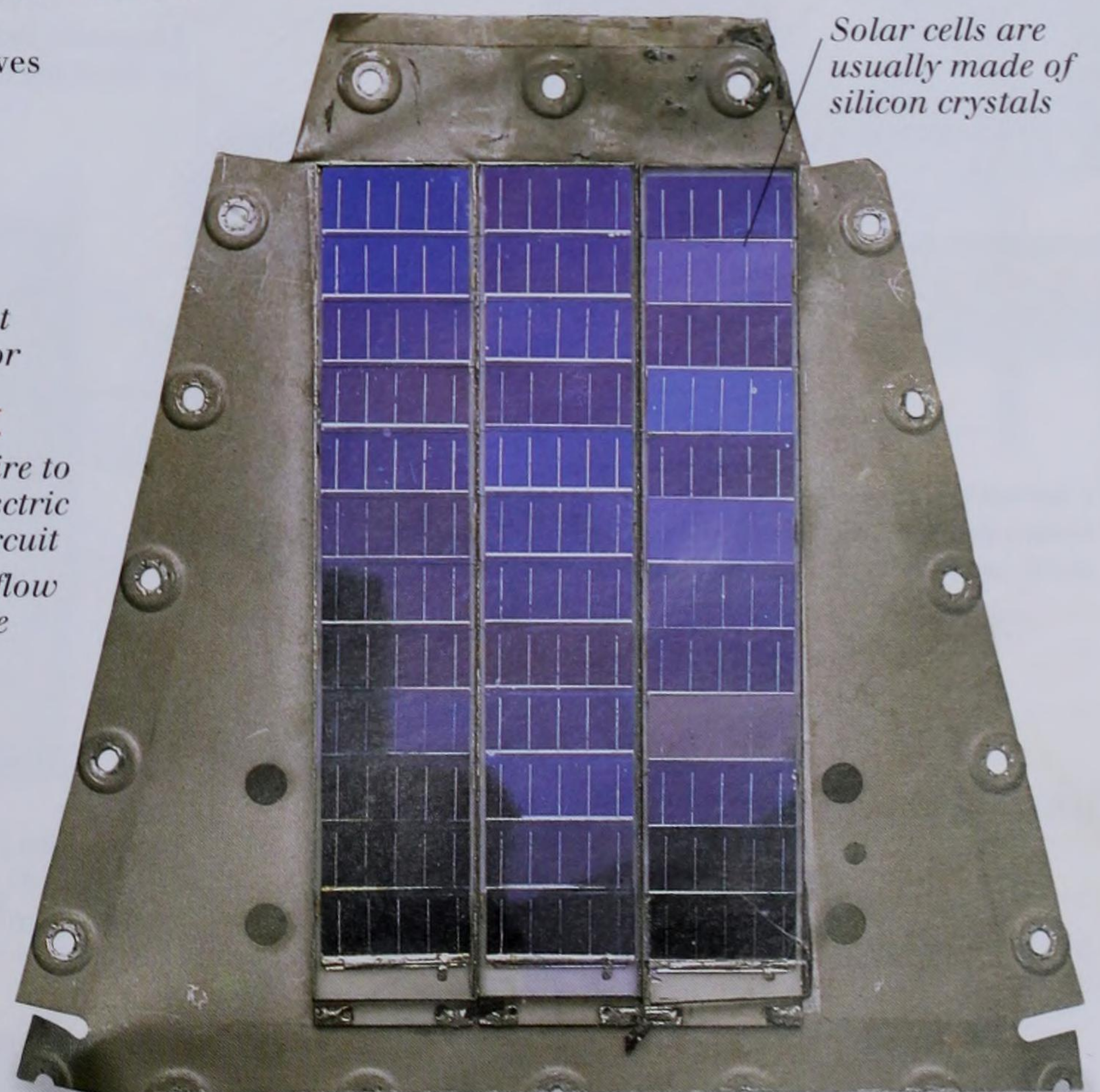
## SOLAR ENERGY

The energy of sunlight produces electricity in solar cells by causing **electrons** to leave the **atoms** in a semiconductor. Each electron leaves behind a gap, or **hole**. Other electrons move into the hole, leaving holes in their atoms. This process continues all the way around a circuit. The moving chain of electrons is an electric current.



MICROSCOPIC VIEW

Solar cells are usually made of silicon crystals

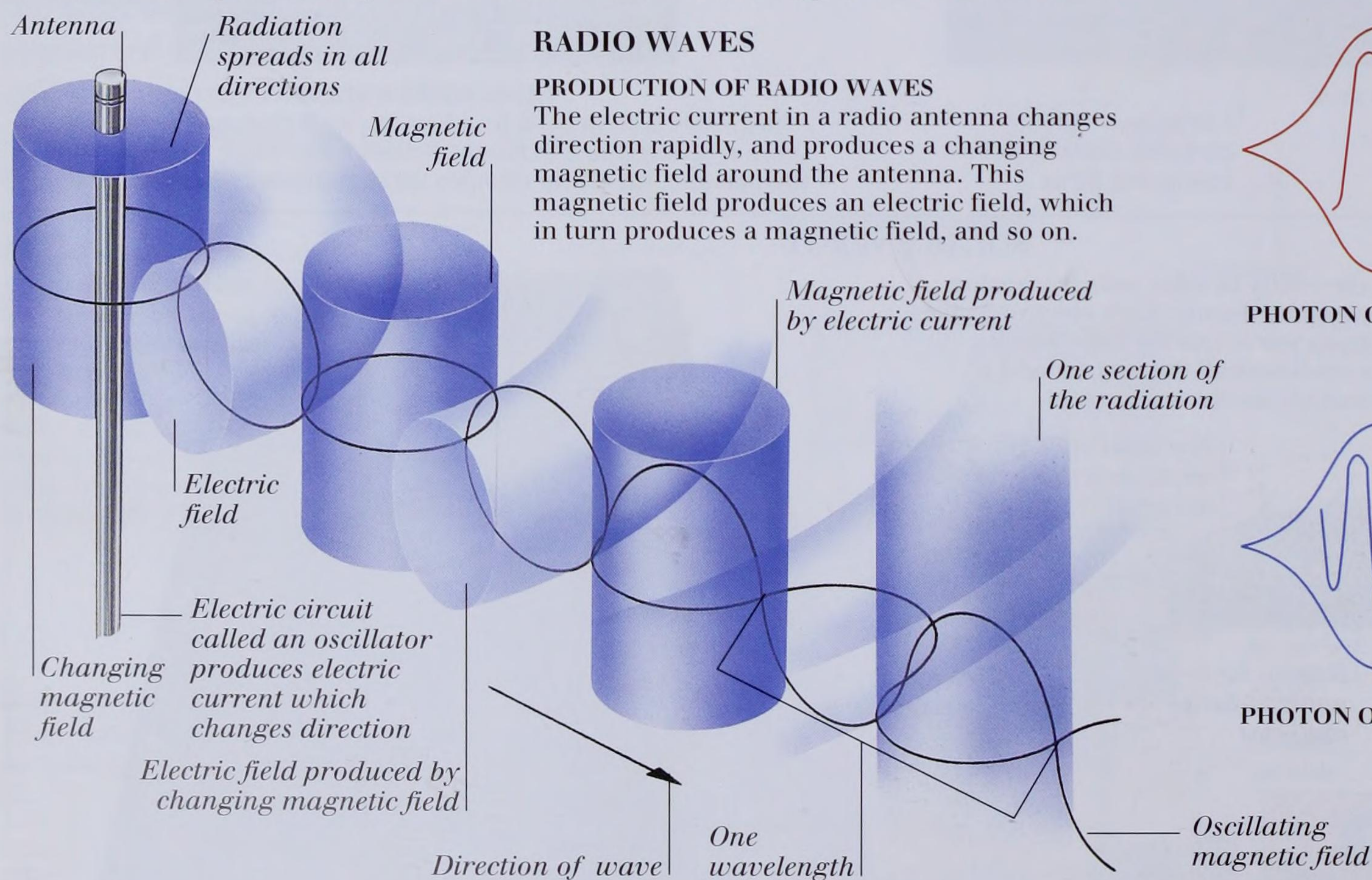


SOLAR CELL



# Electromagnetic radiation

ELECTRICITY AND MAGNETISM ARE DIRECTLY related (see pp. 34–37): a changing **electric field** will produce a changing **magnetic field**, and vice versa. Whenever an electric charge, such as that carried by an **electron**, accelerates, it gives out energy in the form of electromagnetic radiation. For example, electrons moving up and down a radio antenna produce a type of radiation known as radio waves. Electromagnetic radiation consists of oscillating electric and magnetic fields. There is a wide range of different types of electromagnetic radiation, called the electromagnetic spectrum, extending from low-energy radio waves to high-energy, short-wavelength gamma rays. This includes visible light and X-rays. Electromagnetic radiation can be seen as both a wave motion (see pp. 20–21) and as a stream of particles called **photons** (see pp. 48–49). Both interpretations are useful, since they provide a means for predicting the behavior of electromagnetic radiation.



## RADIO WAVES

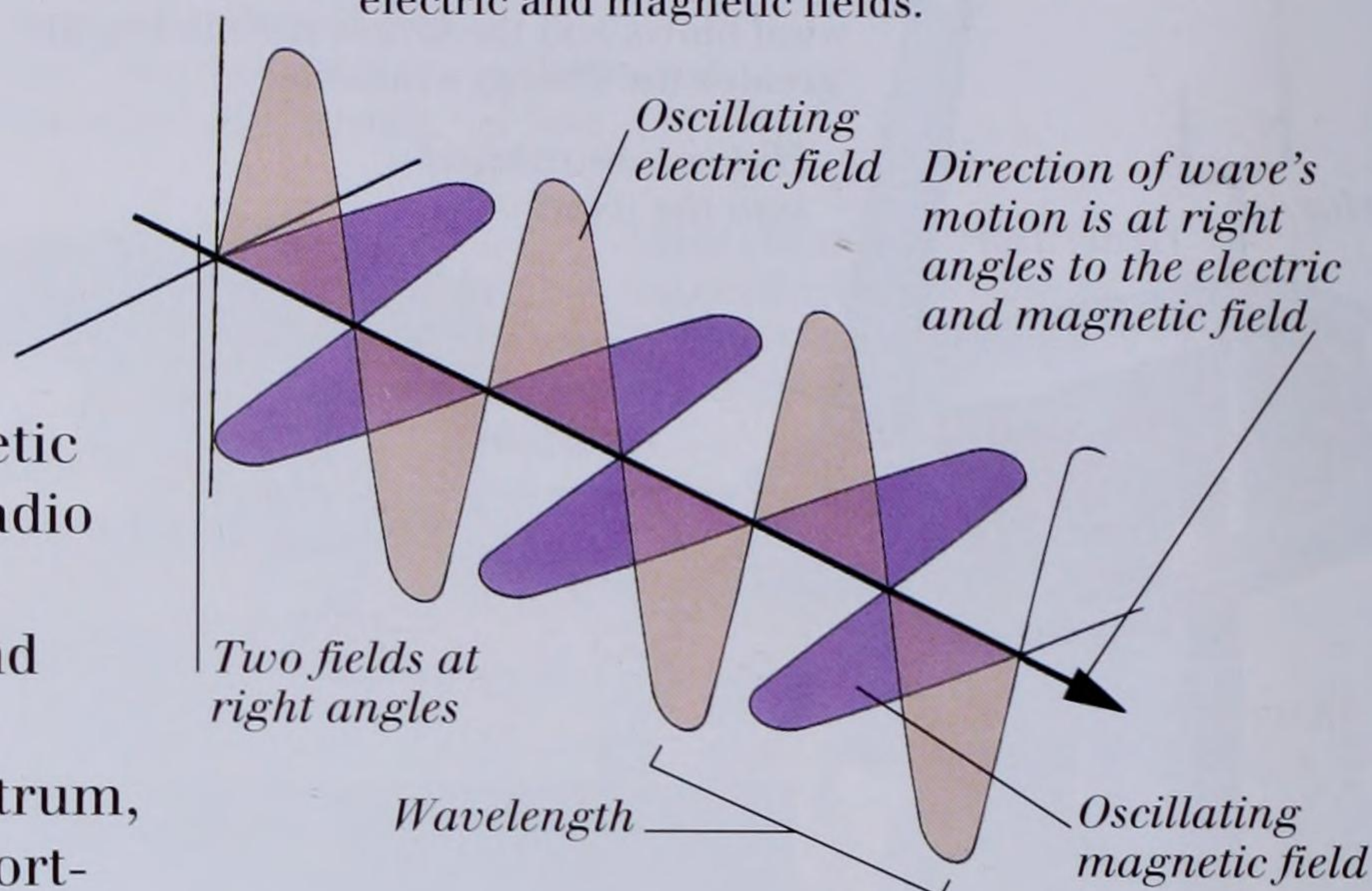
### PRODUCTION OF RADIO WAVES

The electric current in a radio antenna changes direction rapidly, and produces a changing magnetic field around the antenna. This magnetic field produces an electric field, which in turn produces a magnetic field, and so on.

## RADIATION AS PARTICLES AND WAVES

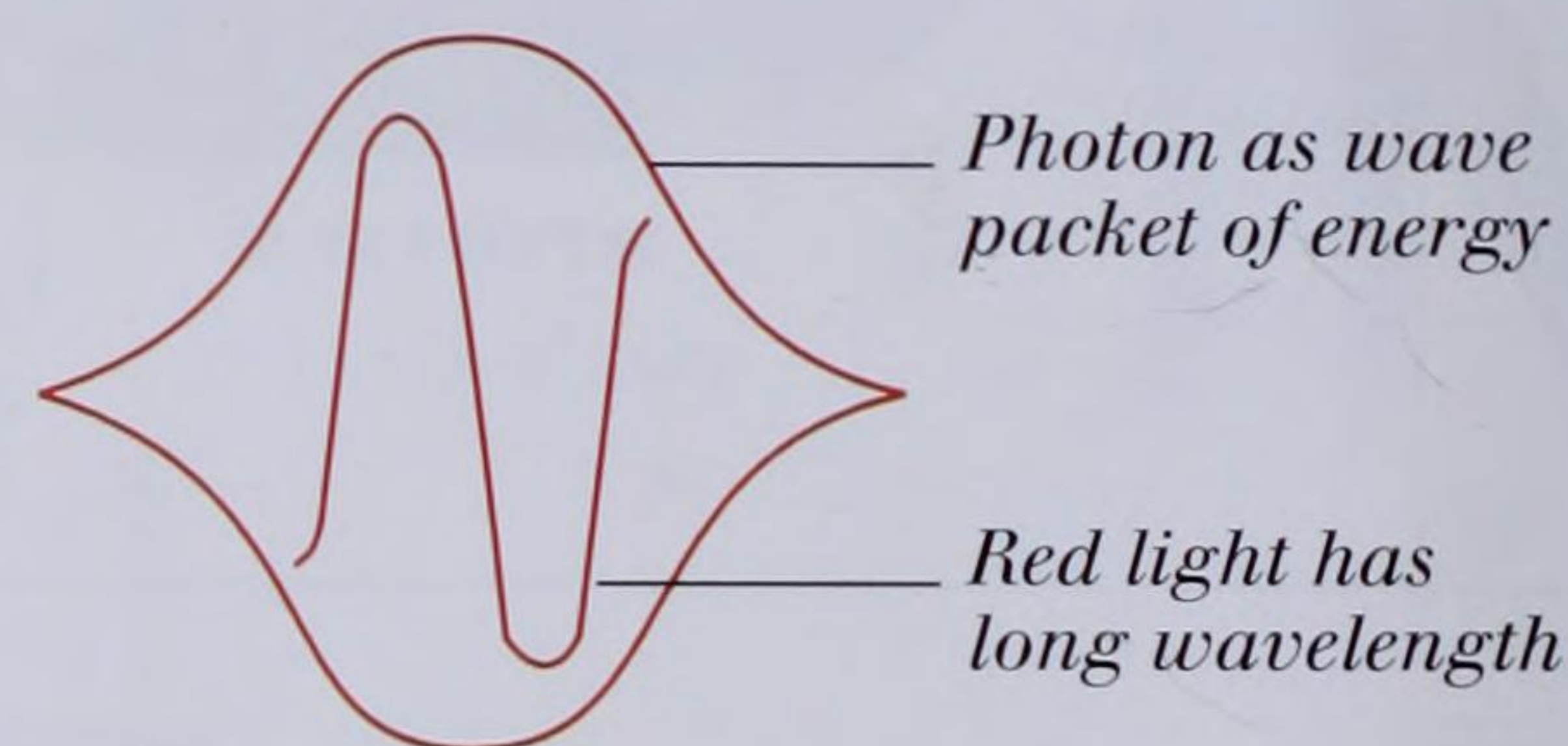
### OSCILLATING FIELDS

All electromagnetic radiation has behavior typical of waves, such as **diffraction** and **interference**. It can be thought of as a combination of changing electric and magnetic fields.

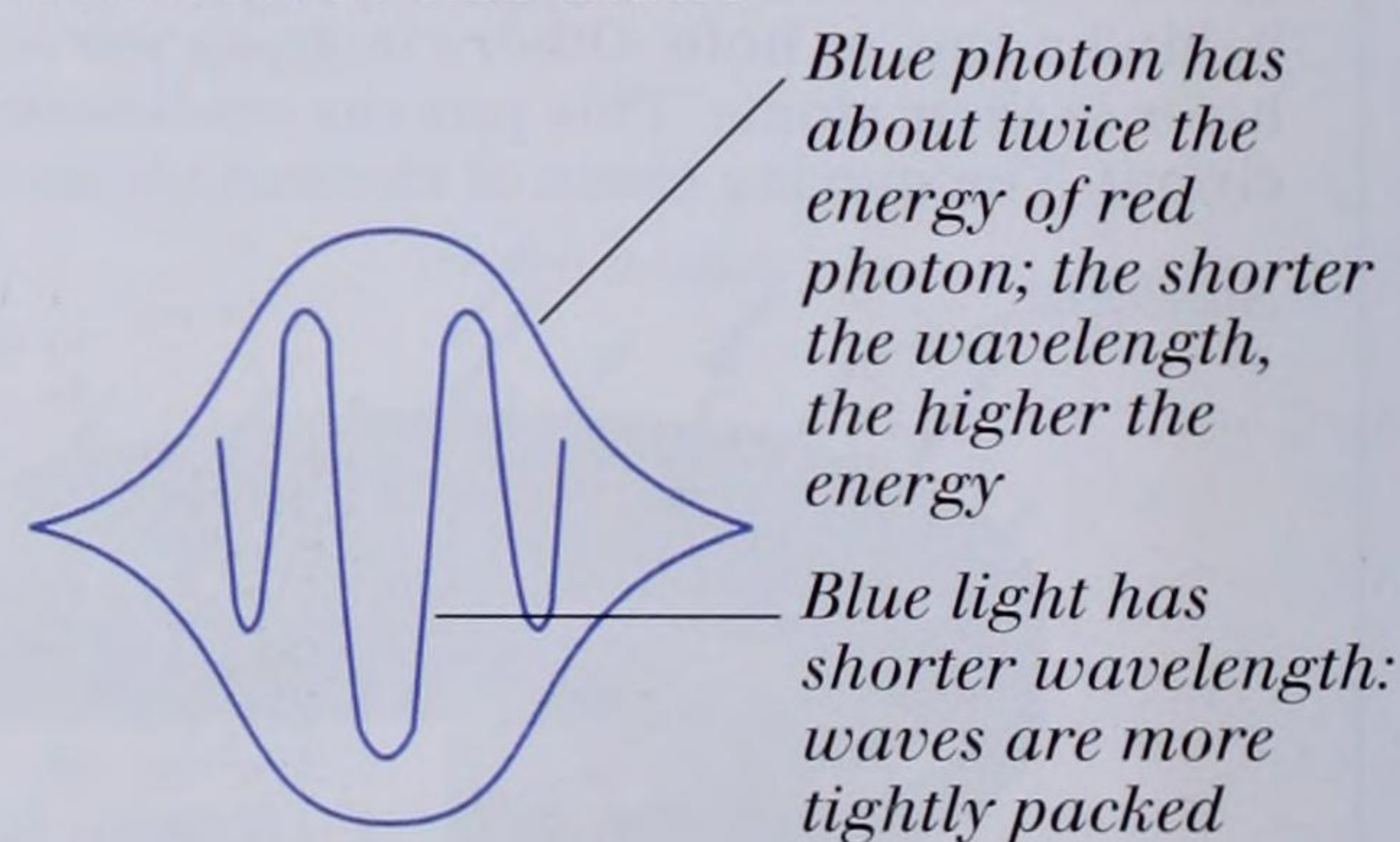


### PHOTONS

All electromagnetic radiation also has behavior typical of particles. For example, its energy comes in individual bundles called photons.

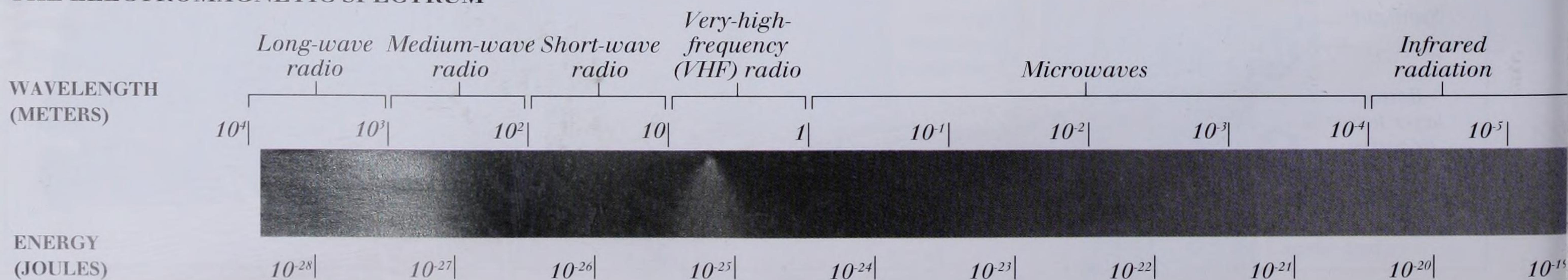


### PHOTON OF RED LIGHT



### PHOTON OF BLUE LIGHT

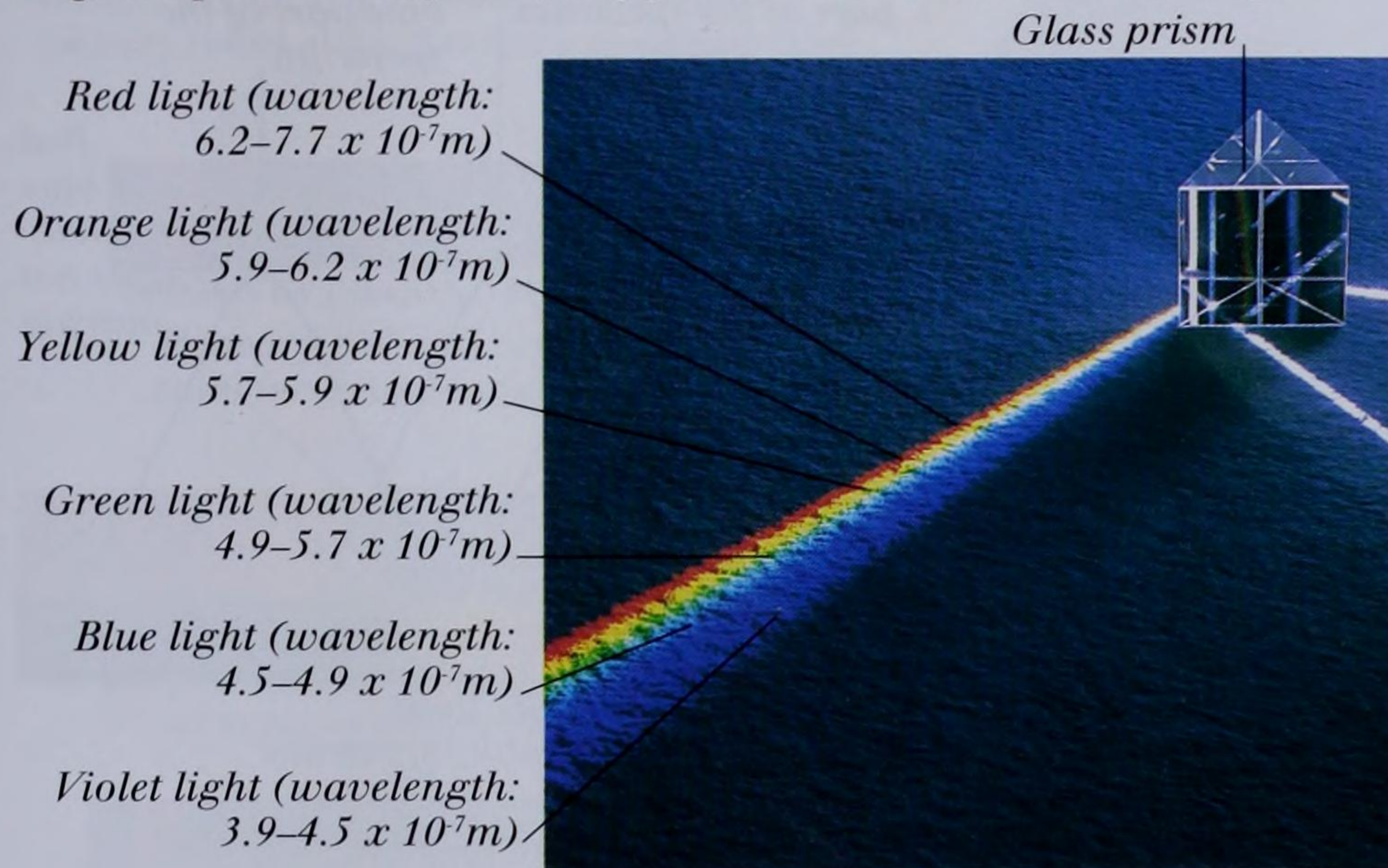
## THE ELECTROMAGNETIC SPECTRUM





## THE WHITE LIGHT SPECTRUM

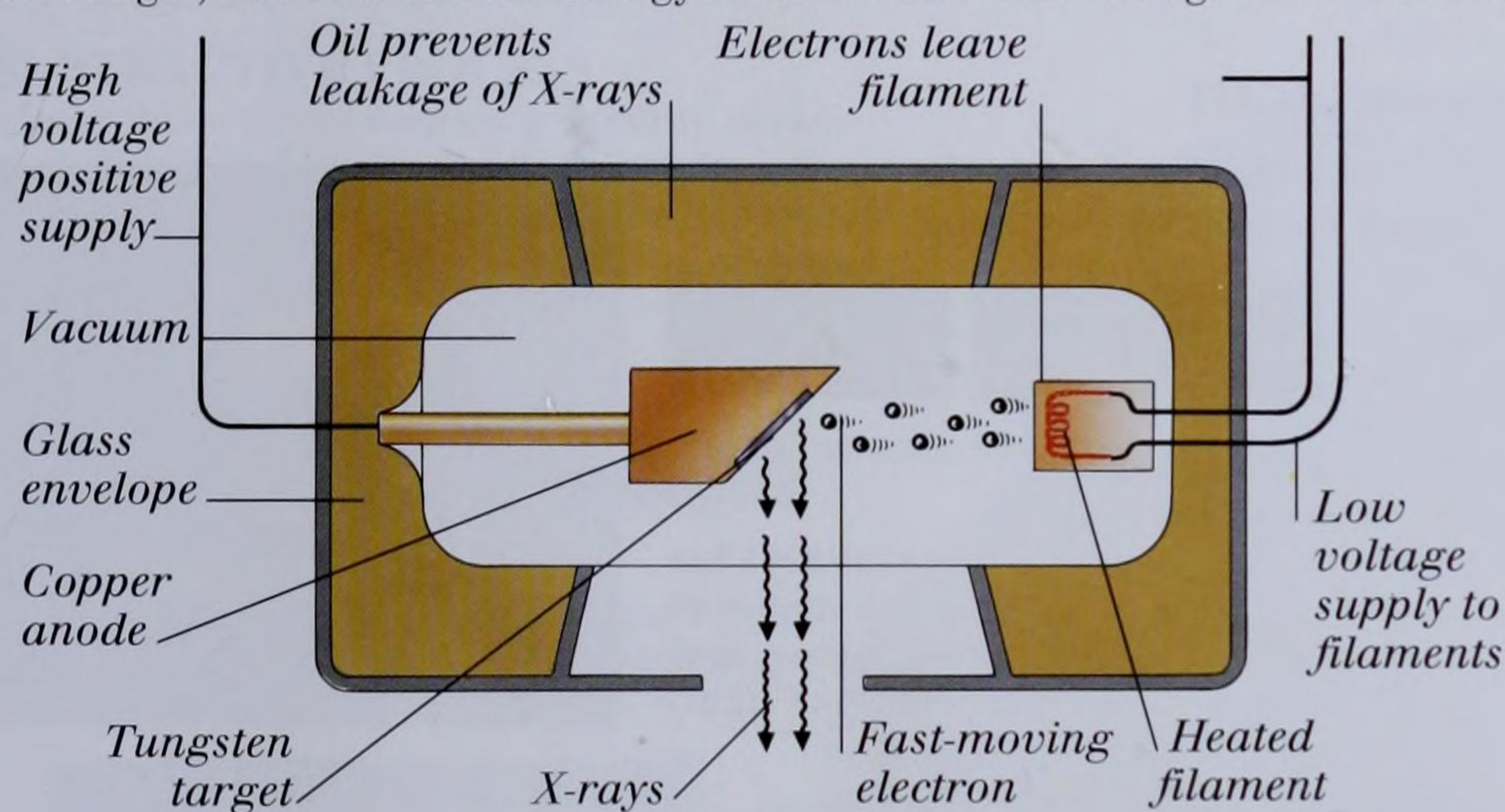
Human eyes can detect a range of wavelengths of electromagnetic radiation, from "red light" to "blue light." When all of the wavelengths within that range are perceived together, they produce the sensation of white light.



## X-RAYS

### PRODUCTION OF X-RAYS

Near the high-energy end of the electromagnetic spectrum come X-rays. In an X-ray tube, electrons are accelerated by a strong electric field. They then hit a metal target, and their kinetic energy is turned into electromagnetic radiation.



### X-RAY PHOTOGRAPH

The main use for X-rays is in medical photography. Radiation from an X-ray tube does not pass through bone, so when an image is recorded on paper sensitive to X-rays, an image of the bone remains. Thus fractures can be investigated without the need for surgery.

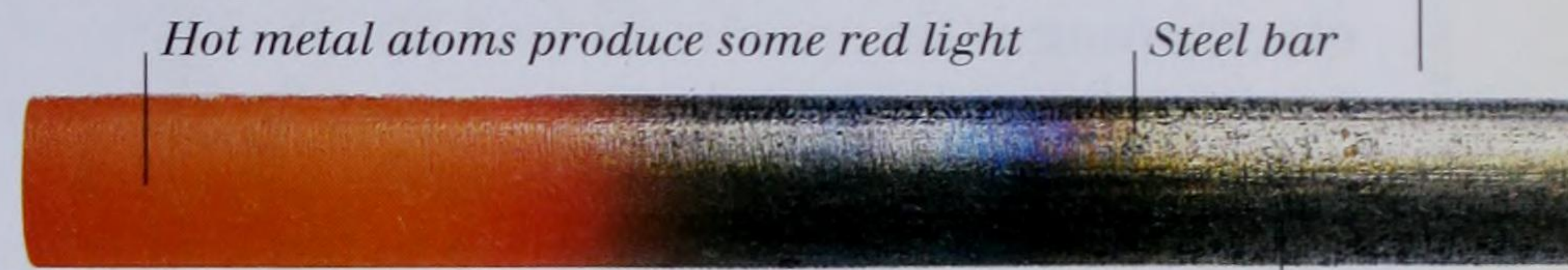
*Bones can be examined for fractures without the need for surgery*



*Image of bone*

## RADIATION FROM HOT OBJECTS

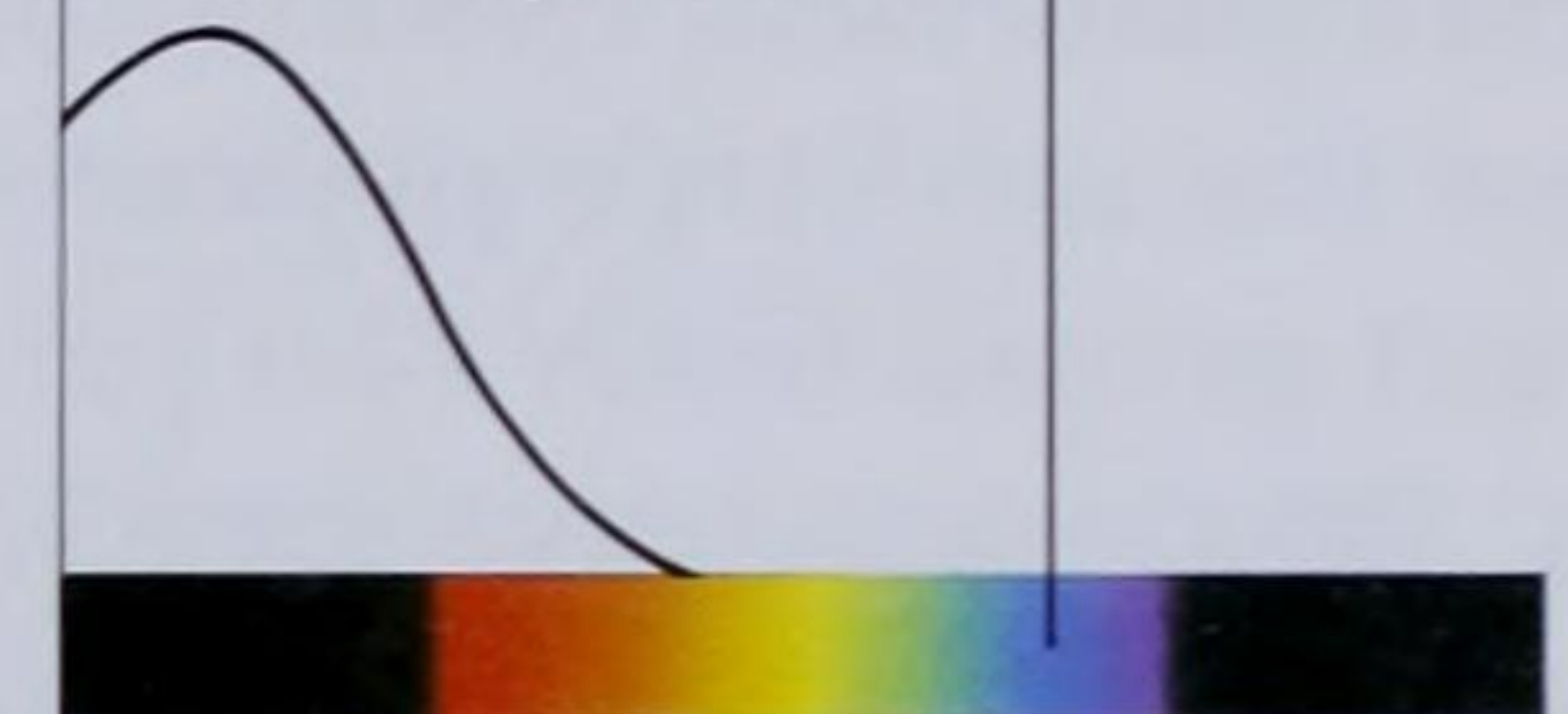
The **atoms** of a solid vibrate (see pp. 22-23). Atoms contain electric charges in the form of **protons** and electrons. Because they vibrate, these charges produce a range of electromagnetic radiation. The rate of vibration, and therefore the wavelengths of radiation produced, depends on **temperature**, as this steel bar shows.



**OBJECT HEATED TO ABOUT 900K (627°C)**  
At 900K, objects give out a range of radiation, mainly infrared. The graph shows how much of each wavelength is radiated.

*Cooler atoms radiate invisible infrared*

*No blue light produced*

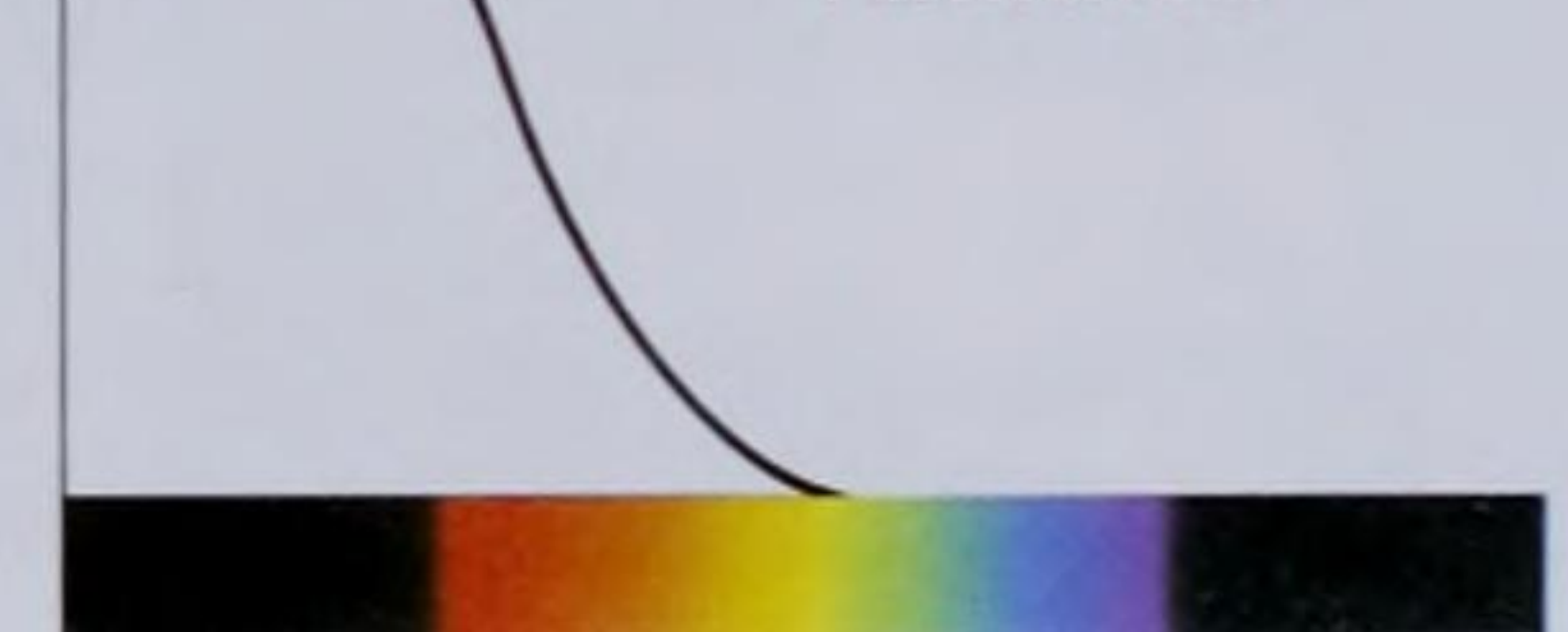


*Radiation now appears yellow*



**OBJECT HEATED TO ABOUT 1,500K (1,227°C)**  
As the metal atoms vibrate more vigorously, the radiation has more energy. It therefore includes more of the visible spectrum.

*More of the spectrum is radiated*



*Radiation now appears white*



**OBJECT HEATED TO ABOUT 1,800K (1,527°C)**  
Near its melting point, the bar produces even more light. The range of light now includes the entire visible spectrum. This is why it looks bright white.

*The complete visible spectrum is radiated*

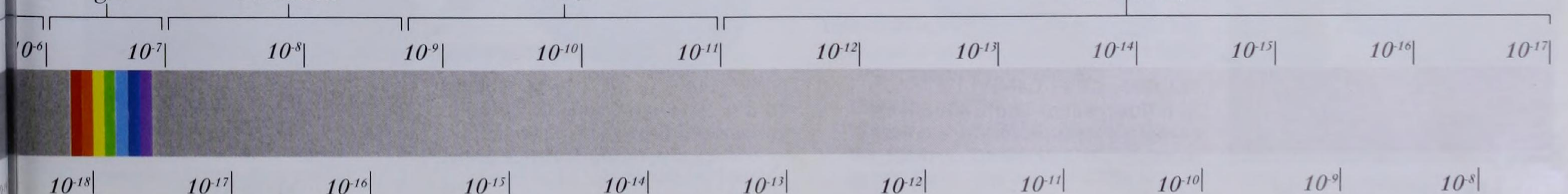


Visible light

Ultraviolet radiation

X-rays

Gamma rays

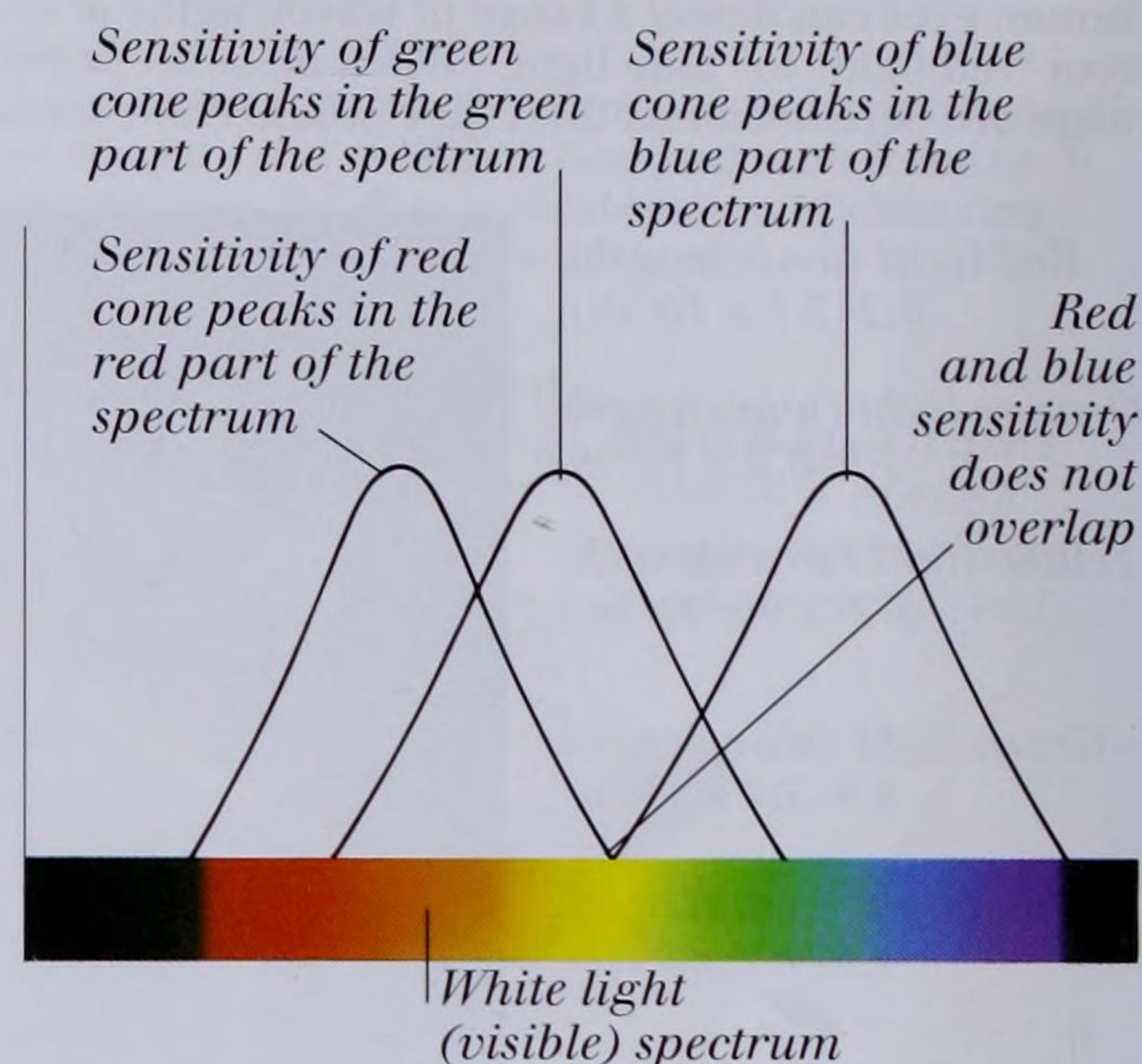




# Color

THE HUMAN EYE CAN ONLY PERCEIVE a small section of the **electromagnetic spectrum** (see pp. 38–39). We call this section “visible light.” Different colors across the spectrum of visible light correspond to different **wavelengths** of light. Our eyes contain cells called **cones**, which are sensitive to these different wavelengths and allow us to see in color. Three different types of cone are affected by light in the red, green, and blue parts of the spectrum. These correspond to the **primary colors**. Different light sources give out different parts of the spectrum, which appear as different colors. When combined, colored lights appear as different colors. This is called the **additive process**. Adding primary light sources in the correct proportions can produce the sensation of other colors in our eyes. When light hits a **pigment** in an object, only some colors are reflected. Which colors are **reflected** and which **absorbed** depends on the pigment. This is the **subtractive process**. Looking at a colored object in colored light may make the colors appear different. This is because pigments can only reflect colors that are present in the incoming light.

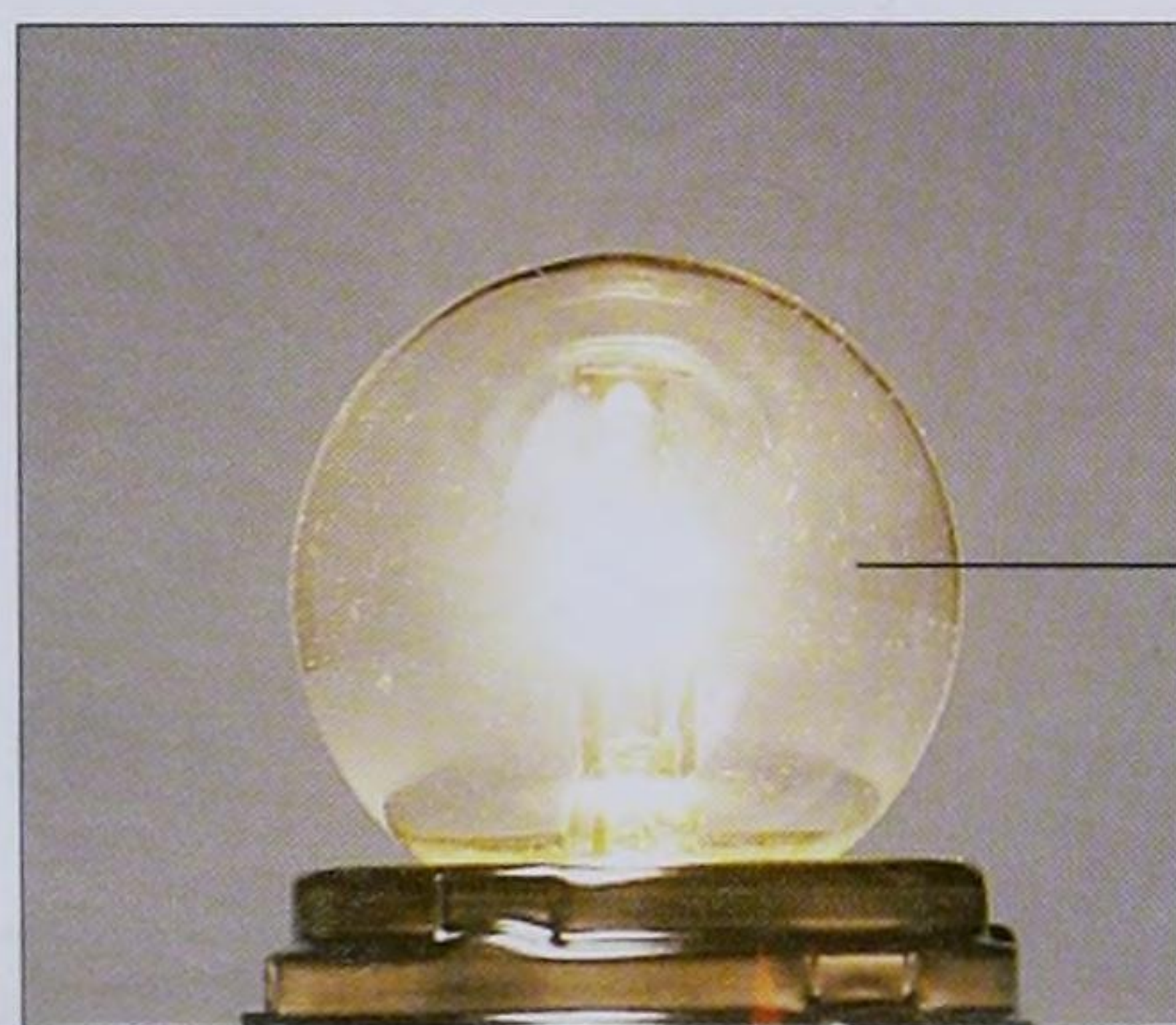
## CONE SENSITIVITY



## COLOR VISION

There are three different types of cone in the normal human eye, each sensitive to a different part of the spectrum. White light stimulates all three types of cone cells.

## SOURCES OF LIGHT



BRIGHT FILAMENT LAMP

This spectrum shows which colors are produced

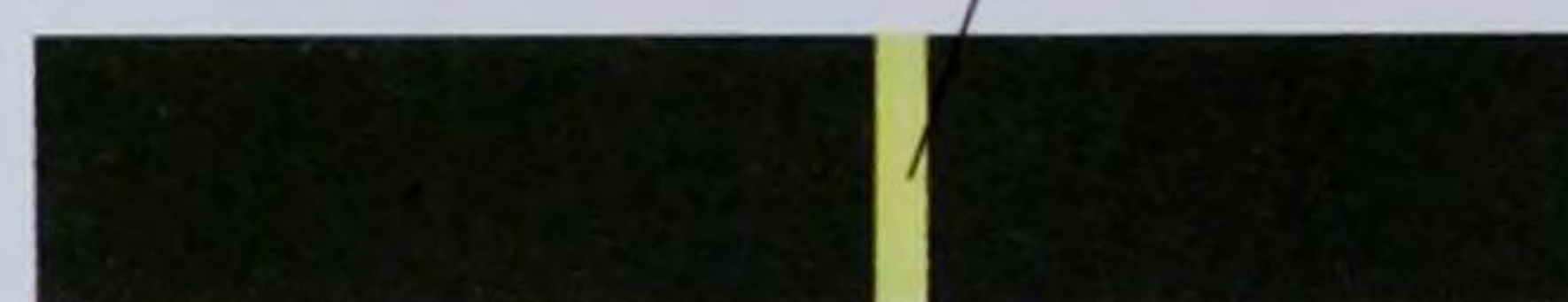


All colors of light together combine to produce white

### BRIGHT FILAMENT LAMP

With a high **electric current**, the whole spectrum of visible light is produced (see p. 39).

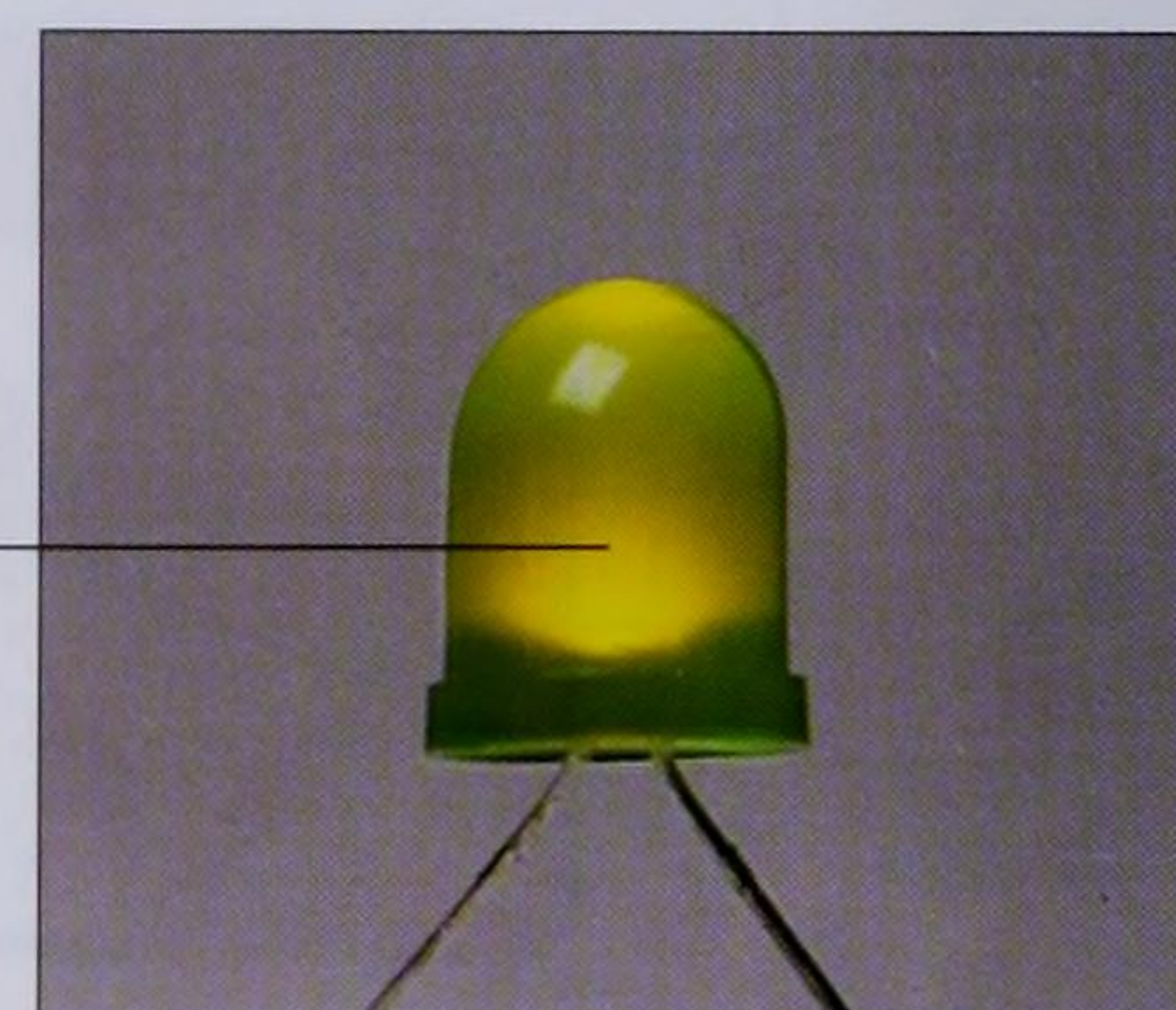
LED produces colors in the green part of the spectrum



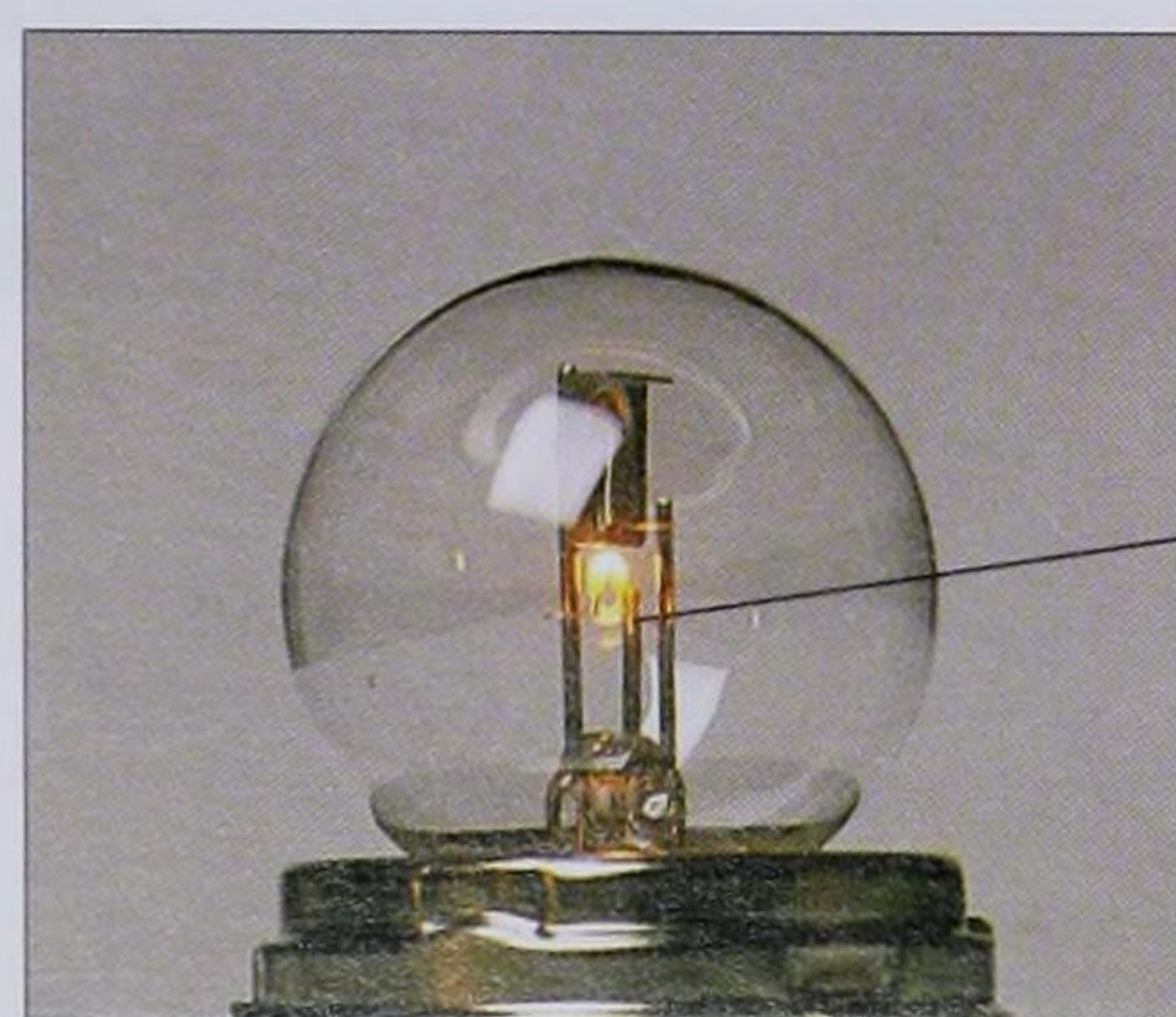
LED appears green

### GREEN LED

An LED (light-emitting diode) is made of a **semiconductor**, and produces certain colors of light.



GREEN LED



DIM FILAMENT LAMP

Red, yellow, and green light combine to produce orange



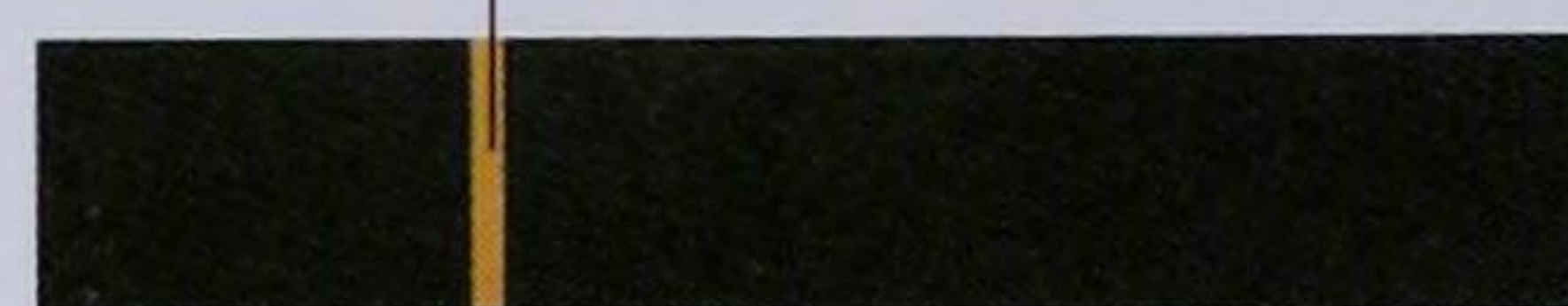
Lamp appears orange

No blue light produced

### DIM FILAMENT LAMP

With a smaller current, the **temperature** of the **filament** (see pp. 32–33) is low.

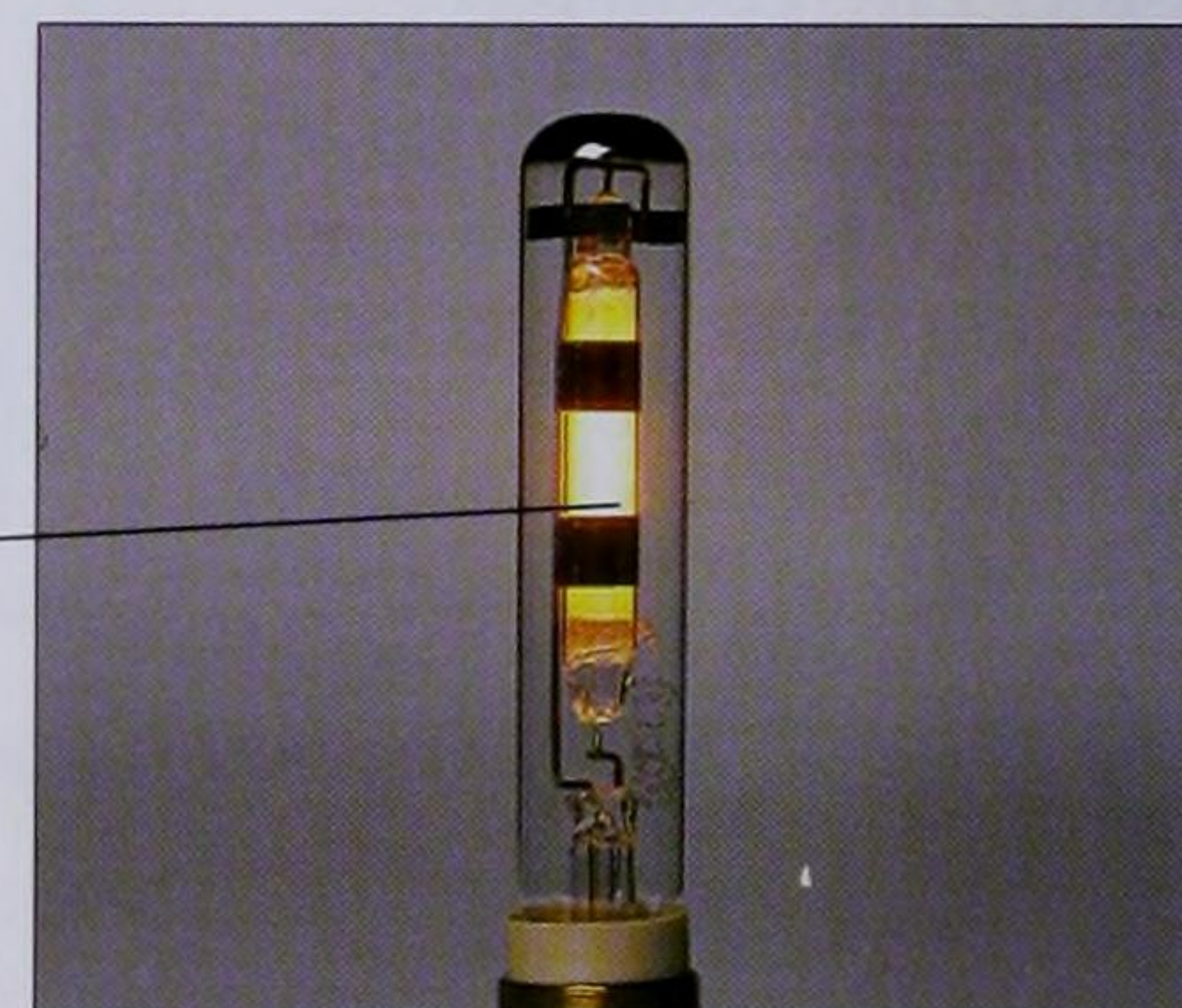
Two colors of light very close together in the orange part of the spectrum are produced



Lamp appears orange

### SODIUM LAMP

In a sodium lamp, an electric current excites electrons in sodium vapour, giving them extra energy. The electrons give the energy out as light.



SODIUM LAMP



FLUORESCENT LAMP

Lamp produces certain colors in each part of the spectrum



All three types of cone are stimulated and lamp appears white

### FLUORESCENT LAMP

In a **fluorescent** lamp, chemicals called **phosphors** produce colors in many parts of the spectrum.

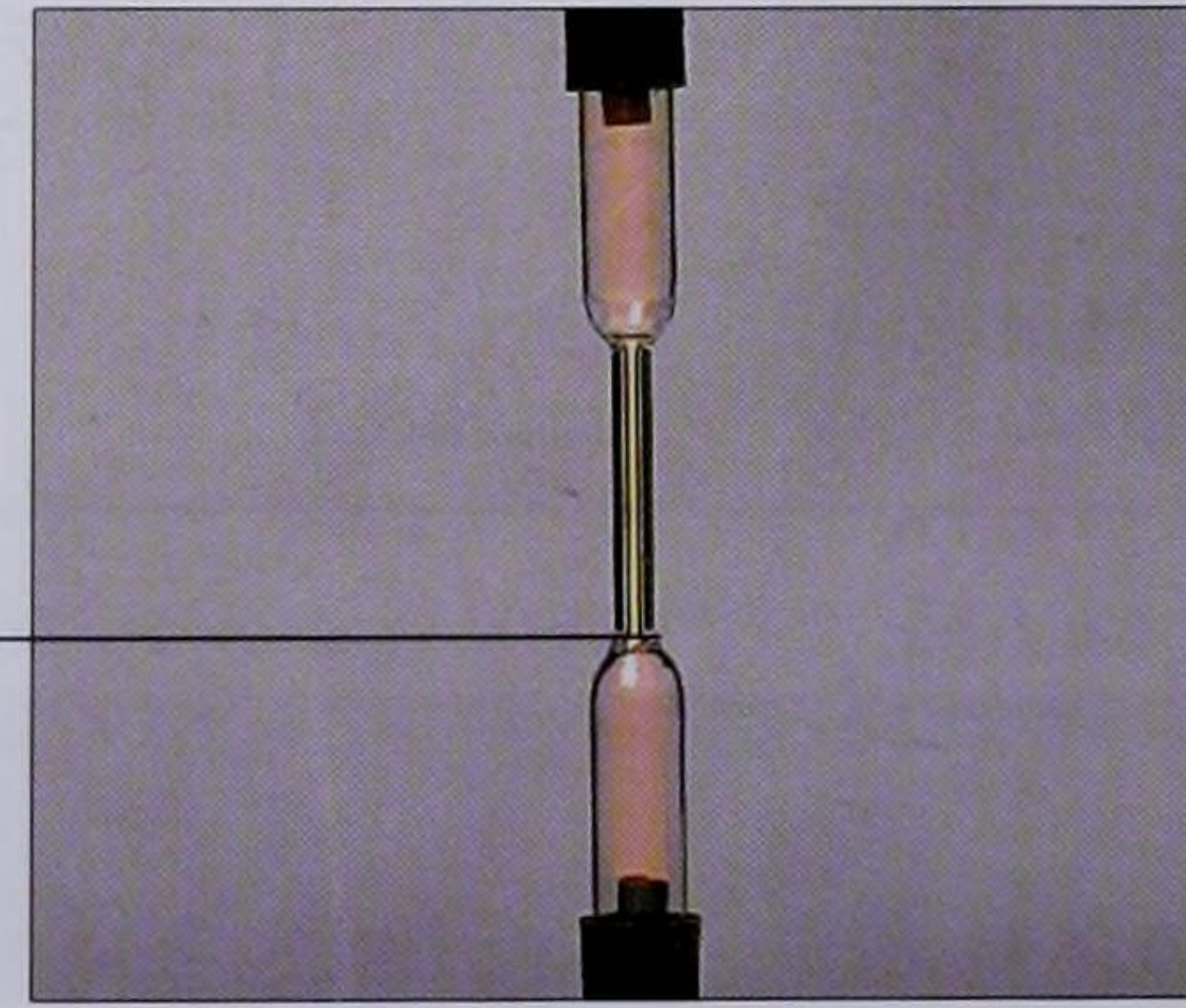
Only certain colors characteristic of neon are produced



Lamp appears orange

### NEON TUBE

In a similar way to a sodium lamp, a neon discharge lamp produces a characteristic orange glow.



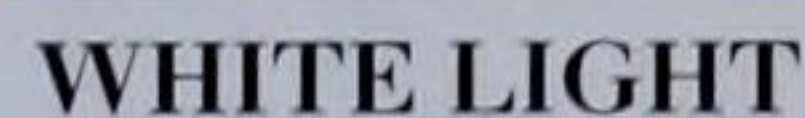
NEON TUBE



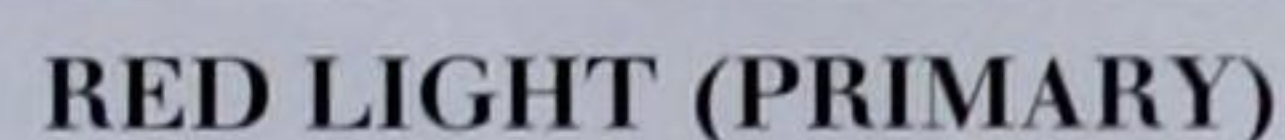
Adding red, green, and blue light in the correct proportions can create the illusion of any other color. These three colors are called primary colors. A color made from adding any two primary colors alone is called a secondary color.



*Primary red and primary blue  
combine to appear as magenta*

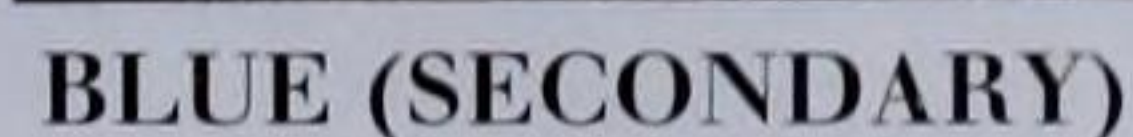


*All the primary colors together stimulate  
all types of cone and appear white*

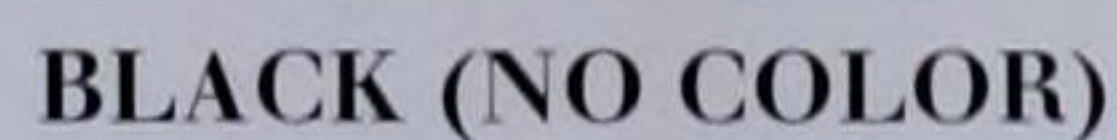


*Primary red light stimulates the red cone*

These three filters contain pigments which absorb some of the colors in the white light passing through them from a light beneath. By mixing primary pigments together, all colors except true white can be produced.



*Magenta and cyan filters together  
allow blue light through*

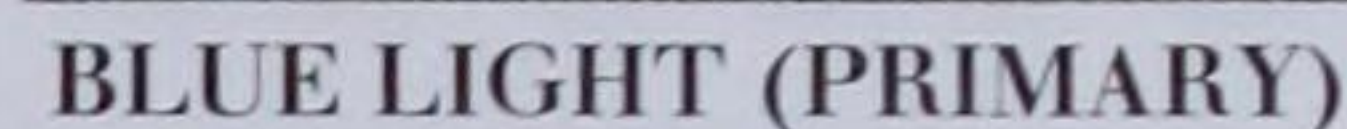


*Where all three filters overlap, they absorb all colors, and appear black*

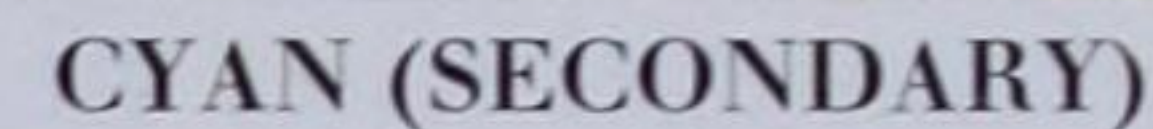


*A primary magenta filter will absorb all light except red and blue*

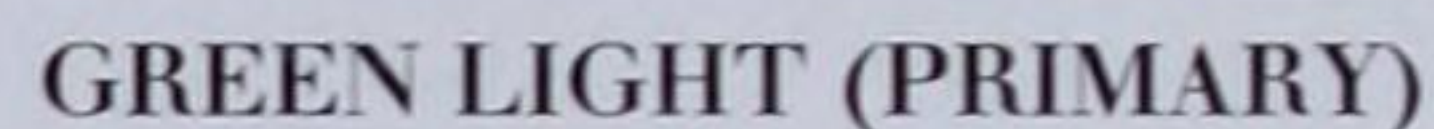
## PRIMARY COLORS FOR THE ADDITIVE PROCESS



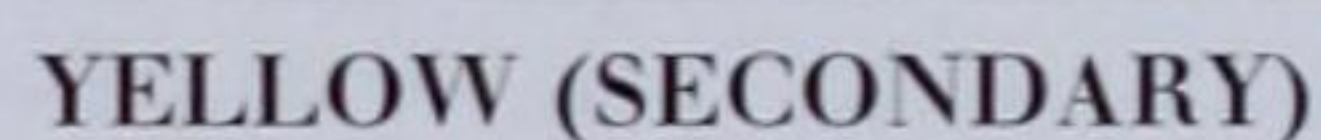
*Primary blue light stimulates the blue cone*



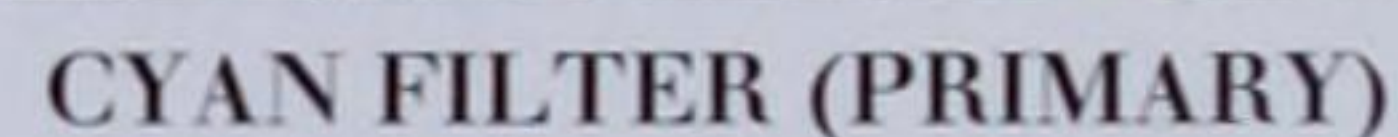
Primary green and primary blue  
combine to appear as cyan



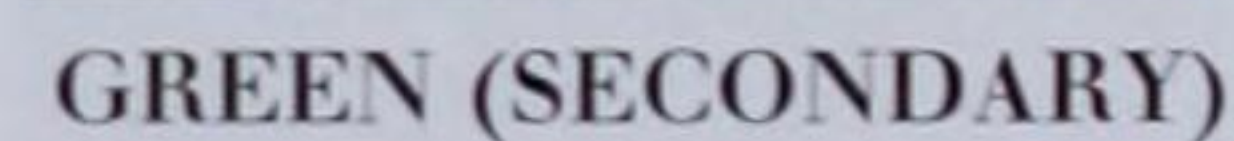
Primary green light stimulates the green cone



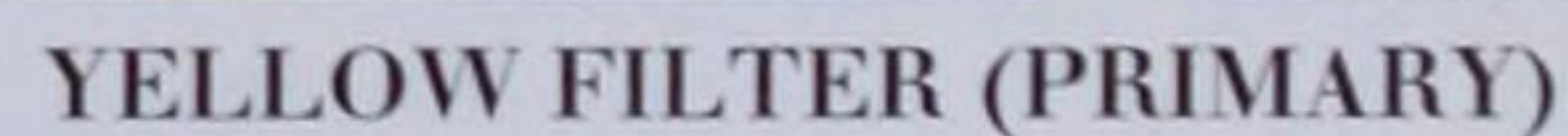
Primary red and primary green  
combine to appear as yellow



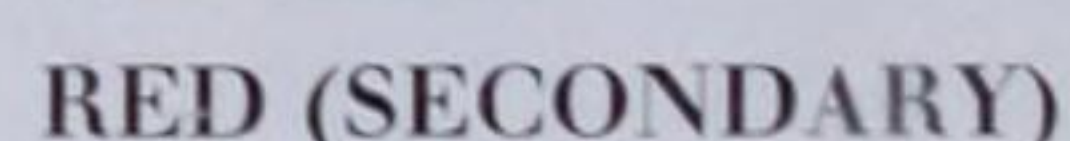
*A primary cyan filter will absorb all light except blue and green*



*Cyan and yellow filters together  
only allow green light through*



*A primary yellow filter will absorb all light except red and green*

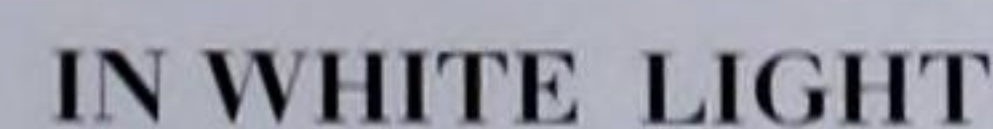


*Magenta and yellow filters together will allow red light through*

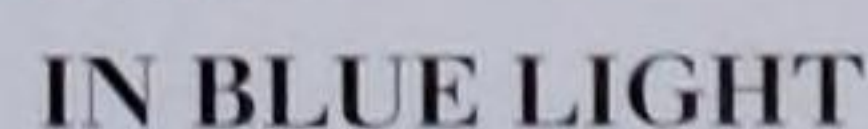
Green pot appears black

Red pot appears red

Green pot appears black



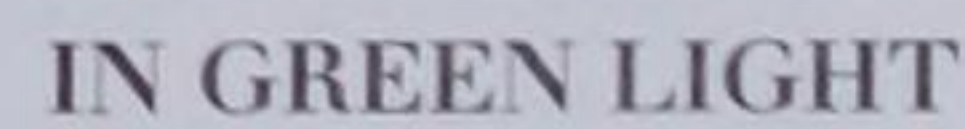
The green pot only reflects the green part of the spectrum, absorbing the other colors.



When only blue light is available, the green pigment can reflect no green light, and appears black.



When only red light is available, the green pigment can reflect no green light, and appears black.



When only green light is available, the green pigment reflects green light and appears green.

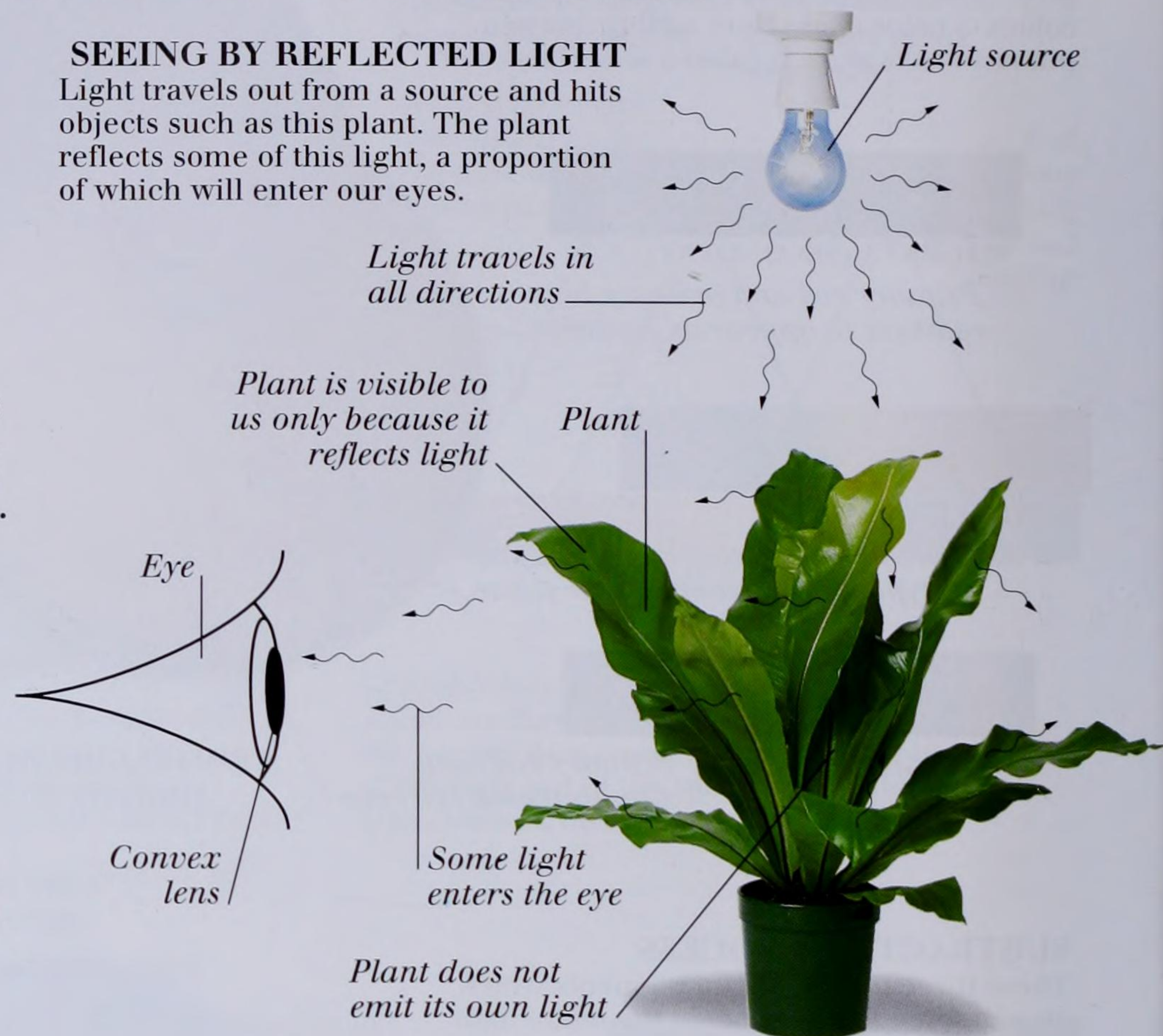


# Reflection and refraction

LIGHT IS A FORM OF **electromagnetic radiation** (see pp. 48–49). In free space, it travels in a straight line at 300 million meters per second. When a beam of light meets an object, a proportion of the rays may be reflected. Some light may also be absorbed, and some transmitted. Without reflection, we would only be able to see objects that give out their own light. Light always reflects from a surface at the same angle at which it strikes it. Thus parallel rays of light meeting a very flat surface will remain parallel when reflected. A beam of light reflecting from an irregular surface will scatter in all directions. Light that passes through an object will be **refracted**, or bent. The angle of refraction depends on the angle at which the light meets the object, and the material from which the object is made. **Lenses** and mirrors can cause light rays to **diverge** or **converge**. When light rays converge, they can reach a point of focus. For this reason, lenses and mirrors can form images. This is useful in binoculars and other optical instruments (see pp. 44–45).

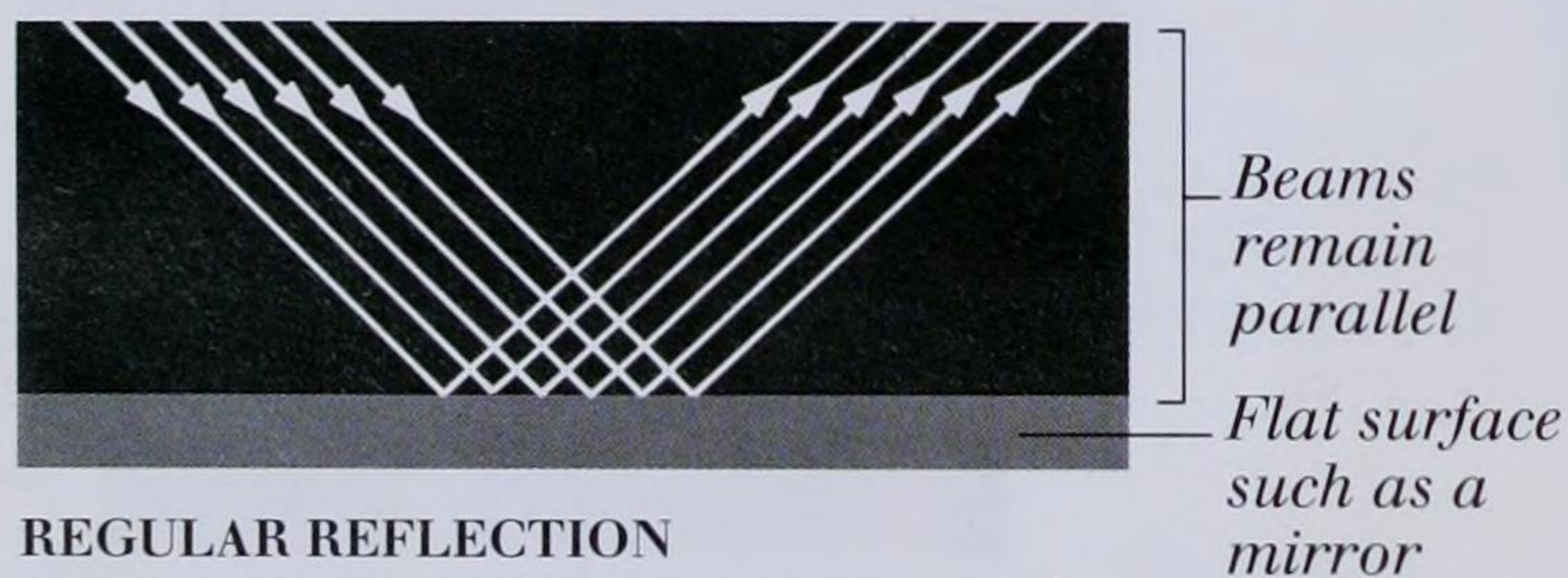
## SEEING BY REFLECTED LIGHT

Light travels out from a source and hits objects such as this plant. The plant reflects some of this light, a proportion of which will enter our eyes.

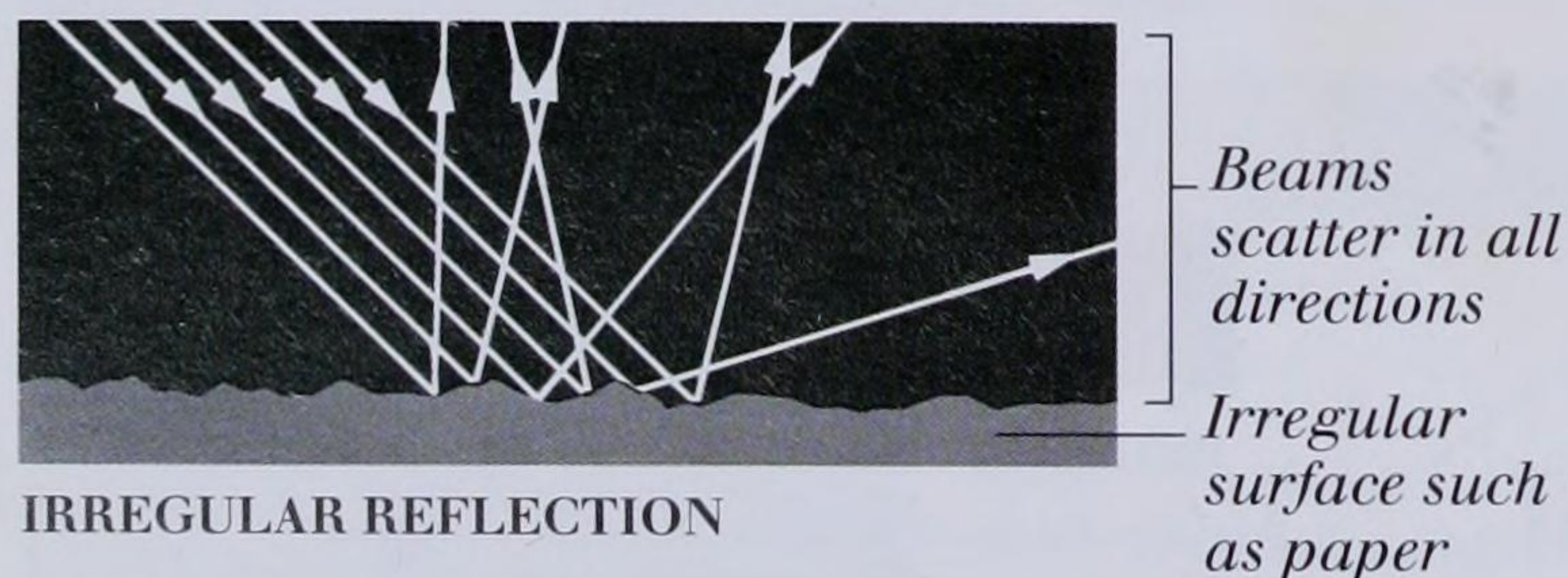


## REFLECTING AND REFRACTING

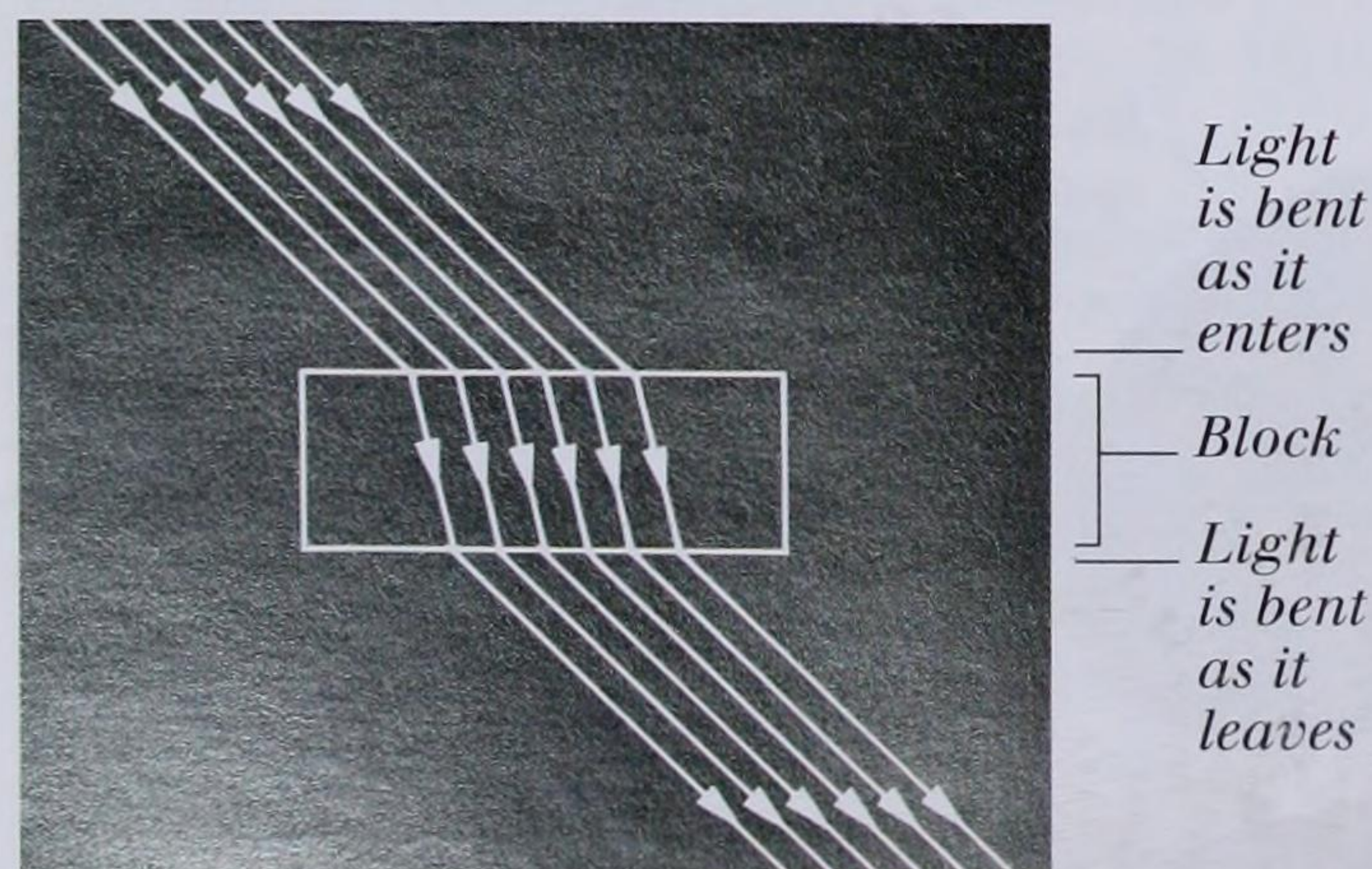
The illustrations below show what happens when parallel beams of light reflect regularly and irregularly and when they refract.



REGULAR REFLECTION



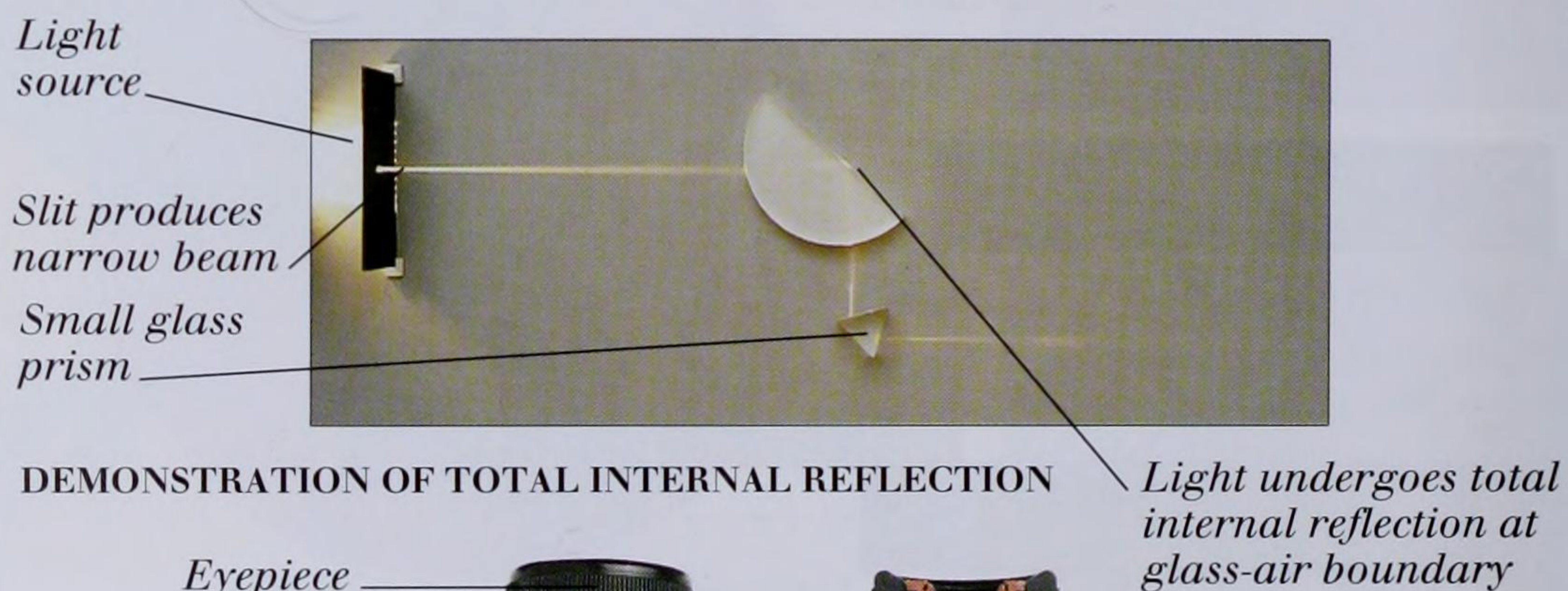
IRREGULAR REFLECTION



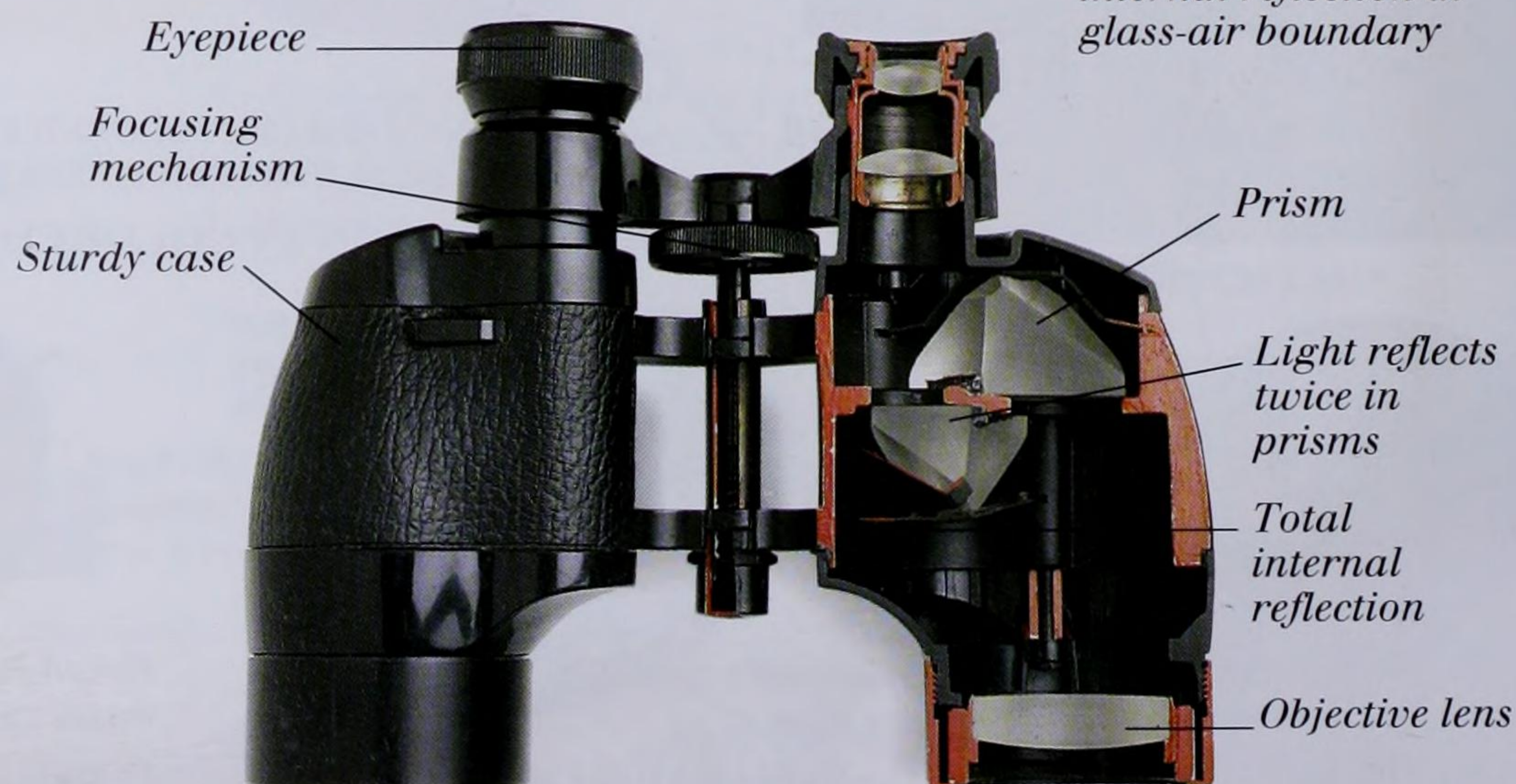
REFRACTION IN A GLASS BLOCK

## TOTAL INTERNAL REFLECTION

When light moves from one medium to another, for example from glass to air, some of the light will normally be reflected. When the light striking the boundary reaches a certain angle—the **critical angle**—all of the light reflects back. This is called **total internal reflection**. It is put to use in binoculars, where the light path is folded by prisms so that it can be contained within a compact case.



DEMONSTRATION OF TOTAL INTERNAL REFLECTION



BINOCULARS

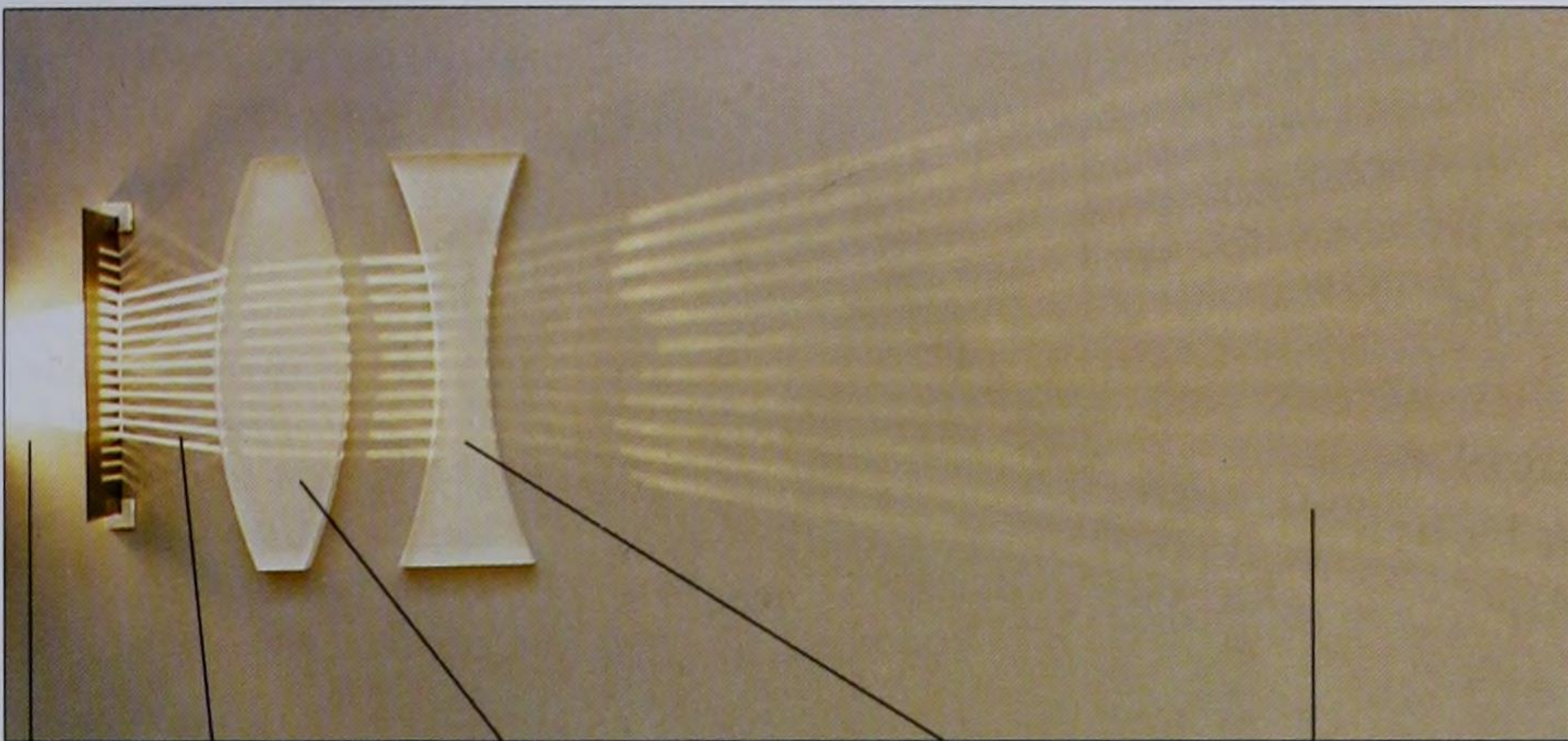


## LENSES AND MIRRORS

The images below show how beams of light from a bulb are affected by **concave** and **convex** mirrors and lenses. Convex lenses and mirrors

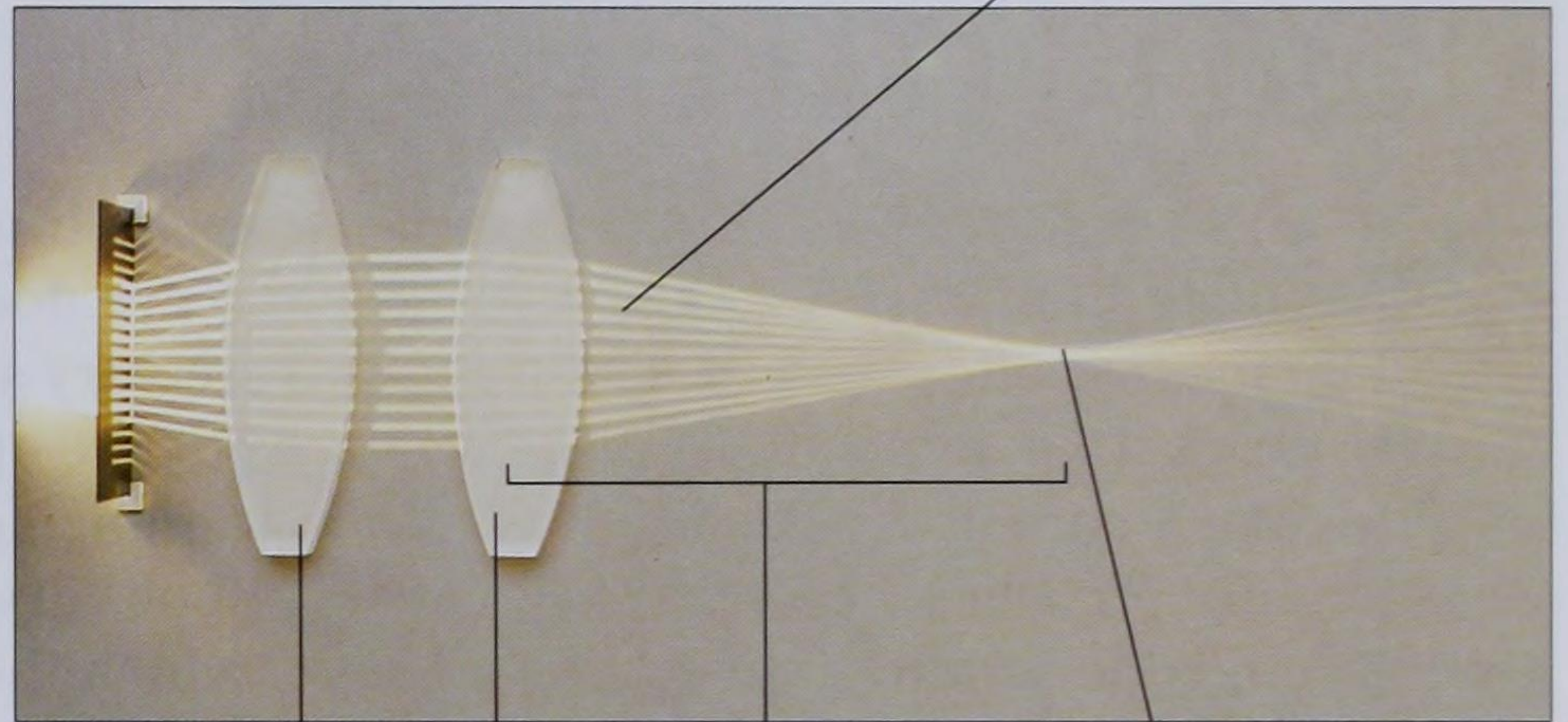
have surfaces that curve outward at the center, while concave lenses curve inward and are thicker at the edges.

### CONCAVE LENS (BENDS LIGHT OUTWARD)



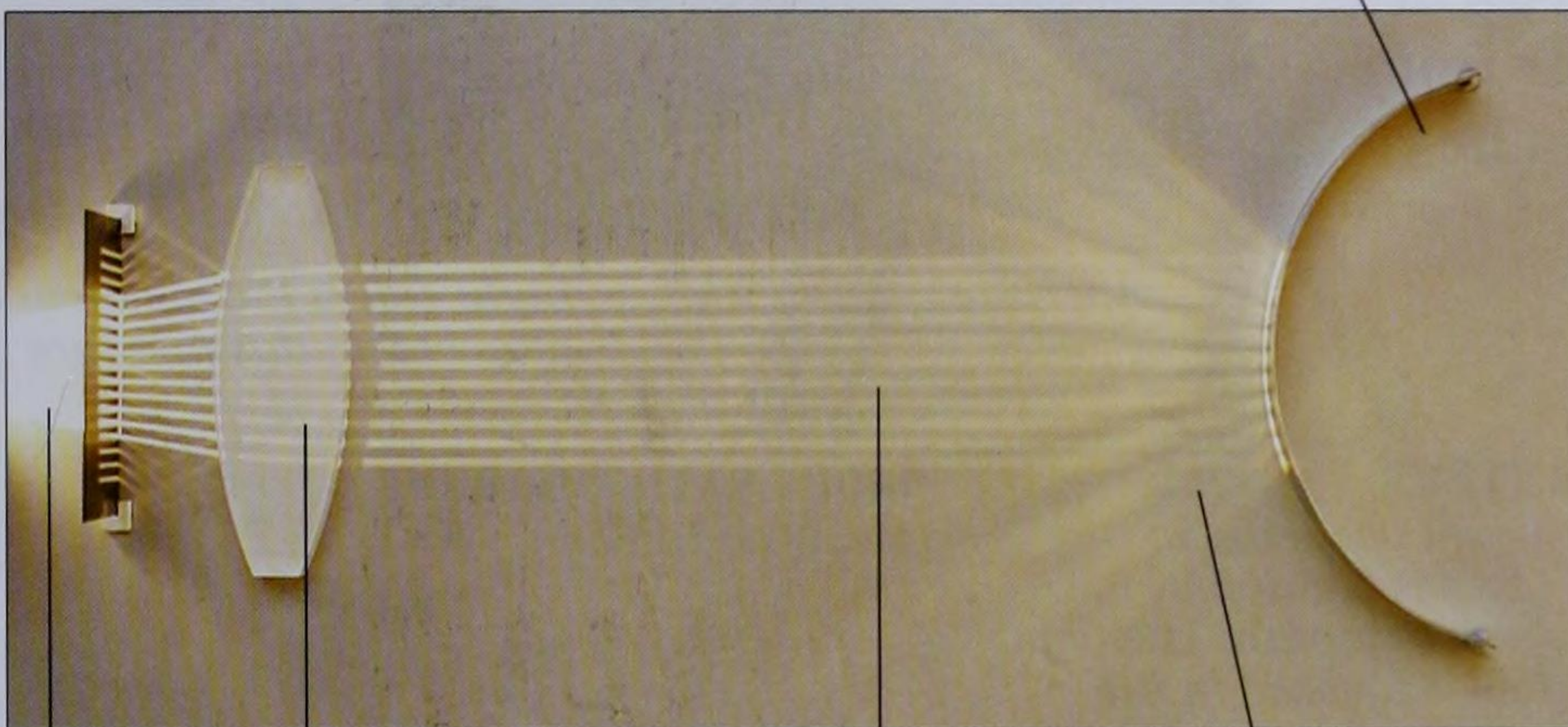
Light source  
Light rays travel out in straight lines  
Convex lens bends diverging rays into straight beam  
Concave lens  
Light rays diverge  
Convex mirror

### CONVEX LENS (BENDS LIGHT INWARD)



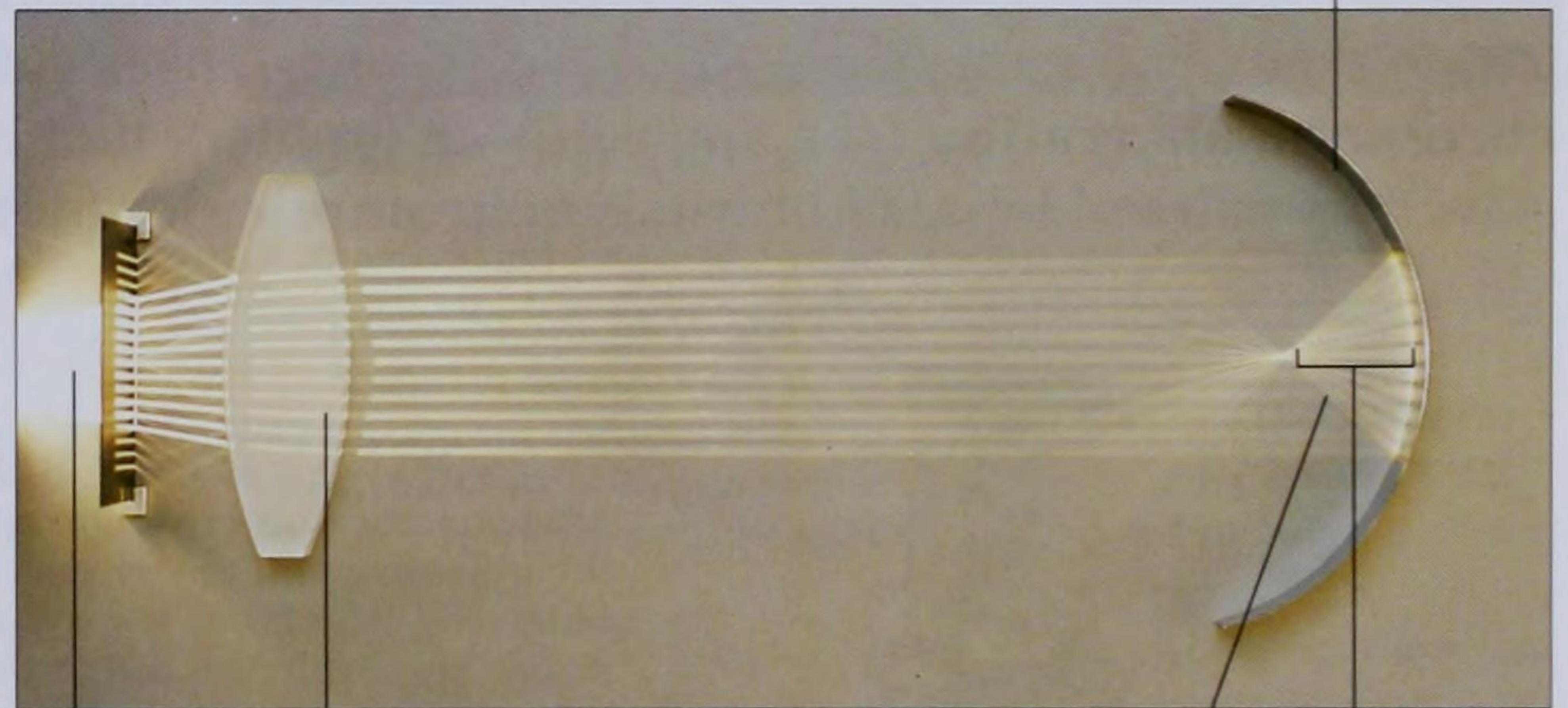
First convex lens produces parallel beam  
Convex lens  
Focal length  
Light focused to a point  
Concave mirror

### CONVEX MIRROR (REFLECTS LIGHT OUTWARD)



Light source  
Convex lens bends diverging rays into a straight beam  
Parallel light rays  
Light rays diverge as they reflect

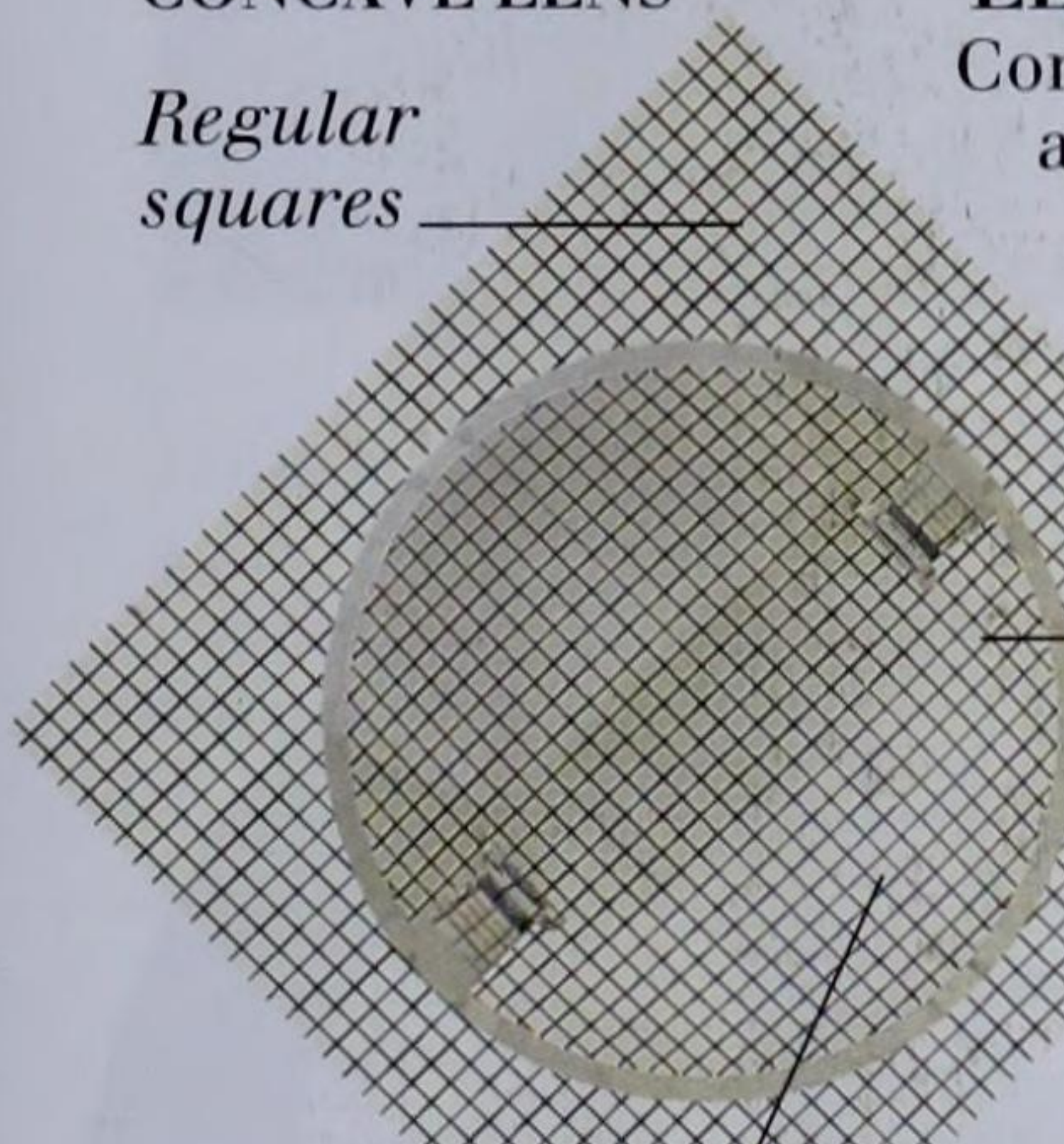
### CONCAVE MIRROR (REFLECTS LIGHT INWARD)



Light source  
Convex lens bends diverging rays into straight beam  
Light rays converge as they reflect  
Focal length

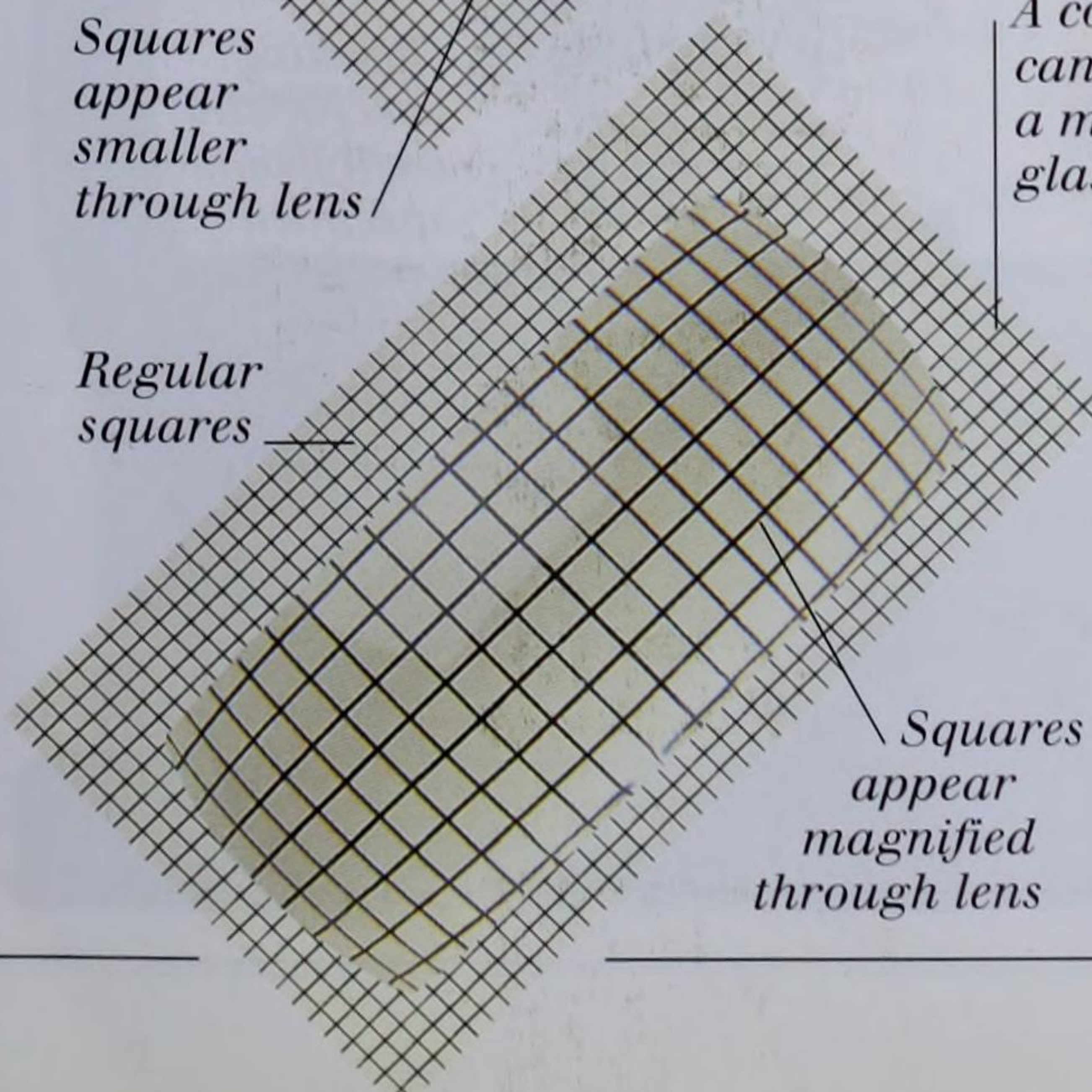
### CONCAVE LENS

Regular squares



Squares appear smaller through lens

Regular squares



Squares appear magnified through lens

### LENSES

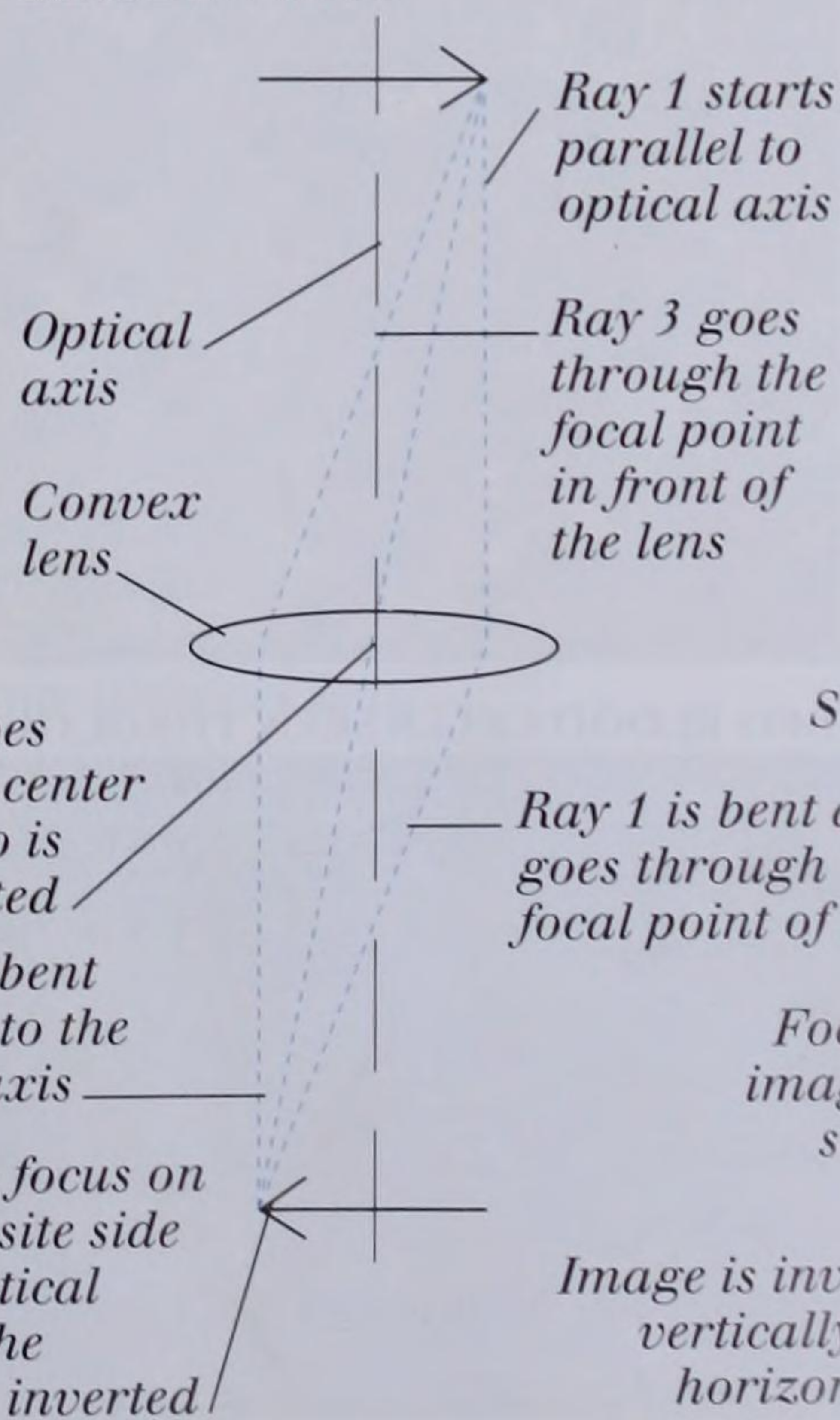
Concave lenses make objects appear smaller, and allow a larger field of vision. Objects lying within the focal length of a convex lens appear larger.

A concave lens is often fitted to the rear window of a vehicle to improve a driver's field of vision

### CONVEX LENS

A convex lens can be used as a magnifying glass

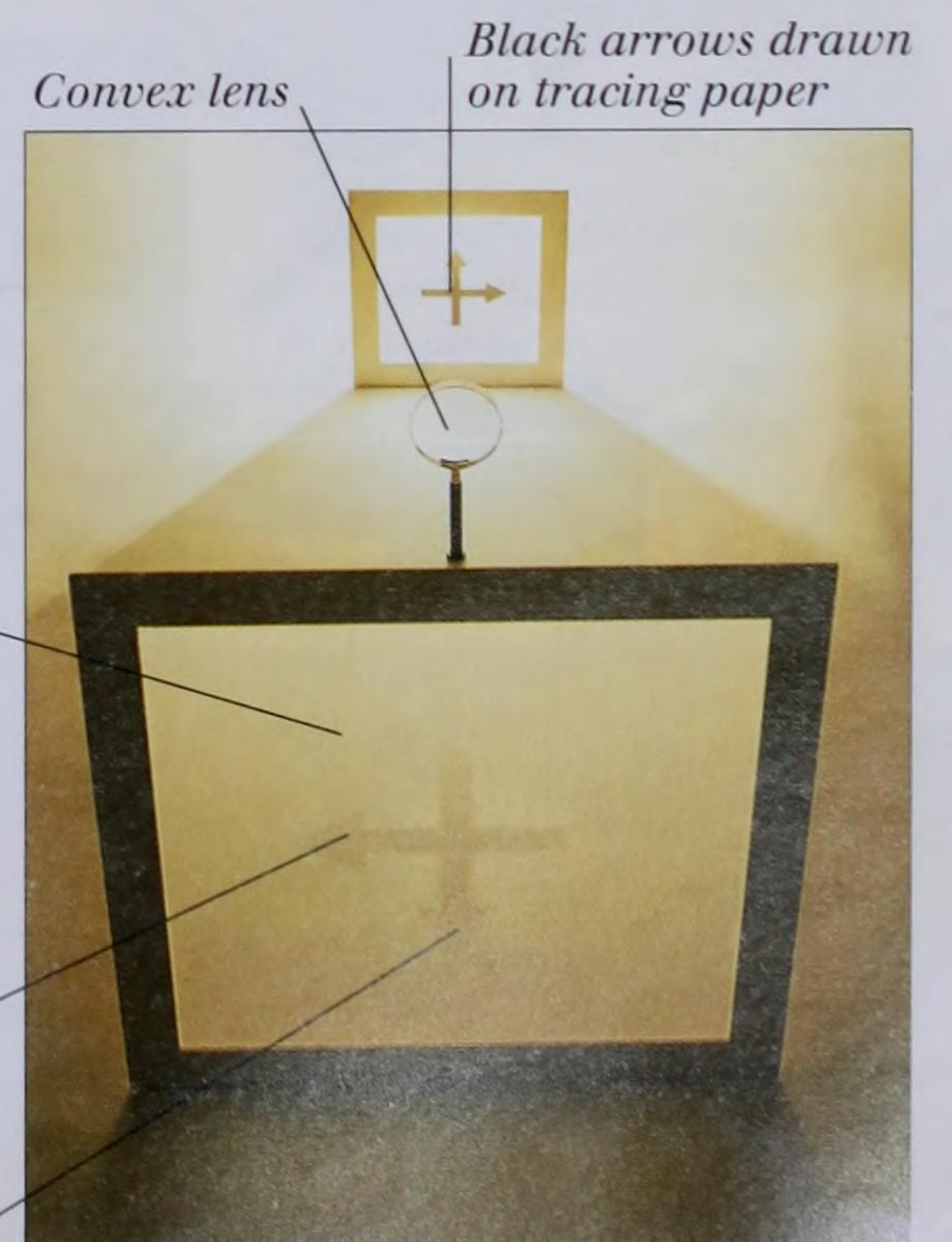
### IMAGE INVERTS



### IMAGE FORMATION

Because they focus light, convex lenses can be used to project images onto a screen. The screen must be placed at a point where the rays focus in order for a clear image to be produced. Only objects that lie within a range of distances from the lens, called the **depth of field**, will be in focus at any one time.

### PROJECTED IMAGE



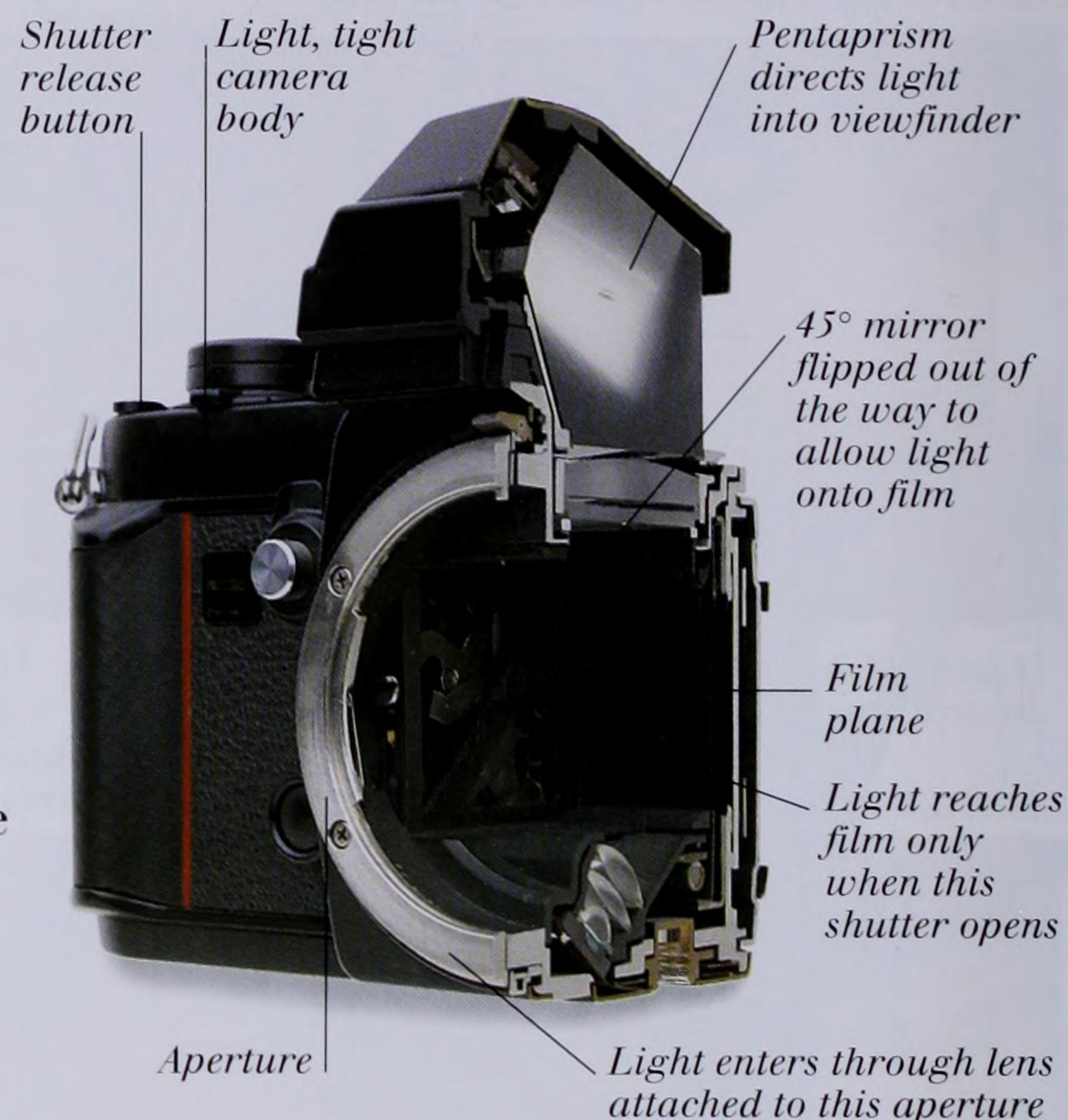


# Optical instruments

THE HUMAN EYE CONTAINS A LENS that produces an **image** by focusing the light that passes through it. But the eye does not record images, or allow us to see objects that are very small, or very far away. To achieve these tasks, we need to use optical instruments. A camera, for example, records an image on light-sensitive film. To see objects that are very small or very far away, we need to produce a magnified image, which the eye can then observe. By using a compound microscope, light from a very small object can be made to produce a magnified image. A telescope produces a magnified image in a similar way to a microscope, using lenses to focus light. There are limits to the use of optical instruments. Even the most precise lenses suffer from **chromatic aberration** (see opposite)—a problem that can be solved by pairs of lenses known as **achromatic doublets**.

## CAMERA

The cutaway view below shows the main features of a single lens reflex (SLR) camera. The light is focused onto film at the back of the camera by a lens or a combination of lenses.



## COMPOUND MICROSCOPE

Together, the lenses in a microscope produce a magnified image

## MICROSCOPES

A basic optical microscope consists of two lenses, the objective and the eyepiece. Light is focused by the objective, which has a very small **focal length**. The light passes up inside the body tube, and is focused by the eyepiece into the eye.

## LIGHT PATH THROUGH MICROSCOPE

Light enters eye

Eyepiece lens focuses light

Light passes through body of microscope

Objective lens focuses light from specimen

Prepared specimen placed on slide

Mirror directs light through the specimen

Eyepiece

Focusing knob

Different objective lens gives different magnification

Objective lens

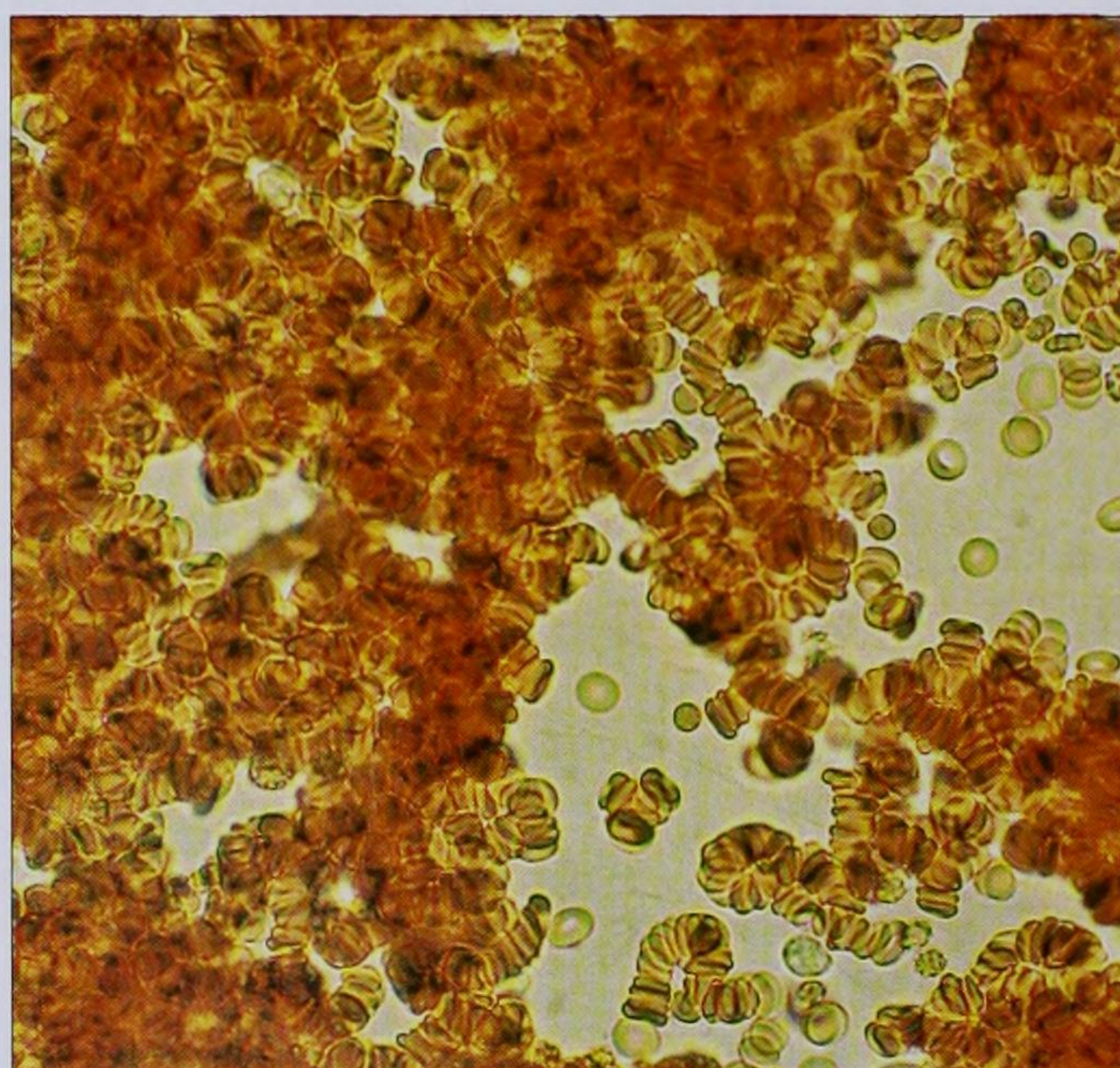
Three-objective turret

Glass slide

Specimen Stage

Angled mirror

Base



RED BLOOD CELLS SEEN THROUGH MICROSCOPE



## TELESCOPES

At the front of a **refracting** telescope is an objective lens that collects light and focuses it, producing an image in the telescope's tube. The eyepiece greatly magnifies this image. A **reflecting** telescope uses a mirror instead of an objective lens.

*Small viewfinder telescope used to position larger telescope*

*Equatorial mount allows user to track stars as they move across the sky*  
*Focusing knob*  
*Prism bends light into eyepiece*  
*Eyeiece*

SMALL REFRACTING TELESCOPE

*Eyeiece*  
*Concave eyepiece*

*Objective lens*

*Telescope tube*

*Equatorial mount allows user to track stars as they move across the sky*

*Counterbalance*

*Tripod*



### TELESCOPE IMAGE OF THE MOON

In a telescope, the greater the difference between focal lengths of objective and eyepiece, the greater the magnification. A larger objective lens will yield a brighter image for a given magnification.

### REFRACTING TELESCOPE

A **convex** objective lens focuses light from a distant star. The eyepiece is fixed in a tube that can be moved in and out of the telescope tube.

*Two rays show the path of light through the telescope*

*Convex objective lens*



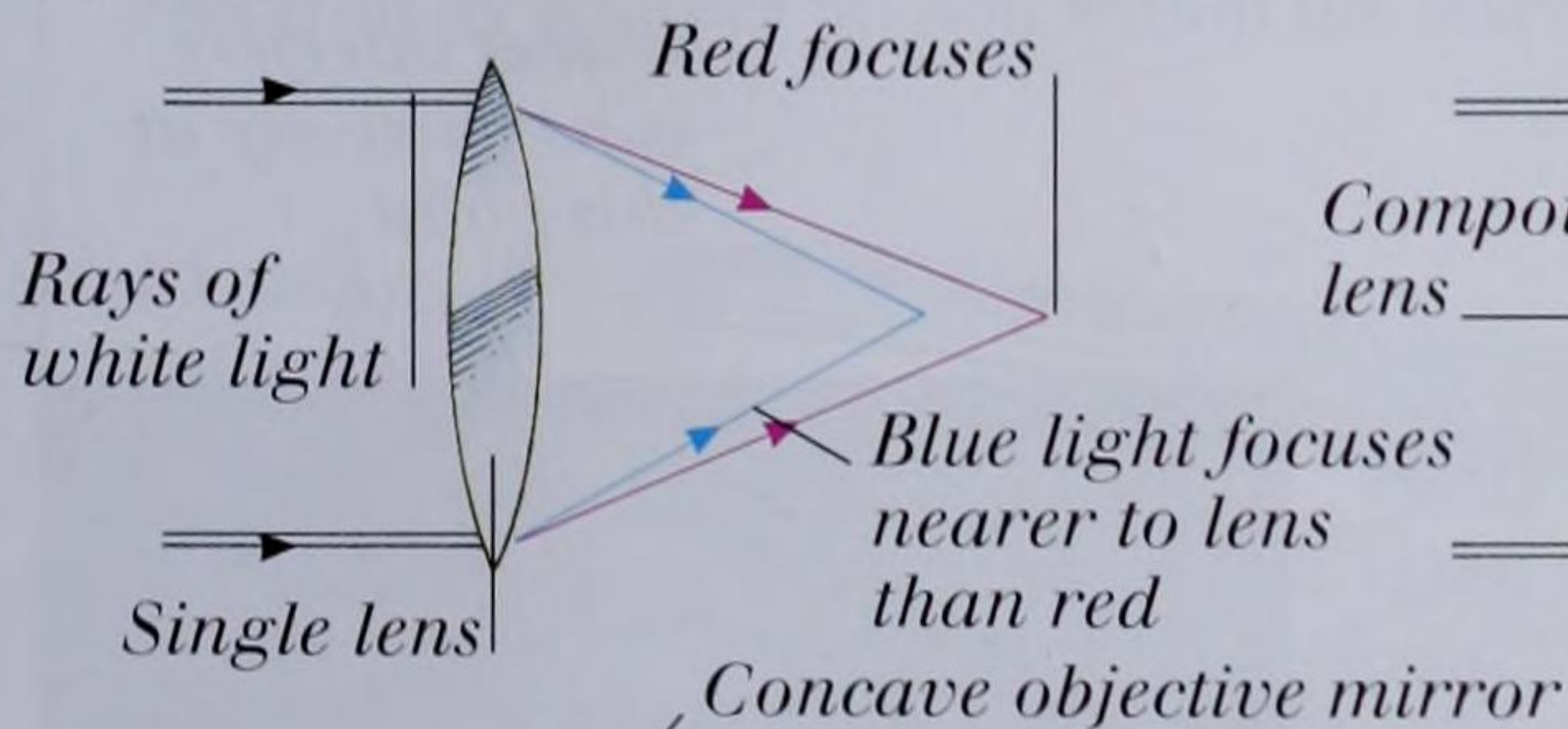
*Focal point of objective lens*

*Telescope tube*

*Light from the object enters through the objective lens*

### ACHROMATIC DOUBLET

A single lens will refract light of different wavelengths by different amounts. For example, red light focuses at a different point from blue light. This chromatic aberration can be eliminated by using two lenses made of different types of glass.



*All colors focus at the same point*

*Light from object reflects from objective lens*

*Eyeiece*

*Eyeiece lens*

*Field lens*

*Flat mirror*

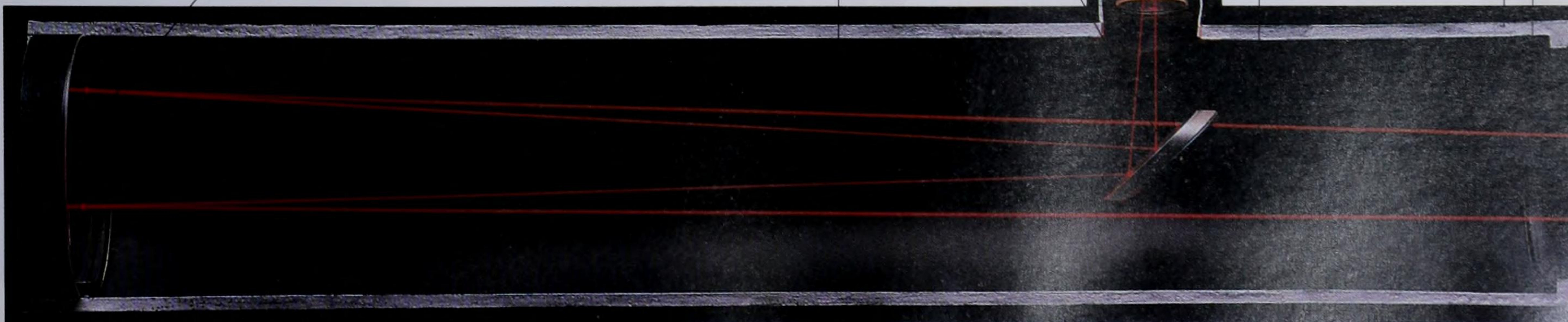
### REFLECTING TELESCOPE

In this case, a **concave** objective mirror focuses the light. A small plane (flat) mirror directs light through a tube off the side of the telescope to the eyepiece.

*Light from objective would focus here*

*Two rays show the path of light*

*Telescope tube*





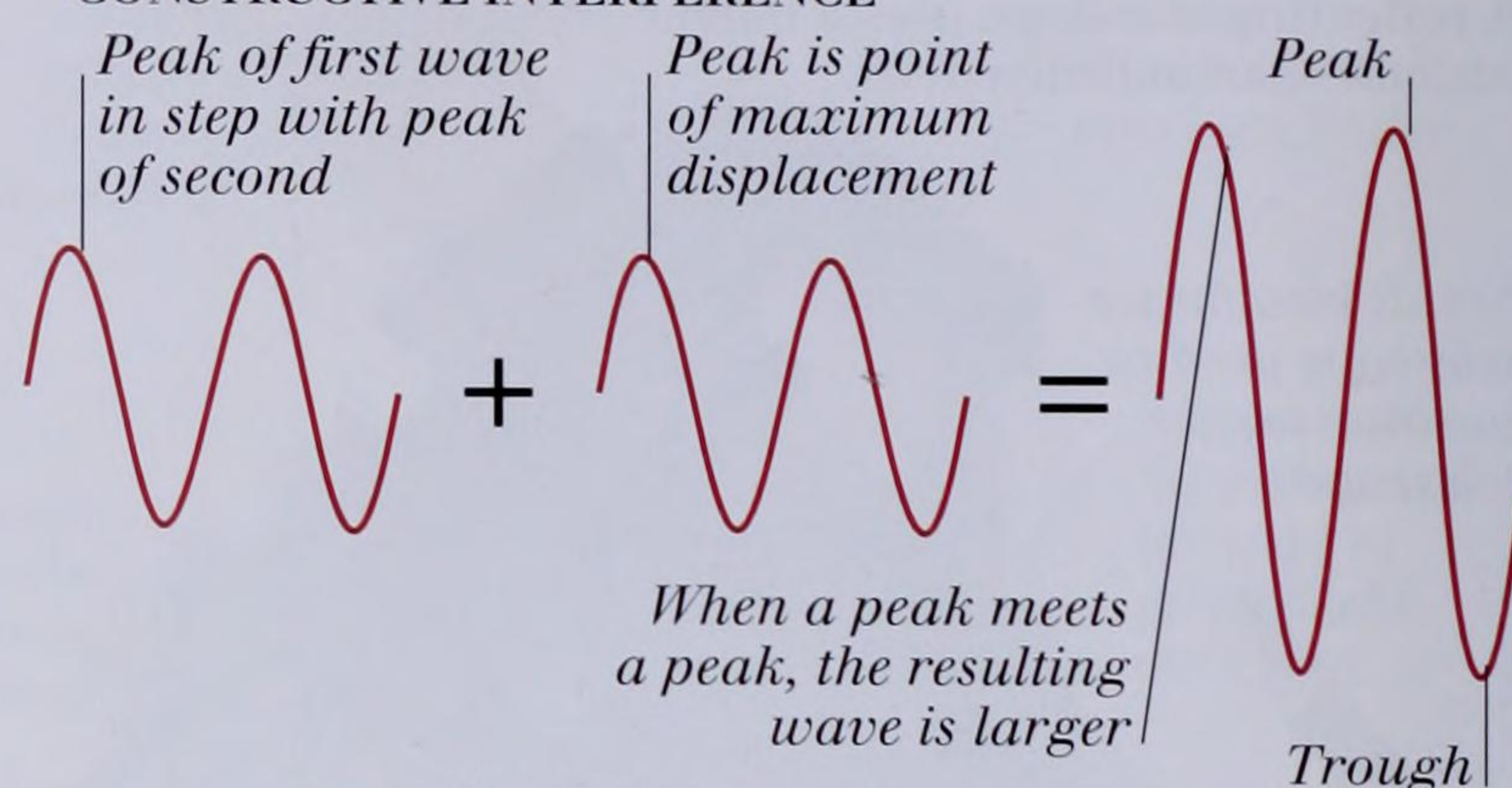
# Wave behavior

ALL TYPES OF WAVE CAN COMBINE OR INTERFERE. If two waves are in step, so that the peaks coincide, the **interference** results in a wave that will be larger than the original one (**constructive interference**). If the waves are out of step, the peak of one wave will cancel out the trough of another (**destructive interference**). Where the waves are equal in size, they can cancel out entirely. As waves pass around objects or through small openings, they can be **diffracted**, or bent. Diffraction and interference can be observed in water waves, using a ripple tank. The colors seen in soap bubbles are the result of some colors being removed from the white light spectrum by destructive interference. Light reflected off the front and back surfaces of the film interferes. One source of very pure light waves is a laser. The light produced by a laser is coherent. This means that all of the waves are in step and of exactly the same wavelength. The light is produced by a process called stimulated emission. To understand this process, light must be thought of in terms of particles called **photons** (see p.48), as well as waves.

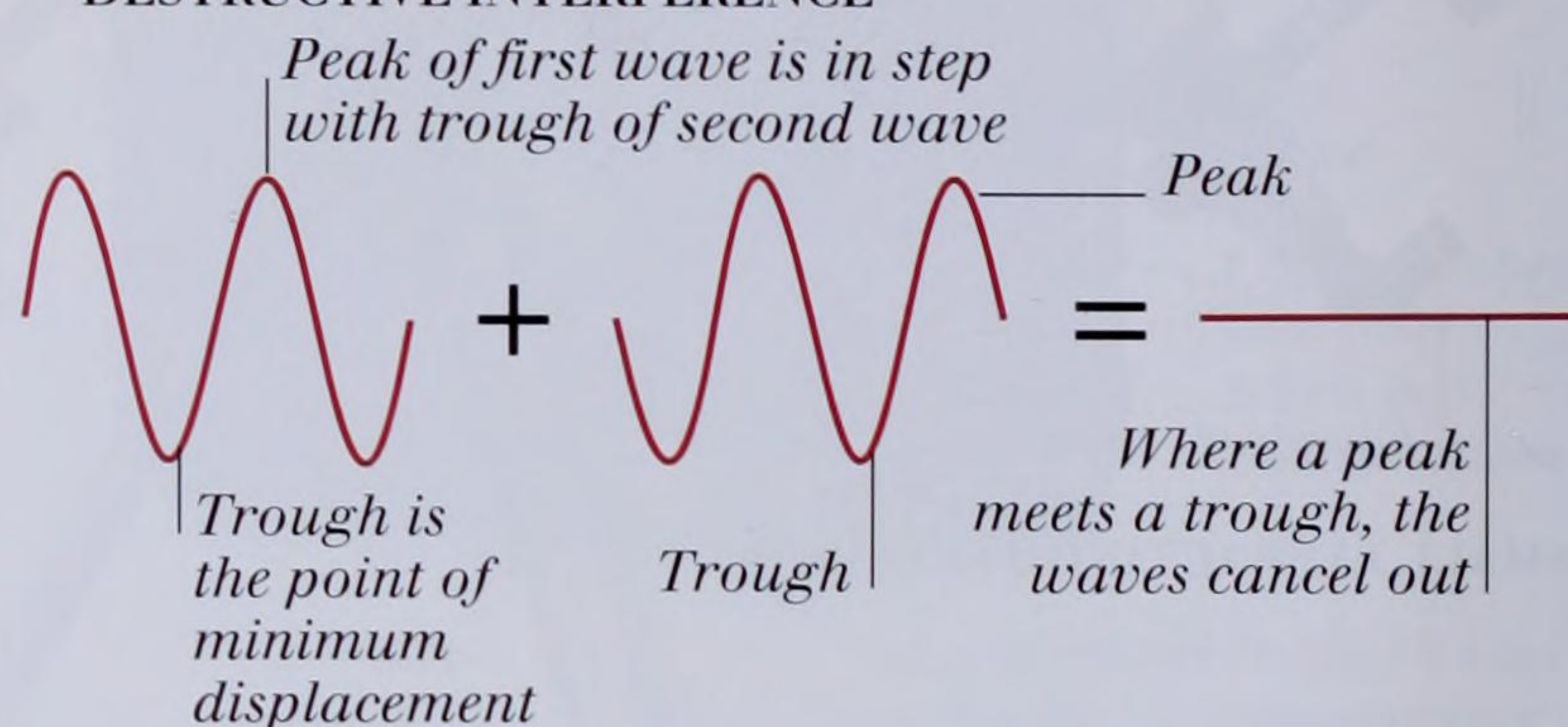
## PRINCIPLE OF SUPERPOSITION

When two waves meet, they add up or interfere. This is called the **Principle of Superposition**, and is common to all types of wave.

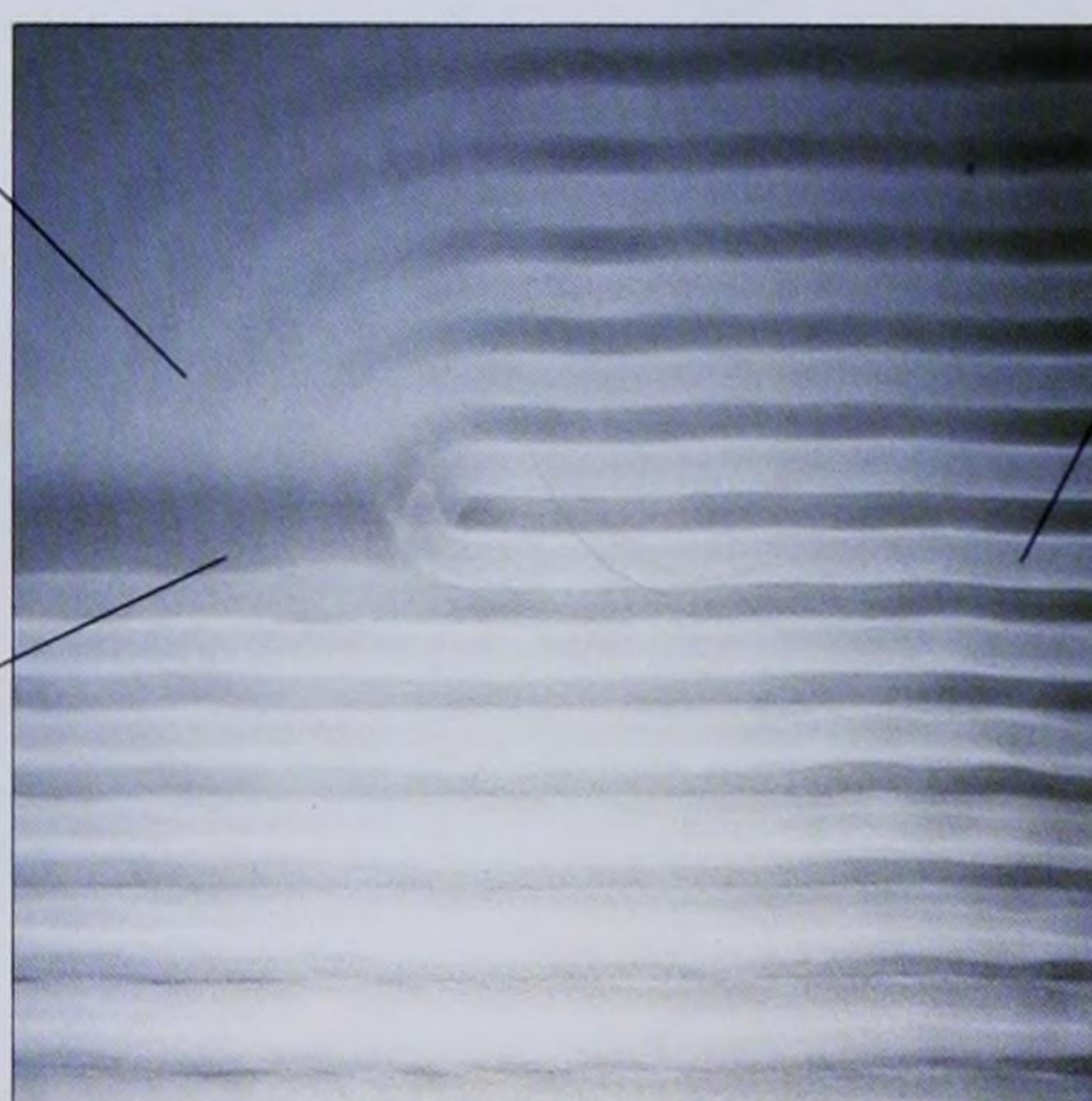
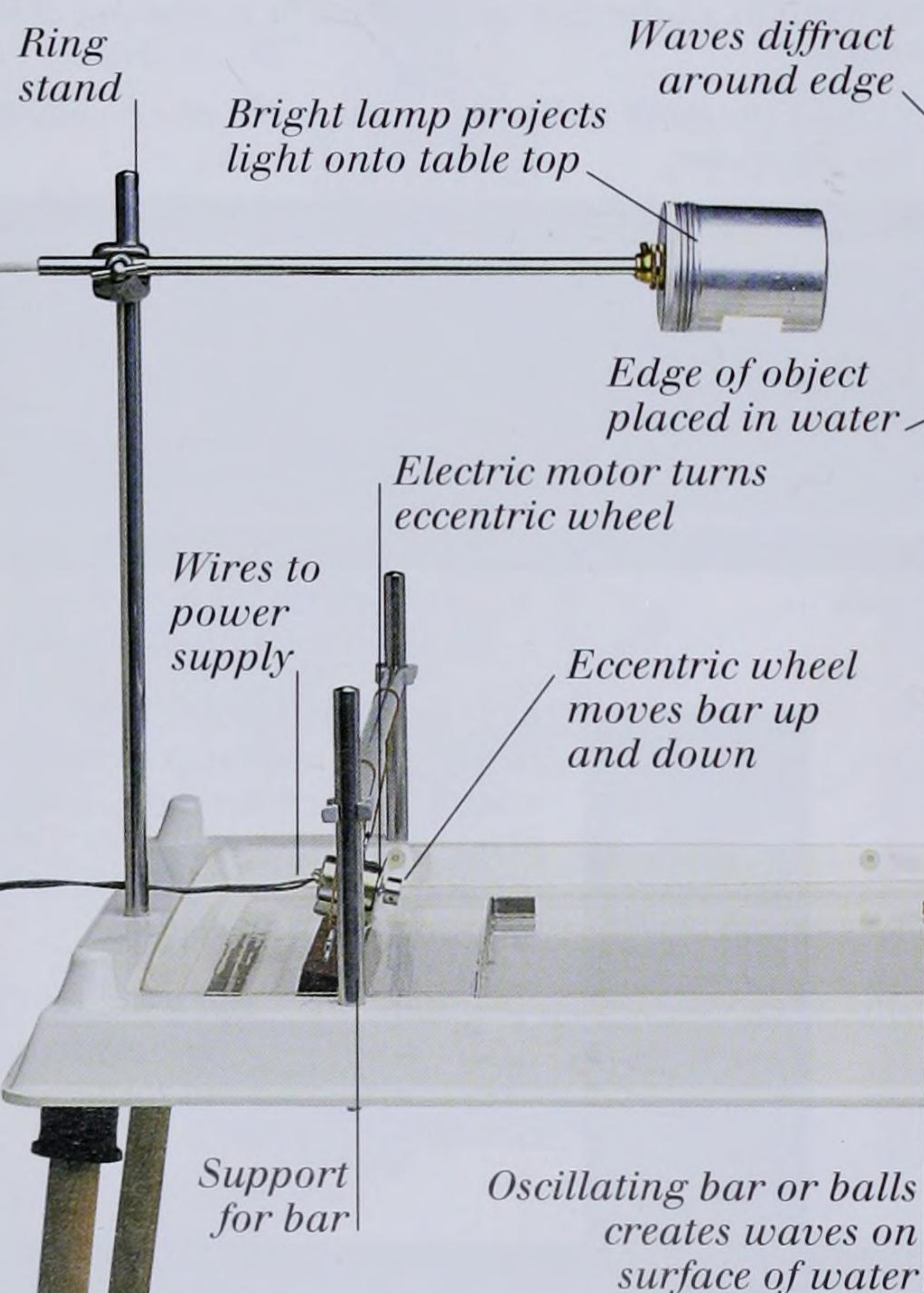
### CONSTRUCTIVE INTERFERENCE



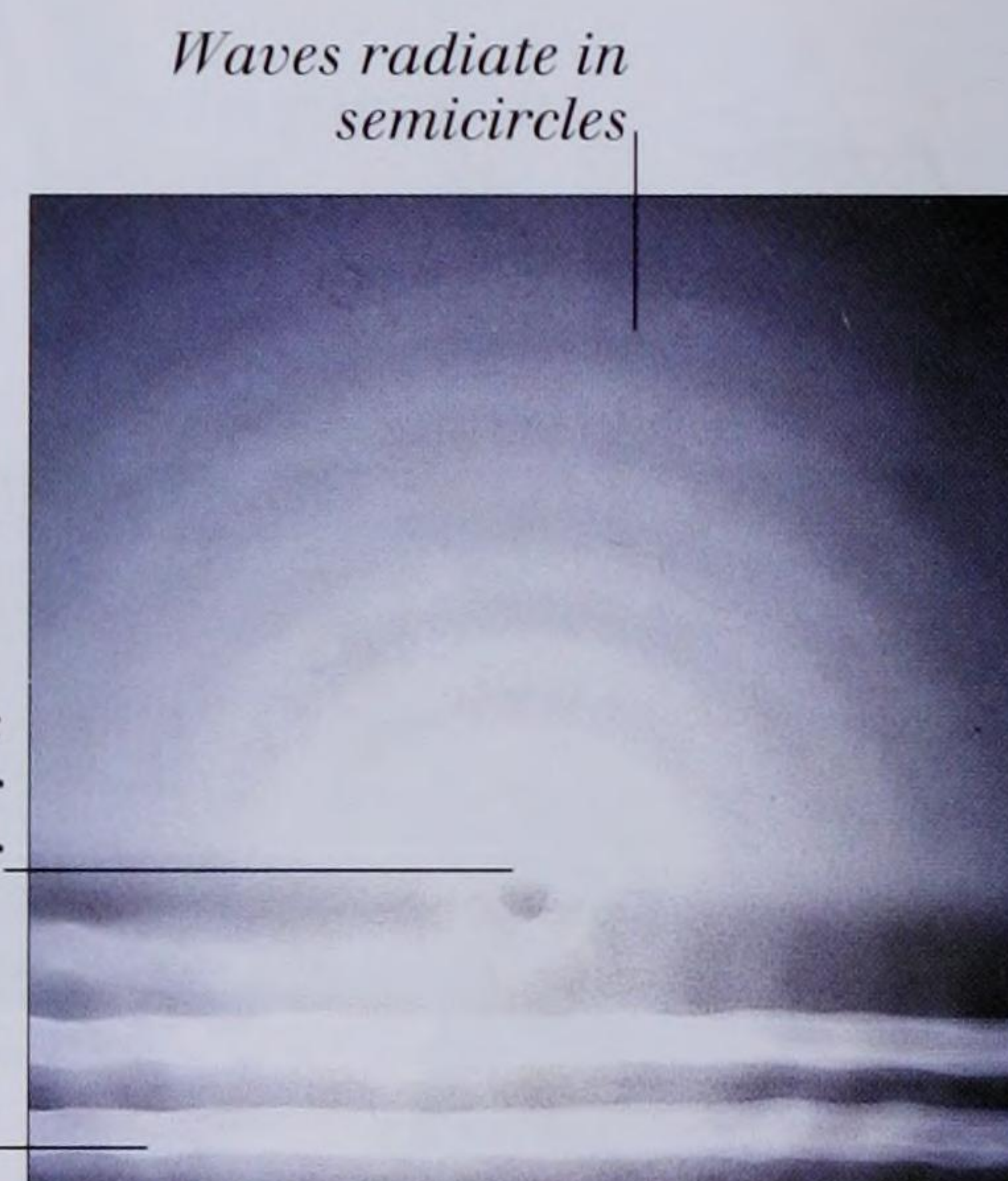
### DESTRUCTIVE INTERFERENCE



## DIFFRACTION AND INTERFERENCE



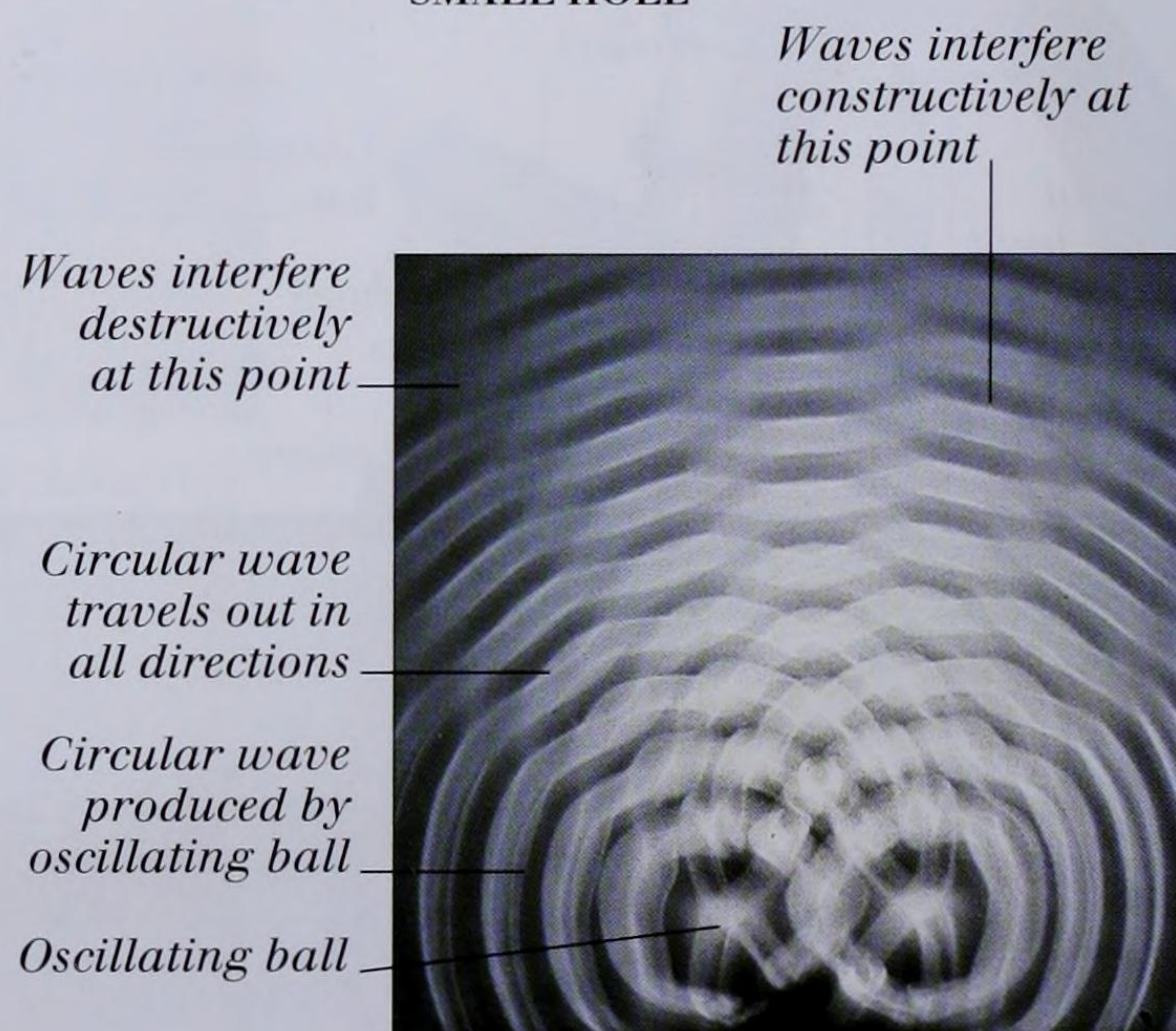
EDGE DIFFRACTION



DIFFRACTION THROUGH SMALL HOLE

### RIPPLE TANK

Diffraction and interference are probably best observed using a ripple tank. A bar moving up and down (oscillating) creates **plane waves** in shallow water. These waves bend around edges and produce semicircular waves after passing through a small hole.

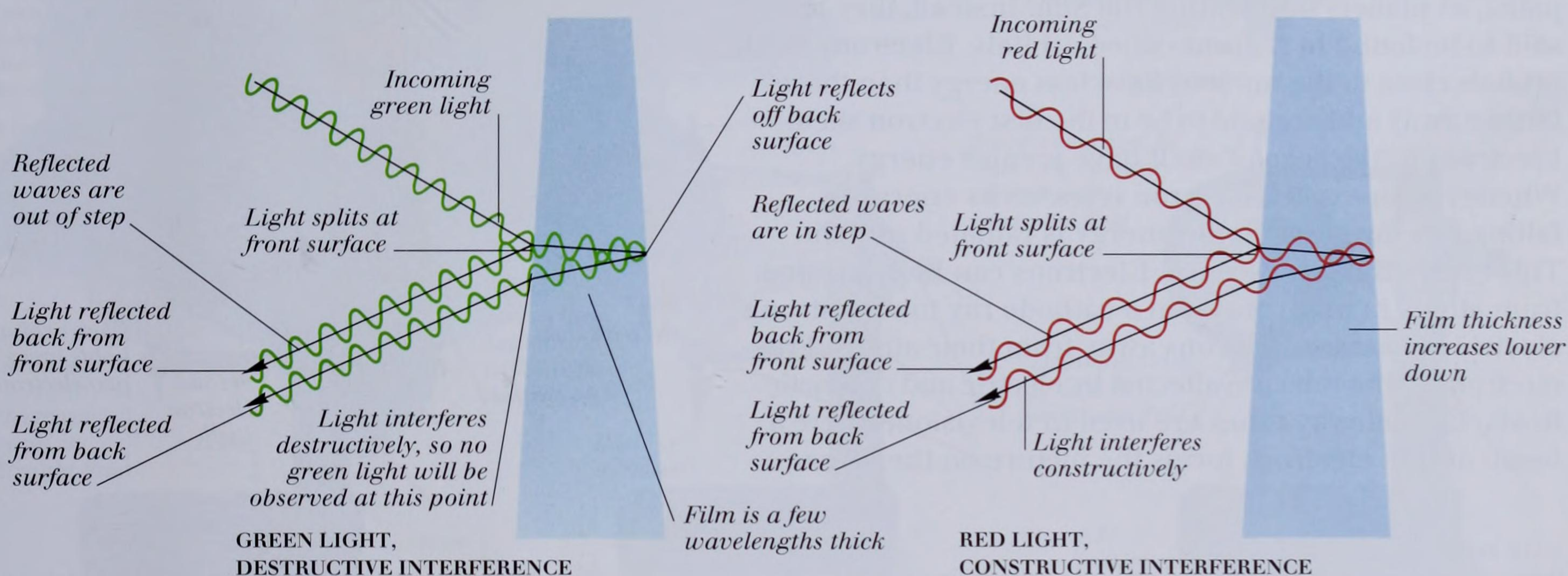
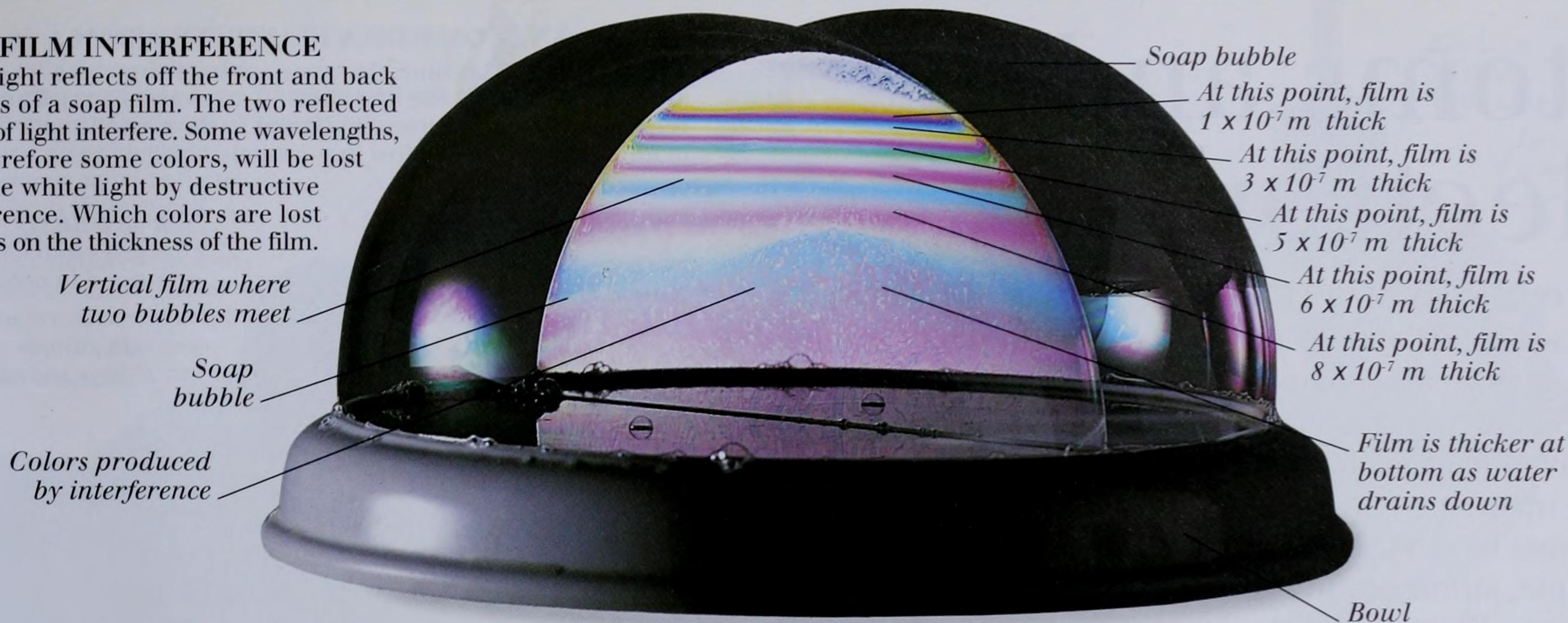


INTERFERENCE



## THIN FILM INTERFERENCE

White light reflects off the front and back surfaces of a soap film. The two reflected beams of light interfere. Some wavelengths, and therefore some colors, will be lost from the white light by destructive interference. Which colors are lost depends on the thickness of the film.



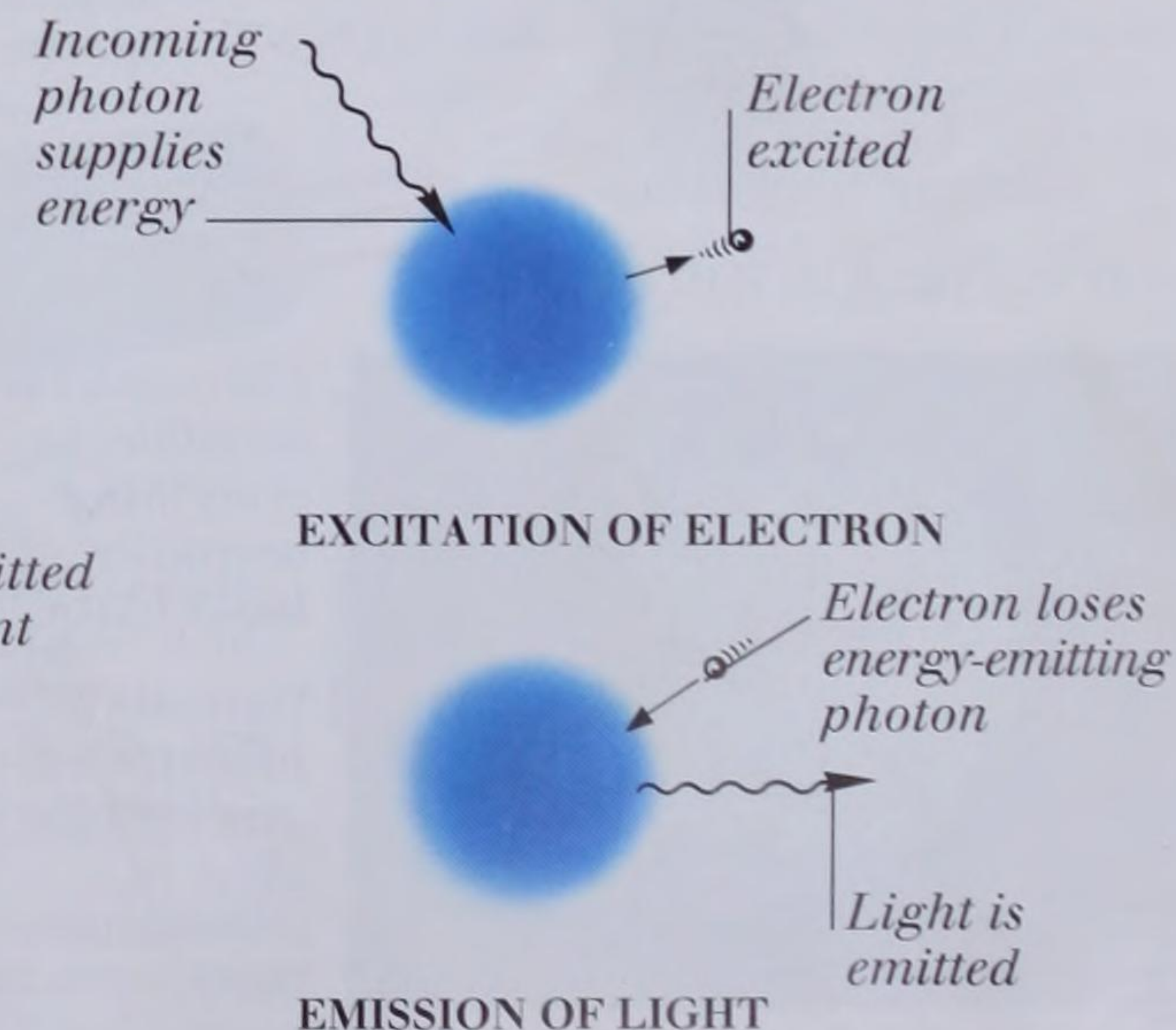
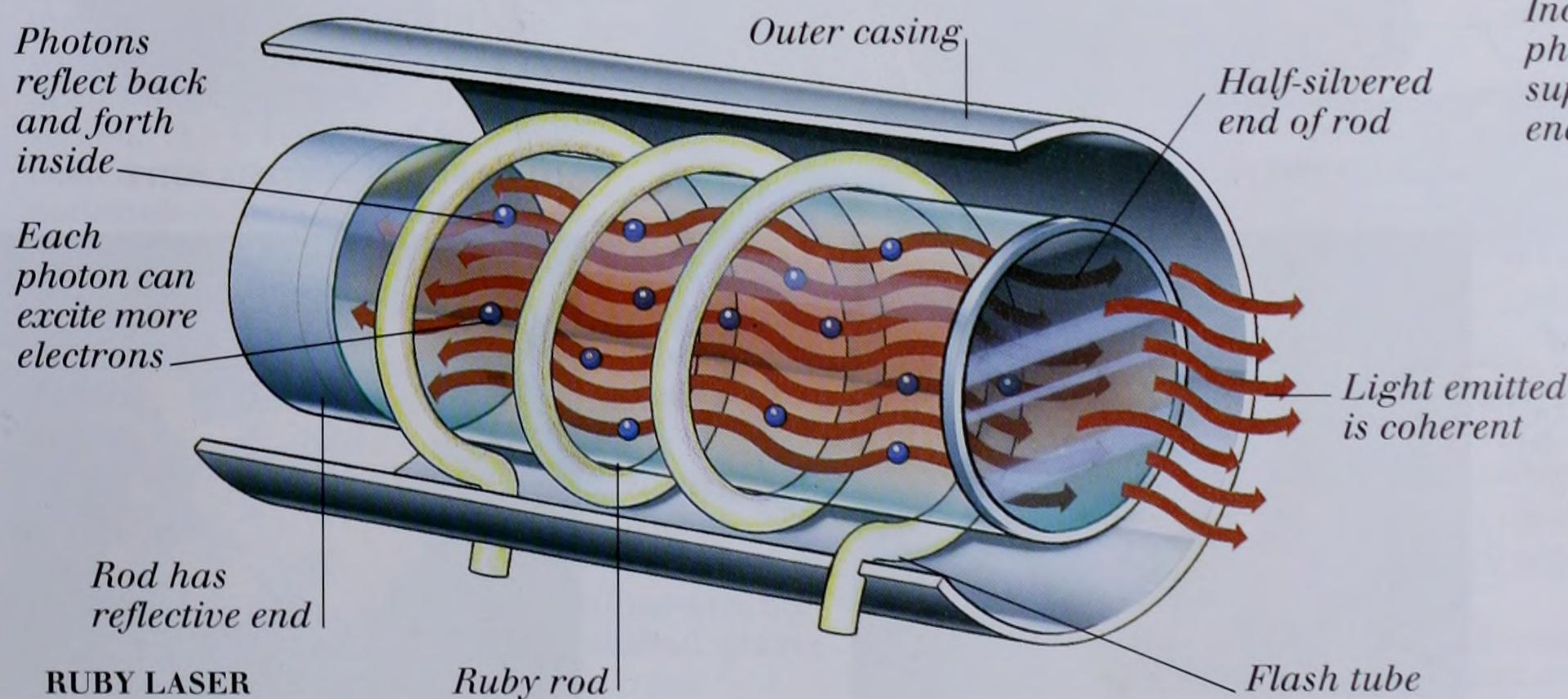
## LASERS

### STIMULATED EMISSION

Light behaves as waves and particles. Laser light exhibits all of the behavior common to waves, including **interference** and **diffraction**. But to understand the operation of a laser, light must be thought of as being composed of particles called **photons**. Each photon is emitted as the result of the stimulation of an excited **electron** by another photon within the laser.

### LASER SURGERY

Lasers have found many applications in medicine. Here, a laser is being used to destroy a cataract in a patient's eye.



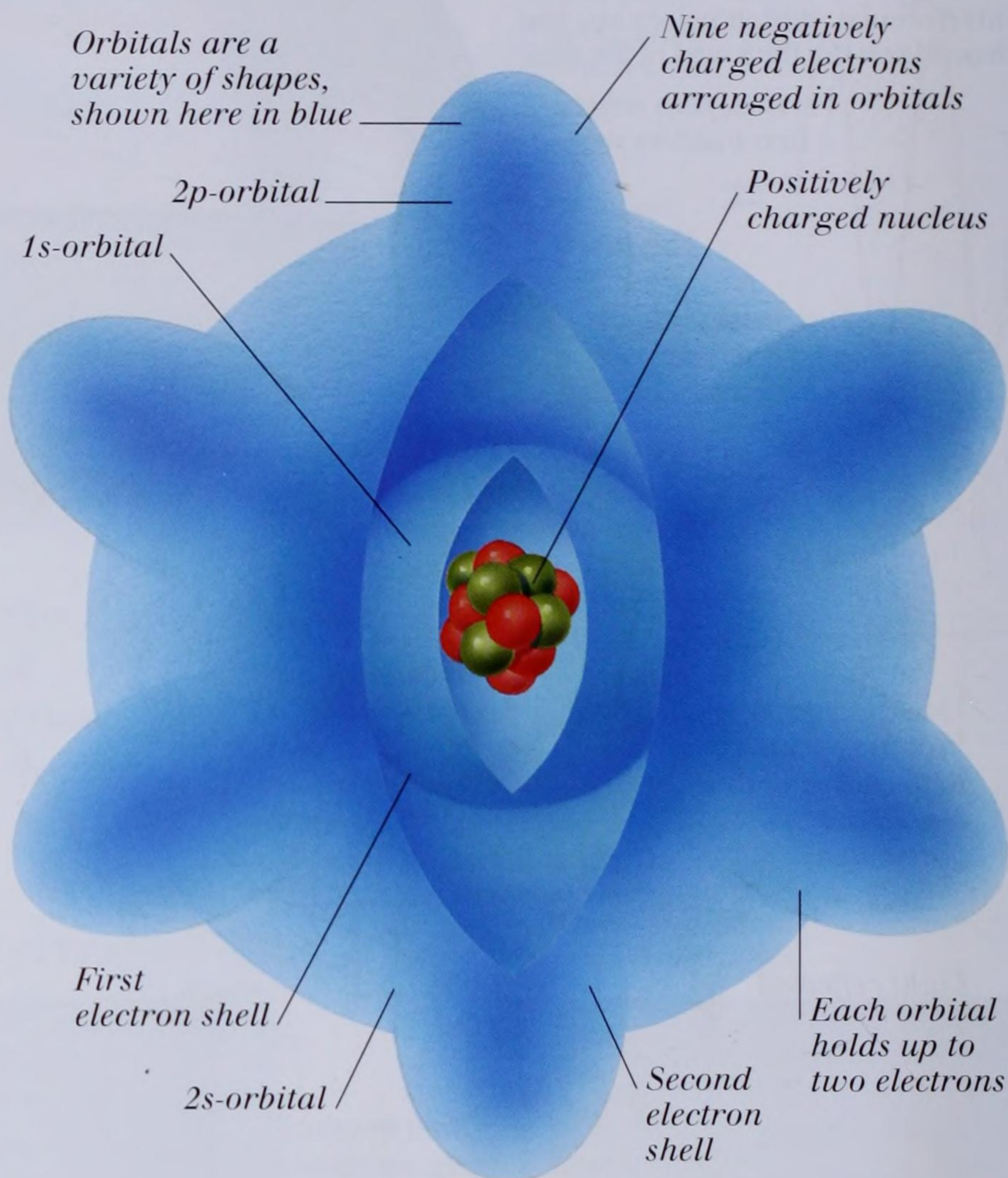


# Atoms and electrons

ALL ORDINARY MATTER (see pp. 6–7) consists of tiny particles called atoms. There are 92 naturally occurring types of atom. Each consists of a central, positively charged nucleus (see pp. 50–51), surrounded by negatively charged electrons. An element is a substance made up of one type of atom only. Atoms of different elements have different numbers of electrons. For example, atoms of the element fluorine have nine electrons. Electrons in the atom do not follow definite paths, as planets do, orbiting the Sun. Instead, they are said to be found in regions called orbitals. Electrons in orbitals close to the nucleus have less energy than those farther away and are said to be in the first electron shell. Electrons in the second shell have greater energy. Whenever an excited electron releases its energy by falling to a lower shell, the energy is radiated as light. This is called luminescence. Electrons can be separated from atoms in many ways. In a cathode ray tube, a strong electric field tears electrons away from their atoms. Free electrons in the tube are affected by electric and magnetic fields. Cathode ray tubes are used in television, where a beam of free electrons forms the picture on the screen.

## ANATOMY OF A FLUORINE ATOM

A fluorine atom has nine electrons around its nucleus. There are two electrons in the first shell, in an s-orbital (1s). The remaining seven electrons are found in the second shell, two in an s-orbital (2s) and five in p-orbitals (2p).

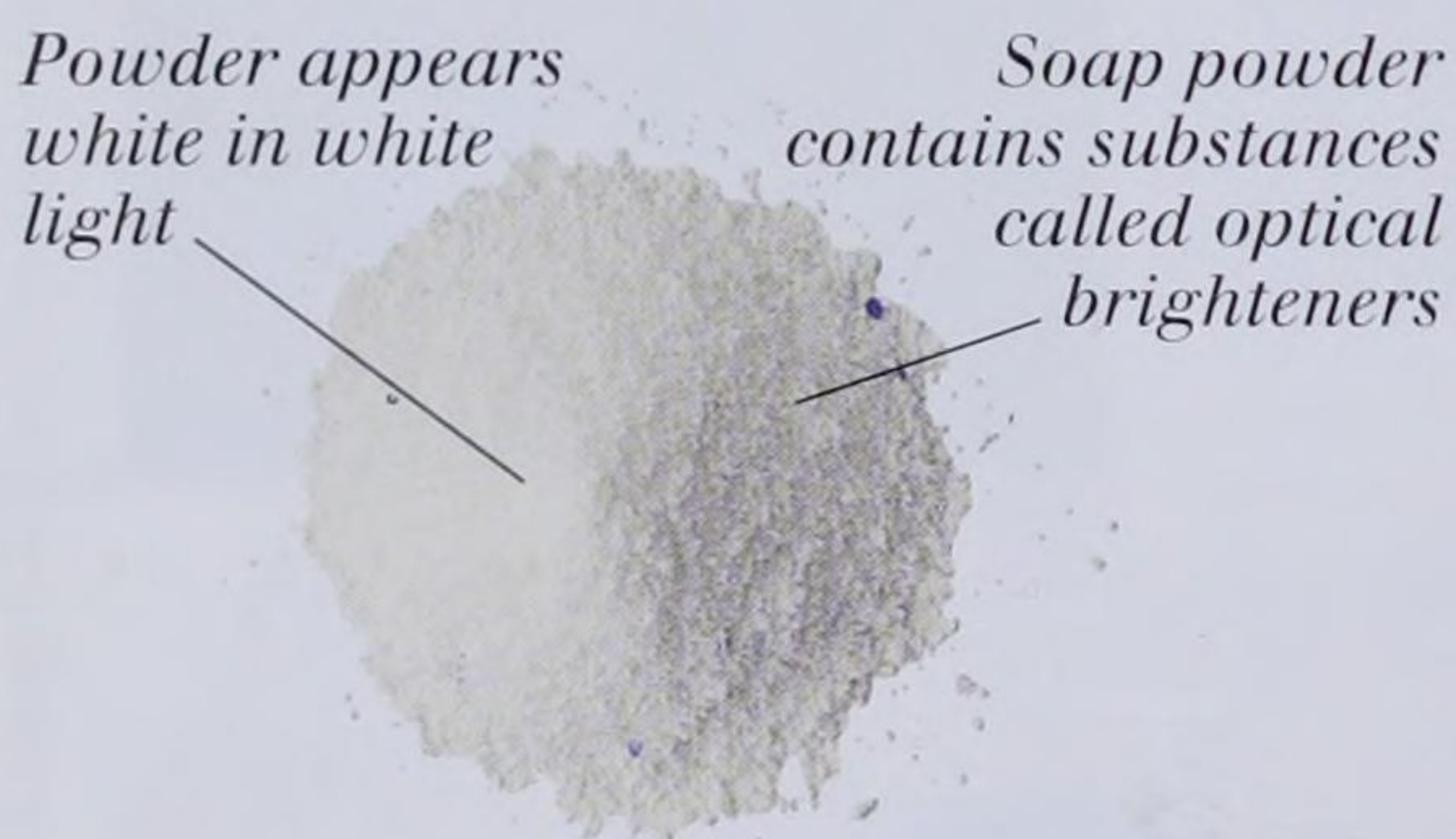


## FLUORESCENCE

### SOAP POWDER

How does soap powder make clothes appear so bright? Soap powder exhibits a form of luminescence (see above) called **fluorescence**. Electrons around atoms in the powder are excited into high-energy shells by incoming energy, in this case invisible ultraviolet light found in

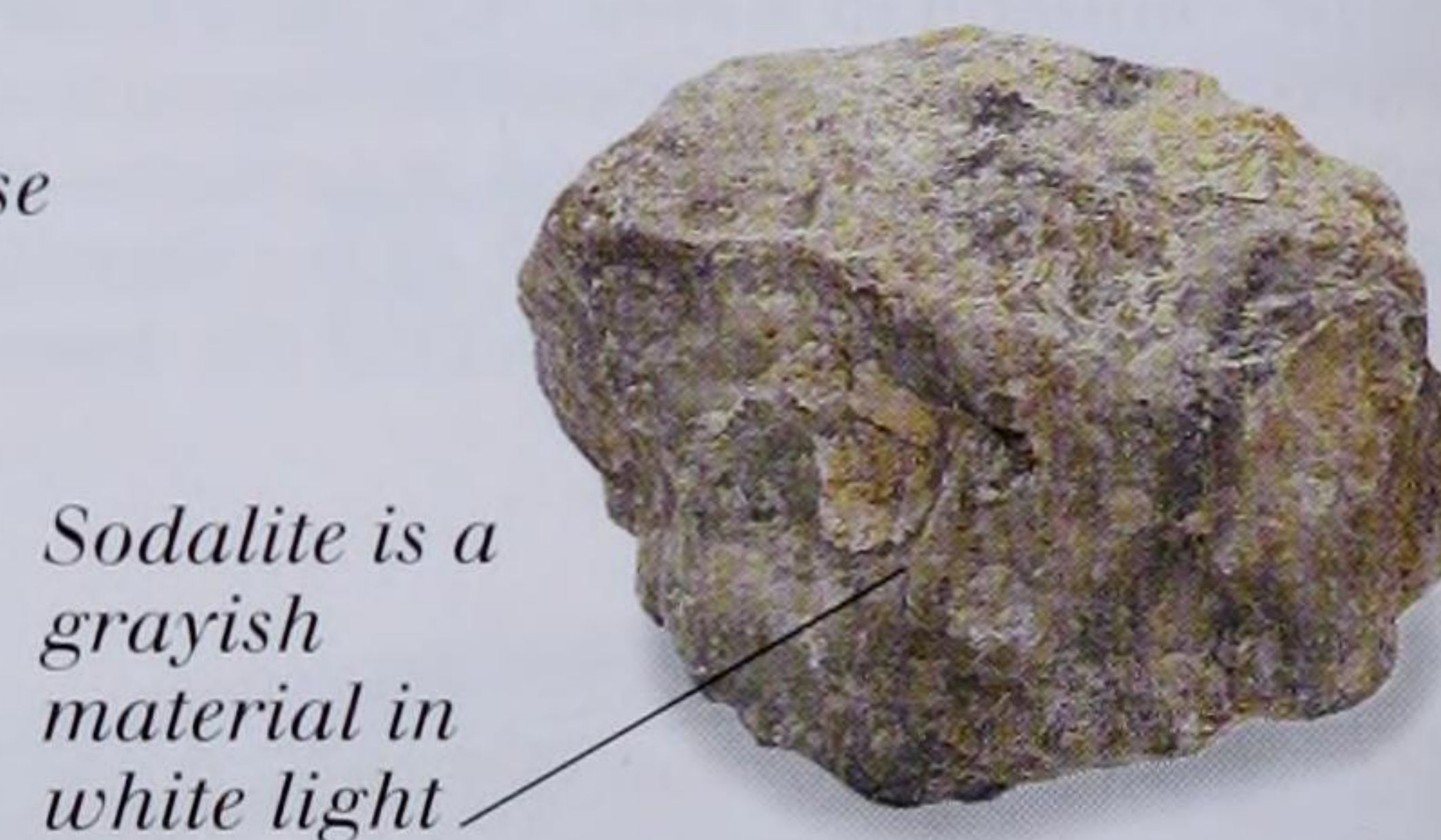
daylight. The electrons fall back to their original shell immediately and reemit the energy as visible light. The color of the light depends upon the difference in energy between the higher and lower shells. This extra light emitted by the powder in the clothes makes them appear bright.



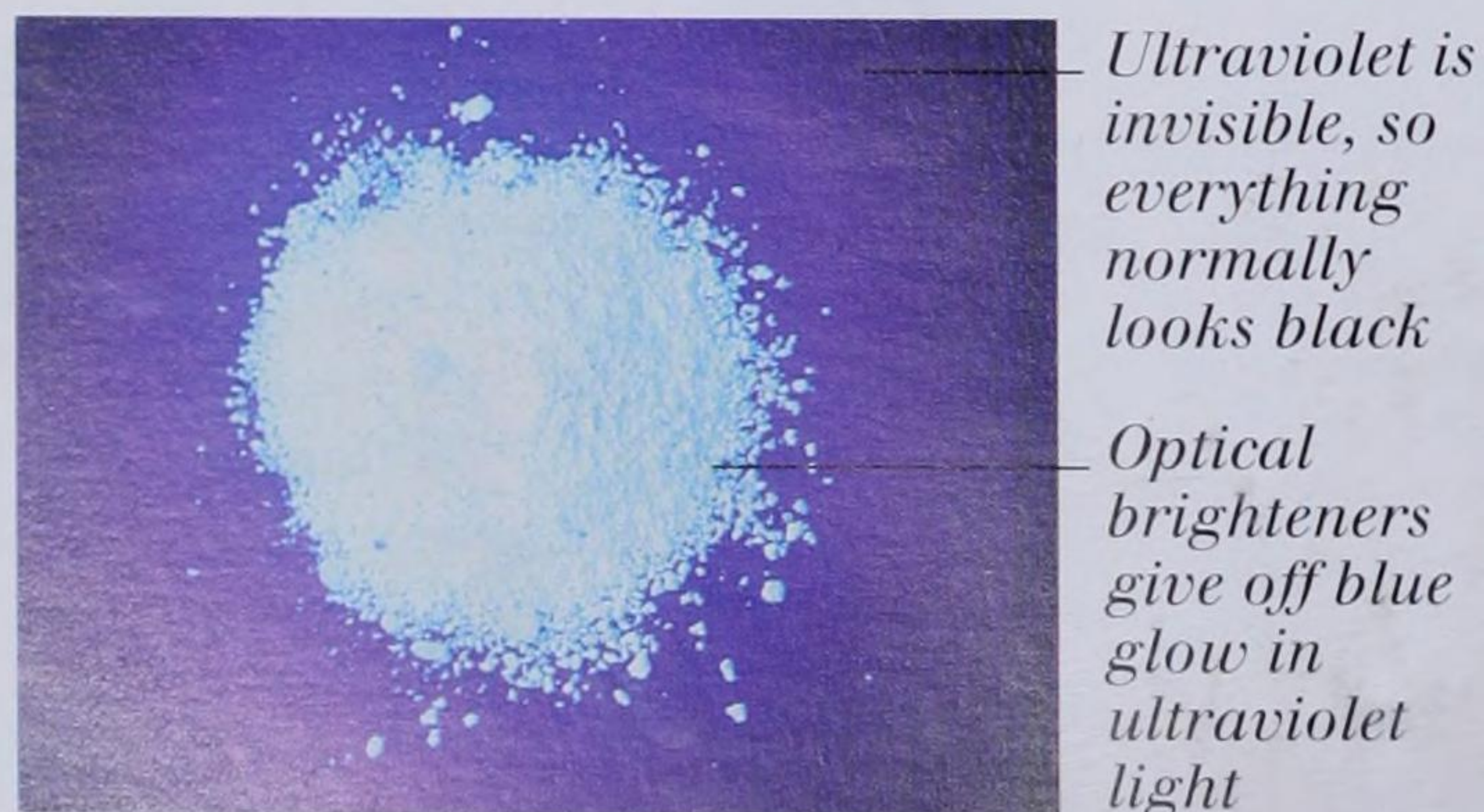
SOAP POWDER IN WHITE LIGHT



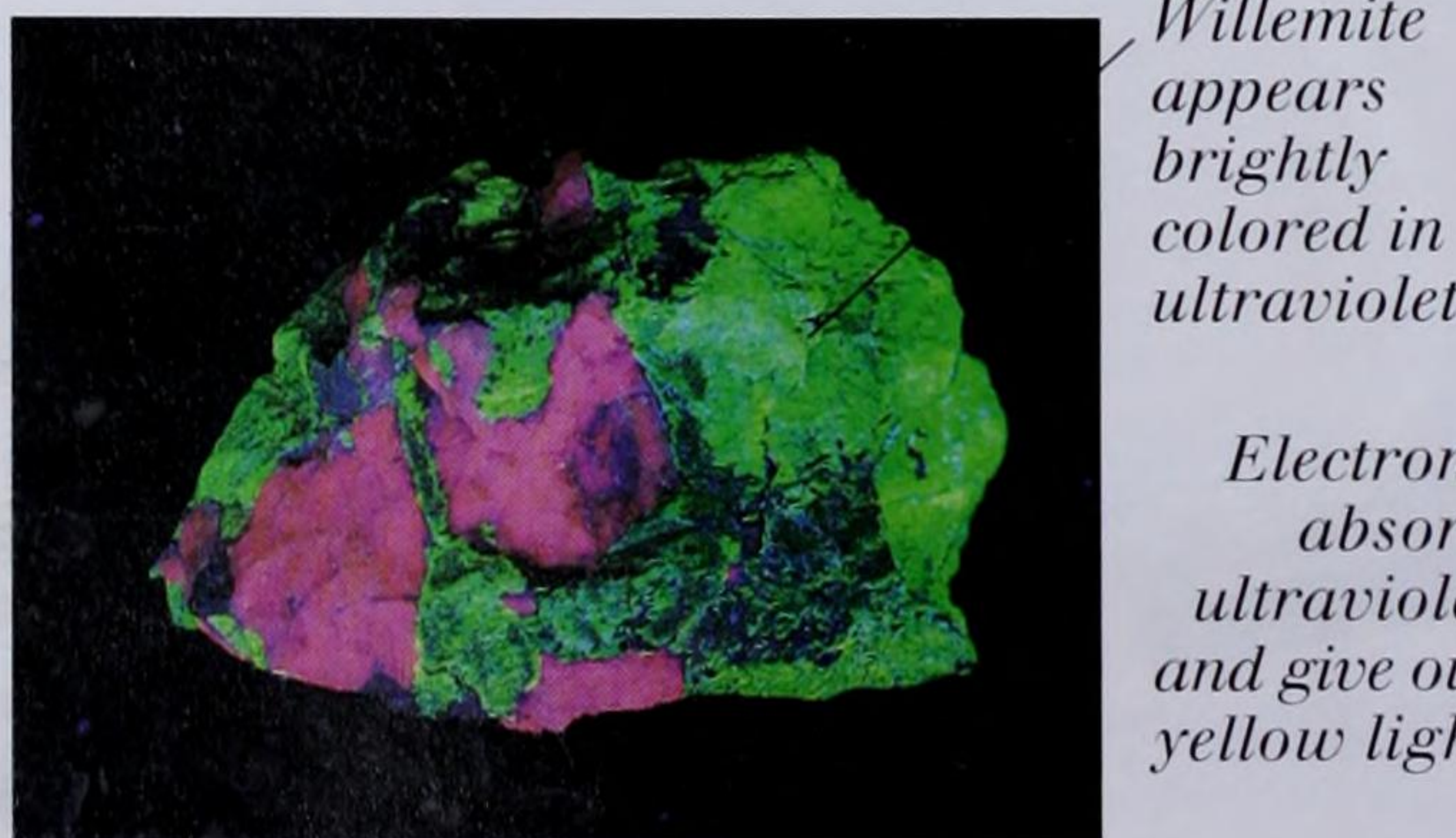
WILLEMITE IN WHITE LIGHT



SODALITE IN WHITE LIGHT



SOAP POWDER IN ULTRAVIOLET LIGHT



WILLEMITE IN ULTRAVIOLET LIGHT



SODALITE IN ULTRAVIOLET LIGHT



## CATHODE RAY TUBE

Inside a cathode ray tube, an **electric current** heats a small filament. The heat generated gives electrons extra energy, moving them farther from their nuclei. A strong electric field then completely removes electrons from their atoms. The free electrons are attracted to the positive **anode**, and pass through it as a cathode ray.

Beam of electrons (cathode ray) made visible by phosphorescent screen

Vacuum  
Phosphorescent screen  
Maltese cross

Glass tube  
Support fork  
Base

## ELECTRON BEAMS

Anode connected to positive supply

Wires connecting heater and cathode to power supply

Wire connecting Maltese cross to positive electrical supply

Wire connecting anode to power supply

Phosphorescent material

Electron shadow of Maltese cross

Green glow where electrons hit screen

Glass tube

SIDE VIEW

FRONT VIEW

## DEFLECTING THE ELECTRONS

Because electrons have electric charge, forces can be applied to them by electric and magnetic fields in the cathode ray tube. The direction of the force depends upon the direction and type of the field.

Screen glows when hit by electrons

Glass tube

Vacuum

Electrons travel in straight line

Base

Anode connected to positive supply

Electrons curve in parabolic path due to electric field

Wires connecting heater and cathode to power supply

Positive terminal

Negative terminal

Electrons travel in part of a circular path due to magnetic field

Helmholtz coil produces magnetic field

4.5V battery

STRAIGHT CATHODE RAY IN TUBE

Wire connecting anode to power supply

DOWNWARD DEFLECTION BY ELECTRIC FIELD

DOWNWARD DEFLECTION BY MAGNETIC FIELD

## HOW A TELEVISION WORKS

### DEFLECTED ELECTRON BEAMS

At the heart of most televisions is a cathode ray tube. Electron beams are produced at the back of the tube. Coils of wire around the tube create magnetic fields, which deflect the electron beams to different parts of the screen. The screen itself is coated with phosphorescent materials called **phosphors**.

Electron beams (cathode rays)

Phosphorescent screen

Picture built up as beams scan across the screen

Red, green, and blue electron guns

Electromagnetic coils are fed with varying electric signal from antenna which builds up a picture from the electron beam

Cathode ray tube

Electronic circuits process and amplify the signal

### PHOSPHORESCENCE

When the cathode rays hit the special coating on a television screen, the screen glows because it is phosphorescent. **Phosphorescence** is a form of luminescence where the incoming energy is not reemitted immediately but is stored and reemitted over a period of time. This means that as the cathode ray quickly scans the picture, the phosphor glows for long enough for a whole picture to form.

Signal received from television antenna consists of a varying electric current

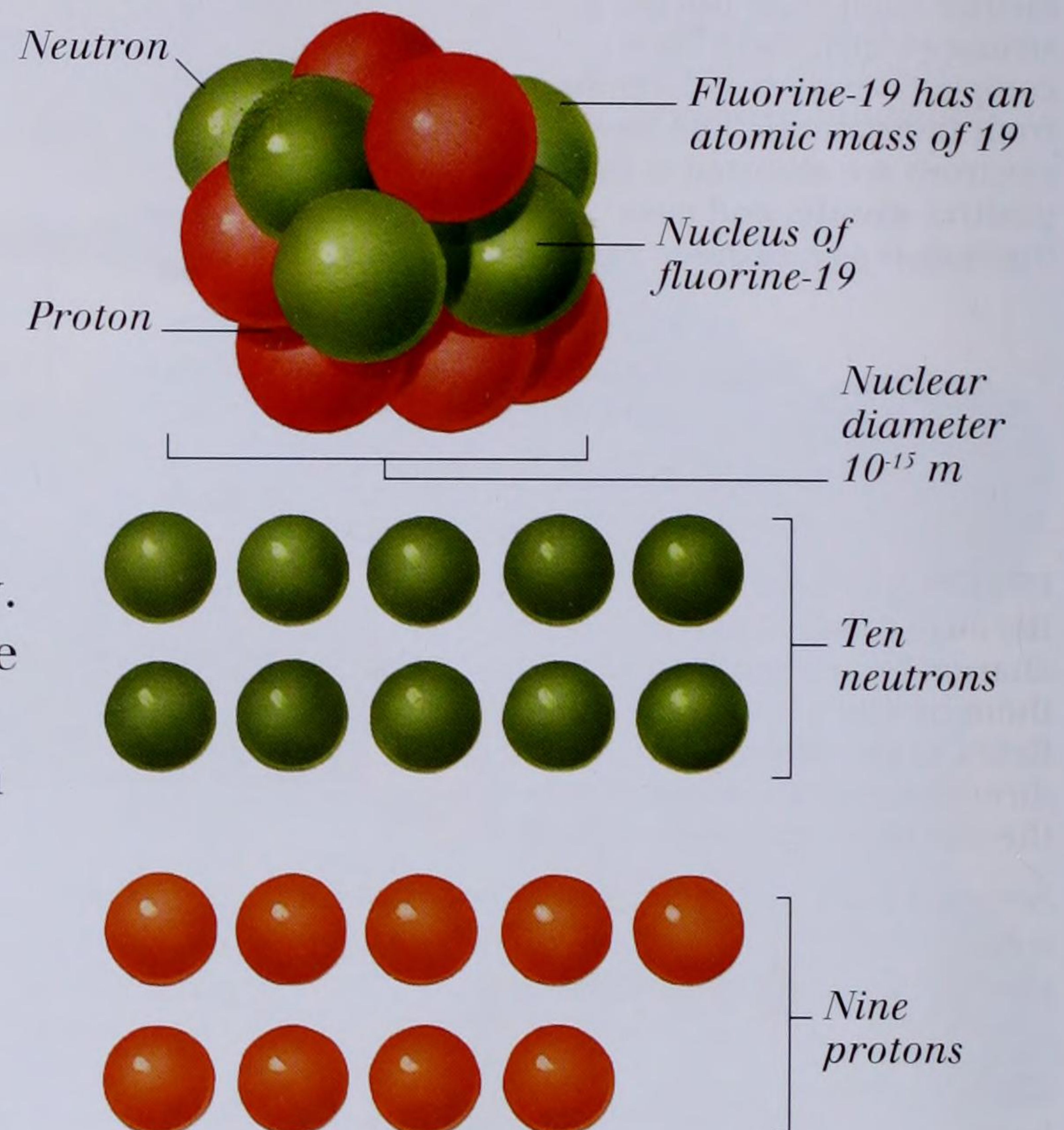


# Nuclear physics

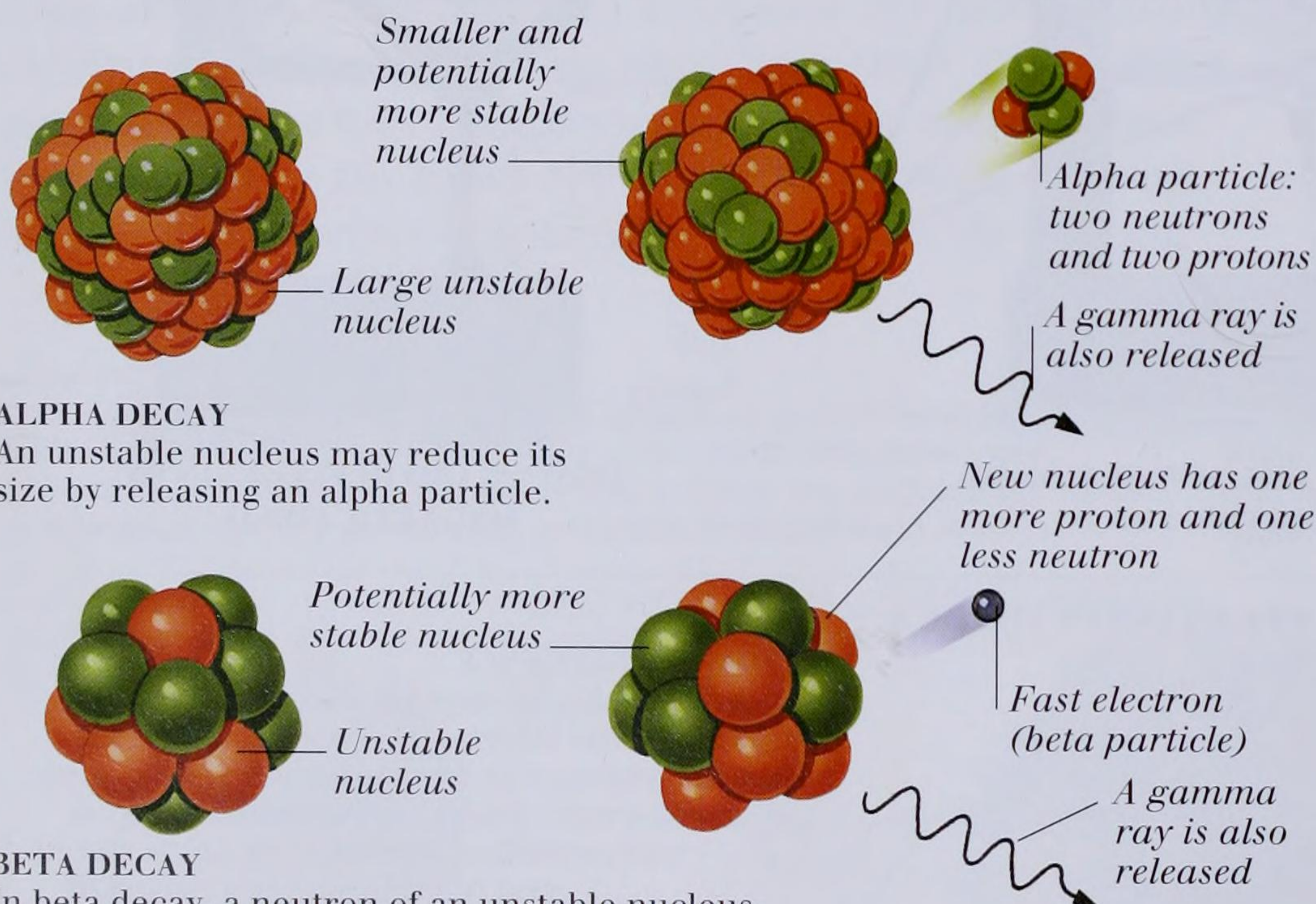
AT THE CENTER OF EVERY ATOM LIES a positively charged **nucleus**. It consists of **protons** and **neutrons**. The number of protons in the nucleus is called the **atomic number**. Because they all have the same **electric charge**, protons repel each other. The nucleus holds together despite this repulsion because of the **strong nuclear force** (see pp. 52–53). The balance between the repulsive force and the strong nuclear force determines whether a nucleus is stable or unstable. On the whole, small nuclei are more stable than larger ones, because the strong nuclear force works best over small distances. An unstable, larger nucleus can break up or decay in two main ways, **alpha decay** and **beta decay**. These produce **alpha** and **beta particles**. In each type of decay, the atomic number of the new nucleus is different from the original nucleus, because the number of protons present alters. Nuclei can also completely split into two smaller fragments, in a process called **fission**. In another **nuclear reaction** called **fusion**, small nuclei join together. Both of these reactions can release huge amounts of energy. Fusion provides most of the Sun's energy, while fission can be used in power stations to produce electricity.

## FLUORINE-19 NUCLEUS

The number of protons in a nucleus defines what **element** the atom is. For example, all fluorine atoms have nine protons. Fluorine has an atomic number of 9. The number of neutrons can vary. Fluorine-19 has ten neutrons, while fluorine-18 has nine.

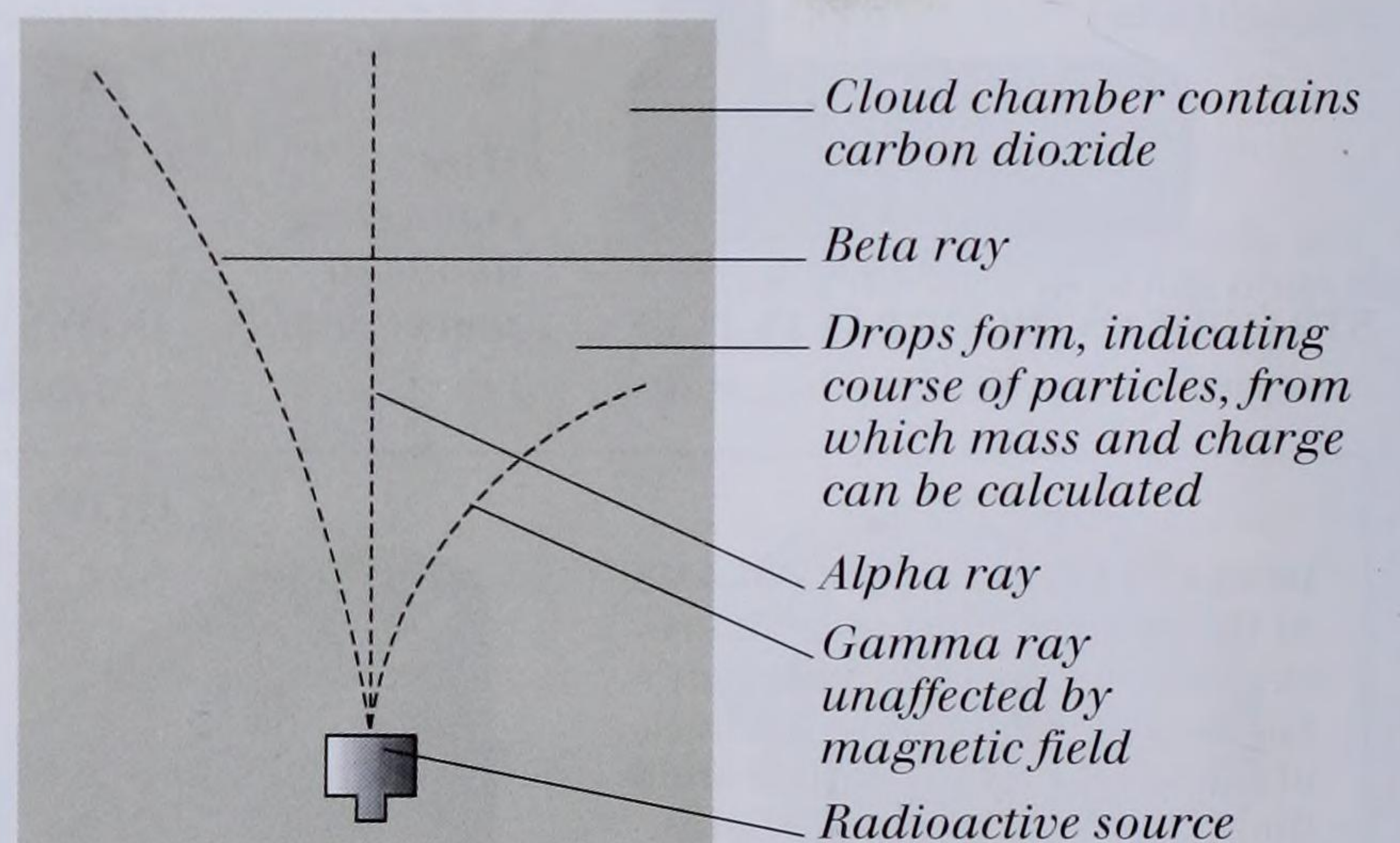


## RADIOACTIVITY



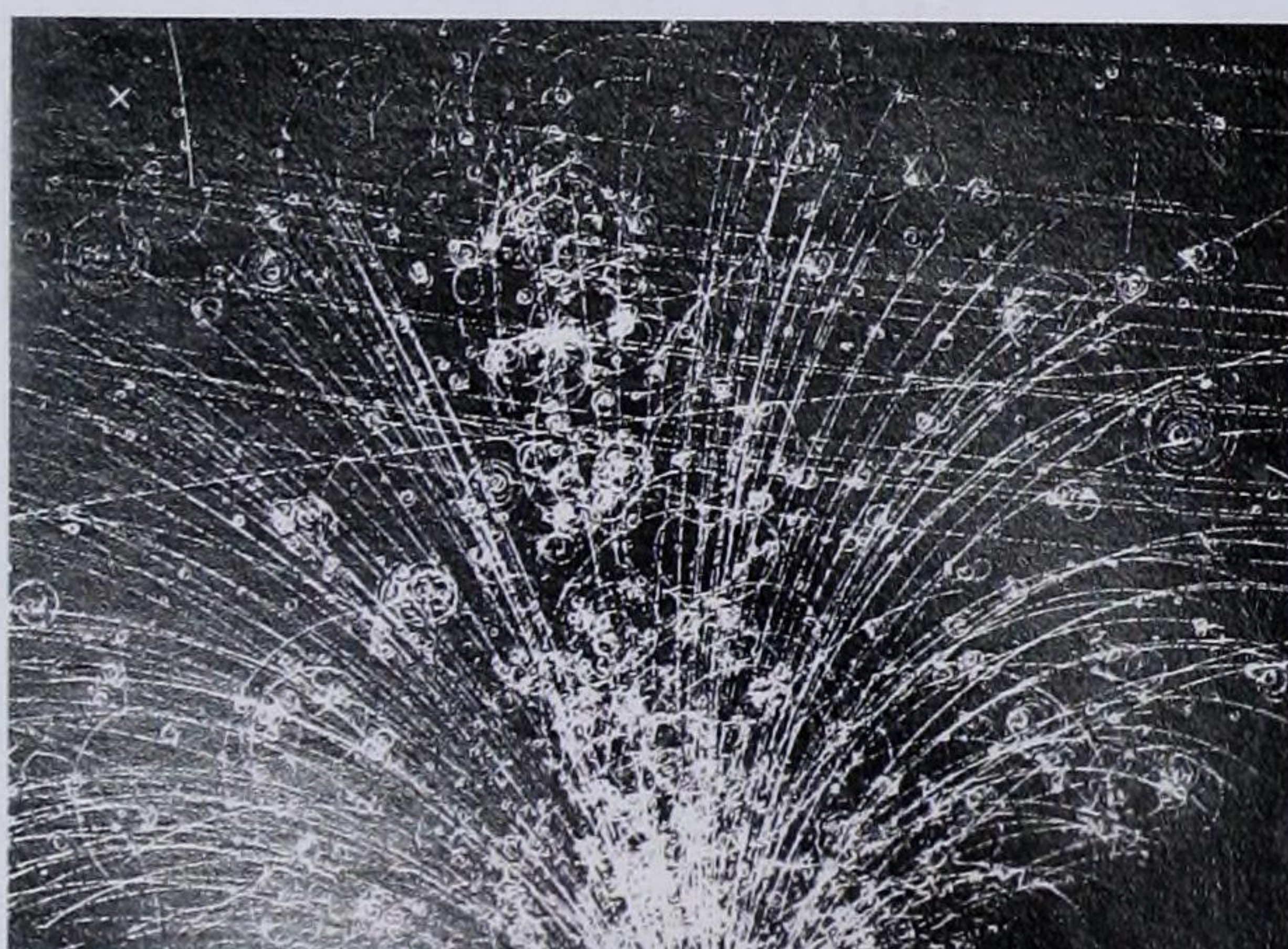
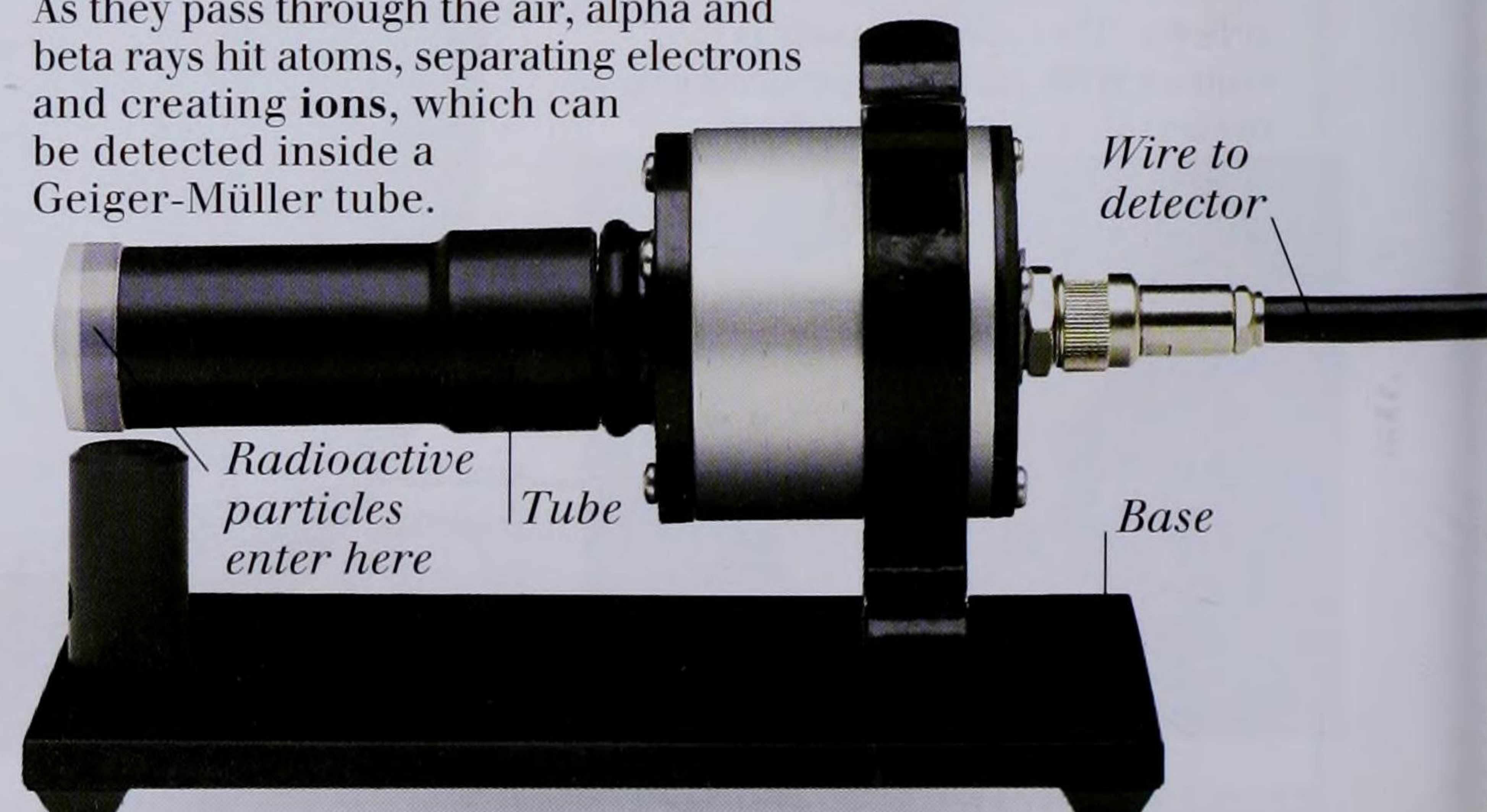
## ANALYZING RADIOACTIVITY

Because of their electric charges, a strong magnetic field will deflect alpha and beta rays in a cloud chamber. Cloud chambers are used to show these paths as in the illustration below.



## GEIGER-MULLER TUBE

As they pass through the air, alpha and beta rays hit atoms, separating electrons and creating **ions**, which can be detected inside a Geiger-Müller tube.



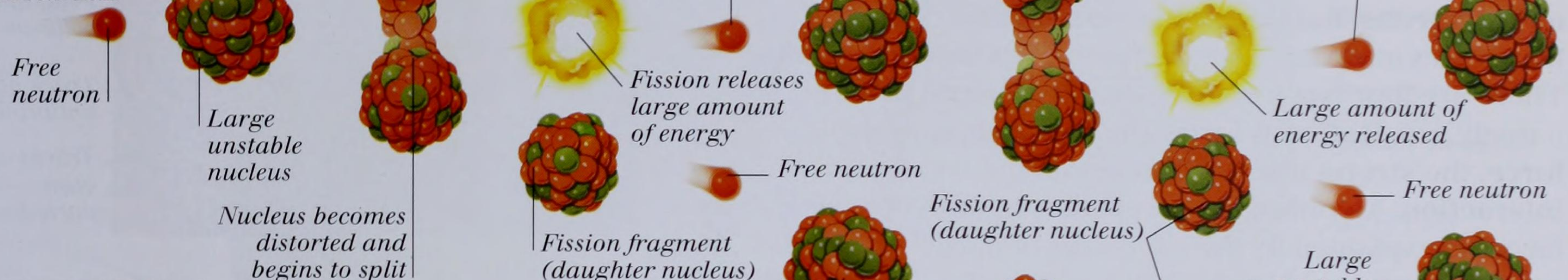
## COSMIC RAYS

The Earth is constantly bombarded by particles from space. Most of these are protons, from atoms of the most abundant element, hydrogen. Occasionally, the protons collide with atoms in the air, producing showers of secondary particles called cosmic rays.

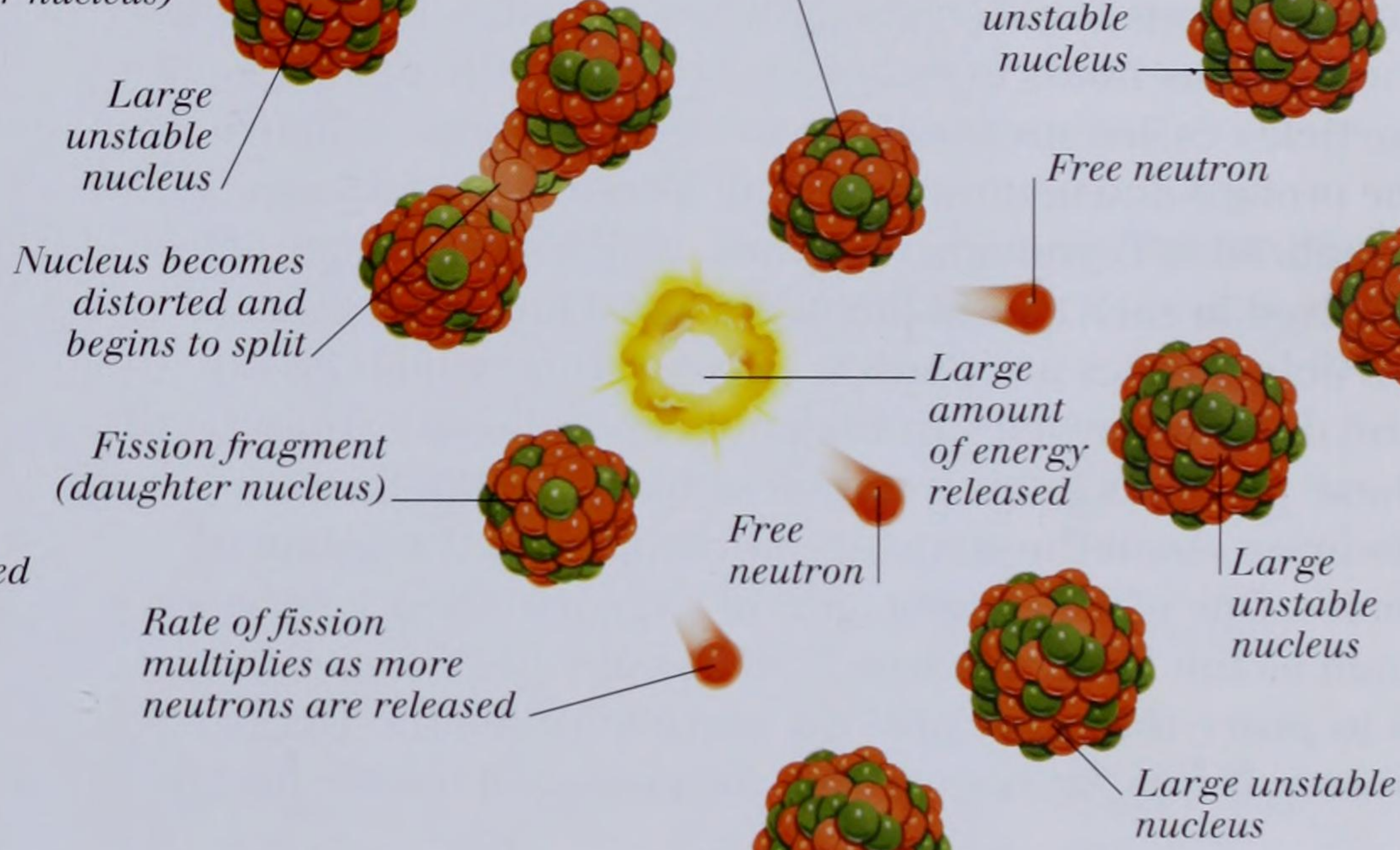
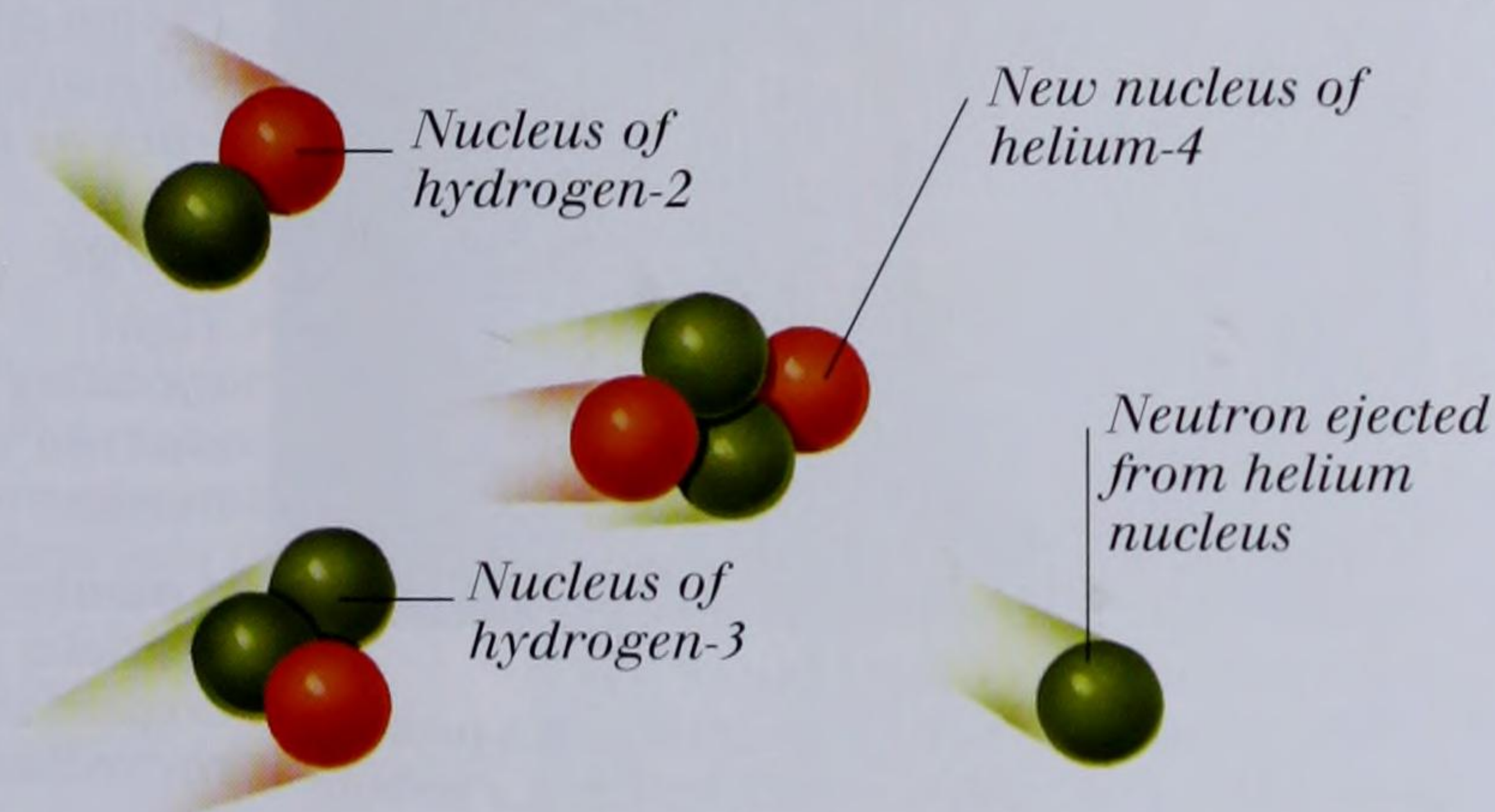
Tracks left by cosmic rays in a bubble chamber



A neutron hitting a large, unstable nucleus may split or fission into two smaller, more stable fragments, releasing large amounts of energy. Often, more free neutrons are produced by this fission, and these neutrons can cause other nuclei to split. The process may continue, involving many nuclei in a **chain reaction**.

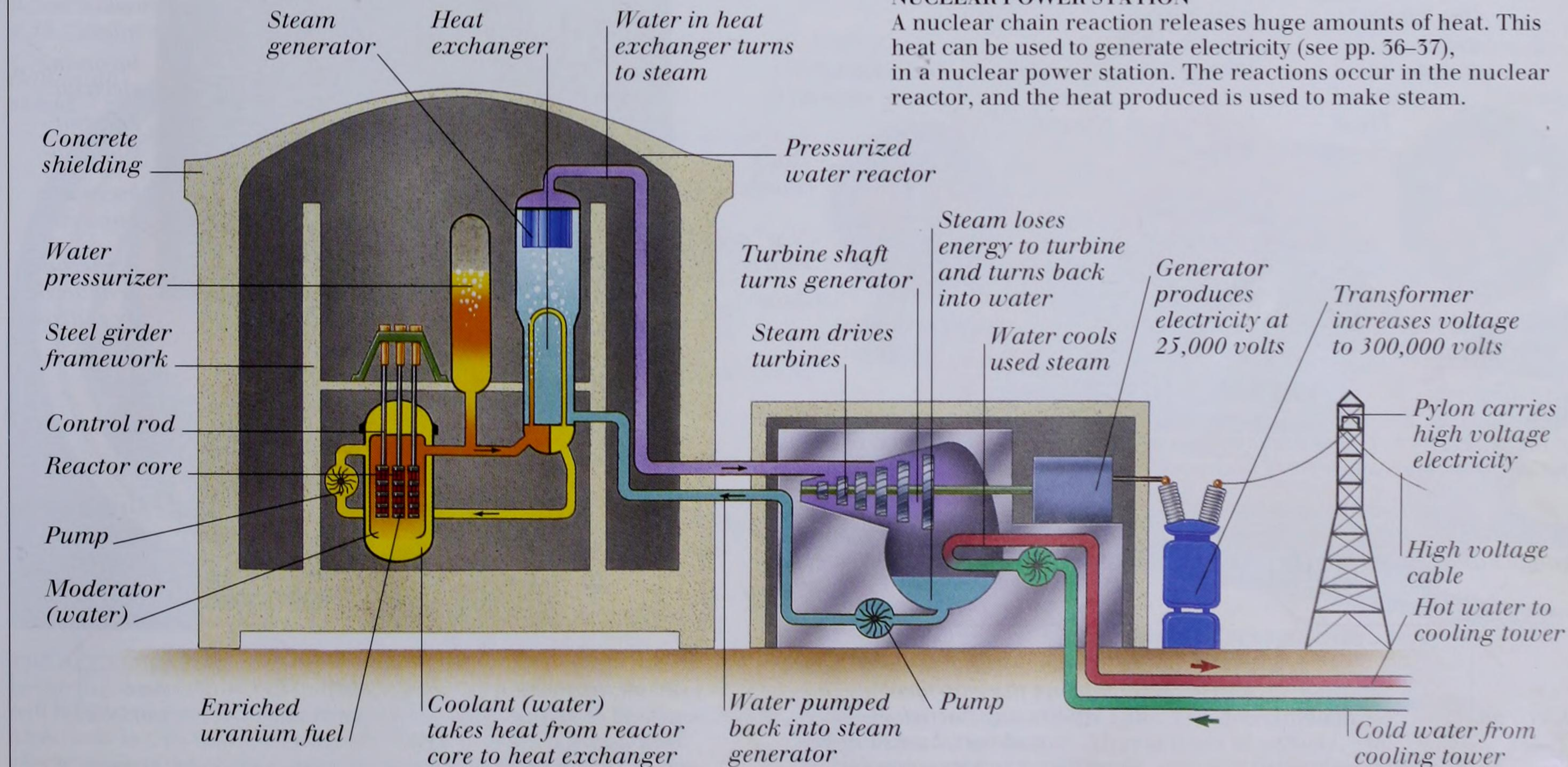


Just as large nuclei can split, so some small nuclei can join together, or fuse. Like fission, fusion can release energy. One of the highest energy fusion reactions involves nuclei of hydrogen, which collide at great speed, forming a nucleus of helium.



## NUCLEAR POWER STATION

A nuclear chain reaction releases huge amounts of heat. This heat can be used to generate electricity (see pp. 56–57), in a nuclear power station. The reactions occur in the nuclear reactor, and the heat produced is used to make steam.



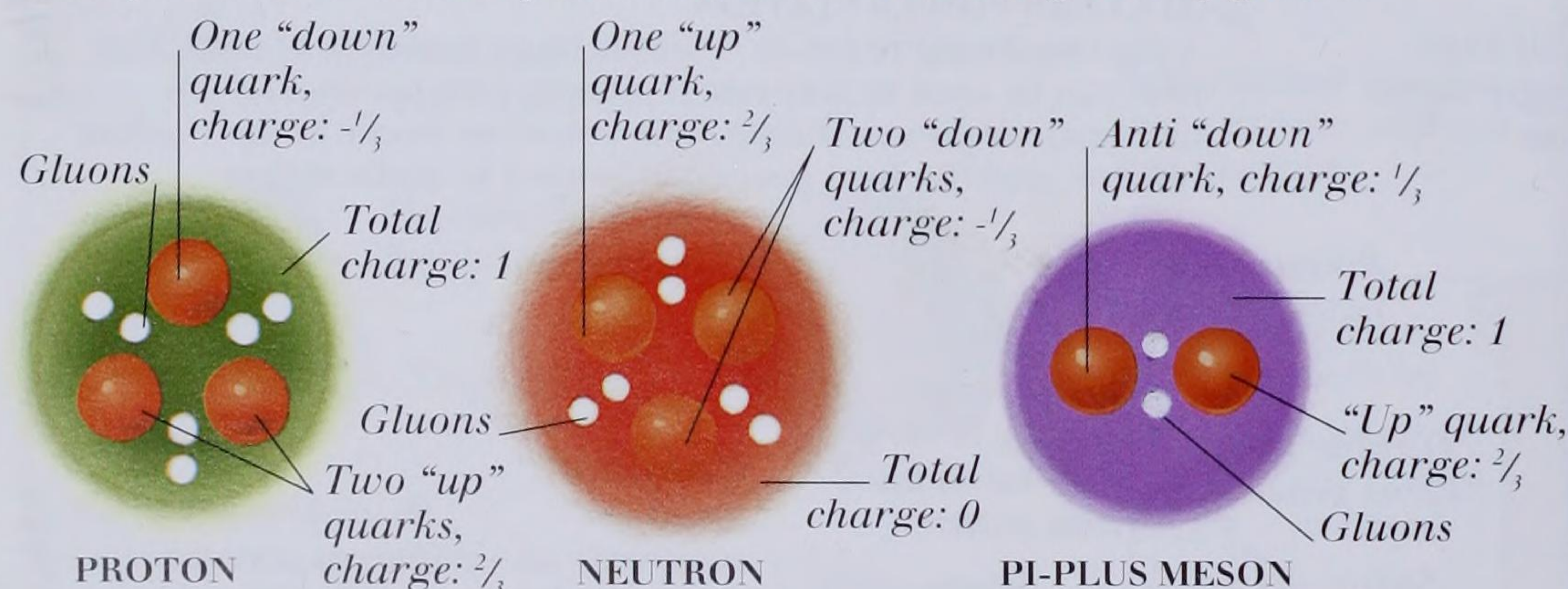


# Particle physics

PARTICLE PHYSICS ATTEMPTS TO EXPLAIN matter and **force** in terms of tiny particles. The **atom**, once thought to be the smallest particle, is actually made of **protons**, **neutrons**, and **electrons**. But the proton and the neutron are themselves made up of smaller particles, known as **quarks**. There are four types of force acting between matter, namely **gravitational force**, the **electromagnetic force**, the **strong nuclear force**, and the **weak interaction**. According to current theory, each of these forces is explained by the exchange of particles called **gauge bosons** between the particles of matter. For example, the **nucleus** holds together as a result of the exchange of particles called **mesons** (a type of gauge boson) between the protons and neutrons present. These exchanges can be visualized in Feynmann diagrams, which show the particles involved in each type of force. The most important tools of particle physics are particle accelerators, which create and destroy particles in high energy collisions. Analysis of these particles helps to prove or disprove the latest theories about the structure of matter and the origin of forces. One of the current aims of large particle accelerators, such as the Large Hadron Collider at **CERN** (opposite), is to prove the existence of a particle called the **Higgs boson**. It may be responsible for giving all matter mass.

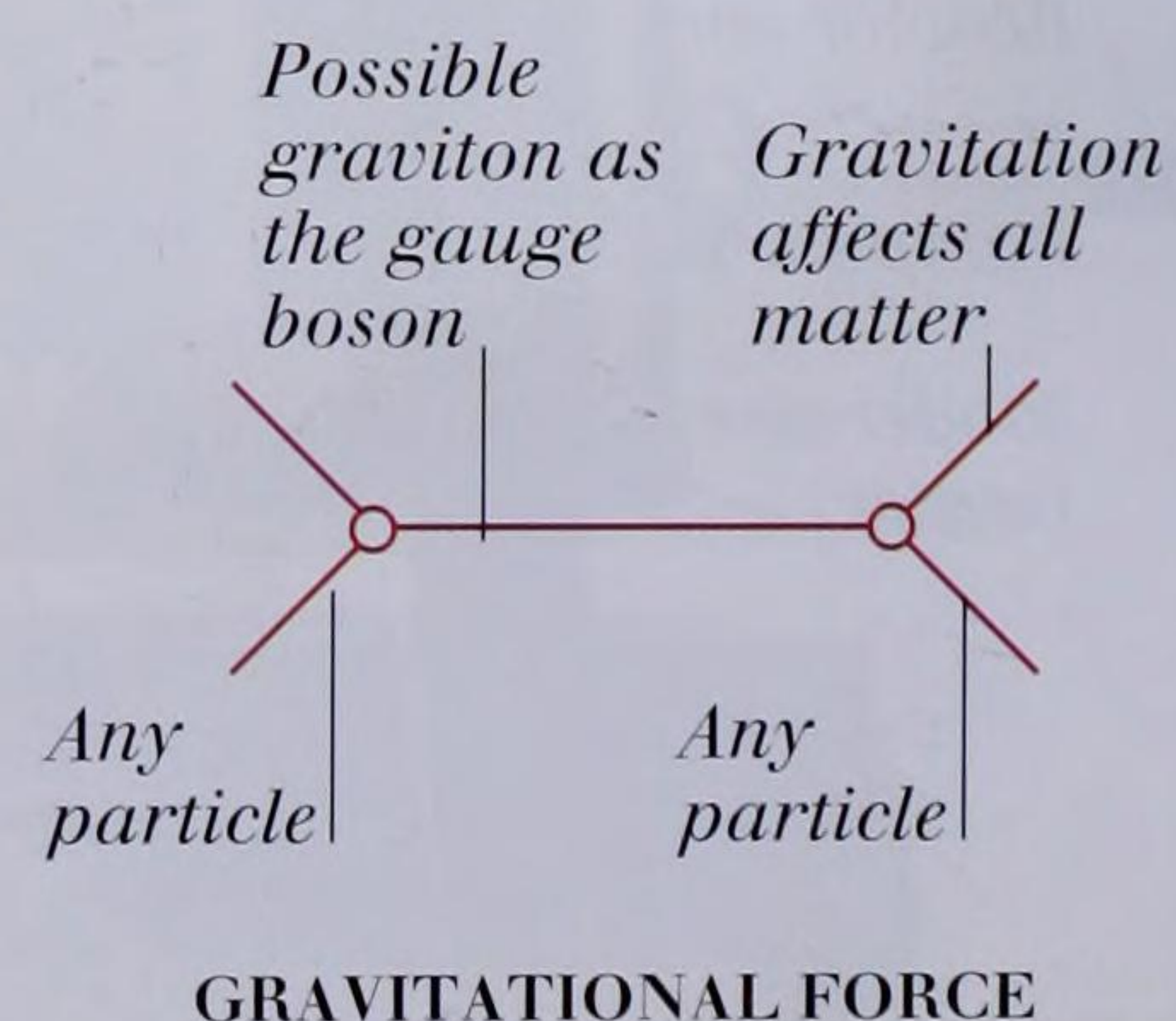
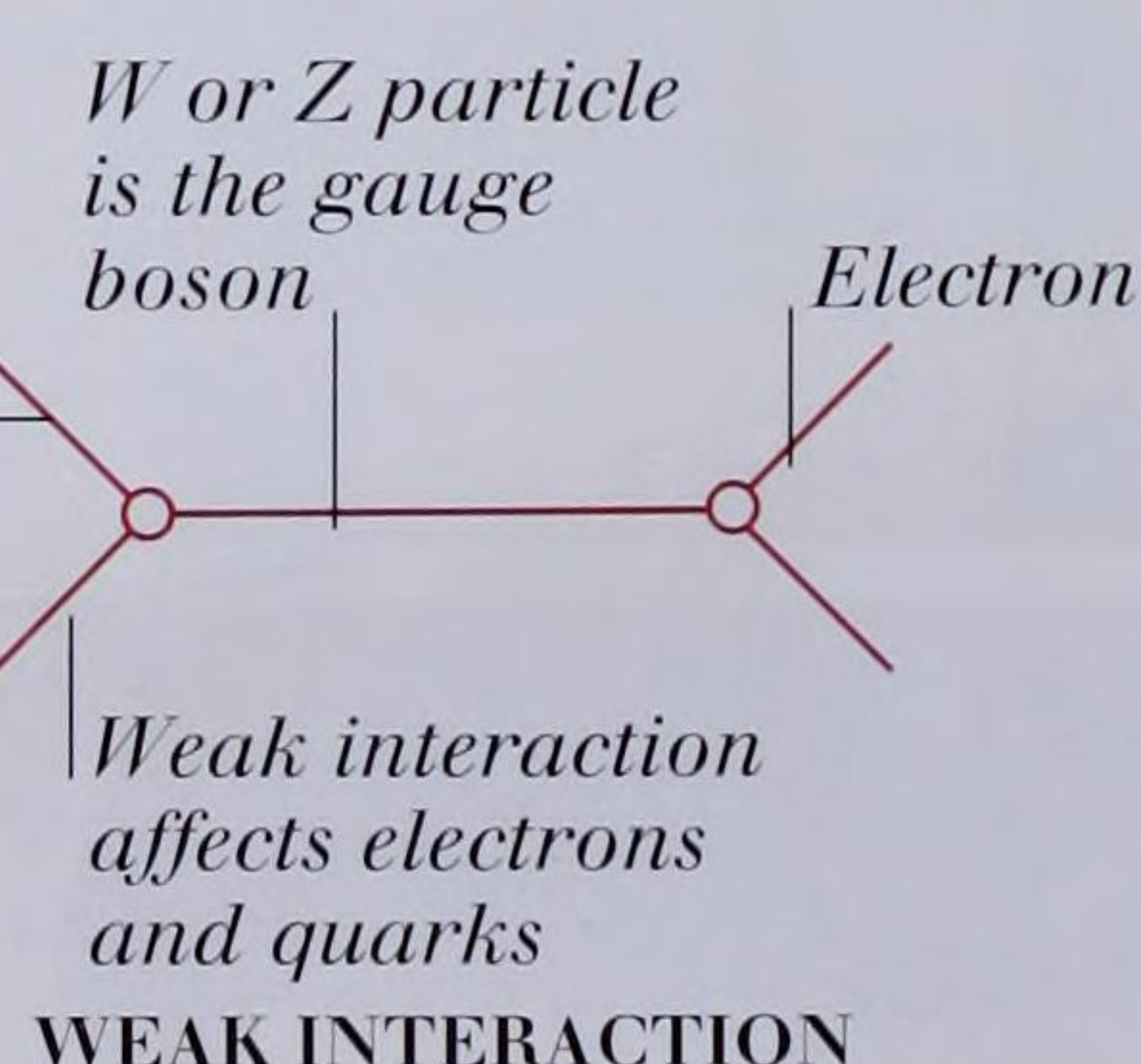
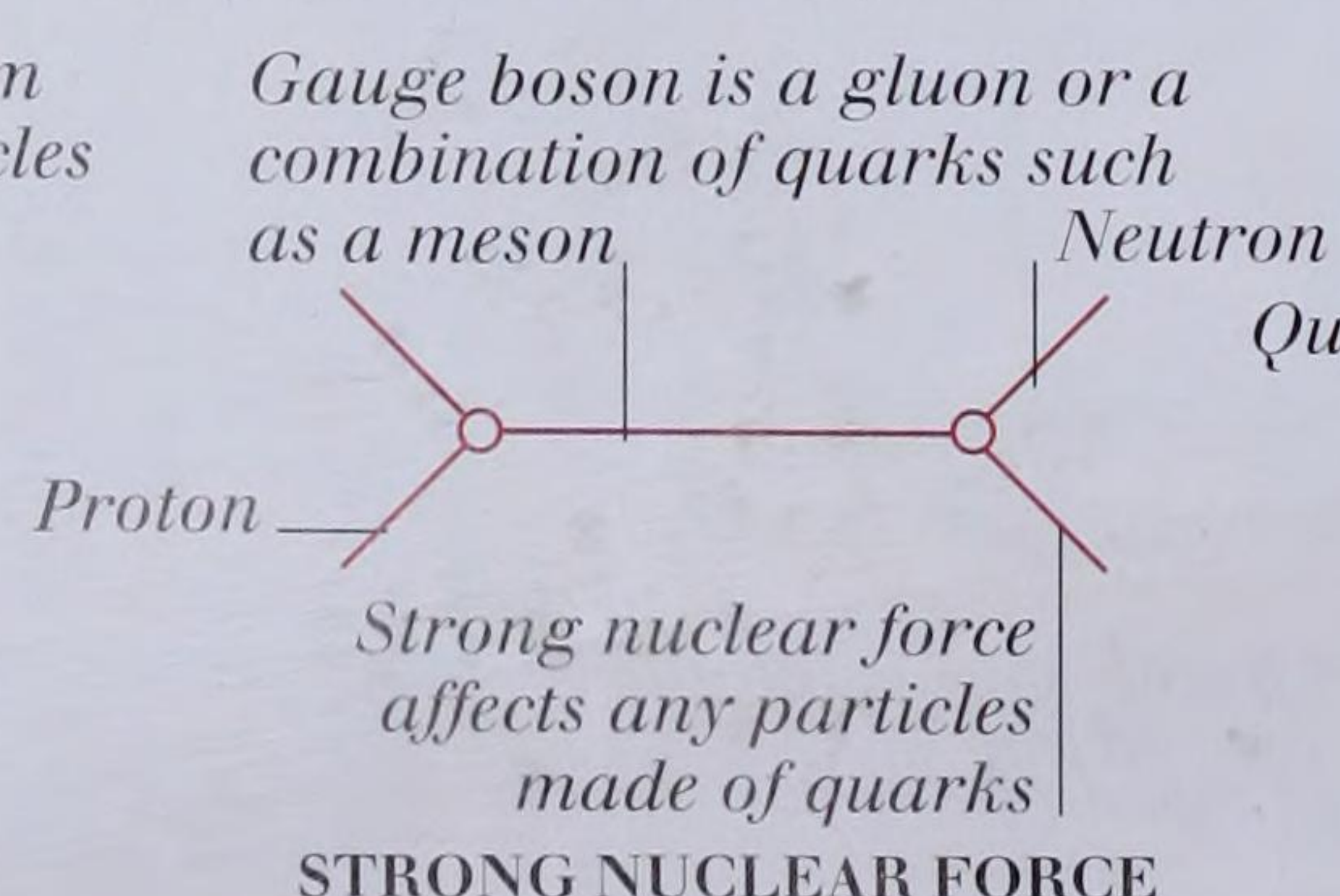
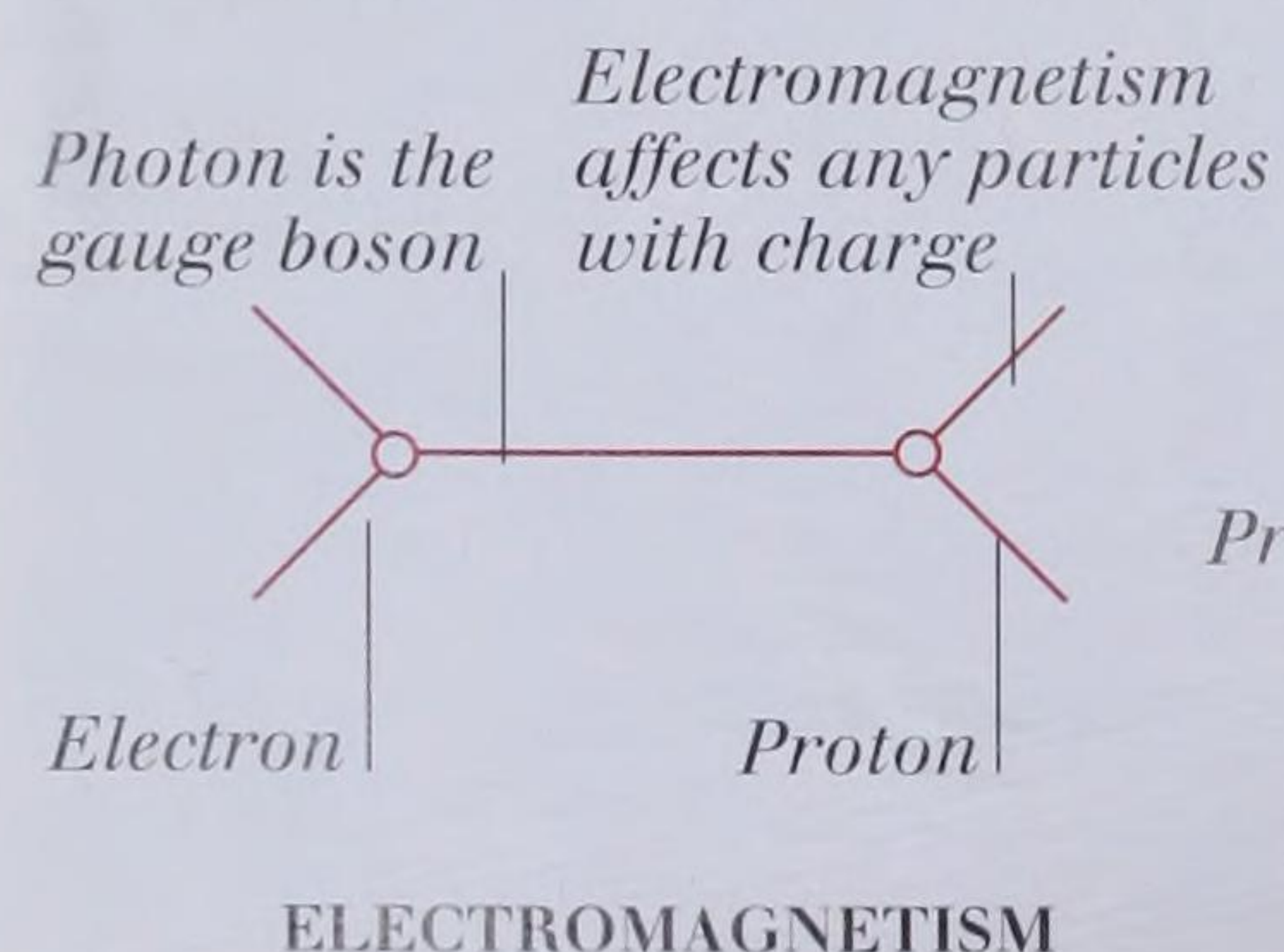
## HADRONS

A **hadron** is a particle consisting of quarks. There are six types of quark, including the "up" and "down" quarks. Protons, neutrons, and mesons are examples of hadrons. The quarks of hadrons are held together by **gluons**.



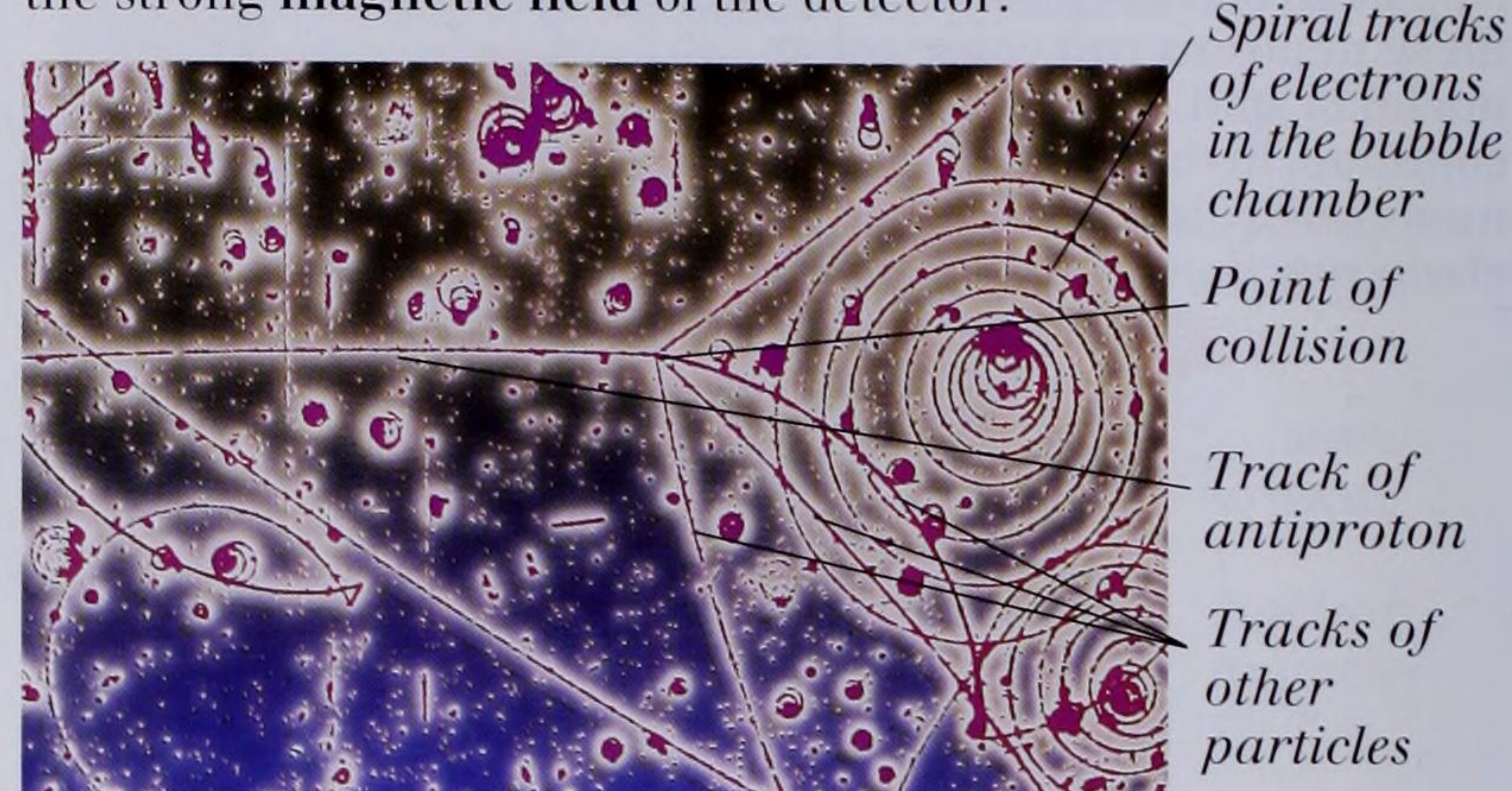
## FEYNMAN DIAGRAMS

These diagrams show which gauge bosons are exchanged to transfer each of the four forces. The horizontal lines represent the gauge boson, whereas the diagonal lines and circles represent the two interacting particles.



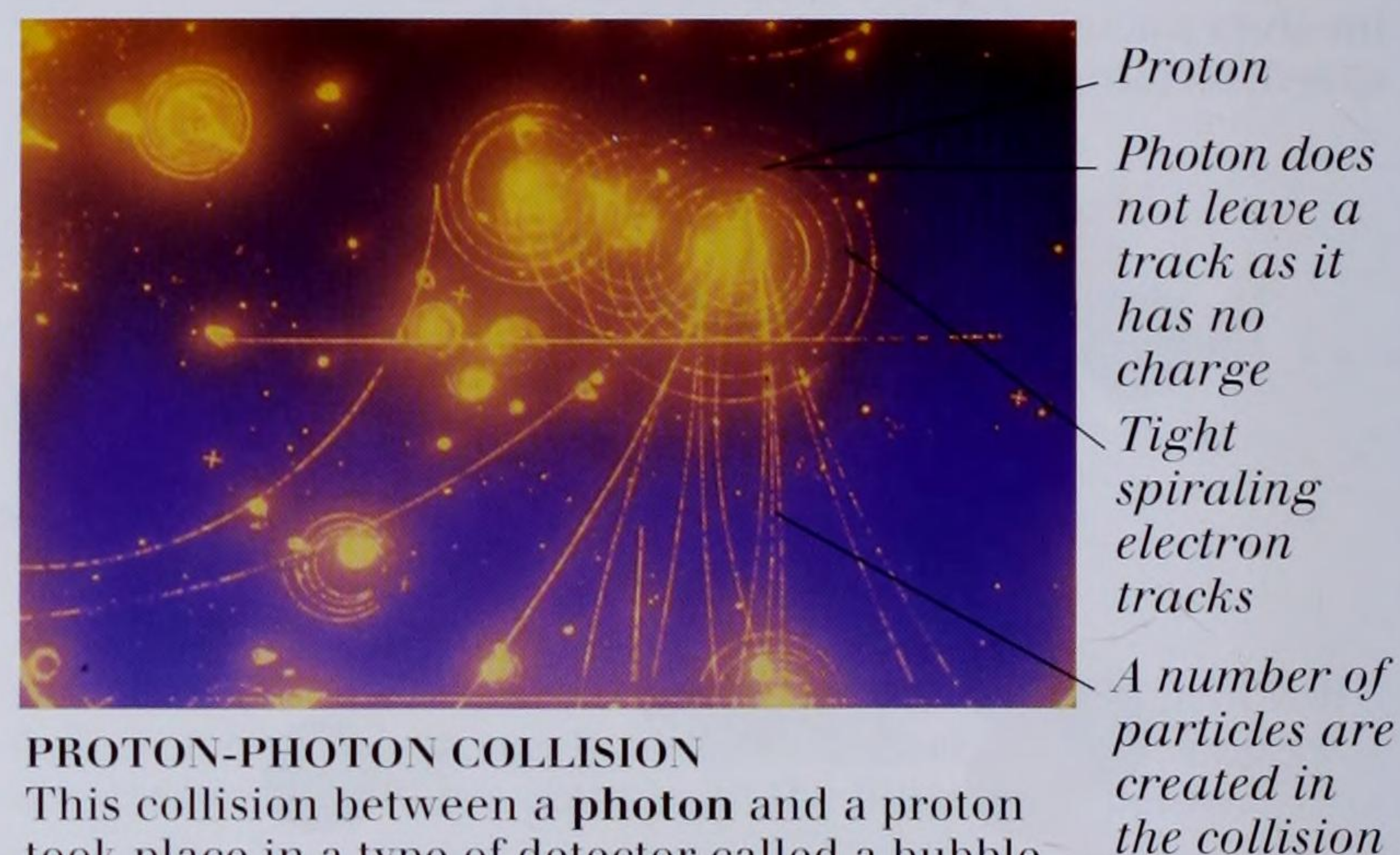
## PARTICLE COLLISIONS

The images below show the results of collisions between particles in particle accelerators. Particles of opposite charge curve in different directions in the strong **magnetic field** of the detector.



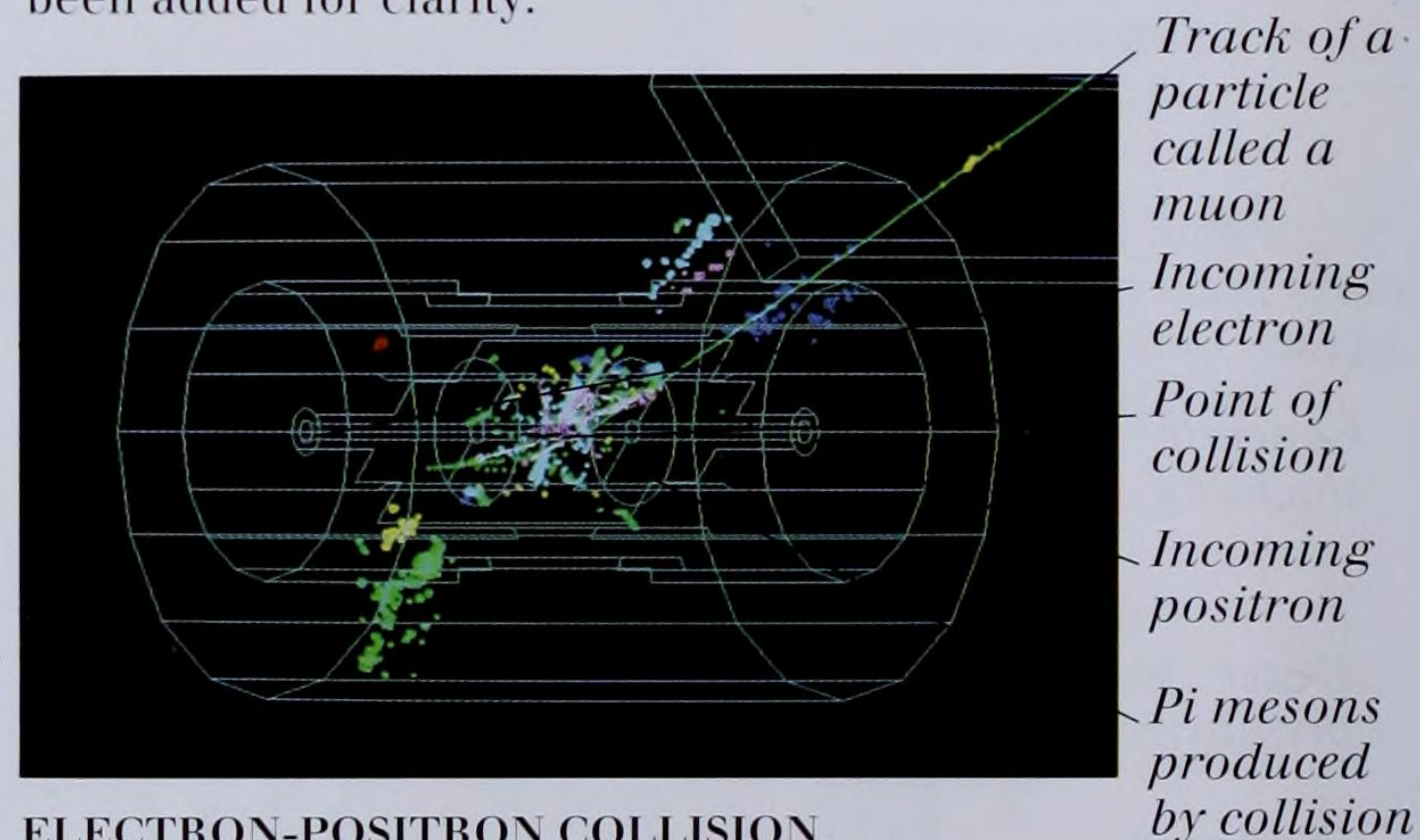
## ANNIHILATION

When a particle and an antiparticle meet, they destroy each other and become energy. This energy in turn becomes new particles.



## PROTON-PHOTON COLLISION

This collision between a **photon** and a proton took place in a type of detector called a bubble chamber. The colors in this photograph have been added for clarity.



## ELECTRON-POSITRON COLLISION

Here, an electron collides with its **antiparticle**, a **positron**. The detector is linked to a computer, which produces this picture of the collision.

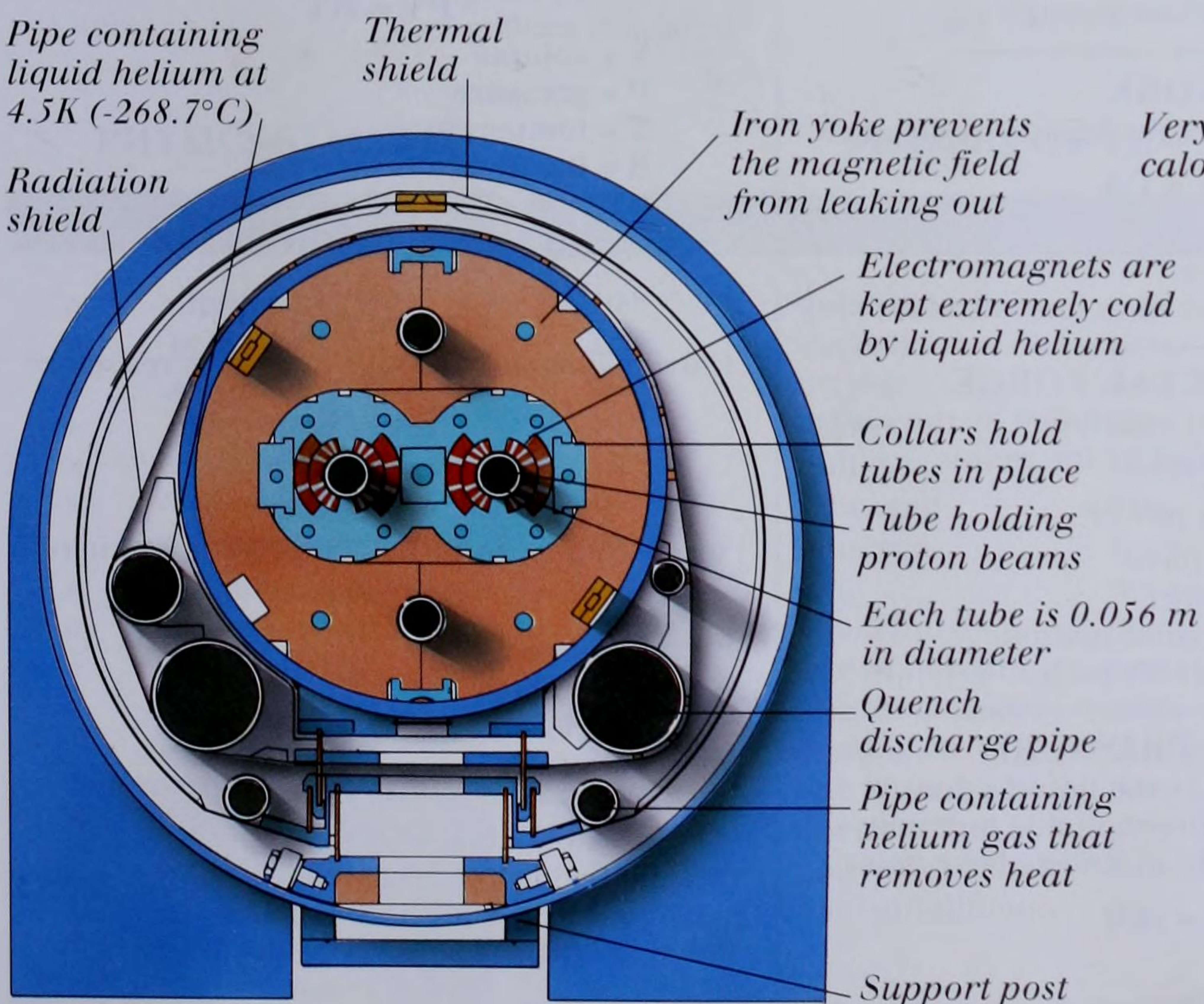


## THE LARGE HADRON COLLIDER

### MAP OF THE SITE

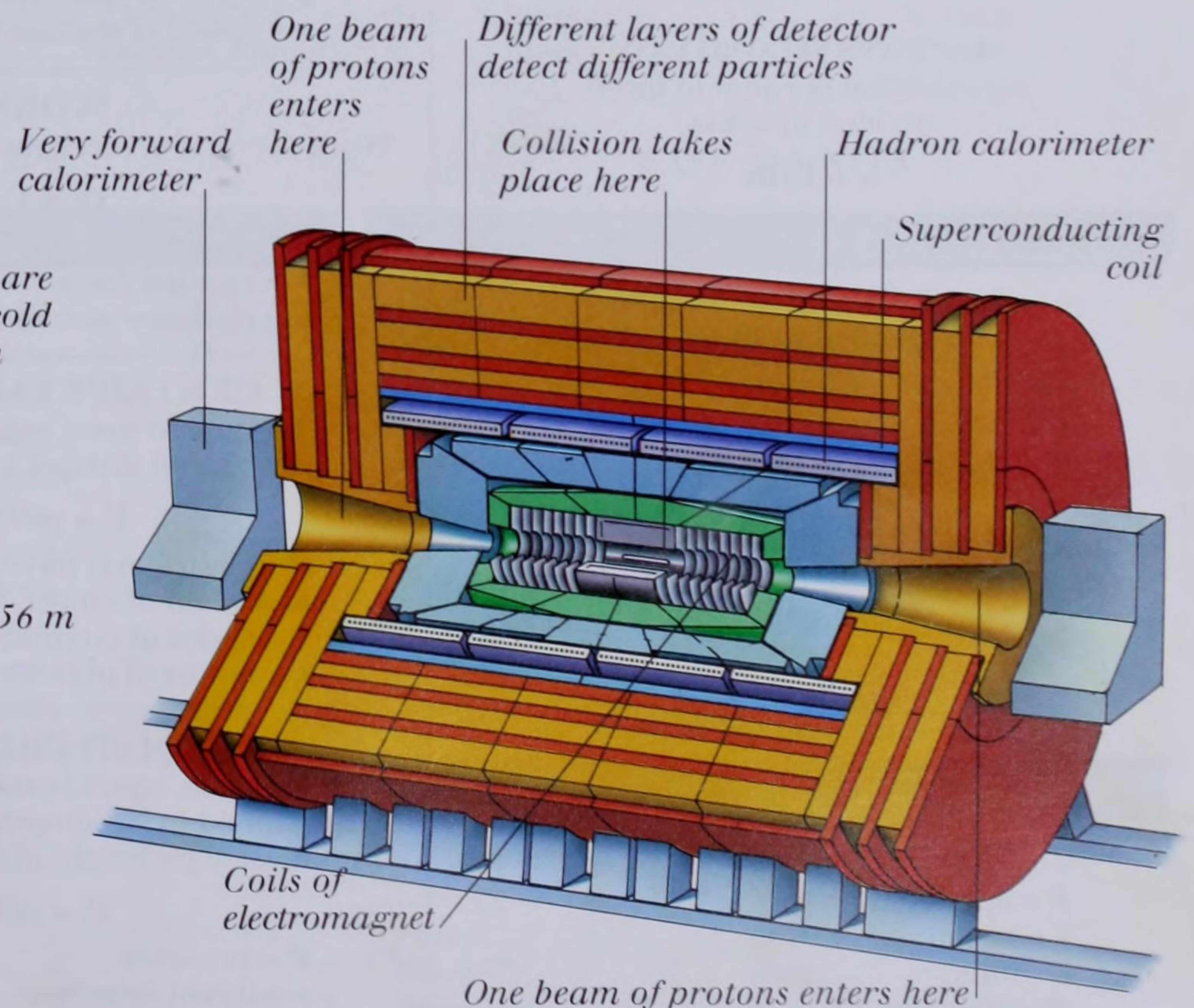
The Large Hadron Collider (LHC), at CERN near Geneva, will be a huge particle accelerator, in a tunnel about 100 meters below ground. The tunnel will be a ring 27 kilometers long, which is already used for

another particle accelerator, the Large Electron Positron (LEP) collider. Two beams of protons will move around in tubes at very high speed, and will be made to collide in detectors, such as the CMS (see below).



### THE ACCELERATOR

In the main experiment of the LHC, protons injected into the ring will be accelerated to nearly the speed of light, traveling in opposite directions in two tubes. **Centripetal force** provided by powerful **electromagnets** keeps the protons moving in a circle.



### THE COMPACT SOLENOIDAL (CMS) DETECTOR

Several detectors will be built for detecting the particles produced by collisions in the LHC. The detectors have different parts that detect different types of particle. The hadron calorimeter, for example, can only detect hadrons.



# Formulas

MANY OF THE PRINCIPLES EXPLAINED IN THIS BOOK can be expressed as formulas. The use of symbols to represent different values enables the physicist to make quick calculations, reducing even complicated physical phenomena to simple mathematical formulas.

## WEIGHT

Weight is equal to mass multiplied by acceleration due to gravity

$$W = mg$$

W = weight

m = mass

g = acceleration due to gravity

## TURNING FORCE

Turning force is equal to force multiplied by distance of applied force from pivot

$$T = Fd$$

T = turning force (moment)

F = applied force

d = distance

## PRESSURE

Pressure is equal to force applied divided by area over which force acts

$$P = F/A$$

P = pressure

F = applied force

A = area over which force acts

## FORCE AND MOTION

### NEWTON'S SECOND LAW

Acceleration is equal to force divided by mass

$$a = F/m$$

### SPEED

Speed is equal to distance divided by time

$$v = d/t$$

### CONSTANT ACCELERATION

Acceleration is equal to change in speed divided by time taken for that change

$$a = (v_2 - v_1)/t$$

### MOMENTUM

Momentum is equal to mass multiplied by speed

$$p = mv$$

F = applied force

v = speed

t = time

d = distance

p = momentum

m = mass

a = acceleration

## GRAVITATION

Gravitational force equals a constant, multiplied by mass one, multiplied by mass two, divided by the distance between the masses squared

$$F = Gm_1m_2/d^2$$

F = gravitational force between two objects

G = gravitational constant

m<sub>1</sub> = mass of object one

m<sub>2</sub> = mass of object two

d = distance between the two objects

## FRICTION

Frictional force between two surfaces is equal to the coefficient of friction multiplied by the force acting to keep the surfaces together

$$F = \mu N$$

F = frictional force

$\mu$  = coefficient of friction; this varies with materials

N = force between two surfaces

## AIR RESISTANCE

Force is proportional to speed

$$F \propto v$$

F = force of resistance due to air

v = speed of motion through air

## WORK

Work is equal to force multiplied by distance

$$W = Fd$$

W = work done

F = applied force

d = distance moved in line with force

## CENTRIPETAL FORCE

Force is equal to mass multiplied by the speed squared divided by the radius

$$F = mv^2/r$$

F = centripetal force

m = mass of object

v = speed of circular motion

r = radius of object's path

## LIQUID PRESSURE

Pressure is equal to the liquid's density multiplied by acceleration due to gravity multiplied by height of water above point

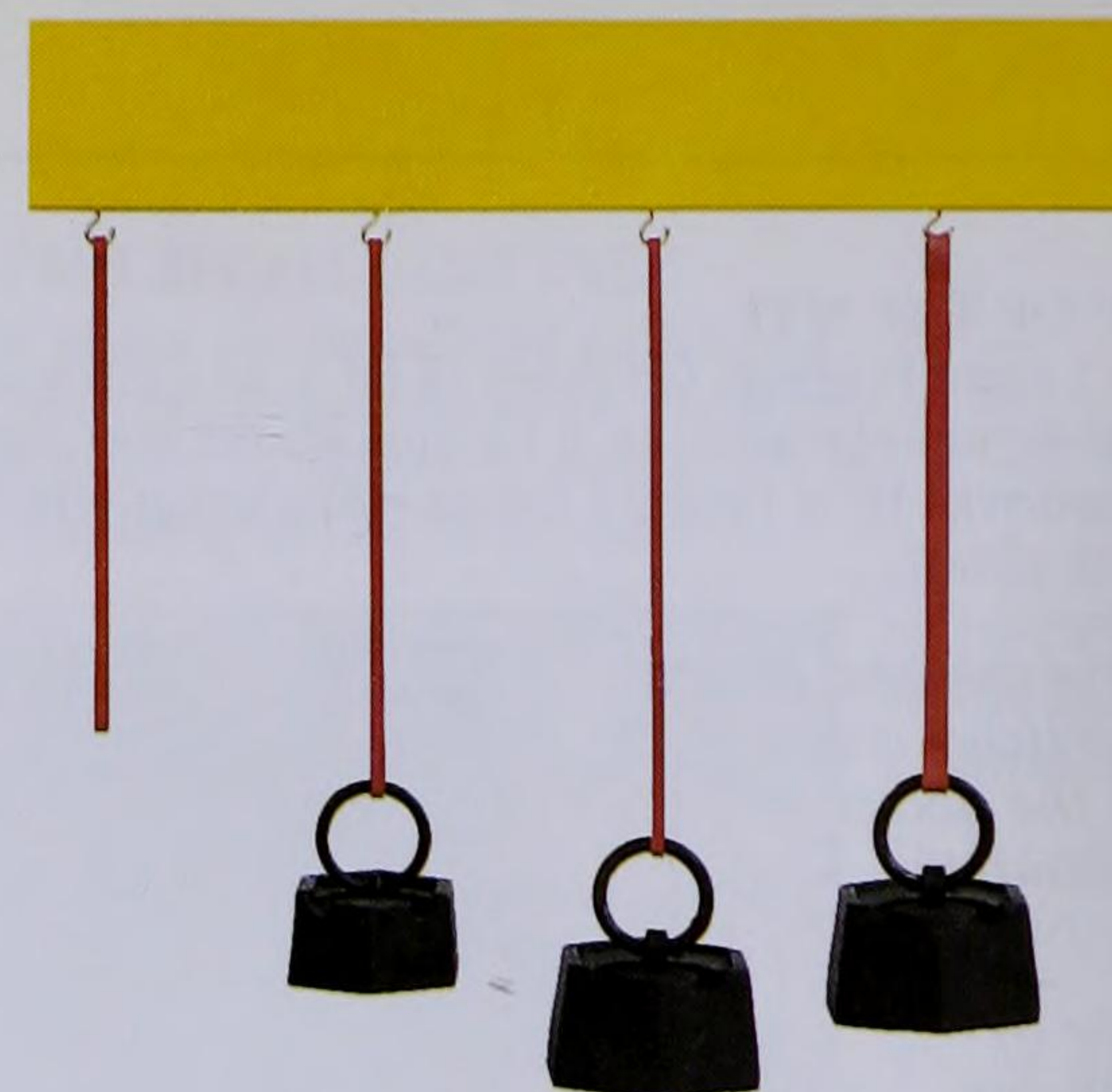
$$P = \rho gh$$

P = pressure

$\rho$  = liquid density

g = acceleration due to gravity

h = height of water above measured point



## ELASTICITY

The extension of a solid is proportional to the force applied to it

$$F \propto x$$

x = extension of solid

F = applied force

## GAS LAWS

### BOYLE'S LAW

Volume is proportional to one divided by pressure

$$V \propto 1/P$$

### CHARLES' LAW

Volume is proportional to temperature

$$V \propto T$$

### PRESSURE LAW

Pressure is proportional to temperature

$$P \propto T$$

### THE IDEAL GAS EQUATION

Pressure multiplied by volume is equal to ideal gas constant multiplied by temperature

$$PV = RT$$

V = volume

P = pressure

T = temperature

R = the ideal gas constant

## ELECTRIC CIRCUITS

### OHM'S LAW

Current is equal to voltage divided by resistance

$$I = V/R$$

### POWER

Power is equal to voltage multiplied by current

$$P = VI$$

I = current

V = voltage

R = resistance

P = power

## IMAGE FORMATION

One divided by the focal length is equal to one divided by the object's distance from lens added to one divided by distance from the lens to the image

$$1/f = 1/u + 1/v$$

f = focal length

u = object's distance from lens

v = distance from lens to image



# Appendix: useful data

PHYSICISTS USE STANDARD UNITS of measurement called SI units (Système International), which include the kilogram, the meter, and the second. In addition to these standard units, there are many other units of measurement. The tables below give details of these.

## TEMPERATURE SCALES

To convert	Into	Equation
Celsius (C)	Fahrenheit (F)	$F = (C \times 9 \div 5) + 32$
Fahrenheit	Celsius	$C = (F - 32) \times 5 \div 9$
Celsius	Kelvin (K)	$K = C + 273$
Kelvin	Celsius	$C = K - 273$
Fahrenheit	Kelvin	$K = ((F - 32) \times 5 \div 9) + 273$

## METRIC - IMPERIAL CONVERSIONS

To convert	Into	Multiply by
<b>Length</b>		
Centimeters	inches	0.3937
Meters	feet	3.281
Kilometers	miles	0.6214
Meters	yards	1.094
<b>Mass</b>		
Grams	ounces	0.03527
Kilograms	pounds	2.205
Metric tonnes	tons	1.102
<b>Area</b>		
Square centimeters	square inches	0.1550
Square meters	square feet	10.76
Hectares	acres	2.471
Square kilometers	square miles	0.3861
Square meters	square yards	1.196
<b>Volume</b>		
Cubic centimeters	cubic inches	0.06102
Cubic meters	cubic feet	35.31
<b>Capacity</b>		
Liters	pints (liquid)	1.760
Liters	gallons (liquid)	0.2200

## IMPERIAL - METRIC CONVERSIONS

To convert	Into	Multiply by
<b>Length</b>		
Inches	centimeters	2.540
Feet	meters	0.3048
Miles	kilometers	1.609
Yards	meters	0.9144
<b>Mass</b>		
Ounces	grams	28.35
Pounds	kilograms	0.4536
Tons	metric tonnes	0.9072
<b>Area</b>		
Square inches	square centimeters	6.452
Square feet	square meters	0.09290
Acres	hectares	0.4047
Square miles	square kilometers	2.590
Square yards	square meters	0.8361
<b>Volume</b>		
Cubic inches	cubic centimeters	16.39
Cubic feet	cubic meters	0.02832
<b>Capacity</b>		
Pints (liquid)	liters	0.5685
Gallons (liquid)	liters	4.546

## PHYSICS SYMBOLS

Symbol	Meaning
$\alpha$	alpha particle
$\beta$	beta particle
$\gamma$	gamma ray; photon
$\epsilon$	electromotive force
$\eta$	efficiency; viscosity
$\lambda$	wavelength
$\mu$	micro-; permeability
$\nu$	frequency; neutrino
$\rho$	density; resistivity
$\sigma$	conductivity
c	speed of light

## POWERS OF TEN

Factor	Name	Prefix	Symbol
$10^{18}$	quintillion	exa-	E
$10^{15}$	quadrillion	peta-	P
$10^{12}$	trillion	tera-	T
$10^9$	billion	giga-	G
$10^6$	million	mega-	M
$10^5$	thousand	kilo-	k
$10^2$	hundred	hecto-	h
$10^1$	ten	deca-	da
$10^{-1}$	one tenth	deci-	d
$10^{-2}$	one hundredth	centi-	c
$10^{-3}$	one thousandth	milli-	m
$10^{-6}$	one millionth	micro-	$\mu$
$10^{-9}$	one billionth	nano-	n
$10^{-12}$	one trillionth	pico-	p
$10^{-15}$	one quadrillionth	femto-	f
$10^{-18}$	one quintillionth	atto-	a

## BASE SI UNITS

Physical quantity	SI unit	Symbol
Length	meter	m
Mass	kilogram	kg
Time	second	s
Electric current	ampere	A
Temperature	kelvin	K
Luminous intensity	candela	cd
Amount of substance	mole	mol
Plane angle	radian	rad
Solid angle	steradian	sr

## DERIVED SI UNITS

Physical quantity	SI unit	Symbol
Frequency	hertz	Hz
Energy	joule	J
Force	newton	N
Power	watt	W
Pressure	pascal (newtons per square meter)	Pa (Nm <sup>-2</sup> )
Electric charge	coulomb	C
Voltage	volt	V
Electric resistance	ohm	$\Omega$



The above terms are those used in the US. In the UK one billion is known as one thousand million and one trillion is known as one billion.



# Glossary

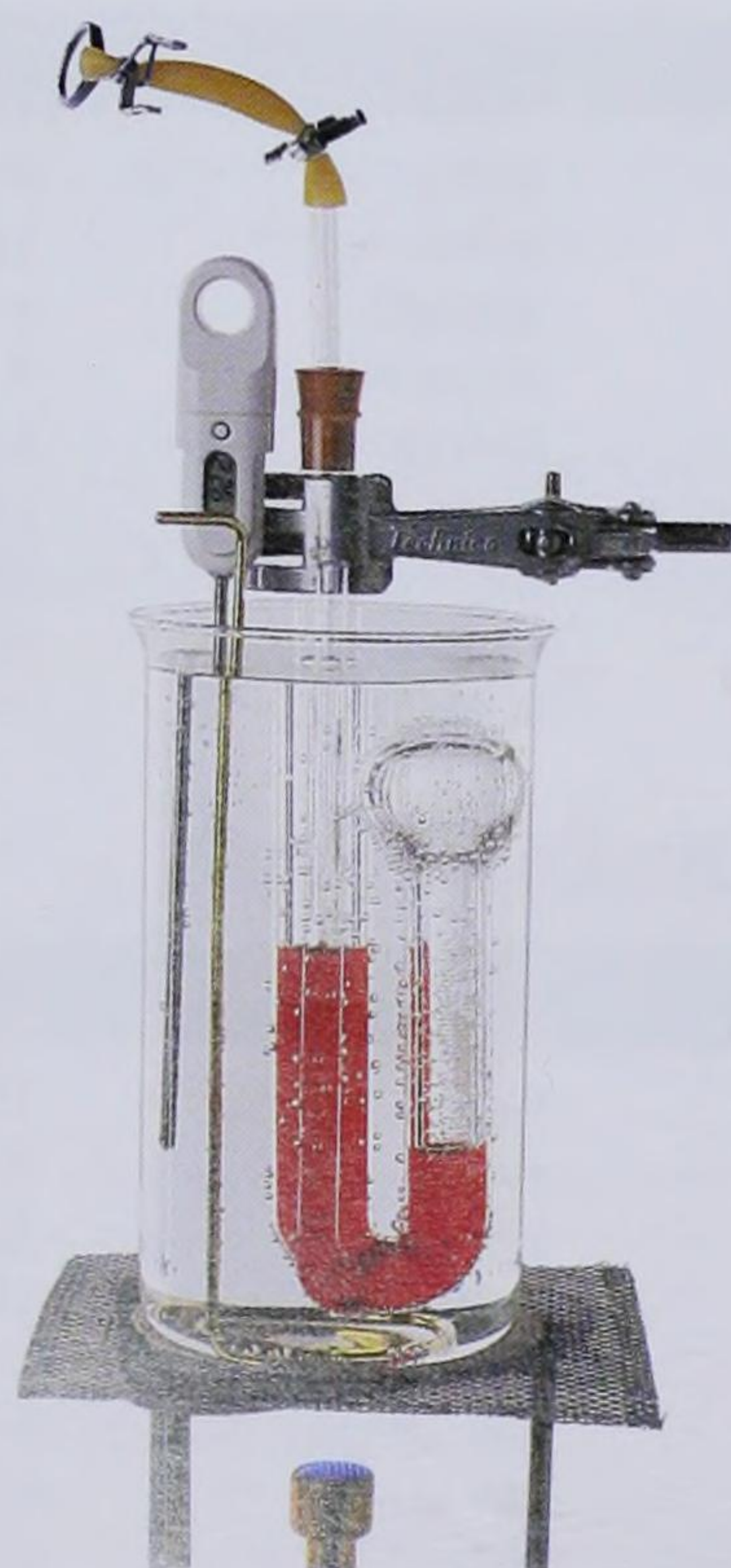
**ABSOLUTE ZERO:** The lowest possible **temperature**. The higher the temperature of matter, the more movement, or **kinetic energy**, its particles possess. At absolute zero the particles do not move at all. Absolute zero is zero **kelvin**,  $-273.15^{\circ}$  **Celsius** or  $-459.67^{\circ}$  **Fahrenheit**.

**ACCELERATION:** A change in the **speed** of an object. A reduction in speed is a negative acceleration, and is often called a deceleration. Acceleration is usually measured in  $\text{ms}^{-2}$  (meters per second per second, or meters per second squared).

**ACHROMATIC DOUBLET:** A system of two **lenses** that eliminates **chromatic aberration**. The two lenses are made of different types of glass.

**ADDITIVE PROCESS:** Combining light of different colors. When light of more than one color enters the eye, the result is a color that is different from each of the initial colors. This is due to human eyes having three types of **cone** cell. The brain combines the signals from each type of cone, and interprets the result.

**ADHESIVE FORCES:** The attractive **forces** between two different types of matter, such as water and glass. The balance between adhesive and **cohesive** forces determines whether the **meniscus** of a liquid will be upward or downward.



CHARLES' LAW

**ALPHA DECAY/PARTICLE:** The breakup of an unstable atomic **nucleus**, resulting in the release of a particle consisting of two **protons** and two **neutrons**—an alpha particle. During alpha decay, the **atomic number** of the nucleus reduces by two and the **atomic mass** by four (see **beta decay**).

**AMORPHOUS SOLID:** Any solid with particles not arranged in a regular, repeating pattern and therefore not composed of **crystals**. Because the particles are not regularly arranged, over a period of time, an amorphous solid can flow, and is often called a supercooled liquid.

**AMPLITUDE:** The intensity of a wave motion. For a sound wave, the amplitude determines how loud the sound will be. For a water wave, the amplitude is the height of the wave, half the distance from the peak to the trough.

**ANGULAR MOMENTUM**  
The product of the speed of rotation and the **moment** of **inertia** of a spinning object. The moment of inertia of an object is a measure of how hard it is to set the object spinning.

**ANODE:** The positive electrode of any electrical apparatus. Because the anode is connected to the positive electrical supply, **electrons** are attracted to it. Anodes are used in X-ray tubes and **cathode ray tubes**.

**ANTIPARTICLE:** A particle that has the same **mass** as another particle, but has one or more of its properties equal and opposite.

**ATOM:** A tiny particle. The building blocks of matter, atoms are the smallest part of an **element**. Atoms are typically  $10^{-10}\text{m}$  (or one ten-millionth of a millimeter) in diameter, and consist of a positively charged **nucleus** surrounded by negatively charged **electrons**.

**ATOMIC MASS:** The total mass of **protons** and **neutrons** in the **nucleus** of an **atom**, expressed in atomic mass units. Protons and neutrons each have a mass of one atomic mass unit. Fluorine-19, with nine protons and ten neutrons, has an atomic mass of nineteen.

**ATOMIC NUMBER:** The number of **protons** present in the **nucleus** of an **atom**. All fluorine atoms have an

atomic number of nine, because they have nine protons.

**BETA DECAY/PARTICLE:** The breakup of an unstable atomic **nucleus**, resulting in the release of a fast-moving **electron**. This electron is called a beta particle. During beta decay, the **atomic number** of the nucleus actually increases by one, because a **neutron** changes into a **proton**, releasing an **electron**. The **atomic mass** is unchanged (see **alpha decay**).

**BROWNIAN MOTION:** The random motion of small solid objects, such as smoke particles, which can be observed under a microscope. The movement is caused by **atoms** and **molecules** of liquid or gas bombarding the solid objects.

**BUBBLE CHAMBER:** A device used to detect particles in



LUBRICATION

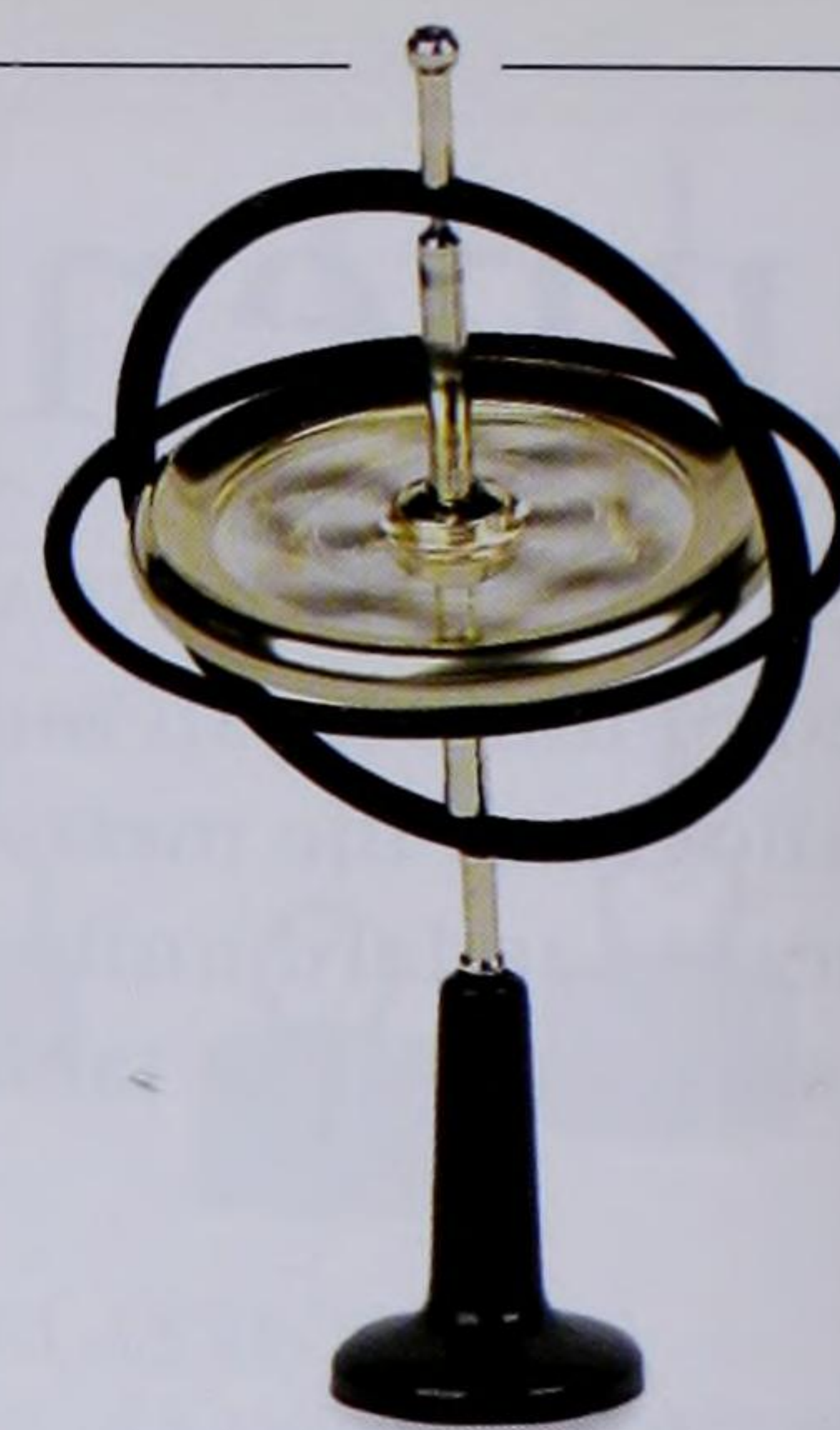
collisions that take place in particle accelerators. The chamber contains a liquid, such as liquid hydrogen, held just below its

boiling point. Any particles that have **electric charge** cause **atoms** in the liquid to become **ions**. The liquid boils around these ions, forming tiny gas bubbles wherever a charged particle passes. A strong **magnetic field** in the chamber causes the particles to travel in curved tracks, and each of the particle types can be identified by their tracks.

**CAPILLARY ACTION:** The rising or falling of a liquid in a narrow tube, above or below the liquid surface, due to surface tension. If **adhesive forces** are stronger than **cohesive forces**, as in the case of water in glass tubes, the liquid will climb up the tube. The narrower the tube, the higher the liquid will rise or fall.

**CATHODE:** The negative electrode of any electrical apparatus. Because the cathode is connected to the negative electrical supply, **electrons** are pushed away from it (see **anode**).

**CATHODE RAY TUBE:** A sealed glass tube used in the display of most televisions. Inside the tube,



GYROSCOPE

**electrons** leave a **cathode**, and are attracted toward the high-voltage **anode**. The electrons form a beam, sometimes called a cathode ray, which can be observed as it touches a **luminescent** screen.

**CELSIUS:** A **temperature** scale on which water freezes at  $0^{\circ}$  and boils at  $100^{\circ}$ . Each degree celsius is equal to one **kelvin**, but the kelvin scale begins at absolute zero ( $-273.15^{\circ}\text{C}$ ). Once called the Centigrade scale, the name was changed in 1948.

**CENTER OF GRAVITY:** The point of an object at which clockwise and counterclockwise **moments** are equal and the object therefore balances.

**CENTRIPETAL FORCE:** The **force** needed to keep an object moving in a circle or an **ellipse**. In the case of circular motion, the force is always directed to the center of the circle, and depends upon the **mass** and **speed** of the object and the radius of the circle.

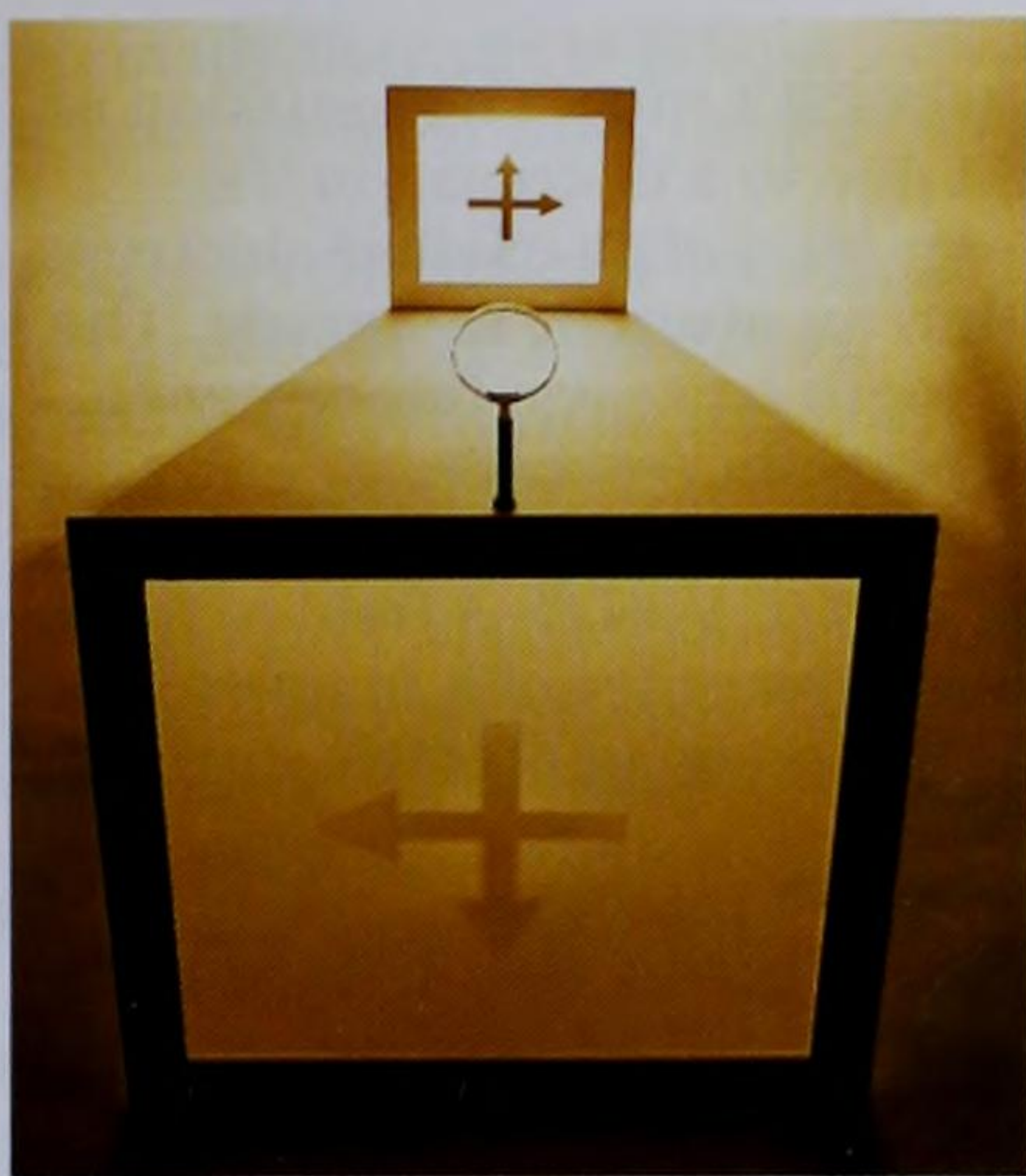
**CERN (CONSEIL EUROPEEN POUR LA RECHERCHE NUCLEAIRE):** The European Laboratory for Nuclear Physics, near Geneva on the Swiss-French border and run by nineteen European nations. It is the site of the Large Hadron Collider.

**CHAIN REACTION:** A process, such as nuclear fission, in which each reaction is in turn the stimulus of a further reaction.

**CHARGE:** See **electric charge**.

**CHROMATIC ABERRATION:** A defect in a **lens**, caused by the fact that different **wavelengths** of light refract by different amounts as they pass through glass. The result of the defect is that different colors of light focus at different points. An image produced by the lens therefore has colored fringes around it, often appearing distorted. The problem can be solved by using an **achromatic doublet**.





#### FOCUSING AN IMAGE

**CLOUD CHAMBER:** A device used to detect and track particles resulting from radioactive decay. It is a sealed unit containing a vapor, usually alcohol, just at the point of condensing to form a liquid. **Alpha** and **beta** particles possess **electric charge**, and for this reason provide sites around which the vapor can condense. The tracks of the particles appear as paths of tiny droplets (see **bubble chamber**).

#### COHESIVE FORCES

The attractive **forces** between atoms or molecules in a liquid, such as water. Cohesive forces are responsible for **surface tension** (see **adhesive forces**).

#### COMPONENT

The effect of a **force** in a particular direction. A force can be thought of as a combination of two or more components.

#### COMPRESSION

The action of squashing a substance, so that it takes up a smaller space. When a **gas** is compressed, its **pressure** increases. When a solid is compressed reaction forces are produced. These forces are responsible for the strength of a solid such as concrete, which is said to be "strong in compression."

**CONCAVE:** Shaped like the inside of a bowl. Concave mirrors make parallel light **converge**. Concave **lenses** make parallel light **diverge**.

**CONDUCTIVE:** Describes a material that allows **electric current** to flow through it easily. A material with a high conductivity allows electricity to flow easily, and is called a conductor. The term can be used to describe heat flow as well as the flow of electricity.

**CONES:** Cells on the surface at the back of the human eye (the retina), which are sensitive to light of particular ranges of color. The cones allow for color vision. There are three types of cone cell: red-, green-, and blue-sensitive.

**CONSTRUCTIVE INTERFERENCE:** The combination of two waves where the waves are "in step." Hence the peaks of one wave correspond to the peaks of the other. The **amplitude** of the resulting combined wave is the sum of the amplitudes of the individual waves.

**CONVERGE:** To come together, as parallel light does, when it comes to a point of focus.

**CONVEX:** Shaped like the outside of a bowl. Convex **lenses** make parallel light **converge**. Convex mirrors make parallel light **diverge**.

**CRITICAL ANGLE:** The angle at or above which light striking the boundary between two different materials undergoes **total internal reflection**.

**CRYOGENIC UNIT:** A cooling device used to reduce the temperature of substances to very low values, often only a few degrees above **absolute zero** (zero kelvin). Cryogenic units are used in particle accelerators, such as those at CERN, to produce liquid helium, which cools **electromagnets** necessary for the operation of the accelerator.

**CRYSTAL LATTICE:** A regular, repeating arrangement of **atoms** or **molecules** in a solid (see **unit cell**).

**CRYSTALS:** Solids whose atoms or molecules are arranged in a **crystal lattice**.

**CURRENT:** See **electric current**.

**DENSITY:** A measure of the concentration of **mass** in a substance. The numerical value for density is calculated by dividing the mass of a given amount of the substance by its **volume**.

**DESTRUCTIVE INTERFERENCE:** The combination of two waves where the waves are "out of step." This means that the peaks of one wave correspond to the troughs of the other wave. The **amplitude** of the combined wave is therefore reduced.

**DIFFRACTION:** The bending of waves around the edge of an object. A small gap consists of two edges, so the waves are bent into concentric curves.

**DIFFUSION:** The mixing of substances, caused by the random motion of their particles. Diffusion

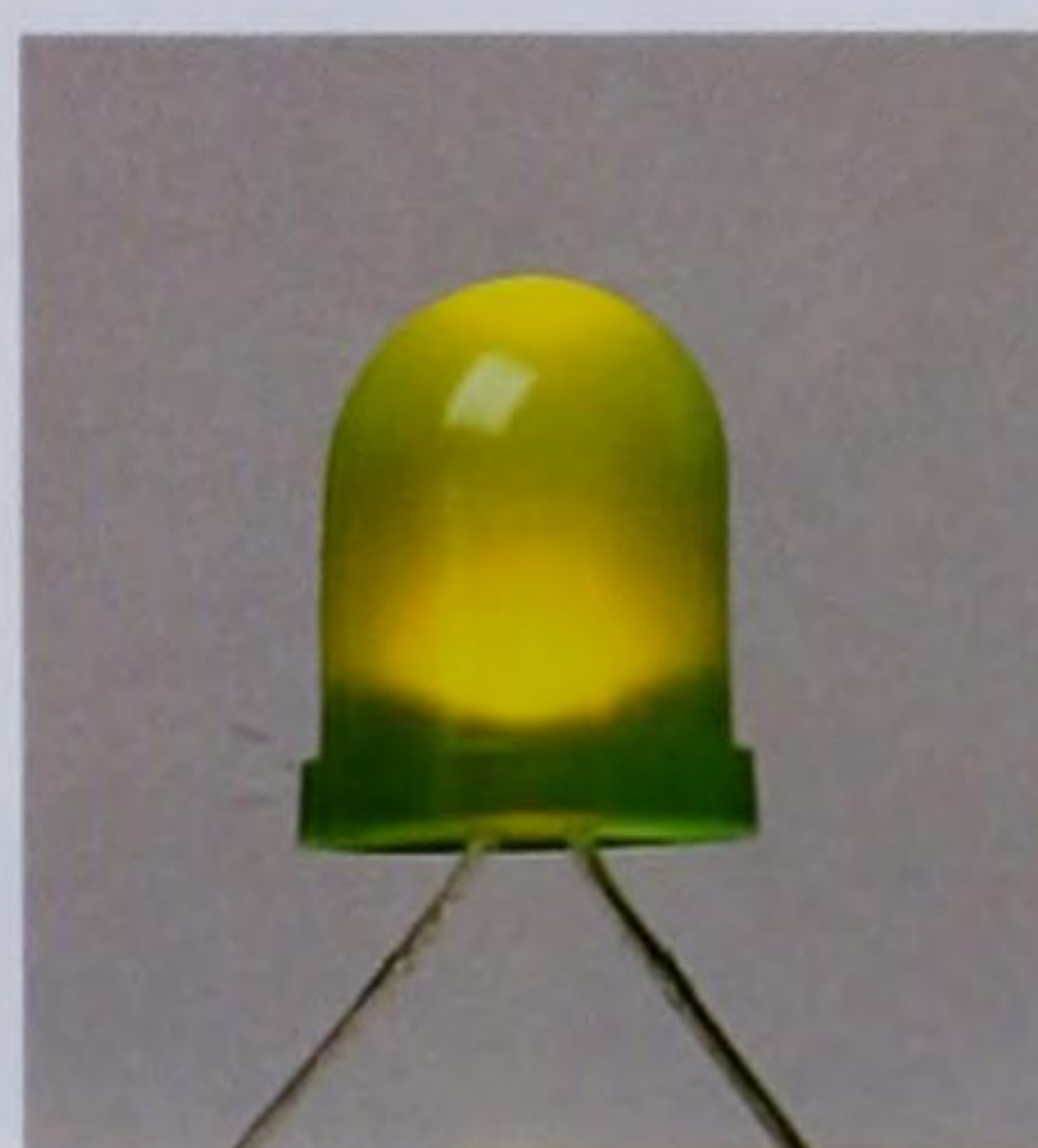
is most noticeable in **gases**, because the movement of the particles is much faster than it is in solids or liquids.

**DISPLACEMENT:** A movement away or the distance of an object from its normal position.

**DIVERGE:** To move apart, as parallel light does when it passes through a concave **lens**.

**DOMAINS:** Tiny regions, between 0.1 and 1 mm in width, within magnetic materials. Every **atom** of a magnetic substance, such as iron, is itself a tiny magnet. All of the atoms within a given domain are lined up, so that each domain is magnetized in a particular direction. In an unmagnetized state, the domains cancel each other out. When a material is magnetized, all of the domains are made to line up with each other.

**ELASTICITY:** The ability of a substance to regain its size and shape after being stretched by **forces** of **tension**. Forces of attraction between **atoms** within the substance are made stronger when the atoms are pulled apart. These forces are responsible for elasticity.



**ELECTRIC CHARGE:** A property of certain particles or substances that results in **electrostatic forces**. There are two types or signs of charge: positive and negative. The numbers of positive and negative charges in matter is normally balanced, giving no overall charge (see **ion**).

**ELECTRIC CURRENT:** The movement of particles with **electric charge**. Most electric currents are the result of moving **electrons**. The movement of electrons is caused by **electrostatic** or **electromagnetic forces**.

**ELECTRIC FIELD:** A region in which a particle with **electric charge** will experience an **electrostatic force**.

**ELECTRODE:** Part of an electrical device that is connected to the electrical supply. The positive electrode is called the **anode**, while the negative electrode is called the **cathode**.

**ELECTROMAGNET:** A device made by winding a continuous coil of wire around an iron core. **Electric current** flowing through the wire

creates magnetism that lines up the **domains** in the iron, turning it into a temporary magnet.

**ELECTROMAGNETIC FORCES:** The forces on **electric charges** moving in a **magnetic field**. The size and direction of the force depends upon the speed, sign, and size of the charge, and on the strength and direction of the magnetic field.

**ELECTROMAGNETIC RADIATION:** A form of energy that travels through space and matter. It is associated with **electric fields** and **magnetic fields**, and can be thought of as a wave motion involving these fields. It can also be thought of as a stream of particles called **photons**. The many types of radiation include light, radio waves, and X-rays.

**ELECTROMAGNETIC SPECTRUM:** The range of different types of **electromagnetic radiation**. Each type of radiation is identical except for its **wavelength** and its energy. Radiation with short wavelengths and high energy include X-rays and gamma rays, while longer-wavelength, lower-energy radiation includes infrared and radio waves.

**ELECTRON:** A particle found in all **atoms**. All electrons have one unit of negative **electric charge**.

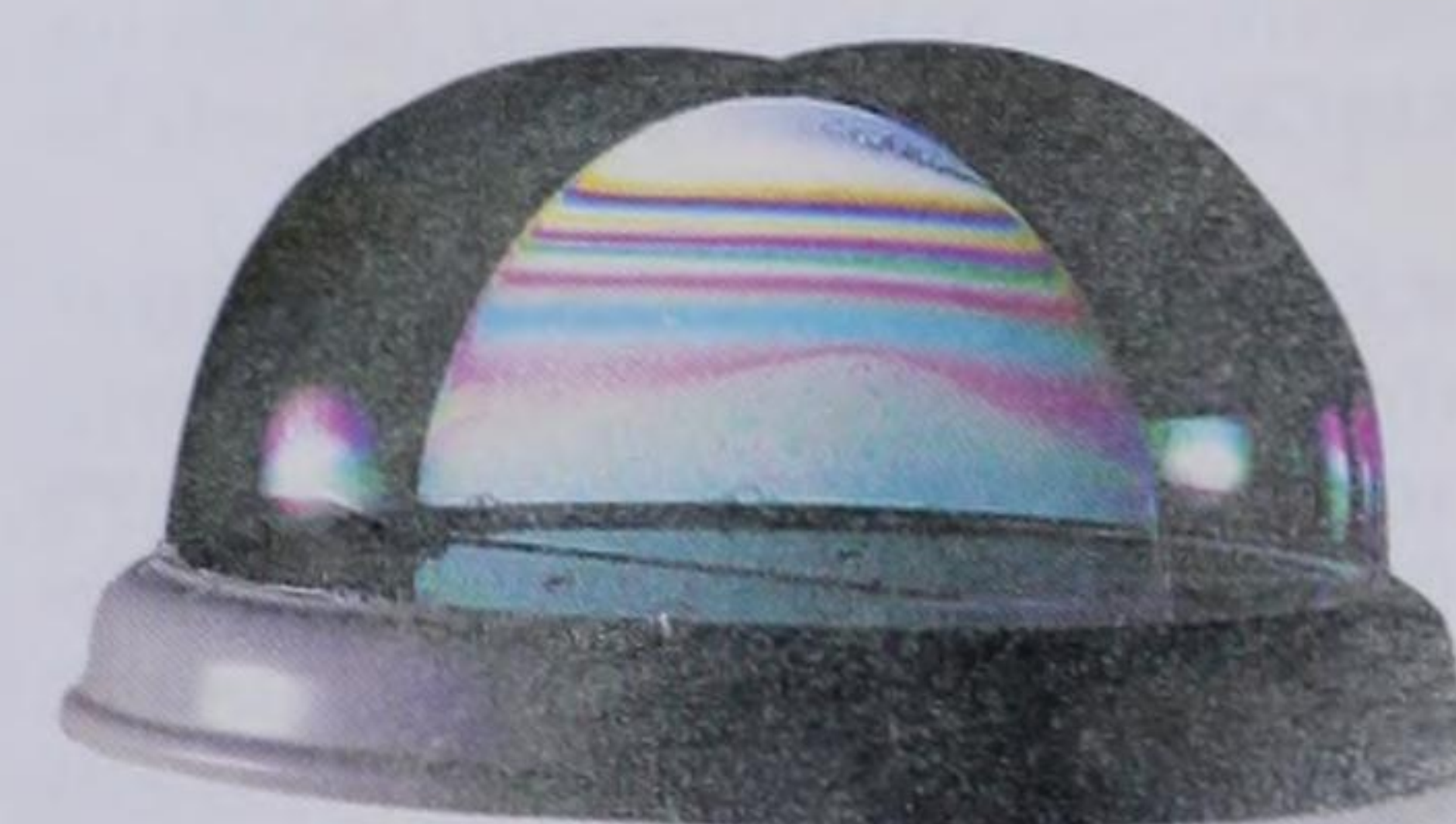
**ELECTROSCOPE:** An instrument for measuring the extent of imbalanced **electric charge** in an object. The most common example consists of a glass box with two pieces of gold foil that are pushed apart as they are charged by **induction**.

**ELECTROSTATIC FORCES:** The forces between electric charges. Two charges of the same sign will push apart, or repel. Charges of different sign pull together, or attract.

**ELEMENT:** A substance consisting of only one type of **atom**. Examples are hydrogen, oxygen, and fluorine.

**ELLIPSE:** A shape that looks like a flattened circle. The orbits of the planets are ellipses.

**EQUILIBRIUM:** A balanced state at which the **resultant** of a number of **forces** on an object is zero.



SOAP BUBBLE





## INDUCTION

**ERROR BAR:** A vertical or horizontal line that is drawn on a **graph** to indicate the margin of accuracy with which a particular measurement was taken.

**EVAPORATION:** The loss of **atoms** or **molecules** from a liquid, as they break free of the liquid to become a vapor. Evaporation takes place below the boiling temperature of the liquid.

**EXCITED:** In possession of extra energy. **Electrons** in **atoms** can be excited by heat or light energy. When they do so, they occupy a new position in the atom, according to their new energy.

**FAHRENHEIT:** Scale of **temperature** on which water freezes at 32 degrees and boils at 212 degrees.

**FILAMENT:** The fine wire in an incandescent light bulb. The filament heats up when electric current flows through it. At high temperatures, the filament glows.

**FISSION:** The splitting of unstable **nuclei** of **atoms**. The process begins as a free **neutron** joins the nucleus, making it more unstable. The nucleus splits into two smaller, more stable nuclei and releases further free neutrons and a large amount of energy in a **chain reaction**. This release of energy is used in atom bombs

**FLUID:** Any substance that flows. Liquids and gases are both fluids.

**FLUORESCENCE:** A type of **luminescence** in which a substance glows with visible light immediately after being **excited** by invisible ultraviolet radiation.

**FOCAL LENGTH:** The distance from a **lens** or curved mirror at which a parallel beam of light comes to a focus.

**FORCE:** A push or a pull.

**FREQUENCY:** The regularity with which something happens. It is most often applied to a wave or vibration. The number of times the complete cycle of a wave happens each second is the frequency of the wave, and is measured in hertz (Hz).

**FULCRUM:** The point about which an object turns. For example, the fulcrum of a lever is its pivot.

**FUSION:** A joining of small **nuclei** of **atoms** to form larger nuclei. In some cases, such as when hydrogen atoms fuse to form atoms of the **element** helium, there is a huge release of energy during the process.

**GAUGE BOSON:** A particle exchanged between two interacting particles. At the sub-microscopic level of the tiniest particles, the exchange is responsible for the four forces: **gravity**, **electromagnetic force**, the **weak interaction**, and the **strong nuclear force**.

**GEIGER-MULLER TUBE:** A device for detecting radioactivity. An **electric current** flows between the wall of the tube and a metal wire at its center when **alpha** or **beta particles** enter.

**GENERATOR:** A machine that produces an electrical **voltage** whenever its rotor is turned. The **kinetic energy** of the rotor becomes electrical energy because of the coils and magnets.

**GLUONS:** Particles responsible for carrying the **strong nuclear force** (see **gauge boson**), according to modern theory.

**GRAPH:** A visual representation of an experiment's results. A graph will highlight any relationships between the various types of data.

**GRAVITY/GRAVITATION:** A force of attraction between all objects with **mass**. The size of the force depends upon the masses of the two objects and the distance between the objects. Modern theory says that gravity is carried by particles known as gravitons (see **gauge boson**).

**GYROSCOPE**  
Usually thought of as a spinning metal disk supported in a metal cage, the word can refer to any spinning object. Gyroscopes have stability as they spin.

**HADRON**  
Any particle that is composed of **quarks**. Examples are the **proton** and the **pi meson**.

**HIGGS BOSON:** Hypothetical particle whose existence would link the **electromagnetic force** and the **weak interaction**, and explain why particles have mass.

**HOLE:** A vacant electron position within the **crystal lattice** of a **semiconductor** that can be thought of as a positive charge.

**IMAGE:** A picture formed by a lens or a curved mirror. Images cast on screens by convex lenses are called real images, while those seen using telescopes or microscopes, which cannot be directly projected, are called virtual images.

**INDUCTION:** 1. The apparent charging of one object by an electrically charged object nearby. The charging is apparent since it is only a shift of **electric charge** within the object. 2. The magnetization of iron objects in the presence of a magnet. The **domains** inside the iron line up with the **magnetic field** of the magnet.

**INERTIA:** The resistance of an object to any change in its motion.

**INFRARED RADIATION:** A type of **electromagnetic radiation**, with a **wavelength** shorter than visible light.

**INTERFERENCE:** The combination of two or more waves.

**ION:** An **atom** with an overall electric charge. The numbers of positively charged protons and of negatively charged electrons in an atom are normally equal. But the removal of one or more electrons leaves the atom with a net positive charge, while extra electrons give a net negative charge.

**KELVIN:** The absolute scale of **temperature**, the Kelvin scale begins at absolute zero, and unlike the **Celsius** and **Fahrenheit** scales, does not rely on fixed points.

**KINETIC ENERGY:** The energy of a moving object, dependent upon the mass and speed of the object.

**LATENT HEAT:** Heat energy that melts a solid or vaporizes a liquid. Latent heat does not raise the temperature of the substance.

**LENS:** A curved piece of glass or other transparent material that refracts light, and can form **images**.

**LIMITING FRICTION:** The **force** which must be overcome to start an object moving when it is in contact with a surface. It is greatest between rough surfaces.

**LUMINESCENCE:** The emission of light due to a decrease in the energy level of an **excited electron** within an **atom** or **molecule**. The two main types are **fluorescence** and **phosphorescence**.

**MAGNETIC FIELD:** A field of **force** around the poles of a magnet or a wire carrying an **electric current**.

**MASS:** The measure of an object's **inertia**. Mass is also defined by **gravitation**. The gravitational force between two objects depends upon their masses.

**MELTING POINT:** The temperature at which a solid substance becomes a liquid. It is dependent upon atmospheric **pressure**.

**MENISCUS:** The curved surface of a liquid where it meets the container. It is caused by a combination of **adhesive** and **cohesive forces**.

**MESONS:** A **hadron** consisting of two quarks. An example is the **pi meson**, which carries the **strong nuclear force** between **protons** and **neutrons** within the **nucleus**.

**MICROMETER:** A device used to measure very small displacements.

**MOLECULE:** The smallest amount of a compound. A water molecule consists of two **atoms** of hydrogen and one of oxygen.

**MOMENT:** The turning effect of a **force**.

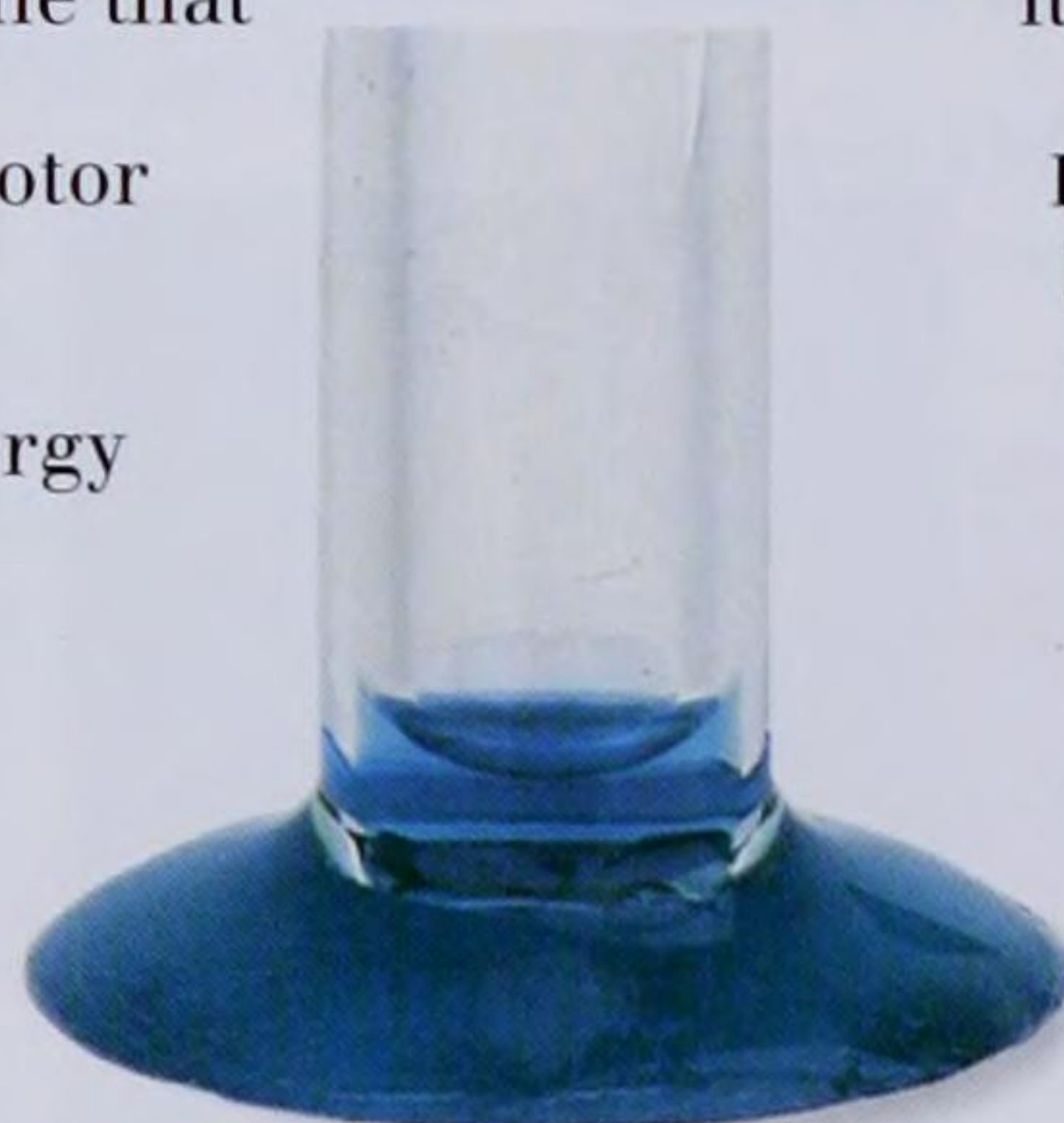
**NEUTRON:** One of the particles in the **nucleus** of an **atom**. It is a **hadron**, and has no **electric charge**.

**NEWTON METER:** A device used to measure **force**. A pointer moves along a scale as a spring inside the meter extends. The extension of the spring depends upon the applied force.

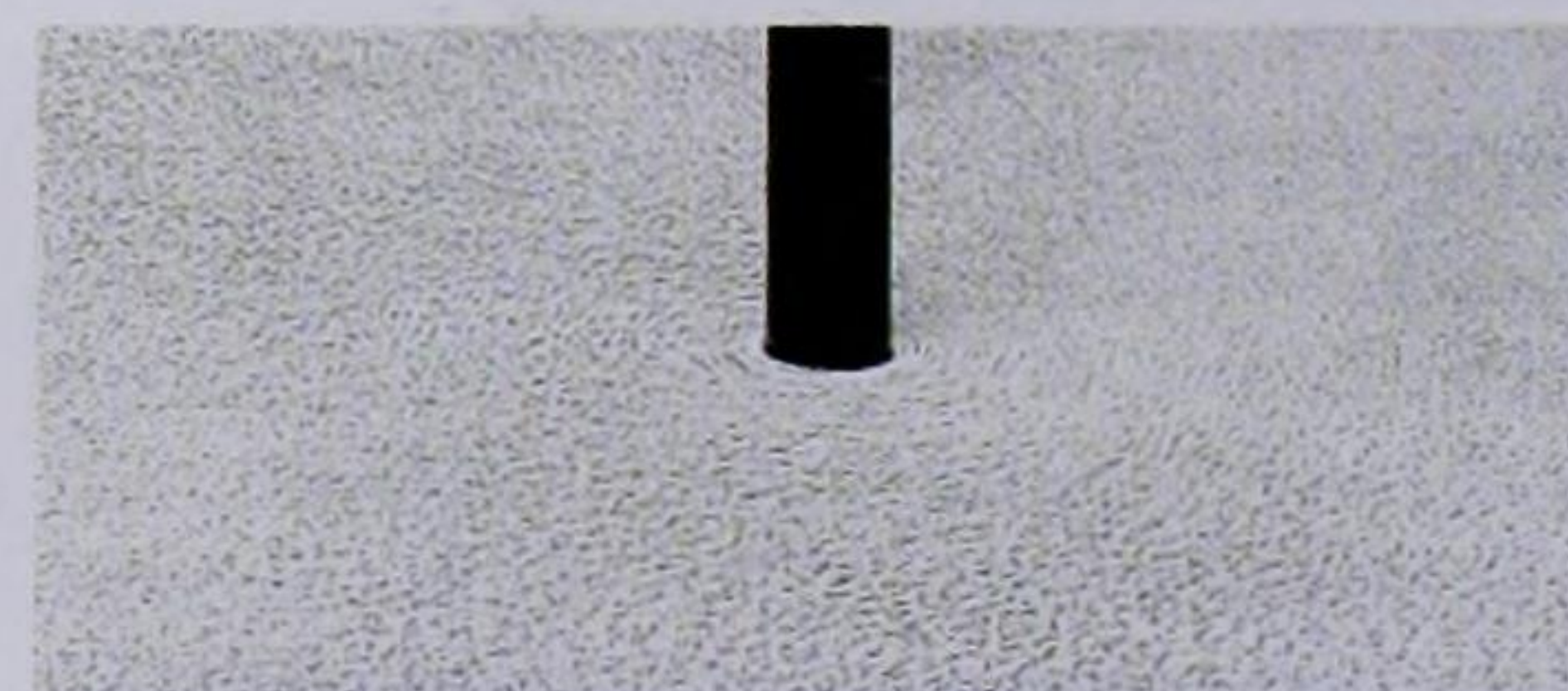
**NUCLEAR REACTIONS:** Changes involving the **nuclei** of **atoms**, such as **fission** and **fusion**.

**NUCLEUS:** The central part of an **atom**. It has a positive **electric charge** because it contains **protons**.

**ORBITALS:** Regions of an **atom** in which **electrons** are found. The name comes from the word "orbit," since electrons were



## MENISCUS



## MAGNETIC FIELD



originally thought to follow definite paths around the **nucleus**.

**OSCILLATOR:** An electric circuit that produces an alternating **electric current**, which repeatedly changes direction.

**PERMANENT MAGNETS:** Objects with a fixed magnetism. The **domains** in a permanent magnet always align to produce a **magnetic field** (see **electromagnet**).

**PHOSPHORESCENCE (PHOSPHOR):** A type of **luminescence** in which a substance glows with visible light some time after being **excited** (see **fluorescence**). A phosphor is any substance exhibiting phosphorescence.

**PHOTON:** A particle of **electromagnetic radiation**. The **energy** of a photon depends only upon the **wavelength** of the radiation. A photon can be thought of as a packet of waves.

**PLANE WAVE:** A wave motion in which the waves are parallel to one another and perpendicular to the direction of the wave's motion.

**POSITRON:** The **antiparticle** of the **electron**. It is identical to the electron in every way, except that it has positive **electric charge**.

**POTENTIAL ENERGY:** Energy that is "stored." An object held in the air has potential energy by virtue of its height and the **gravitational force** pulling it downward.

**PRESSURE:** A measure of the concentration of a **force**. The pressure exerted by a force is equal to the size of the force divided by the area over which it acts. Solids, liquids and gases exert pressure.

**PRIMARY COLORS:** A set of three colors, which, when combined in the correct proportion, can produce any other color. The set of

primaries for the **additive process** is different from that for the **subtractive process**.

**PRINCIPLE OF SUPERPOSITION:** The rules governing the **interference** of waves.

**PRINCIPLE OF THE CONSERVATION OF ENERGY:** Energy can be neither created nor destroyed, it can only change, or transfer, from one form to another.

**PROTON:** A **hadron** found within the **nucleus** of an **atom**. It has a positive **electric charge**.

**QUARKS:** Particles that combine together to form **hadrons**, such as **protons** and **neutrons**. No quark has ever been detected in isolation.

**RAREFACTION:** The lowering of the **density** and **pressure** of a gas; the opposite of **compression**.

**REACTION:** A **force** produced by an object that is equal and opposite to a force applied to the object.

**REFRACTION:** The bending of light or other **electromagnetic radiation** as it passes from one material to another.

**RESISTANCE:** A measure of the opposition to the flow of **electric current**. It is the ratio of **voltage** to current.

**RESULTANT:** The combined effect of two or more **forces**.

**SANKEY DIAGRAM:** An illustration of the energy changes in a process. The diagram consists of a large arrow, which represents the input of energy to the process, and which splits according to the energy changes that occur.

**SEMICONDUCTOR:** A material in which the **electrons** are held only loosely to their **atoms**. The electrons can become free, and the material therefore becomes **conductive**, with only a small input of **energy**.

**SHELL:** An energy level occupied by **electrons** within an **atom**. It is generally true that the lower the energy of electrons in the shell, the closer it is to the **nucleus**.

**SI UNITS (SYSTEME INTERNATIONAL D'UNITES):** A system of units accepted by the worldwide scientific community as the standard. Its seven base units include the kilogram and second.

**SOLENOID:** A long coil of wire, which produces a **magnetic field**

similar to that of a bar magnet. With an iron bar inside the coil, a solenoid becomes an **electromagnet**.

**SOLUTION:** A liquid mixture of substances in which the particles of the substances are uniformly mixed.

**SPEED:** The rate at which an object moves, equal to the distance moved divided by the time taken.

**STATE:** The form of a substance, either solid, liquid, or gas.

**STRONG NUCLEAR FORCE:** The **force** between **hadrons**. It is carried by **gluons**, or by combinations of **quarks** (see **gauge boson**). The strong nuclear force is responsible for holding the **nucleus** together.

**SUBTRACTIVE PROCESS:** The process by which pigments absorb parts of the visible spectrum of light, but reflect others, making objects appear to have color.

**SUPERCOOLED LIQUID:** See **amorphous solid**.

**SURFACE TENSION:** The **resultant force** at the surface of a liquid, due to the **cohesive forces** between the particles of the liquid.

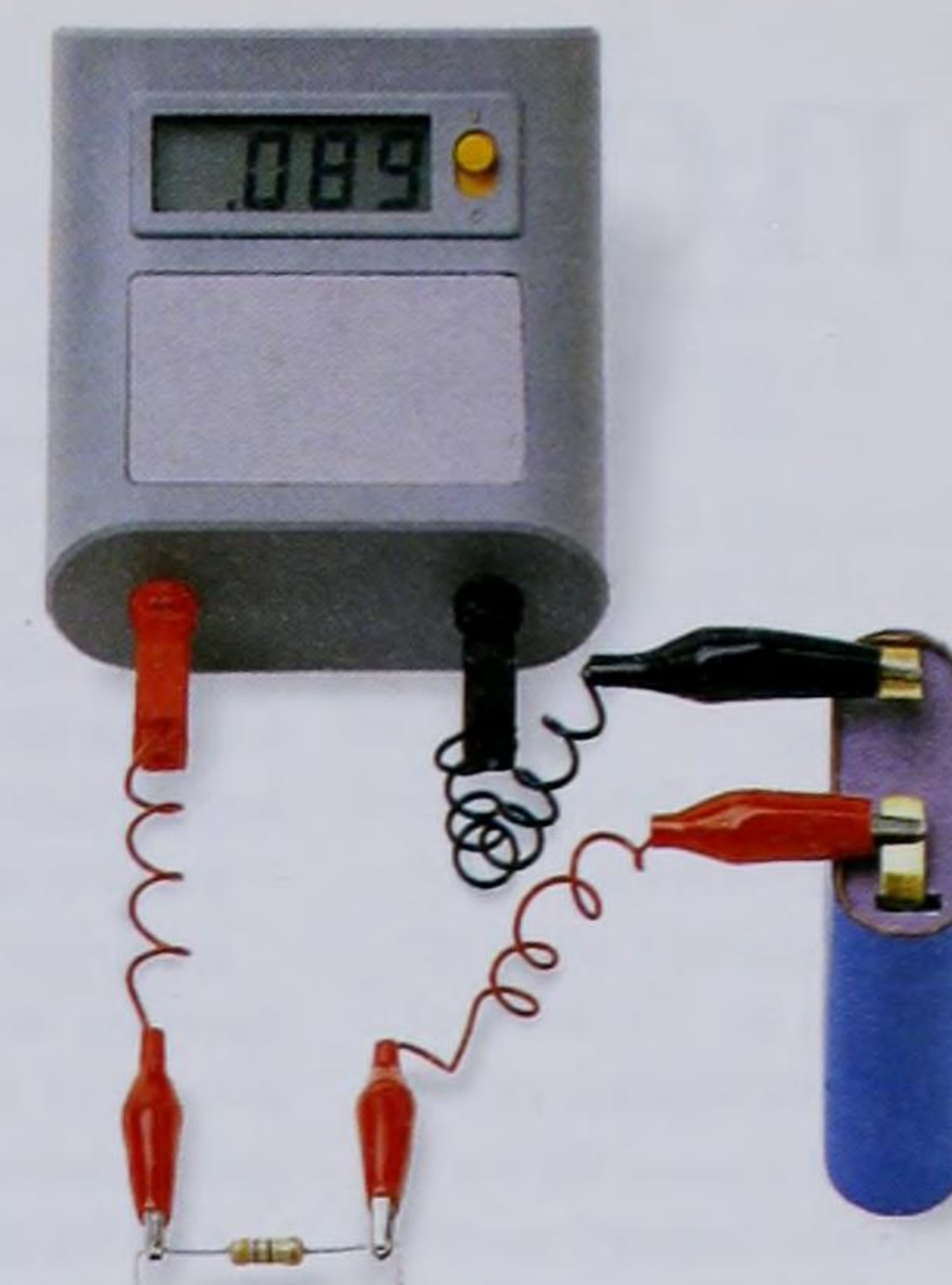
**TEMPERATURE:** How hot or cold a substance is. Temperature relates directly to the **kinetic energy** of a substance's particles. The particles in hotter objects have more kinetic energy.

**TENSION:** A **reaction force** in a solid that is stretched, which pulls the **atoms** of the solid together. It is the opposite of **compression**.

**TERMINAL VELOCITY:** The maximum speed attained by an object falling through a liquid or gas. A parachute falling through air has a relatively low terminal velocity, while that of a ball bearing will be much greater.

**THERMAL EXPANSION:** The expansion of a solid as its **temperature** increases. It is due to the increased vibration of the **atoms** and **molecules** of the solid. This increased vibration occurs at higher temperatures due to the increased **kinetic energy** of the atoms and molecules.

**TOTAL INTERNAL REFLECTION:** The reflection of light from the border between two materials, as the light passes from the denser to the less dense material.



MEASURING RESISTANCE

**TURBINE:** A machine in which a liquid or a gas causes rotation. When attached to a **generator**, the turbine helps to generate electricity.

**UNIT CELL:** The group of **atoms** or **molecules** in a crystal; when repeated, it forms the **crystal lattice**.

**UPTHRUST:** An upward **force** on an object immersed in a liquid or a gas. Upthrust is the **resultant** of the liquid or gas pressure acting on the object.

**VELOCITY:** The **speed** and direction of an object's motion.

**VERNIER SCALE:** A scale attached to instruments such as callipers, which allows very accurate measurements to be taken.

**VOLTAGE:** A measure of the **force** on particles with **electric charge**. The voltage in an **electric circuit** pushes **electrons** around the circuit.

**VOLUME:** The amount of space an object takes up.

**WAVELENGTH:** The distance from one wave peak to another. The wavelength of electromagnetic radiation determines the type of radiation. For example, X-rays have a shorter wavelength than light.

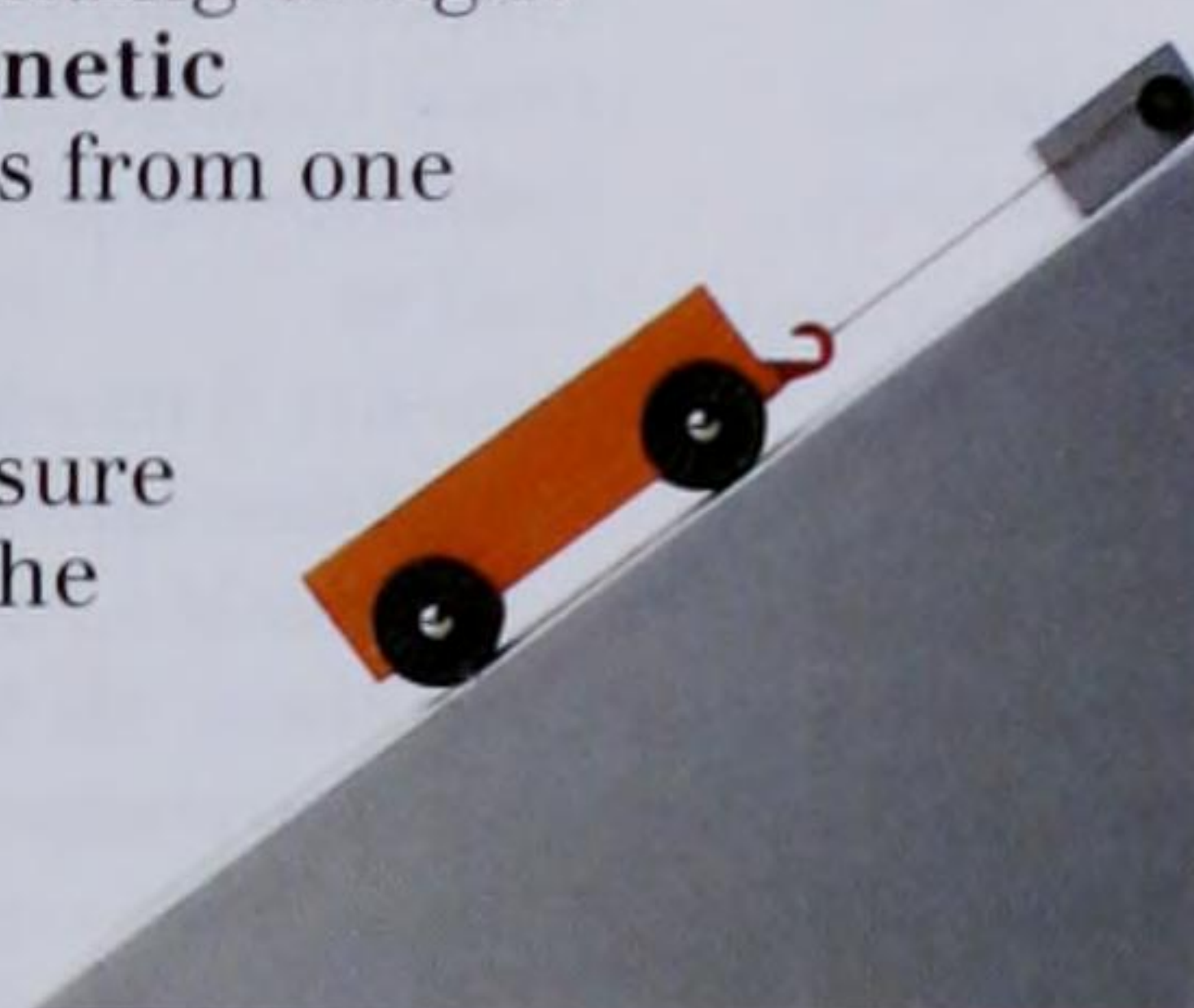
**THE WEAK INTERACTION:** A **force** between some types of particle, including **electrons**. Also involved in the decay of **hadrons**, such as the **beta decay** of **neutrons** in the **nucleus**. The force is carried by W and Z particles (see **gauge boson**).

**WEIGHT:** The **force** of **gravity** on an object, dependent on the mass of the object.

**WORK:** The amount of energy involved in a particular task. For example, work is said to be done when a pulley lifts a load. The amount of work done is equal to the force acting multiplied by the distance moved.



ONE KG MASS



AN INCLINED PLANE



# Index

## A

Absolute zero 22, 56  
 cryogenic unit 57  
 Kelvin scale 58

Absorption  
 color 40  
 light 42

Acceleration 56  
 formula 54  
 measurement 12, 13  
 Newton's Laws 12, 13  
 terminal velocity 14

Accuracy, error bars 58

Achromatic doublet 56  
 chromatic aberration 44-5

Acres 55

Additive process 56  
 color 40, 41  
 primary colors 59

Adhesive forces 56  
 capillary action 26, 56  
 liquids 26  
 meniscus 58

Air  
 diffusion 28  
 liquification 22  
 soap bubbles 26  
 sound waves 20, 21  
 total internal reflection 42  
 upthrust 27

Air molecules  
 Brownian motion 6  
 evaporation 6  
 sound waves 20, 21

Air pressure  
 increasing pressure with  
 depth 27  
 melting point 58  
 sound waves 21

Air resistance 14  
 formula 54

Alcohol 57

Alligator clip 9

Alpha decay 50, 56

Alpha particles 56  
 cloud chamber 57  
 Geiger-Müller tube 58  
 symbol 55

Ammeter  
 electric circuit 33  
 measuring electric  
 current 34  
 measuring resistance 5, 32

Amorphous solid 56  
 atoms 24

Amount of substance 55

Ampere 55

Amplitude 56  
 constructive  
 interference 57  
 destructive  
 interference 57  
 waves 20

Amps 32

Angles  
 critical 57  
 crystal systems 24-5  
 light reflection 42  
 light refraction 42  
 SI units 55  
 total internal reflection  
 42

Angular momentum 19, 56

Anode 56  
 cathode ray tubes 49, 56  
 electrodes 57

Antenna 58

Antiparticle 56  
 particle accelerators 52  
 positrons 59

Aperture 44

Area 55

Atmospheric pressure 27  
 melting point 58  
 sound waves 21

Atomic mass 56

Atomic number 56  
 alpha decay 56  
 beta decay 56  
 nuclear physics 50

Atoms 56  
 Brownian motion 6, 56  
 cohesive forces 57  
 crystal lattice 57  
 crystal systems 24  
 domains 57  
 elasticity 57  
 electric charges 30  
 electric circuits 32  
 electric current 32  
 electromagnetic  
 radiation 39  
 electrons 48-9, 57  
 elements 50, 57  
 equalization of  
 temperature 25  
 evaporation 6, 58  
 excited electrons 58  
 fission 58  
 friction 14  
 fusion 58  
 gases 28  
 heat energy 22  
 ions 30, 58  
 luminescence 58  
 matter 6  
 neutrons 58  
 nuclear physics 50-1  
 nuclear reactions 58  
 nucleus 58  
 orbitals 59  
 particle physics 52  
 protons 59  
 semiconductors 59  
 shell 59  
 solar cells 37  
 solids 24  
 solids under tension 25  
 solutions 6  
 thermal expansion 24, 59  
 unit cell 59

Axes 16

Axis 19

Axles  
 ball bearings 15  
 wheels 17

## B

Ball bearings 15

Balloons, hot air 29

Barrages, tidal power 37

Batteries  
 cars 7

electric circuit 33  
 electric current 32

Bearings  
 ball bearings 15  
 gyroscopes 19

Beta decay 56  
 nuclear physics 50  
 weak interaction 59

Beta particle 56  
 cloud chamber 57  
 Geiger-Müller tube 58  
 symbol 55

Bicycles 17

Billion 55

Binoculars  
 lenses 42  
 total internal reflection 42

Block and tackle 16

Blue light  
 additive process 41  
 color vision 40  
 lenses 45  
 photons 38  
 subtractive process 41  
 white light spectrum 39

Boiling point 22  
 Celsius scale 56  
 Fahrenheit scale 58

Bone 39

Bosons  
 gauge 58  
 Higgs 58  
 particle physics 52

Bourdon gauge 2, 29

Boyle's Law 2, 28  
 formula 54

Brain 56

Brakes  
 cars 7  
 friction 15

Bright filament lamps 40

Bromine 28

Brownian motion 6, 56

Bubble chamber 56  
 cosmic rays 50  
 particle accelerators 52

Bubbles  
 colors 46, 47  
 soap 5  
 surface tension 26

Buildings 11

Bunsen burner 2

Burning 22

## C

Caliper units 15

Calipers 59

Cameras 44

Candela 55

Capacity 55

Capillary action 26, 56

Carbon dioxide 50

Cars  
 brakes 7  
 energy transfer 7  
 engines 12

Cathode 56  
 electrodes 57

Cathode ray oscillograph 21

Cathode ray tube 5, 56  
 anodes 56  
 electrons 48, 49  
 television 48, 49

Cells, unit 24, 59

Celsius scale 22, 56  
 absolute zero 56  
 conversion table 55

Center of gravity 56  
 point of suspension 11

Centigrade scale 56  
 see also Celsius scale

Centimeters, conversion  
 table 55

Centripetal force 18-19, 56  
 formula 54  
 particle accelerators 55

CERN (Conseil Européen  
 pour la Recherche  
 Nucléaire) 56

Large Hadron Collider  
 52, 53

Chain reaction 56  
 nuclear fission 51

Chainwheel 17

Charged atoms 30

Charles' Law 28, 56  
 experiment 29  
 formula 54

Chemical energy 7

Chromatic aberration 56  
 lenses 44, 45

Circles  
 centripetal force 56  
 ellipses 57

Circuit breaker 36

Circular motion 18-19

Circular waves 46

Clockwise moment 11

Cloud chamber 50, 57

CMS detector 53

Coal 56

Cohesive forces 57  
 capillary action 56  
 liquids 26  
 meniscus 56, 58  
 surface tension 59

Coils  
 electric motors 35  
 electromagnets 34, 35  
 generators 36  
 solenoids 35  
 televisions 49

Cold 22, 59

Collisions 52

Color 40-1  
 additive process 40, 41,  
 56  
 chromatic aberration 56  
 destructive interference  
 47  
 fluorescence 48  
 primary colors 40, 59  
 secondary colors 41  
 soap bubbles 46, 47  
 subtractive process 41,  
 59  
 vision 57

Commutator 35

Compass  
 electromagnetism  
 affecting needle 1, 34  
 magnetism 31

Component 57  
 weight 10

Compound microscope 44

Compounds 58

Compression 57  
 solids 24  
 sound waves 21  
 springs 20

Concave lens 43

Concave mirror 43  
 telescopes 45

Condensation 22

Conduction 22

Conductivity 55

Conductors  
 electric circuits 32  
 semiconductors 59

Cones (eyes)  
 additive process 56  
 color vision 40, 57

Conservation of energy 7  
 Principle of 59

Constant acceleration 54

Constructive interference 46,  
 57

"Conventional current" 32

Converge 57

Converging light rays 42

Conversion tables 55

Convex lenses  
 eyes 42  
 images 58

Convex mirrors 45

Cooling 57

Cooling towers 51

Corkscrew 16

Corners 19

Cosmic rays 50

Coulomb, symbol 55

Counterclockwise moment  
 11

Critical angle 57  
 total internal reflection 42

Cross-sectional area 25

Crown wheel 7

Cryogenic unit 57

Crystal lattice 57  
 holes 58  
 unit cells 24, 59

Crystallization 6  
 states of matter 22

Crystals 57  
 cubic system 24  
 external features 24  
 hexagonal system 25  
 monoclinic system 25  
 orthorhombic system 24  
 particles 6  
 solids 24  
 tetragonal system 24  
 triclinic system 25  
 trigonal system 25  
 unit cells 24, 59

Cubic centimeters 55

Cubic feet 55

Cubic inches 55

Cubic meters 55

Cubic system 24

Cyan 41

## D

Data  
 graphs 8, 58  
 interpreting 9

Daughter nucleus 51

Decay  
 alpha decay 50, 56  
 beta decay 50, 56

Deceleration 56

Deflection 49

Degree Celsius 22

Degree Fahrenheit 22

Density 57  
 rarefaction 59  
 solids 24  
 symbol 55  
 total internal reflection 59

Depth 27

Depth of field 43

Destructive interference 46,  
 57

Diffraction 57  
 electromagnetic  
 radiation 38  
 waves 46

Diffusion 57  
 gases 28

Digital thermometer 9

Dim filament lamp 40

Disk brakes 15

Displacement 57  
 oscillation 20

Dissolving 6

Distance  
 measurement 8  
 speed 15, 59

Diverge 57

Diverging light rays 42

Domains 57  
 induction 58  
 magnetism 30, 31  
 permanent magnets 59

Double pulley 17

"Down" quarks 52

Downward meniscus 26

Drawing pin 5, 11

## E

Earth  
 centripetal force 18  
 cosmic rays 50  
 geothermal power 37  
 gravity 8, 12, 13  
 magnetic field 31  
 temperature 22

Efficiency 55

Effort 16  
 levers 17

Elasticity 57  
 formula 54  
 solids 24  
 solids under tension 25

Electric charge 30, 57  
 bubble chambers 56  
 electromagnetic  
 radiation 38  
 electrons 57  
 electroscopes 57  
 induction 58  
 nuclear physics 50  
 nucleus 58  
 positrons 59  
 protons 59  
 SI units 55  
 voltage 59

Electric circuits 32-3  
 formulas 54  
 oscillators 58-9  
 resistance 32  
 voltage 59

Electric current 57  
 cathode ray tubes 49  
 circuits 32-3  
 conductivity 57  
 electromagnetism 34  
 magnetic field 58  
 Ohm's Law 32  
 oscillators 58-9  
 resistance 59  
 SI units 55  
 solar cells 37

Electric field 57  
 cathode ray tubes 48, 49



- electrostatic forces 30
  - radio waves 38
  - Electric motor
    - electromagnetism 34, 35
    - energy transfer 7
    - kinetic energy 7
  - Electric resistance 5
    - SI units 55
  - Electricity 30-1
    - car batteries 7
    - energy 6
    - loudspeakers 21
    - nuclear physics 50
    - Ohm's Law 2
    - oscilloscope 21
    - photovoltaic cell 7
    - resistance 5
    - static 30
    - Van de Graaff generator 5, 31
  - Electricity generation 36-7, 58
    - geothermal power 36, 37
    - nuclear power stations 51
    - renewable sources 36-7
    - solar energy 37
    - turbines 59
    - water power 36
    - wind power 37
  - Electrode 57
    - anodes 56
    - batteries 32
    - cathodes 56
  - Electromagnetic force 57
    - gauge bosons 58
    - Higgs bosons 58
    - particle physics 52
  - Electromagnetic radiation
    - 38-9, 57
    - infrared 58
    - light 42
    - photons 59
    - refraction 59
    - Sun 7
    - wavelength 59
  - Electromagnetic spectrum
    - 38-9, 57
    - visible light 40
  - Electromagnetism 34-5
    - compass needle 1, 34
    - electric motors 35
    - force 10
    - freefall experiment 9
    - particle physics 52
    - solenoids 35
  - Electromagnets 35, 57
    - electricity generation 36
    - loudspeakers 21
    - particle accelerators 55
  - Electromotive force 55
  - Electron 48-9, 57
    - anodes 56
    - batteries 32
    - beta particles 56
    - cathode ray tube 5, 49, 56
    - cathodes 56
    - electric circuits 32
    - electric current 32, 57
    - electromagnetic radiation 38, 39
    - excited 48, 58
    - fluorescence 48
    - ions 58
    - lasers 47
    - light bulbs 33
    - luminescence 58
    - negative electric charge 30
  - nuclear physics 50
    - orbitals 48, 59
    - particle physics 52
    - positrons 59
    - semiconductors 59
    - shell 59
    - sodium lamps 40
    - solar cells 37
    - static electricity 30
    - Van de Graaff generator 31
    - voltage 59
    - weak interaction 59
    - X-rays 39
  - Electron gun 49
  - Electronic charge 55
  - Electroscope 30, 57
  - Electrostatic forces 57
    - electric charges 30
    - electric circuits 32
  - Element 57
    - atoms 48, 56
    - protons 50
  - Ellipse 57
    - centripetal force 56
    - planetary orbits 19
  - Energy 6-7
    - cathode ray tubes 49
    - conduction 22
    - electricity generation 36-7
    - electromagnetic radiation 38
    - electrons 48
    - fusion 58
    - generators 58
    - heat 22-3
    - kinetic 6, 58
    - latent heat 23, 58
    - nuclear fission 51
    - nuclear fusion 51
    - nuclear physics 50
    - nuclear power stations 51
    - particle physics 52
    - phosphorescence 49
    - photons 59
    - potential 59
    - Principle of the Conservation of Energy 7, 59
    - Sankey diagram 59
    - semiconductors 59
    - SI units 55
    - Sun 7
    - transfers 7
    - waves 20
    - work 16, 59
  - Equalization of temperature 23
  - Equations
    - ideal gas equation 54
    - temperature conversion 55
  - Equilibrium 10, 57-8
    - oscillation 20
  - Error bar 58
  - Evaporation 58
    - states of matter 6, 22
  - Excited electrons 48, 58
  - Expansion
    - solids 25
    - thermal 24, 59
  - Expansion joints 24
  - Experiment 8-9
    - freefall experiment 9
    - graphs 58
  - Eyes
    - color vision 40
  - cone cells 56, 57
  - laser surgery 47
  - lenses 44
  - seeing by reflected light 42
- ## F
- Fahrenheit scale 22, 58
    - absolute zero 56
    - conversion table 55
  - Falling 9
  - Feet 55
  - Feynman diagram 52
  - Field of vision 43
  - Filament 58
    - cathode ray tubes 49
    - light bulbs 33
  - Filament lamp 40
  - Film
    - cameras 44
    - thin film interference 47
  - Filters 41
  - Fission 58
    - chain reaction 56
    - nuclear physics 50, 51
  - Floating 26, 27
  - Fluid 58
    - lubrication 14, 15
  - Fluorescence 48, 58
  - Fluorescent lamp 40
  - Fluorine 57
    - atomic mass 56
    - atomic number 50, 56
    - atoms 48
  - Focal length 58
    - concave mirrors 45
    - convex lenses 43
    - formula 54
    - microscopes 44
    - telescopes 45
  - Focal point 45
  - Focusing an image 57
    - human eye 44
    - image formation 43
    - microscopes 44
  - Force 10-13, 58
    - adhesive 26
    - centripetal 18-19, 56
    - cohesive 26, 57
    - component 57
    - compression 24
    - electric fields 50
    - electromagnetism 10, 57
    - electrostatic 30, 57
    - equilibrium 10
    - formula 54
    - friction 14-15
    - gravity 10, 13
    - limiting friction 58
    - magnetism 31
    - moment 58
    - Newton meters 58
    - Newton's Laws 12-13
    - oscillation 20
    - particle physics 52-3
    - pressure 11, 59
    - reaction 10, 59
    - resultant 10, 59
    - SI units 55
    - simple machines 16
    - strong nuclear force 59
    - tension 24, 25
    - turning 11
    - upthrust 27
    - weak interaction 59
    - work 16
  - "Force lines" 34
  - Formulas 54
  - Foundations 11
  - Fractures 39
  - Francis turbine 36
  - Freefall experiment 9
  - Freezing point
    - mercury 22
    - states of matter 22
    - water 22, 56, 58
  - Frequency 58
    - SI units 55
    - sound waves 21
    - symbol 55
  - Friction 14-15
    - between surfaces 2
    - centripetal force 19
    - formula 54
    - limiting 14, 58
    - Newton meter 3
  - Fulcrum 58
    - levers 17
    - scales 8
  - Fusion 58
    - nuclear physics 50, 51
- ## G
- Gallons 55
  - Gamma rays
    - electromagnetic spectrum 38, 39
    - radioactivity 50
    - symbol 55
  - Gases 28-9
    - Boyle's Law 28
    - Charles' Law 29
    - compression 57
    - diffusion 28, 57
    - evaporation 6
    - fluids 58
    - formulas 54
    - natural gas 22
    - particles 22
    - pressure 28-9
    - Pressure Law 29
    - rarefaction 59
    - temperature 22, 28-9
    - turbines 59
    - upthrust 59
  - Gasoline 7
  - Gauge boson 58
    - particle physics 52
  - Gears
    - electric motors 7
    - wind turbines 37
  - Geiger-Müller tube 50, 58
  - Generator 58
    - electricity generation 36-7
    - hydroelectricity 36
    - nuclear power stations 51
    - turbines 59
    - Van de Graaff 5, 51
    - wind power 37
  - Geothermal power 36, 37
  - Glass
    - capillary action 26
    - flow 22
    - lenses 58
    - refraction 42
    - total internal reflection 42
  - Glass prism 59
  - Gluons 58
    - particle physics 52
    - strong nuclear force 59
  - Gold leaf electroscope 30
  - Grams 55
  - Graphs 8, 58
    - error bars 58
    - freefall experiment 9
  - Gravitons 58
  - Gravity and gravitation 58
    - center of gravity 56
    - centripetal force 18
    - force 10
    - formula 54
    - freefall experiment 9
    - gauge bosons 58
    - gyroscopes 19
    - mass 58
    - oscillation 20
    - particle physics 52
    - planets 19
    - potential energy 59
    - solar system 12, 13
    - tidal power 37
    - turning force 11
    - weight 8, 59
  - Green light
    - additive process 41
    - color vision 40
    - destructive interference 47
    - LED (light-emitting diode) 40
    - subtractive process 41
    - white light spectrum 39
  - Gyroscope 3, 58
    - angular momentum 19
    - centripetal force 18
- ## H
- Hadron 58
    - mesons 58
    - neutrons 58
    - particle accelerators 55
    - particle physics 52
    - protons 59
    - quarks 59
    - strong nuclear force 59
    - weak interaction 59
  - Hammer thrower 18
  - Headlamps 7
  - Heat
    - cars 7
    - conductivity 22, 57
    - electromagnetic radiation 39
    - energy 6, 22-3
    - equalization of temperature 23
    - heating 23
    - latent 2, 22, 58
    - nuclear power stations 51
    - radiation 22
    - temperature 59
  - Hectare 55
  - Helium
    - cryogenic unit 57
    - nuclear fusion 51, 58
    - particle accelerators 55
  - Hertz 21
    - symbol 55
  - Hexagonal system 25
  - Higgs boson 58
    - particle physics 52
  - High pressure 21
  - High-voltage anodes 56
  - Hole 58
  - Hot-air balloons 29
  - Hot objects 39
  - Hundred 55
  - Hydroelectric power 36
- ## I
- Hydrogen 57
    - bubble chambers 56
    - cosmic rays 50
    - nuclear fusion 51, 58
    - water molecules 58
  - Ideal gas equation 54
  - Image 58
  - Image formation
    - convex lenses 45
    - focusing 57
    - formula 54
    - human vision 44
    - lenses 58
  - Immersed objects 27
  - Imperial units 55
  - Inches 55
  - Inclined plane 16
  - Induction 57-8
    - charged atoms 30
    - electrical charge 58
    - electroscopes 57
    - magnetization 58
  - Inertia 12, 58
    - angular momentum 56
    - mass 58
  - Infrared radiation 58
    - electromagnetic spectrum 38
    - hot objects 39
    - transfer of heat 23
  - Interference 58
    - constructive 46, 57
    - destructive 46, 57
    - electromagnetic radiation 38
    - Principle of Superposition 59
    - wave behavior 46-7
  - Interpreting data 9
  - Inverted images 45
  - Ion 30, 58
    - bubble chambers 56
    - Van de Graaff generator 51
  - Iron
    - Earth's magnetic field 31
    - electromagnetism 34, 35
    - magnetism 30, 31
    - melting point 22
- ## J
- Joule 6
    - symbol 55
- ## K
- Kelvin symbol 55
  - Kelvin scale 22, 58
    - absolute zero 56
    - conversion table 55
  - Kilogram 8
    - conversion table 55
    - symbol 55
  - Kilogram metres
    - per second 13
  - Kilometers 55
  - Kinetic energy 6, 58
    - brakes 15
    - energy transfer 7
    - generators 36, 58
    - heat 22
    - heat transfer 23
    - solids 24
    - temperature 59



# L

Lamps 40

Large Electron Positron (LEP) collider 53

Large Hadron Collider (LHC) 52, 53

Lasers 46, 47

Latent heat 2, 58

  melting a substance 22, 23

Lead 22

Length

  conversion table 55

  measurement 8

  SI units 55

Lens 42-3, 58

  achromatic doublets 44, 45, 56

  binoculars 42

  cameras 44

  chromatic aberration 44, 45, 56

  concave 57

  convex 57

  eyes 42, 44

  focal length 58

  formula 54

  images 58

  microscopes 44

  telescopes 44, 45

Levers

  fulcrum 58

  three classes 17

Light

  additive process 41

  cameras 44

  chromatic aberration 56

  color 40

  color vision 57

  constructive interference 47

  converging 57

  destructive interference 46, 47

  divergence 57

  electromagnetic radiation 38

  electromagnetic spectrum 39

  electrons 48

  energy 6

  fluorescence 48, 58

  focal length 58

  lasers 46, 47

  lenses 43, 58

  luminescence 58

  mirrors 43

  optical instruments 44

  phosphorescence 59

  radiation from hot objects 59

  reflection 42-3

  refraction 42, 59

  speed of 42

  subtractive process 41, 59

  telescopes 45

  total internal reflection 59

  velocity symbol 55

  wavelength 59

Light bulbs 33

  color 40

  electric circuit 33

  filaments 58

Light-emitting diode (LED) 57

  color 40

  digital thermometer 9

Lightning

  lightning conductor 37

  temperature 22

Limiting friction 14, 58

Liquid pressure 54

Liquids 26-7

  capillary action 26, 56

  cohesive forces 57

  crystallization 6

  dissolving 6

  drops and bubbles 26

  evaporation 6, 58

  fluids 58

  latent heat 58

  melting point 58

  meniscus 26, 58

  particles 22

  pressure 26, 27

  solutions 59

  supercooled 22, 56

  temperature 22

  turbines 59

  upthrust 27, 59

Liquification 22

Liters 55

Load 16

  levers 17

Long-wave radio 38

Longitudinal waves 20

Loudspeakers 21

Low pressure 21

Lubrication 14, 15

Luminescence 58

  electrons 48

  fluorescence 48, 58

  phosphorescence 49, 59

Luminescent screens 56

Luminous intensity 55

# M

Machines 16-17

Magenta

  additive process 41

  subtractive process 41

Magnetic field 58

  bubble chambers 56

  cathode ray tubes 48, 49

  “force lines” 34

  particle accelerators 52

  permanent magnets 59

  radio waves 38

  solenoids 35, 59

Magnetism 30-1

  domains 57

  electromagnetic radiation 38-9

  electromagnetism 34-5

  induction 58

Magnets

  compasses 31, 34

  electricity generation 36

  electromagnets 35, 57

  loudspeakers 21

  permanent 59

Magnification 44-5

Magnifying glasses 43

Manganese 48

Mass 58

  antiparticles 56

  atomic 56

  centripetal force 18, 56

  conversion table 55

  density 57

  gravity 58

  Higgs bosons 58

  kinetic energy 58

  lifting with pulleys 17

  measurement 8

  momentum 13

  Newton’s Laws 12

  oscillation 20

  particle physics 52

  SI units 55

  weight 59

Matter 6-7

  adhesive forces 56

  atoms 48-9

  Brownian motion 6

  crystallization 6

  dissolving 6

  evaporation 6

  heat energy 22

  mass 8

  particle physics 52-3

  states of 22

  temperature 22

Measurement 8-9

  acceleration 12, 13

  distance 8

  electric charge 57

  electric current 32

  error bars 58

  force 10

  frequency 21

  gas pressure 29

  heat 8

  imperial-metric conversions 55

  latent heat 2

  length 8

  limiting friction 14

  liquid pressure 27

  mass 8

  metric-imperial conversions 55

  micrometer 58

  momentum 13

  Newton meters 58

  pressure 11

  radioactivity 50

  resistance 5, 59

  SI units 59

  speed 12, 13

  temperature 22, 55

  thermal expansion 24

  time 8

  turning force 11

  Vernier scale 59

  weight 8

  width 8

Mechanical advantage 16

Mechanical energy 7

Medical photography 39

Medium-wave radio 38

Melting

  melting point 23, 58

  states of matter 22

Meniscus 2, 26, 58

  adhesive and cohesive forces 56

  downward 20

  upward 26

Mercury

  meniscus 2, 26

  thermometers 9

Mesons 58

  particle physics 52

Metals

  electric circuits 32

  electromagnetic radiation 39

  welding 22

Meters 8

  conversion table 55

  symbol 55

Meters per second 12, 13

Meters per second per second 12, 13

Micro- 55

Micrometer 58

  measuring thermal expansion 24

Microphone 21

Microscope

  compound 44

  images 58

  traveling 8

Microwaves 38

Miles 55

Million 55

Mirrors 42, 43

  cameras 44

  concave 57

  convex 57

  focal length 58

  images 58

  microscopes 44

  telescopes 45

Mole 55

Molecule 58

  Brownian motion 6, 56

  capillary action 26

  cohesive forces 57

  crystal lattice 57

  evaporation 6, 58

  gases 28

  heat energy 22

  induction 30

  luminescence 58

  matter 6

  thermal expansion 59

  unit cell 59

Moment 58

  center of gravity 56

  of inertia 56

  turning force 11

Momentum 15

  angular 19, 56

  formula 54

  oscillation 20

Monoclinic system 25

Moon

  gravity 12, 13

  telescopic images 45

  tidal power 37

Motion

  Brownian 6, 56

  circular 18-19

  displacement 57

  energy of 7

  formula 54

  friction 14

  gas particles 28

  inertia 58

  Newton’s Laws 12-13

  oscillation 20

  precession 19

  speed 59

Motorcycles

  brakes 15

  centripetal force 18, 19

# N

Naphthalene 2, 23

Natural gas 22

Navigation 31

Negative electric charge 30

  electric current 30, 32

Neon lamp 40

Neutrino 55

Neutron 58

  atomic mass 56

  beta decay 56

  chain reaction 56

  mesons 58

  nuclear fission 51, 58

  nuclear physics 50

  particle physics 52

  quarks 59

Newton, Sir Isaac 12

Newton meter 58

  measuring force 10

  measuring limiting friction 3, 14

  measuring resultant force 4

  measuring turning force 11

  measuring weight 8

Newton meters 11

Newtons 8

  measuring force 10

  symbol 55

Newton’s Laws 12-13

  formula 54

Newtons per square meter 11

Nichrome wire 2

North seeking pole 31

Nuclear force 59

Nuclear physics 50-1

Nuclear power stations 51

Nuclear reactions 58

  Sun 7

Nucleus 58

  atomic mass 56

  atoms 48

  beta decay 56

  chain reaction 56

  fission 58

  fusion 58

  neutrons 58

  particle physics 52

  protons 59

  strong nuclear force 59

  weak interaction 59

# O

Objective lenses

  binoculars 42

  microscopes 44

  telescopes 45

Oceans 37

Ohms 32

  symbol 55

Ohm’s Law 2, 32

  formula 54

Oil

  electricity generation 36

  lubrication 15

Optical brighteners 48

Optical instruments 44-5

  binoculars 42

  cameras 44

  microscopes 44

  telescopes 45

Orange light

  neon lamps 40

  sodium lamps 40

  white light spectrum 39

Orbitals 48, 59

Orbits

  centripetal force 18

  planetary 19

Orthorhombic system 24

Oscillation 20-1

  electromagnetic radiation 38

  wave behavior 46

Oscillator 58-9

Oscilloscopes 21

Ounces 55

Oxygen 57

  water molecules 58

# P

Paper

  burning 22

  induction in 30

Particle accelerators 52-3

  bubble chambers 56

  cryogenic unit 57

Particle physics 52-3

Particles

  absolute zero 56

  amorphous solids 56

  atoms 48-9, 56

  Brownian motion 6, 56

  bubble chambers 56

  cloud chamber 57

  crystallization 6

  diffusion 57

  dissolving 6

  electric current 57

  electric field 57

  electromagnetic radiation 38, 57

  evaporation 6

  gases 22, 28

  gauge bosons 58

  gluons 58

  gravitons 58

  hadrons 58

  kinetic energy 23

  light 47

  liquids 22, 26

  matter 6

  melting 23

  nuclear physics 50

  quarks 59

  solids 22

  solutions 59

  supercooled liquid 22

  temperature 22

  weak interaction 59

Pascal 55

Pedals 17

Pendulum 20

Penstock 36

Pentaprism 44

Permanent magnets 59

  loudspeakers 21

Permeability symbol 55

Phosphor 59

  fluorescent lamps 40

  television screens 49

Phosphorescence 59

  luminescence 58

Phosphorescent screens

  cathode ray tubes 49

  television 49

Photography

  cameras 44

  X-rays 39

Photon 59

  electromagnetic radiation 38, 57

  light 46, 47

  particle accelerators 52

  particle physics 52



- symbol 55
  - Photovoltaic cell 7
  - Pi mesons 52
    - hadrons 58
    - mesons 58
  - Pigments
    - color 40
    - subtractive process 41, 59
  - Pints 55
  - Pistons 15
  - Pivot 58
    - turning forces 11
  - Plane 16
  - Plane angle 55
  - Plane wave 59
    - wave behavior 46
  - Planets
    - centripetal force 18
    - gravity 19
    - orbits 19
  - Plexiglass 14
  - Point of suspension 11
  - Poles 31
  - Pollution 36
  - Polyethylene rod 30
  - Positive electric charge 30
    - electric current 30, 32
  - Positive ions 31
  - Positron 59
    - particle accelerators 52
  - Potassium permanganate
    - crystallization 6
    - dissolving 6
  - Potential energy 6, 7, 59
    - cars 7
    - hydroelectricity 36
  - Pounds 55
  - Power
    - formula 54
    - SI units 55
  - Power stations
    - generators 36
    - hydroelectricity 36
    - nuclear power 50, 51
  - Precession 19
  - Pressure 11, 59
    - Boyle's Law 28
    - Charles' Law 29
    - compression 57
    - formula 54
    - gases 28-9
    - liquids 26, 27
    - melting point 58
    - rarefaction 59
    - SI units 55
    - upthrust 59
  - Pressure Law 28, 29
    - formula 54
  - Primary colors 59
    - additive process 41
    - vision 40
  - Principle of the
    - Conservation of Energy 6, 59
    - Principle of Superposition 46, 59
  - Prisms
    - binoculars 42
    - telescopes 45
    - white light spectrum 39
  - Proton 59
    - atomic mass 56
    - atomic number 56
    - cosmic rays 50
    - electromagnetic radiation 39
    - hadrons 58
    - ions 58
    - mesons 58
    - nuclear physics 50
    - nucleus 58
    - particle accelerators 52, 53
    - particle physics 52
    - positive electric charge 30
    - quarks 59
    - weak interaction 59
  - Proton synchrotron (PS) ring 53
  - Pull 10
  - Pulleys 5, 16, 17
  - Pumps 51
  - Push 10
  - Pylons 51
- Q**
- Quadrillion 55
  - Quadruple pulley 17
  - Quarks 59
    - hadrons 58
    - mesons 58
    - particle physics 52
    - strong nuclear force 59
  - Quartz 48
  - Quintillion 55
- R**
- Radian 55
  - Radiation
    - electromagnetic 7, 38-9, 57
    - electromagnetic spectrum 38-9
    - heat transfer 22
    - infrared 23, 58
  - Radio waves 38
  - Radioactivity 50
    - analyzing 50
    - cloud chamber 57
    - Geiger-Müller tube 58
    - geothermal power 37
  - Rails 24
  - Rarefaction 59
    - sound waves 21
    - springs 20
  - Reaction 10, 59
    - Newton's Laws 12
    - tension 59
  - Reactors 51
  - Real images 58
  - Red light
    - additive process 41
    - color vision 40
    - constructive interference 47
    - lenses 45
    - photons 38
    - subtractive process 41
    - white light spectrum 39
  - Reflecting telescope 45
  - Reflection
    - color 40
    - light 42-3
    - total internal reflection
  - Rocks 37
  - Rollers
    - ball bearings 15
    - reducing friction 2, 14, 15
  - Rotation, turbines 59
  - Rotor
    - electric motor 35
    - generators 36, 58
  - Rubber 25
  - Ruby lasers 47
- S**
- Sandpaper 2, 14
  - Sankey diagram 7, 59
  - Scales 8
  - Scrapyard electromagnets 35
  - Screens
    - cathode ray tubes 49
    - projecting images onto 43
    - television 49
  - Screw 16
  - Seas 37
  - Second 8
    - symbol 55
  - Secondary colours 41
  - Semiconductor 59
    - holes 58
    - LED (light-emitting diode) 40
    - solar cells 36, 37
  - Shell 59
    - electrons 48
  - Ships 27
  - Short-wave radio 38
  - Shutters 44
  - SI units (Système International d'Unités) 8, 59
    - conversion table 55
  - Silicon 37
  - Simple machines 16-17
  - Simple pulley 17
  - Single lens reflex (SLR)
    - cameras 44
  - Slopes 10
  - Smoke particles 6
  - Soap bubbles 5
    - colors 46, 47
    - surface tension 26
  - Soap powder 48
  - Sodalite 48
  - Sodium lamp 40
  - Solar cell 7
    - electricity generation 36, 37
  - Solar energy 36, 37
  - Solar System 19
  - Solenoid 35, 59
  - Solid angle 55
  - Solids 24-5
    - amorphous 24, 56
    - compression 57
    - crystal lattice 57
    - crystallization 6
    - crystals 24-5
    - dissolving 6
    - electromagnetic radiation 39
    - latent heat 58
    - melting 22, 58
    - particles 22
    - solutions 59
    - temperature 22
    - tension 24, 25, 59
    - thermal expansion 24, 59
  - Solution 6, 59
  - Sound energy 7
  - Sound waves
    - amplitude 56
    - frequency 21
    - vibrations 20, 21
    - wavelength 21
    - waves 20, 21
  - South-seeking pole 31
  - Space 59
  - Spectrum
    - colors 40
    - electromagnetic 38-9, 57
    - white light 39
  - Speed 13, 59
    - acceleration 12, 13, 56
    - air resistance 14
    - centripetal force 18, 56
    - formula 54
    - kinetic energy 58
    - of light 42
    - measurement 12, 13
    - terminal velocity 59
    - velocity 12, 59
  - Spinning object
    - angular momentum 56
    - gyroscopes 18, 19, 58
  - Springs 20
  - Square centimeters 55
  - Square feet 55
  - Square inches 55
  - Square kilometers 55
  - Square meters 55
  - Square miles 55
  - Square yards 55
  - Stability 19, 58
  - States of matter 59
    - gases 22
    - liquids 22
    - solids 22
    - supercooled liquid 22
    - and temperature 22
  - Static electricity 30
  - Stator 36
  - Steam
    - electricity generation 36
    - geothermal power 37
    - nuclear power stations 51
  - Steel 39
  - Steel rails 24
  - Steradian 55
  - Stimulated emission 46, 47
  - Strong nuclear force 59
    - gauge bosons 58
    - gluons 58
    - mesons 58
    - nuclear physics 50
    - particle physics 52
  - Structure 6
  - Sublimation 22
  - Substances
    - SI units 55
    - state of matter 22
  - Subtractive process 59
    - color 40
    - filters 41
    - primary colors 59
  - Sun
    - energy 7
    - gravity 12, 13
    - nuclear physics 50
    - orbit of planets 19
    - solar power 36, 37
    - temperature 22
  - Super proton synchrotron (SPS) ring 55
  - Supercooled liquids 56
    - particles 22
  - Surface tension 59
    - bubbles 26
    - cohesive forces 26, 57
  - Surfaces 14
  - Surgery 47
  - Suspension 11
- T**
- Switch gear 36
  - Switches 33
  - Symbols 55
  - Tailrace 36
  - Telescopes 44, 45
    - images 58
    - lenses 45
    - reflecting 45
    - refracting 45
  - Television 48, 49
  - Temperature 22-3, 59
    - absolute zero 56
    - Boyle's Law 28
    - Celsius scale 56
    - Charles' Law 29
    - conversion table 55
    - cryogenic unit 57
    - electromagnetic radiation 39
    - equalization of 23
    - Fahrenheit scale 58
    - filament lamps 40
    - gases 28-9
    - Kelvin scale 58
    - latent heat 58
    - measurement 22
    - melting point 58
    - Pressure Law 29
    - SI units 55
    - solids 24
    - thermal expansion 59
  - Tension 59
    - centripetal force 18
    - elasticity 57
    - oscillation 20
    - solids 24, 25
    - surface 26, 59
  - Tension force 16
  - Terminal velocity 14, 59
  - Tetragonal system 24
  - Theories 8
  - Thermal expansion 24, 59
  - Thermometer
    - digital 9
    - measuring heat 8
    - measuring latent heat 2
    - mercury 9
  - Thin film interference 47
  - Thousand 55
  - Tidal power 37
  - Time
    - freefall experiment 9
    - measurement 8
    - SI units 55
    - speed 59
  - Tonnes 55
  - Tons 55
  - Total internal reflection 42, 59
    - binoculars 42
    - critical angle 57
  - Transfer of energy 7
  - Transformer
    - hydroelectricity 36
    - nuclear power stations 51
  - Transmission 42
  - Transverse waves 20
  - Traveling microscope 8
  - Triclinic system 25
  - Trigonal system 25
  - Trillion 55
  - Tubes 26
  - Tuning fork 21
  - Turbine 59
  - geothermal power 37
  - hydroelectricity 36
  - nuclear power stations 51
  - tidal power 37
  - wind power 37
- U**
- Turning force 11
    - formula 54
    - moment 58
    - rollers 14, 15
    - screw 16
  - Ultraviolet radiation
    - electromagnetic spectrum 39
    - fluorescence 48, 58
  - Unit cell 24, 59
  - Units of temperature 22
  - "Up" quarks 52
  - Upthrust 59
    - hot-air balloons 29
    - immersed objects 26, 27
  - Upward meniscus 26
  - Uranium 51
- V**
- Vacuum 49
  - Van de Graaff generator 5, 30-1
  - Vapor 58
  - Variable resistors 54
  - Velocity 12, 59
    - light 55
    - momentum 13
    - terminal 14, 59
  - Venus 19
  - Vernier calipers 8
  - Vernier scale 59
    - traveling microscope 8
  - Very high frequency (VHF) radio 38
  - Vibration
    - electromagnetic radiation 39
    - frequency 21, 58
    - oscillation 20
    - sound 20, 21
    - thermal expansion 24, 59
  - Viewfinder
    - camera 44
    - telescope 45
  - Violet light 39
  - Virtual images 58
  - Viscosity 55
  - Vision
    - color 40, 57
    - seeing by reflected light 42
  - Voice coil 21
  - Voltage 59
    - electric circuits 52
    - formulas 54
    - generators 36, 58
    - Ohm's Law 32
    - resistance 59
    - SI units 55
    - Van de Graaff generator 5
    - wind power 37
  - Voltmeters 2
    - electric circuit 55
  - Volts 32
    - symbol 55
  - Volume 59
    - Charles' Law 29
    - conversion table 55



density 57	tidal power 37	constructive interference 57	air resistance 14	Willemite 48	wavelength 59
gases 28-9	upthrust 27	destructive interference 57	component 10	Wind farms 37	
	wave behavior 46	diffraction 57	formulas 54	Wind power 36-7	<b>Y</b>
<b>W</b>	waves 56	electromagnetic radiation 38	lifting with pulleys 17	Wire	Yards 55
W particles 52	Watts 55	frequency 58	measurement 8	electricity generation 36	Yellow light
weak interaction 59	Wavelength 20, 59	interference 46, 58	resultant force 10	electromagnetism 34-5, 57	additive process 41
Water	chromatic aberration 56	light 47	upthrust 27	solenoids 59	sodalite 48
boiling point 22, 56, 58	colors 40	plane 46, 59	Welding metals 22	Wood 22	subtractive process 41
capillary action 26, 56	electromagnetic radiation 38-9	Principle of Superposition 46, 59	Wheels	Work 59	white light spectrum 39
electricity generation 36	electromagnetic spectrum 57	sound 20	axles 17	formula 54	
evaporation 6	infrared radiation 58	in springs 20	ball bearings 15	simple machines 16	<b>Z</b>
freezing point 22, 56, 58	photons 59	Weak interaction 59	White light	Worm gear 7	Z particles 52
geothermal power 37	sound waves 21	gauge bosons 58	additive process 41		weak interaction 59
meniscus 26	symbol 55	Higgs bosons 58	color vision 40	<b>X</b>	Zinc 48
molecules 58	white light spectrum 39	particle physics 52	destructive interference 46-7	X-rays 39	
nuclear power stations 51	Waves 20-1	Wedges 16	radiation from hot objects 39	electromagnetic radiation 38	
pressure 27	amplitude 56	Weight 59	spectrum 39	electromagnetic spectrum 39	
solutions 6	behavior 46-7		Width 8		
surface tension 26					

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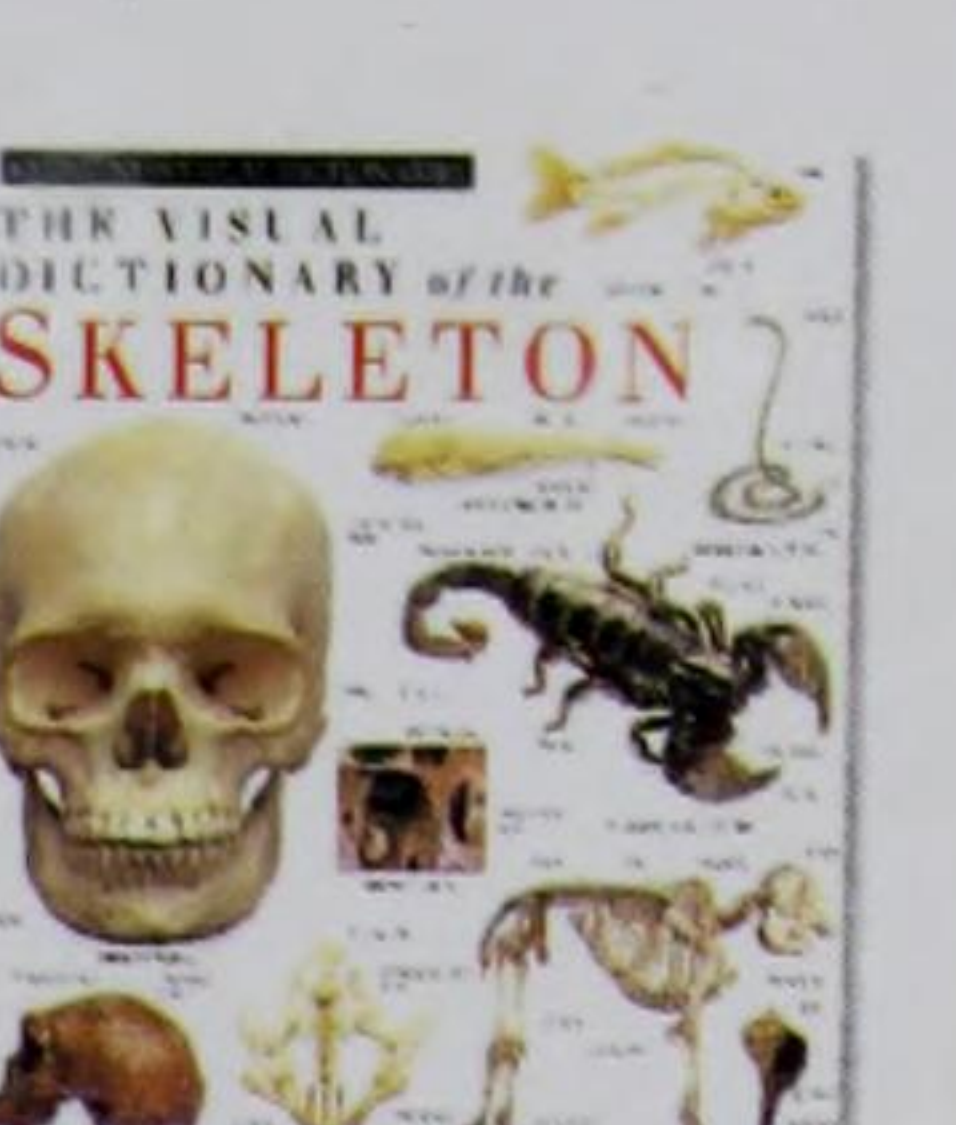
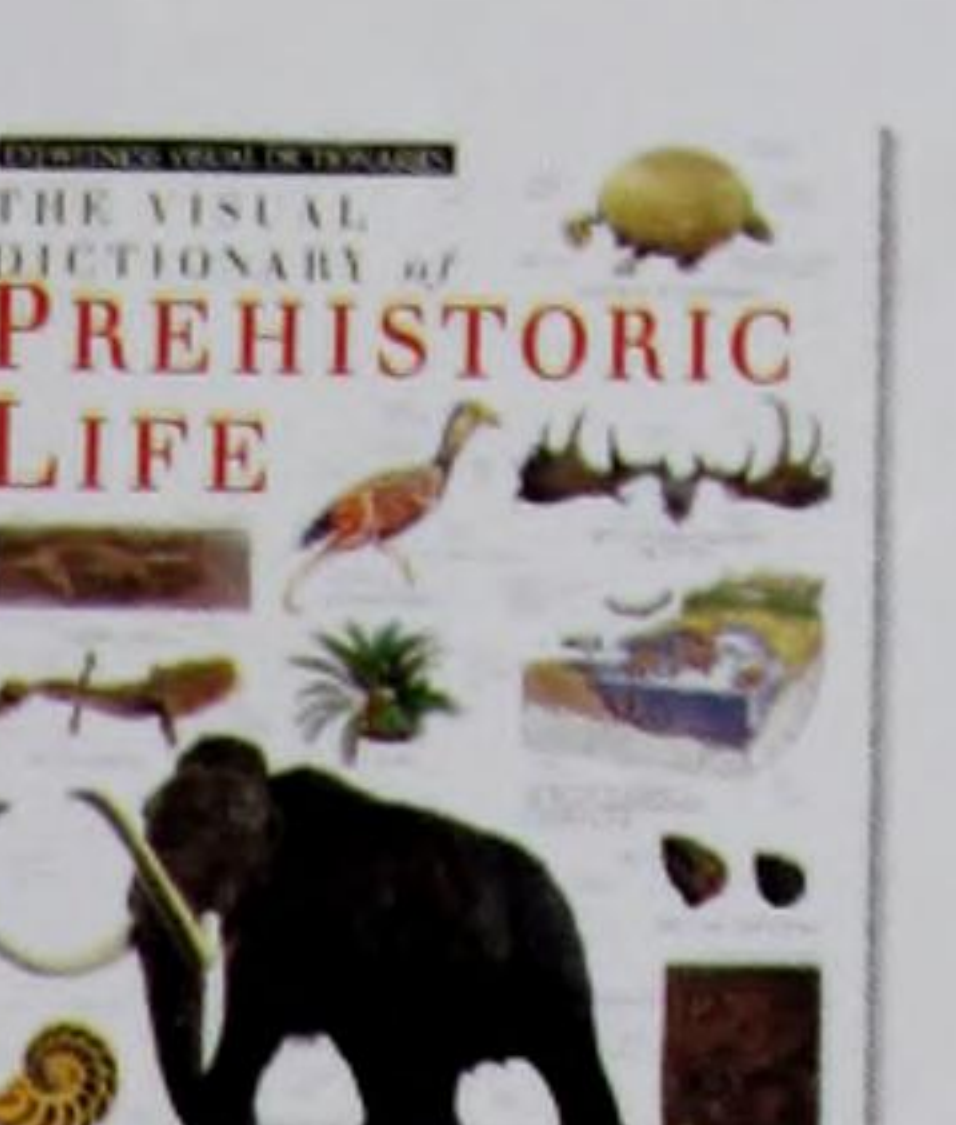
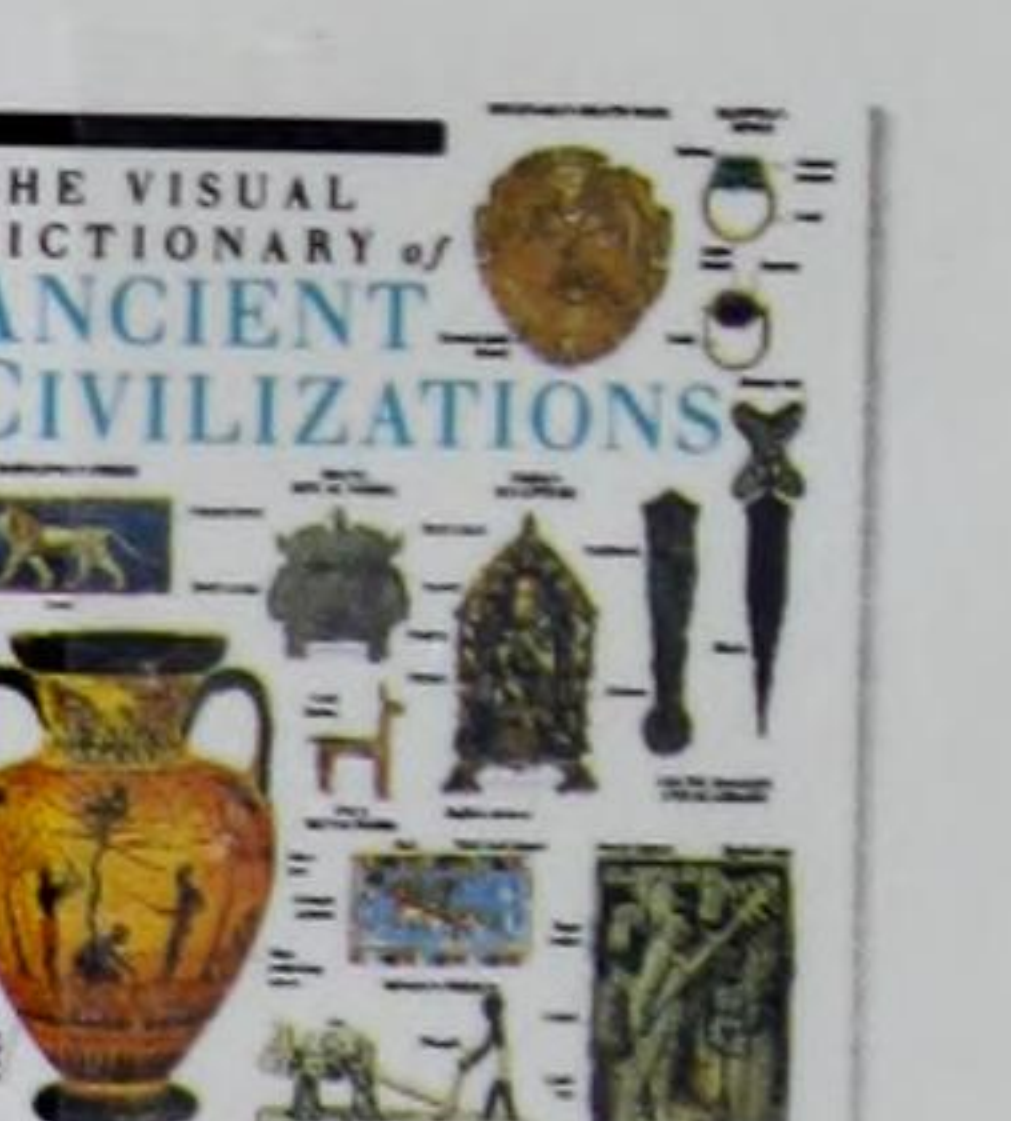
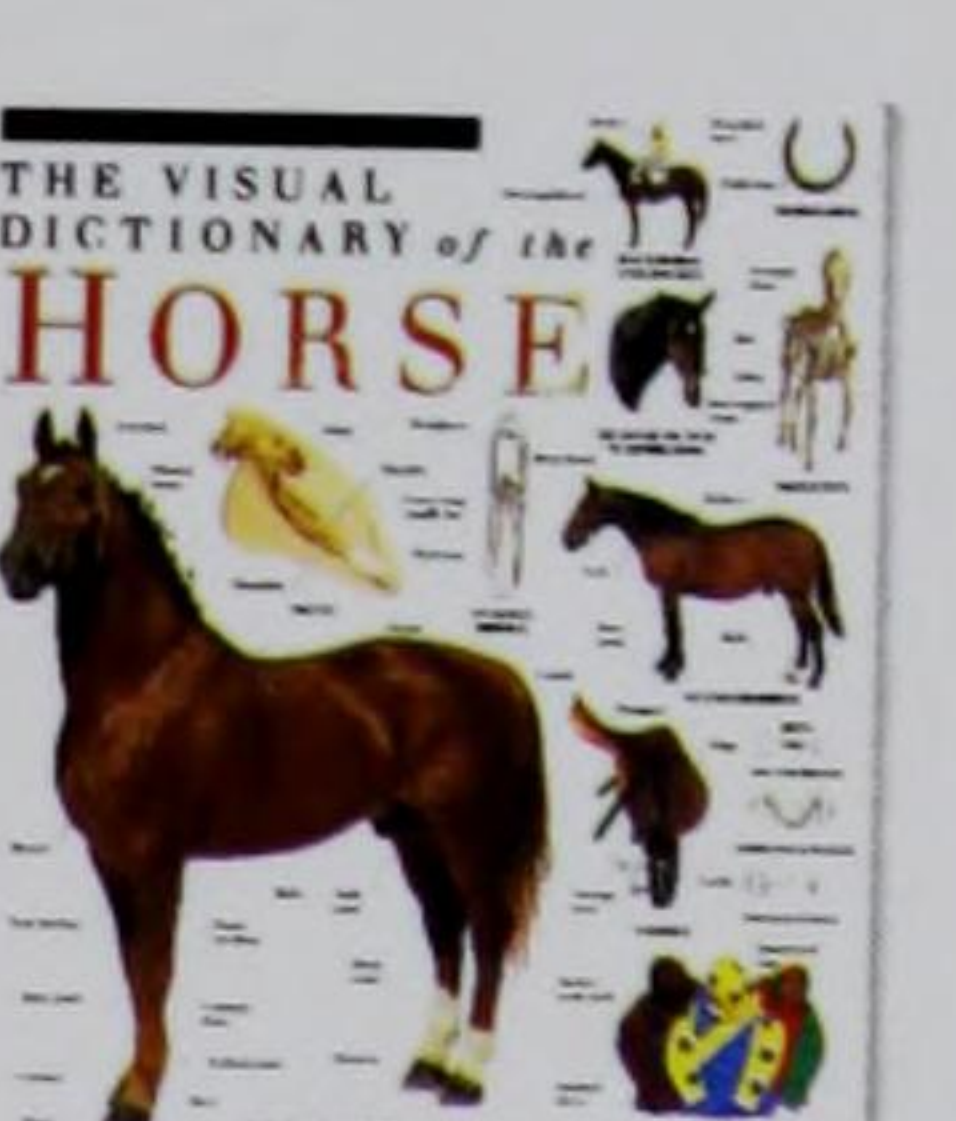
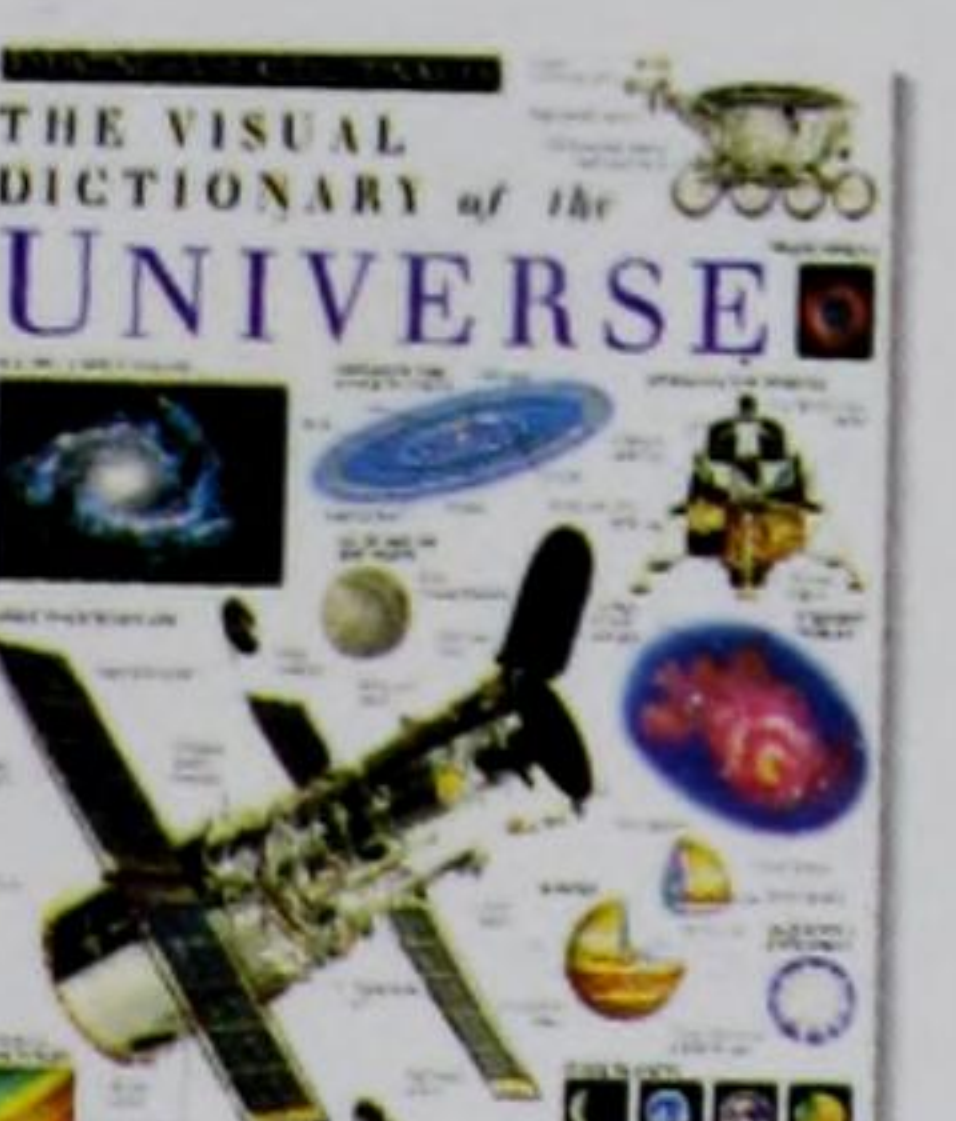
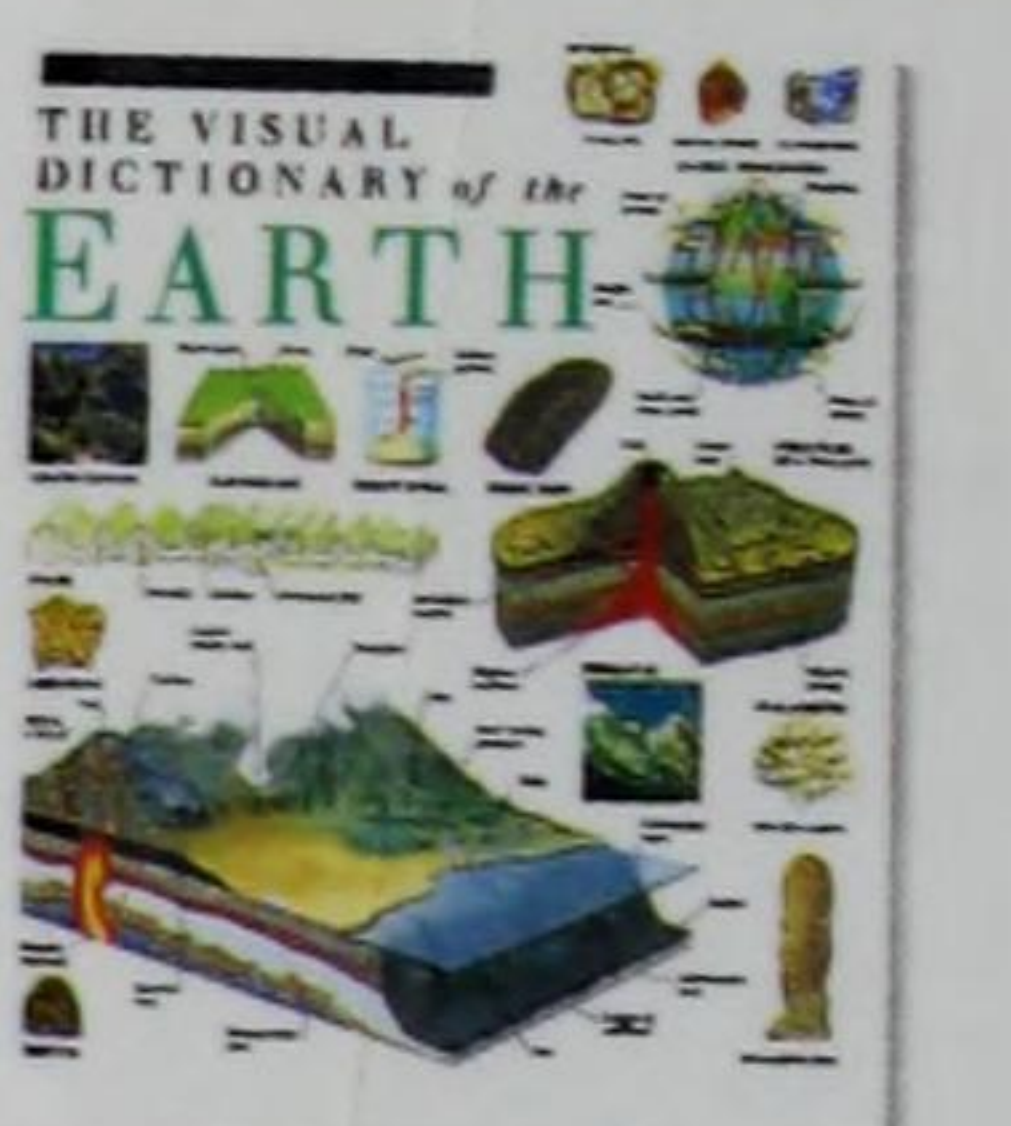
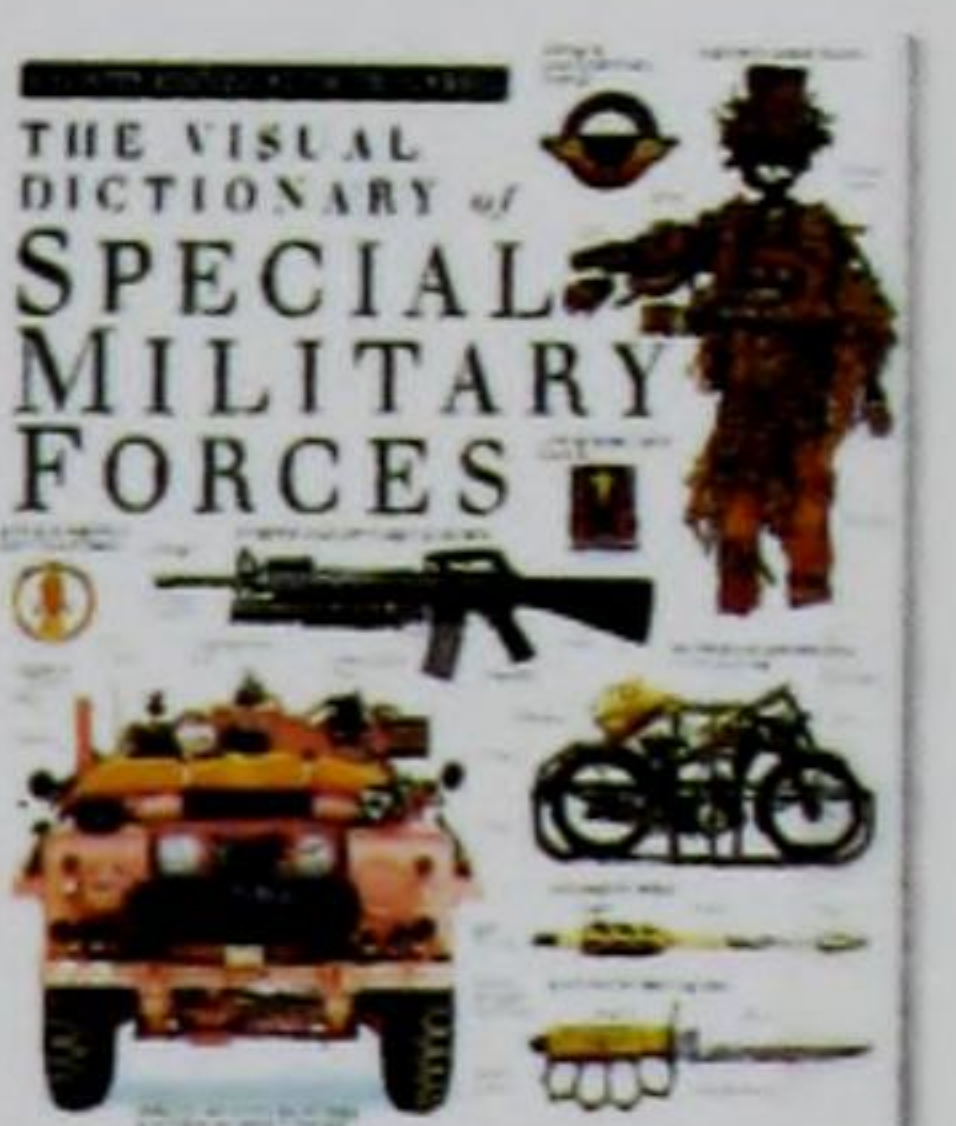
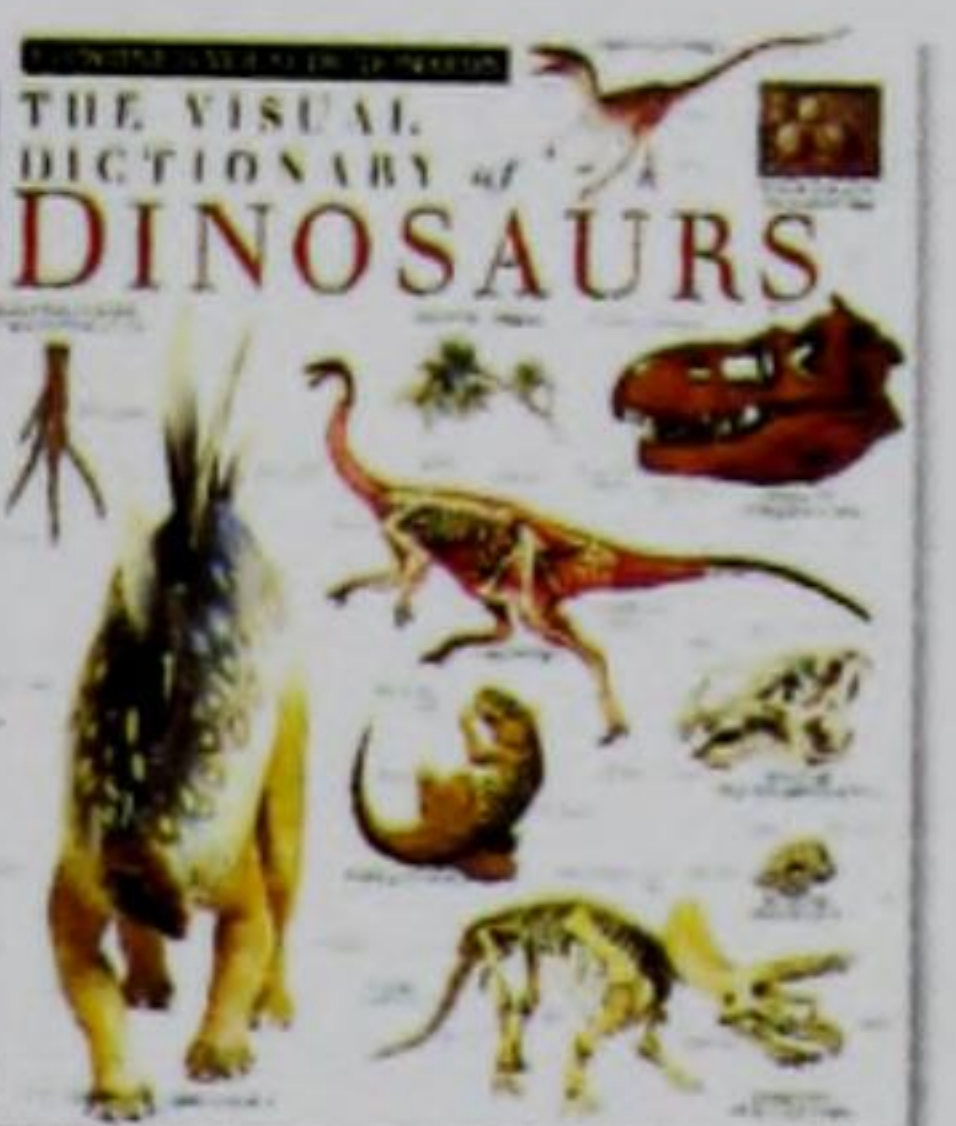
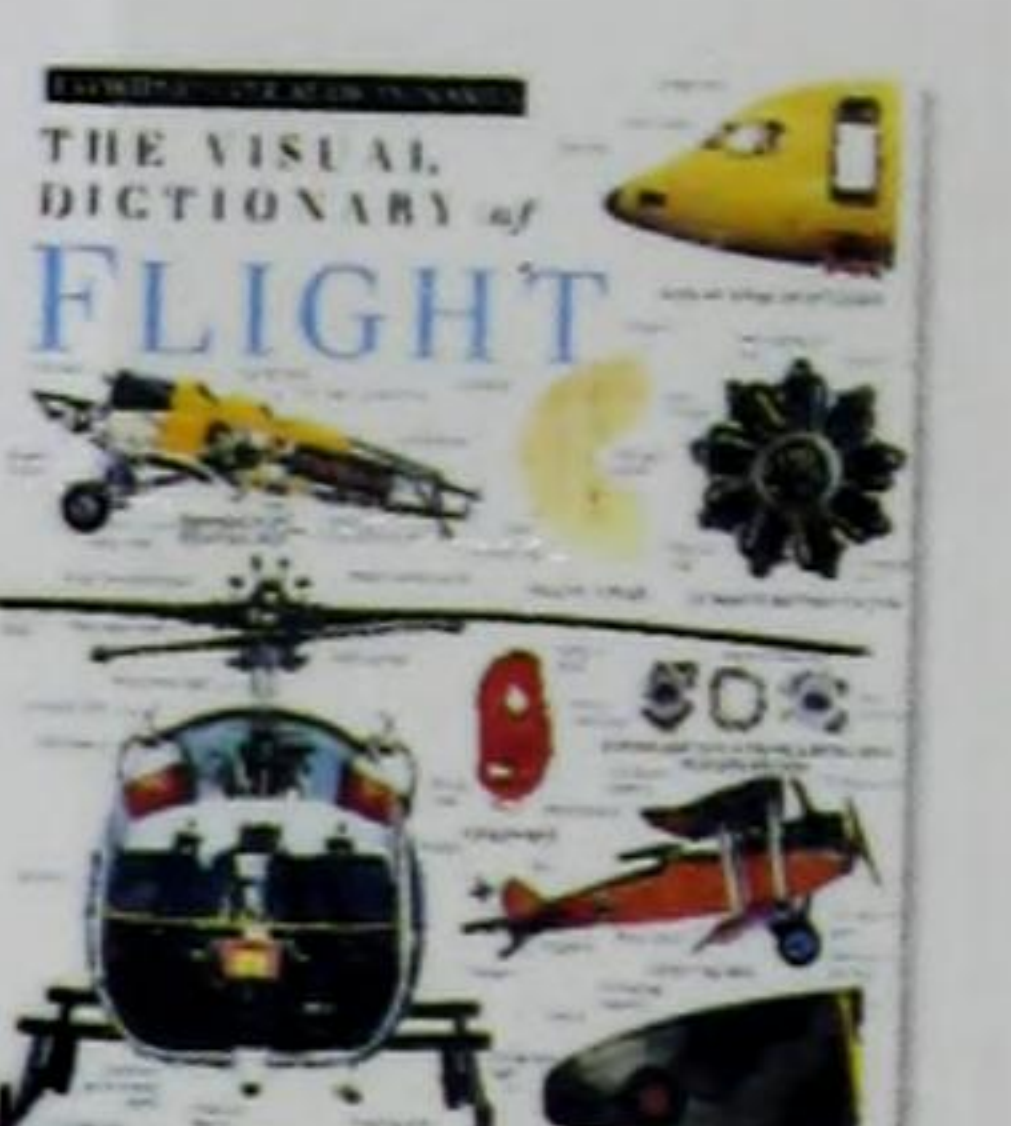
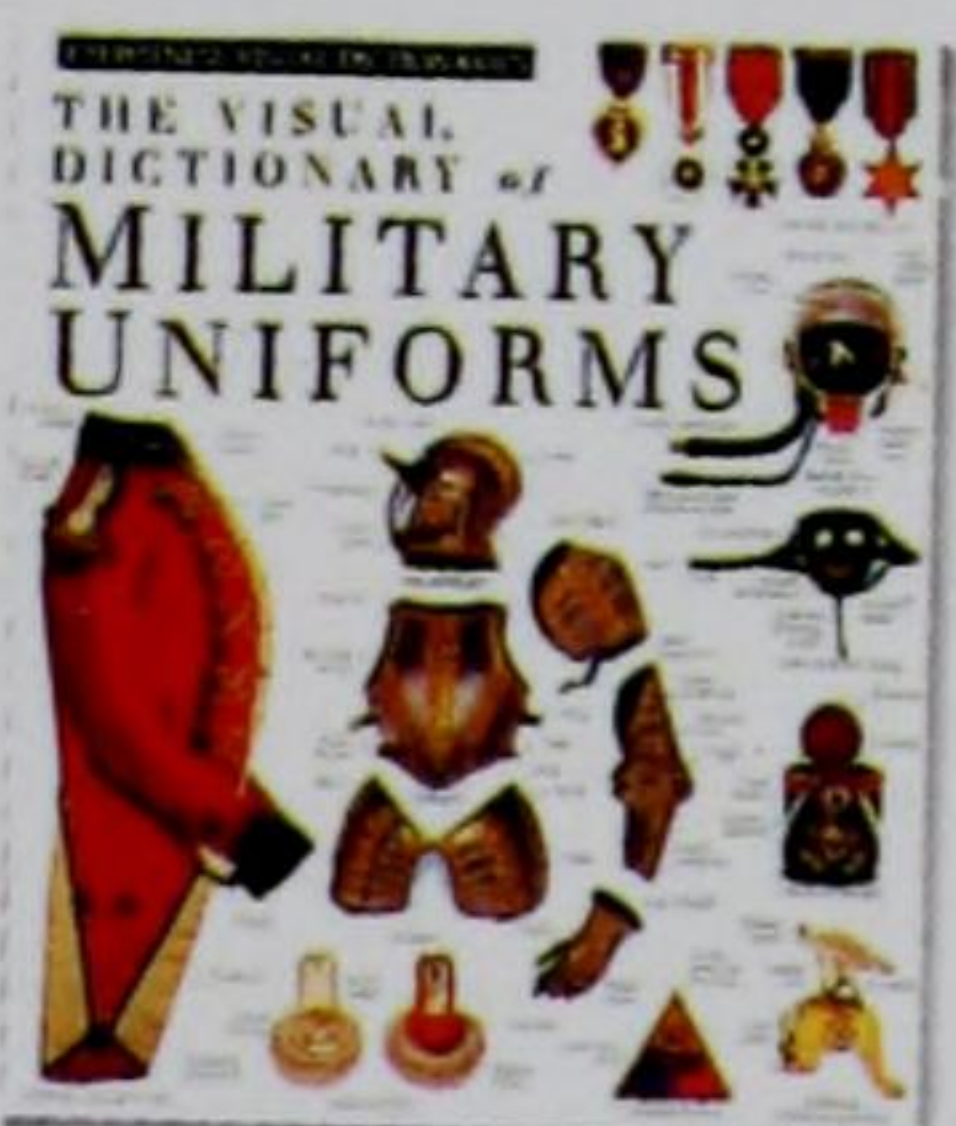
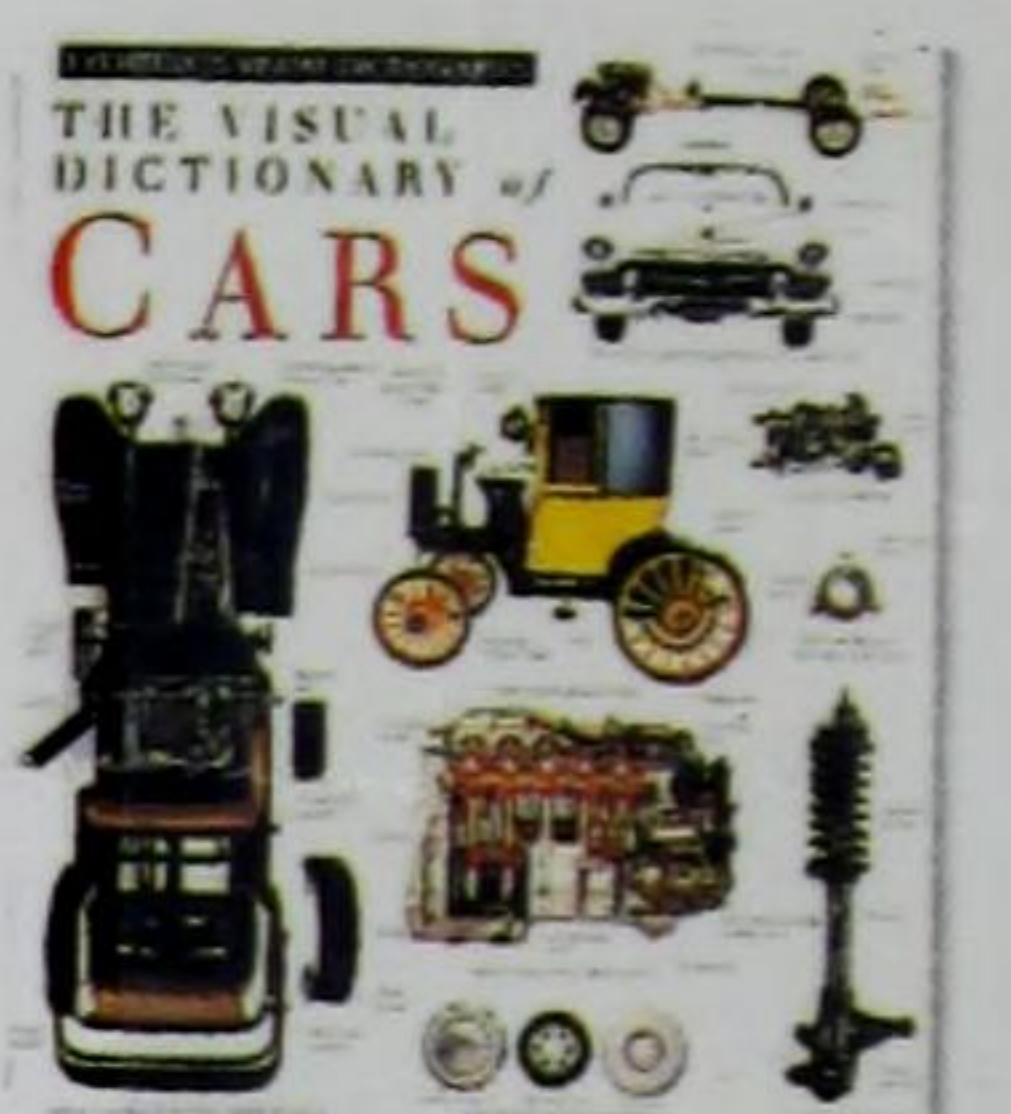
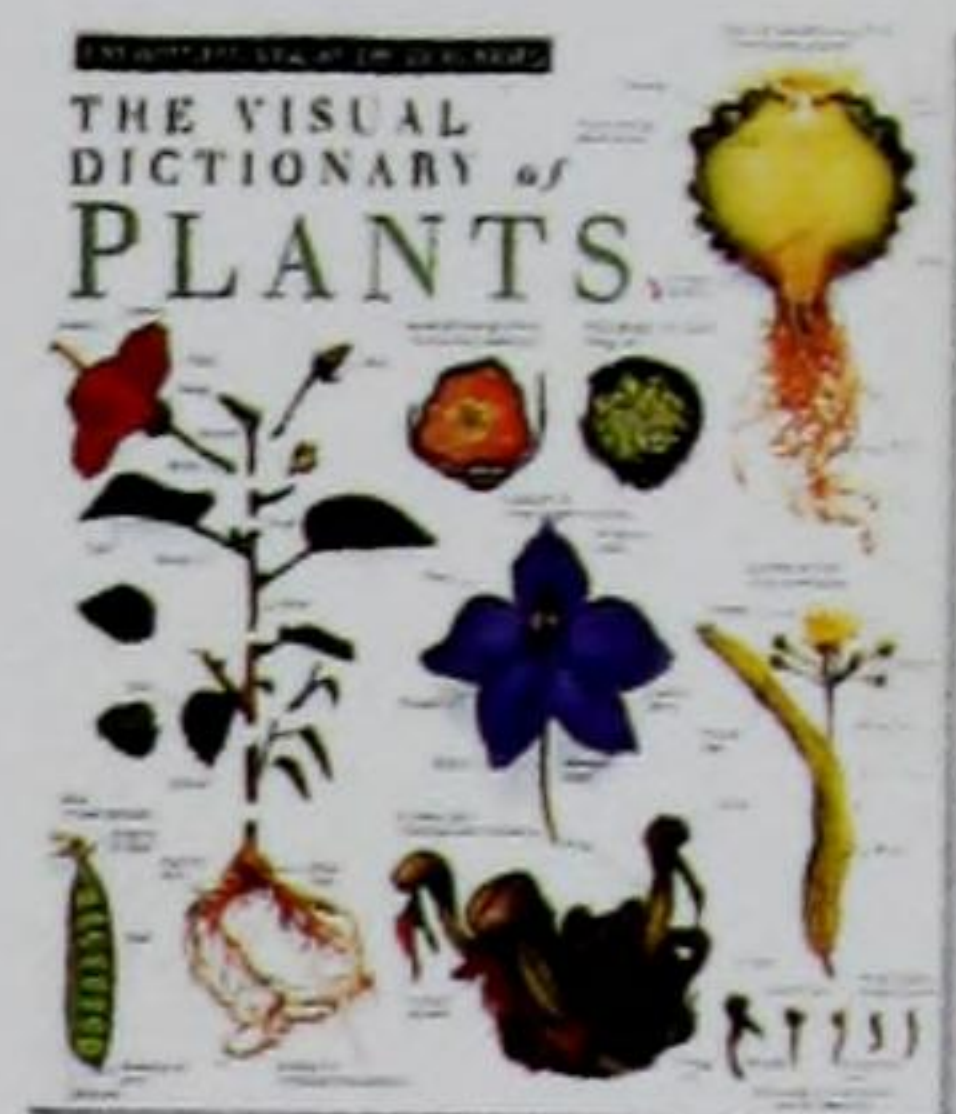
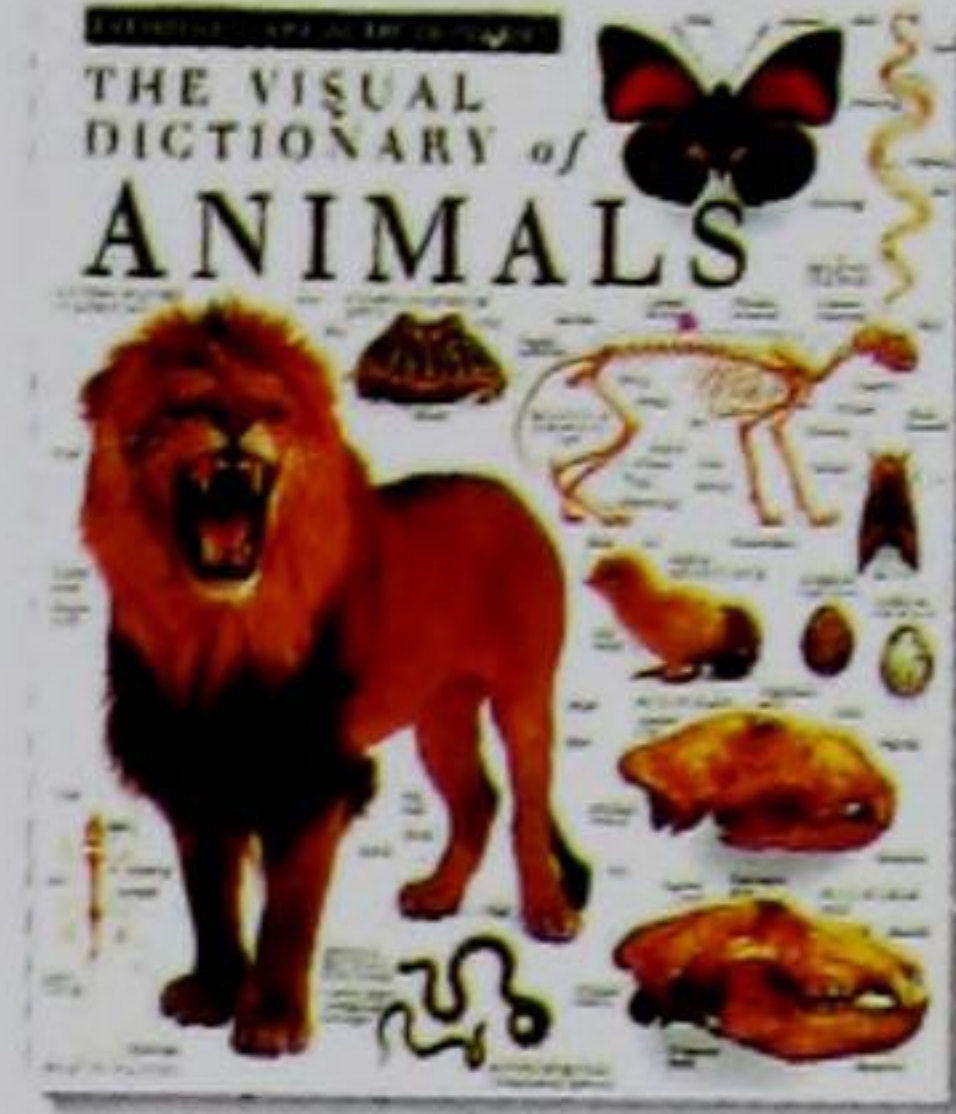
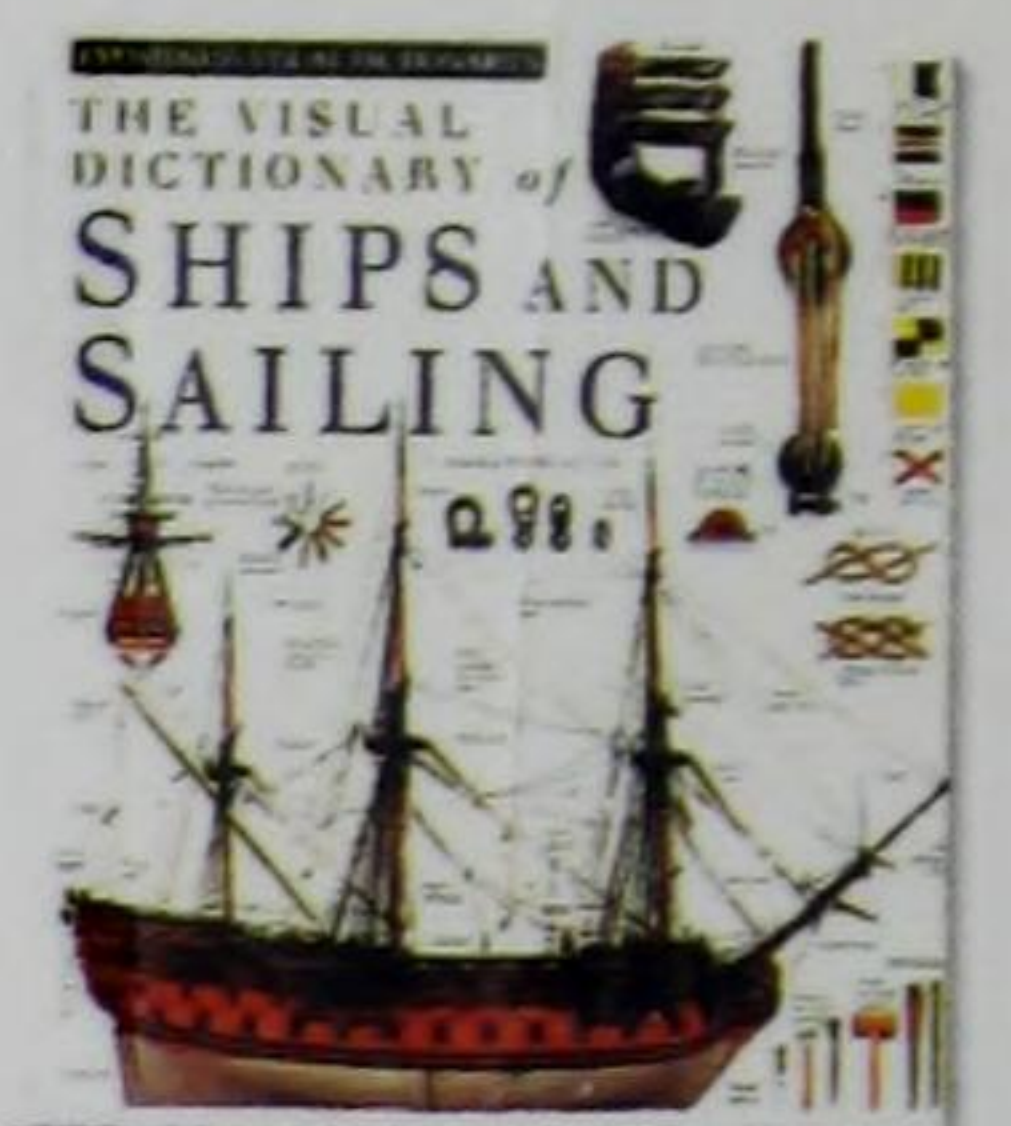
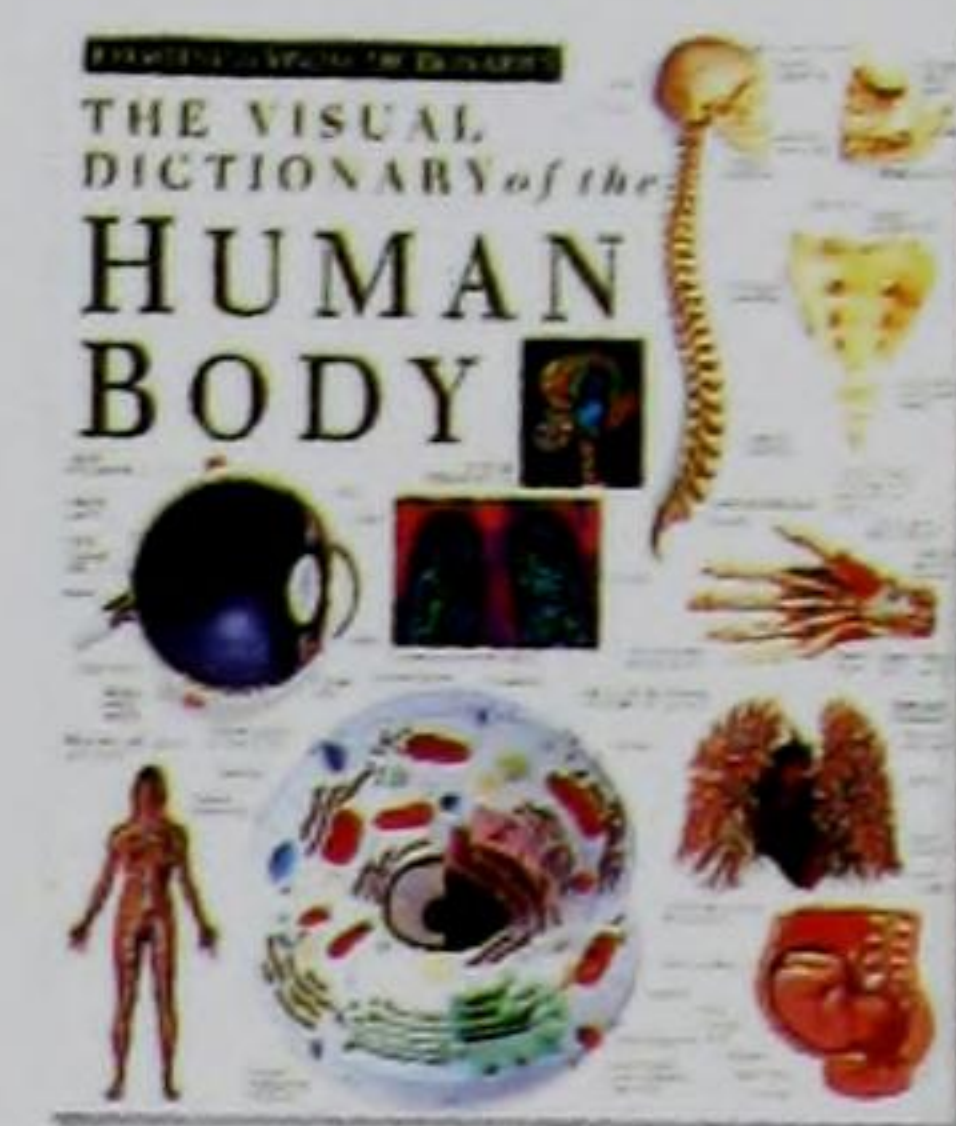
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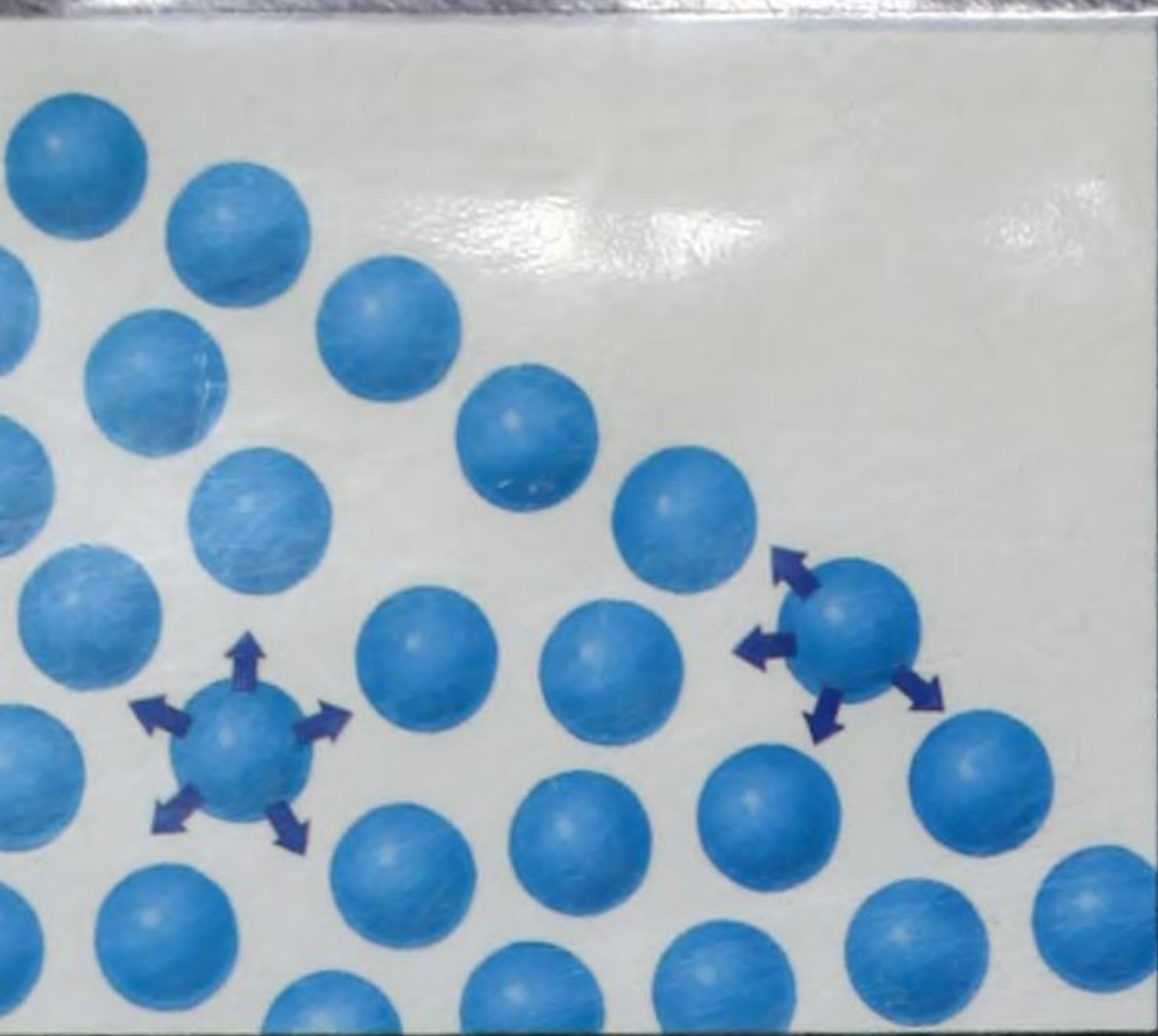
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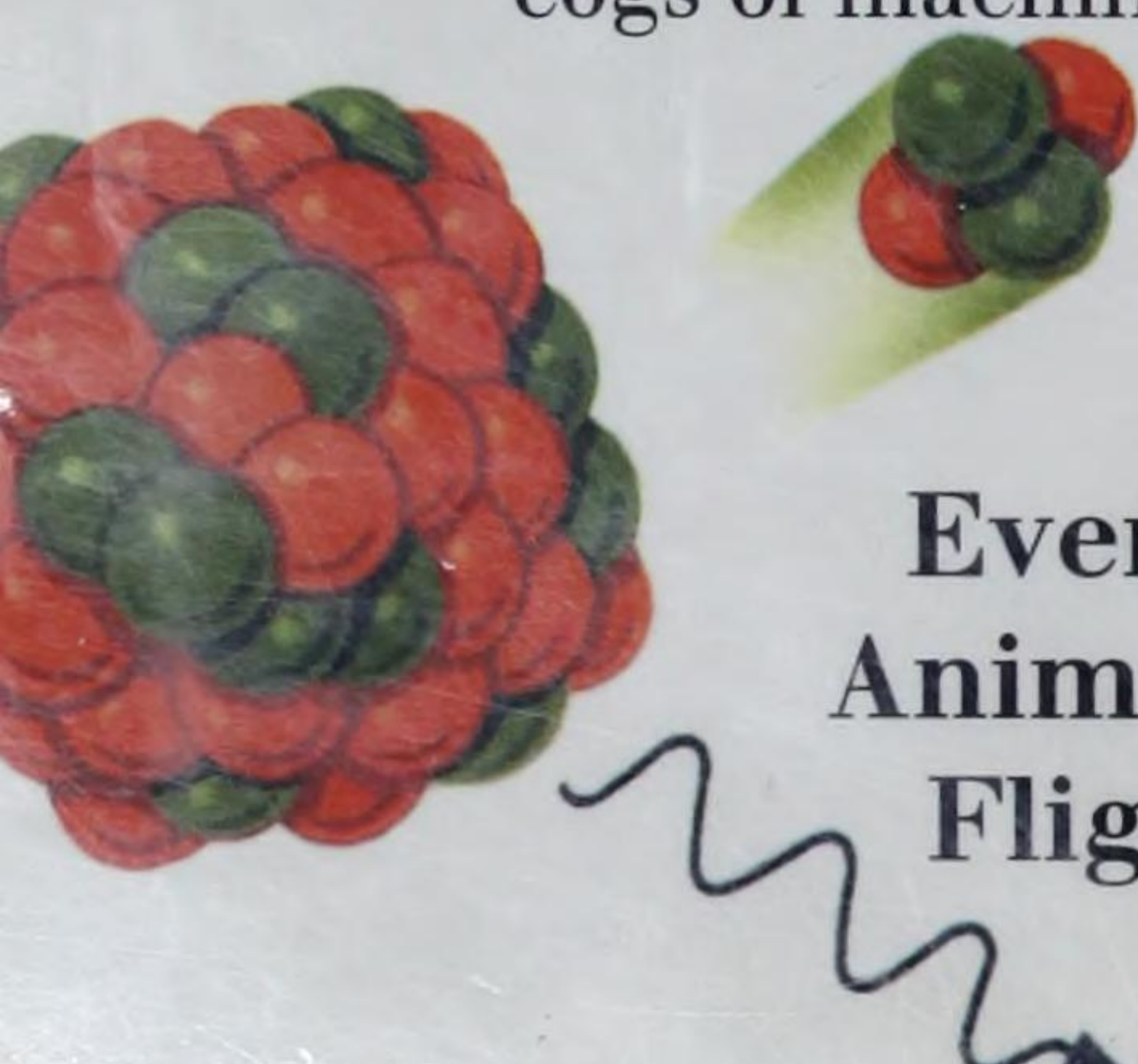
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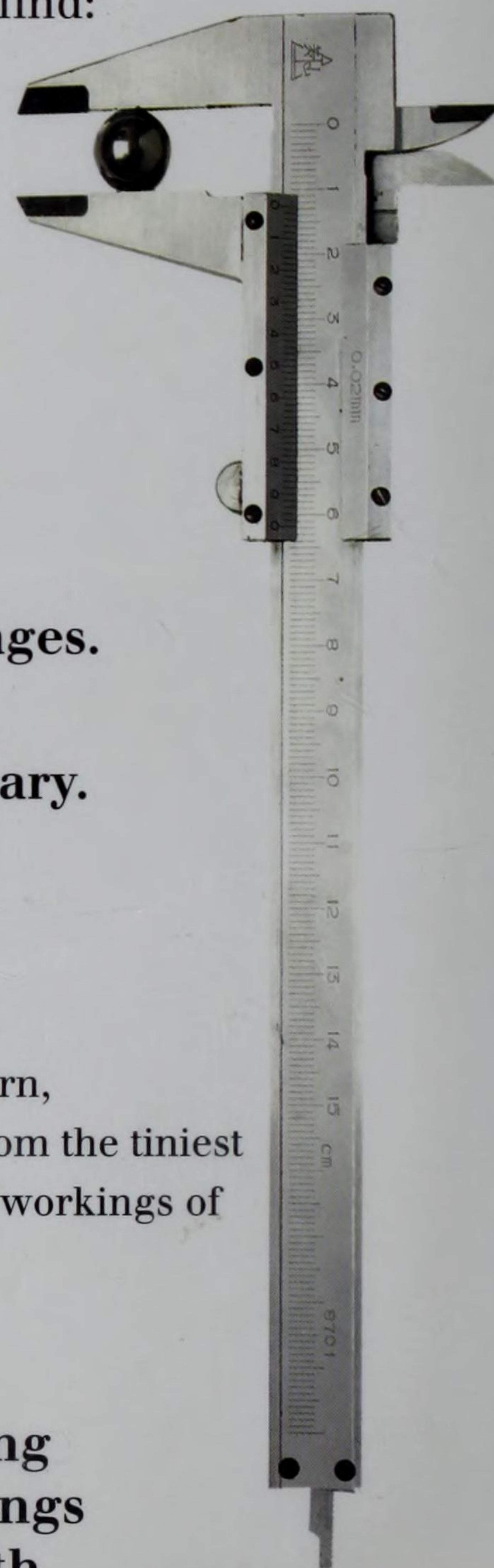
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